

Appendix A

Scoping Report

Troy Mine Revised Reclamation Plan
Environmental Assessment

Scoping Report

March 10, 2008

Contents

Section 1 Introduction

Section 2 Background

2.1	Troy Mine History	2
2.2	Project Management Plan	4
2.3	Public Involvement Process	4

Section 3 Implementation of the Scoping Process

3.1	Press Release.....	5
3.2	Scoping Document.....	5
3.3	Open House and Public Meeting	6

Section 4 Results of the Scoping Process

4.1	Issues Raised During the Scoping Process	7
4.1.1	Water Quality	7
4.1.2	Cover Source Materials	8
4.1.3	Impoundment Stability	8
4.1.4	Buried Drums	8
4.1.5	Subsidence	8
4.1.6	Other	8
4.2	Issues Considered But Not Recommended For Further Evaluation	9
4.3	Issues to be Considered in the EA	10
	Mine Water	10
	Groundwater Protection	10
	Long Term Prognosis/Monitoring	10
	Water Treatment	11
	Cover Source Materials	11
	Impoundment Stability	11

Section 5 Identification of Data Gaps

5.1	Water Treatment Alternatives.....	12
5.2	Copper Attenuation Process	12
5.3	Surface Water Quality	13

Section 6 References

Appendices

<i>Appendix A</i>	Press Release
<i>Appendix B</i>	Scoping Document
<i>Appendix C</i>	Mailing List for Open House/Public Meeting
<i>Appendix D</i>	Newspaper Advertisements for Open House/Public Meeting
<i>Appendix E</i>	Sign in Sheets for Open House/Public Meeting
<i>Appendix F</i>	Written Comments Received During the Scoping Process

Section 1

Introduction

CDM was contracted by the Montana Department of Environmental Quality (DEQ) in October 2007 to begin the first phases of an Environmental Assessment (EA) for the proposed reclamation at the Troy Mine near Troy, Montana. The contract includes four task orders:

- Task Order 1 –Reclamation Plan Review and Public Scoping Activities
- Task Order 2 – Administrative Record and Project Management Plan
- Task Order 3 – Data Gap Analysis
- Task Order 4 – Alternatives Development, Initial Evaluations, and Environmental Assessment

Task Order 1 includes two subtasks: Review Permit and Proposed Reclamation Plan and Public Scoping and Public Input Activities. CDM has completed the first task order and is submitting this scoping report in accordance with the deliverables list provided in Task Order 1.

Public scoping is the first step in conducting an EA. It is a process for determining the scope of issues to be addressed and identifying the significant issues related to a proposed action. The scoping process helps in determining what will be reviewed in the EA and the level of analysis, in part, through the collection of written and verbal comments from the public. Scoping helps agencies identify environmental issues associated with the project and aids in the development of reasonable reclamation alternatives. It also allows effective public and stakeholder involvement prior to the submittal of the draft EA.

This report describes the results of the scoping activities conducted between October and December 2007 as part of the EA. It includes:

- Introduction
- Background
- Implementation of the Scoping Process
- Results of the Scoping Process
- Identification of Data Gaps

Section 2

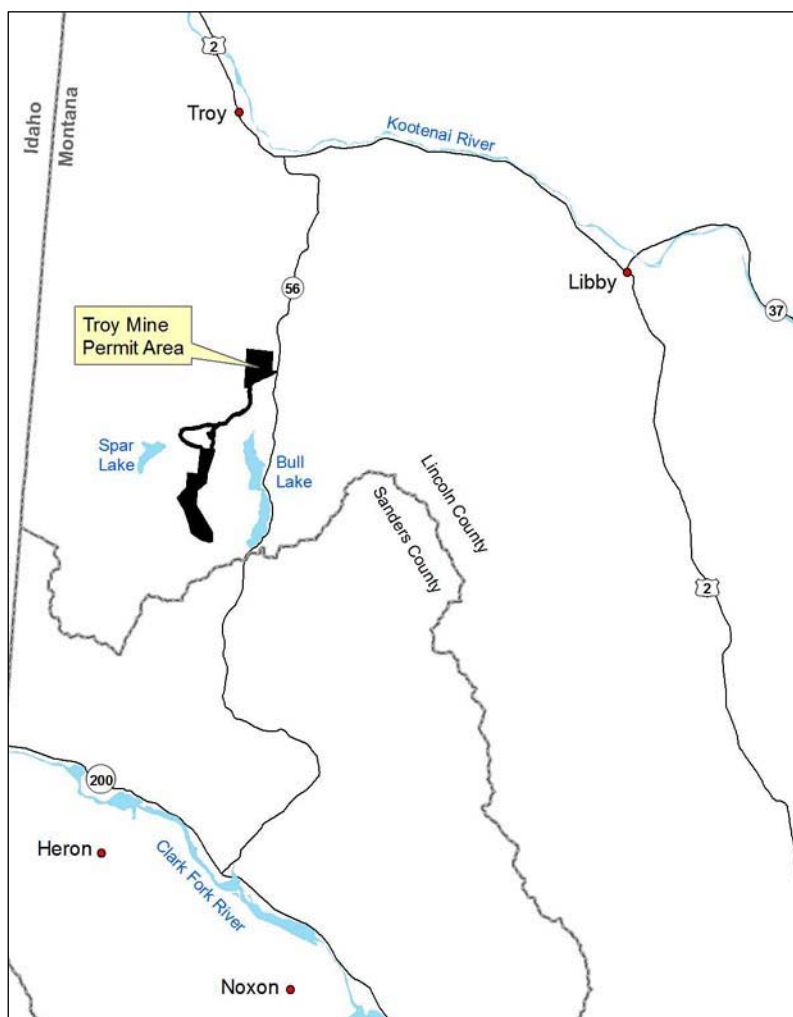
Background

2.1 Troy Mine History

The Troy Mine is located in Lincoln County, Montana, about 15 miles south of Troy. The mine permit area covers 2,635 acres of public and private land off MT Highway 56 (see figure below). It consists of an underground mine, mill, and office facilities; tailings and water pipelines; power line; tailings impoundment; and associated support facilities.

The ASARCO, Inc. Troy Mine project was permitted in 1978 by the U.S. Forest Service (USFS) and State of Montana as a 20,000 ton-per-year underground copper/silver mine. The ore was and is mined using the “room-and-pillar method.” ASARCO began full production in 1982 but put the mine on standby mode in 1993 due to depressed metal prices. During the shut-down period, ASARCO sold its interest to Sterling Mining Company, which became Revett Silver Company (Revett). Production resumed in 2005 under Genesis (a wholly owned subsidiary of Revett) and is projected to continue at least 5 years until the ore body is depleted.

The Troy Mine is now permitted and operated by Genesis. Genesis currently has posted a \$10.5 million reclamation bond for the project. The reclamation bond is being updated to \$12.9 million. Genesis submitted the Revised Reclamation Plan in March of 2006. The revised plan describes proposed reclamation elements for final closure of the Troy Mine following cessation of mining activities.



Over the course of mine operation, knowledge has been gained through studies, data collection, revegetation test plots, prior reclamation plan development and reviews, and bonding determinations. In addition, some conditions have changed since the 1979 planning

and permitting documents were prepared. One changed condition is the presence of toe ponds at the perimeter of the tailings facility dam. Another is the change in state water discharge standards which do not allow direct discharge of produced mine water into Stanley Creek. The Revised Reclamation Plan attempts to incorporate and address all of the above issues and changes.

The Revised Reclamation Plan addresses facility removal for many developmental and operational features including:

- Roads
- Waste rock dumps
- Mine portals
- Underground facilities
- Mill and plant site
- Soil borrow sites
- Power supply – transmission line, generators, fuel tanks, etc.
- Water supply and pump facilities
- Sewage treatment facility
- Pipelines
- Tailings impoundment
- Monitoring wells
- Fences

The Revised Reclamation Plan proposes to use concrete plugs in the service and conveyor adits. These plugs would be designed to avoid the possibility of large-scale mine discharge or overflow from one of the adits. Other water treatment options include piping the mine water discharge to the tailings decant ponds for direct groundwater discharge.

The Revised Reclamation Plan also contains reports and plans including:

- Mine Flooding Report
- Revegetation Status Report
- Assessment of Fate and Transport of Copper in Decant Pond Disposal System
- Soil Investigation Results
- Weed Control Plan
- Water Monitoring Plan
- Revegetation Plan
- Geotechnical Monitoring Plan
- Tailing Facility Stability Report
- Conceptual Portal Plug Study

In addition, a report has been prepared that discusses the cause of two surface subsidence events and potential future subsidence issues.

2.2 Project Management Plan

The project management plan serves as a guide for the Troy Mine Revised Reclamation Plan EA process. It establishes responsibilities, schedules, and procedures for the project team. It includes a description of the project team and potential cooperating agencies, contact information, project tasks, a task schedule, known data gaps, issues of concern, the public involvement strategy, a preliminary draft of the proposed action and alternatives, the format for the EA, and a description of how the Administrative Record will be compiled.

2.3 Public Involvement Process

Early attention to consensus building generally makes the project move more smoothly by ensuring that the public has an opportunity to voice their concerns and to be part of the overall decision making process. CDM and DEQ will gather public input using various components of the scoping process.

The public involvement program also includes the preparation of a scoping document to be distributed before the open house and the preparation of this scoping report. CDM will also assist DEQ in compiling important EA-related documents for an information repository. This will include fact sheets and newspaper articles, as well as copies of the draft and final EA.

Section 3

Implementation of the Scoping Process

The following scoping activities were completed between October and December 2007 as part of the scoping process for the Troy Mine Revised Reclamation Plan EA:

- Press Release
- Scoping Document
- Open House / Public Meeting

The scope of each of these elements and their implementation is discussed below. The issues identified during the scoping process are provided in Section 4.

3.1 Press Release

A press release was published by the DEQ on October 11, 2007. The press release provided notice of the scoping period for the Troy Mine Revised Reclamation Plan EA process. The press release is shown in Appendix A.

3.2 Scoping Document

A scoping document (Appendix B) was prepared and distributed prior to the open house and public meeting. The scoping document included the following topics:

- Opportunities for public involvement
- Troy Mine history
- Components of the Revised Reclamation Plan
- Initial issues of concern
- Overview of the EA process
- EA schedule and deliverables
- Sources of additional information

The scoping document was reviewed and approved by DEQ and USFS prior to being finalized. It was distributed by mail to approximately 88 individuals on the mailing list provided by DEQ. The mailing occurred on October 15, 2007. Approximately 20 individuals and 68 government organizations, elected officials, or private interest groups were included on the mailing list (Appendix C).

3.3 Open House and Public Meeting

The public meeting and open house were held at the Kootenai Senior Citizens Center in Troy, Montana on Tuesday, October 30, 2007. The public meeting began at 6:30 pm with a half-hour presentation. An open house followed the presentation and continued until 8:30 pm when the meeting was closed.

An advertisement was prepared by USFS for publication in four area newspapers: the *Western News* on October 24, 2007, the *Sanders County Ledger* on October 25, 2007, the *Daily Inter Lake* on October 21, 2007, and the *Bonner County Daily Bee* on October 24, 2007. These four ads are included in Appendix D.

The format for the public meeting was agreed upon with DEQ prior to the event. CDM started the meeting with a 20-minute presentation that introduced the EA team and discussed the scope of the EA, history of the mine, and potential remedial alternatives. Gwen Pozega of CDM led this presentation.

Seven tables, each representing a particular topic, were set up in a large meeting room. Each table was staffed by one or more CDM, AMEC, USFS, or DEQ employees. Two copies of the revised reclamation plan were available for review. Individual topic areas and their respective representatives were:

- MEPA/NEPA - Kathy Johnson (DEQ) and John Gubel (USFS)
- Engineering/Geotechnical - Charlie Freshman (DEQ) and Darrel Stordahl (CDM)
- Hydrogeology/Water Quality - Wayne Jepson (DEQ), Jim Castro (DEQ), Rebecca Ridenour (DEQ), John McKay (USFS), and Kent Whiting (CDM)
- Reclamation/Soils/Vegetation - Patrick Plantenberg (DEQ), Lisa Boettcher (DEQ), and Steve Prieve (USFS)
- Permit Amendment Process - Herb Rolfes (DEQ)
- Wildlife/Fisheries - Pat Mullen (AMEC) and John Carlson (USFS)
- Troy Mine/Genesis Inc.- Genesis Inc. Representative

Thirty-three people registered on the sign in sheet for the open house (Appendix E). Attendees were encouraged to move freely from table to table, depending upon their interests. During discussion with the technical representative at each table, notes of the discussions were recorded on a flip chart for use in summarizing the event.

Section 4

Results of the Scoping Process

The comments received during the scoping process have been summarized into major issues. Copies of the written comments are provided in Appendix F. This section presents the major issues and identifies those that will be carried forward for further evaluation in the EA. Issues that will not be carried forward were eliminated because of technical impracticability. Although issues may have been eliminated from further evaluation, they will still have an impact on the EA process because of the background information and community insight they provided.

4.1 Issues Raised During the Scoping Process

Issues raised during the scoping process are summarized by major topic area in Sections 4.1.1 through 4.1.6. The topic areas are:

- Water Quality
- Cover Source Materials
- Impoundment Stability
- Buried Drums
- Subsidence
- Other

4.1.1 Water Quality

Water quality issues and concerns identified during the scoping activities include:

- Mine Water – Will there be increased flows to the impoundment area? What are the potential impacts of increased flows on water quality, the toe ponds, etc? What is the level inside the mine that water would reach (at a “sustainable” rate) using portal plugs? Why not replicate the decant ponds at the mill site and remove the seven-plus miles of tailings pipelines?
- Groundwater Protection – How does the Reclamation Plan address long-term seepage/groundwater discharge to Lake Creek? Why is groundwater an issue at this time, after 20 years of water being diverted to the decant ponds? Sealing mine openings is not reasonable; may pose future threat of releasing contaminated water should structural failure occur. Toe ponds leak directly into Lake Creek. Discharges from the mine have begun to appear in the vicinity of Ross Creek. An MPDES permit should be required for discharge of mine water.

- Long Term Prognosis/Monitoring – What will copper concentrations be in the future? Will copper reach Lake Creek? If so, when? The monitoring proposed in the Revised Reclamation Plan is inadequate in every respect; monitoring “in perpetuity” is impossible.
- Water Treatment - Can mine water be treated in the service adit prior to discharge? Is there any treatment of discharged mine water? Several comments stated a preference for treatment water now rather than long-term (indefinitely). Perpetual discharge of polluted water is a poorly conceived idea.

4.1.2 Cover Source Materials

- The impacts would be lessened if the topsoil stockpiles at the toe ponds are not disturbed, but where would cover material come from?
- Comment to use topsoil stockpiles at toe ponds instead of disturbing a new area.

4.1.3 Impoundment Stability

- How will continued discharge of mine water affect the stability of the tailings impoundment?
- Stability of the tailings impoundment is tied to its dewatering; sinkholes and recent field reports indicate inherent instability features that should be investigated.
- Instability of the impoundment and failure of the dam may negatively impact operation of the Northern Lights dam; this should be assessed and mitigated.

4.1.4 Buried Drums

- Are the buried drums addressed in the revised Reclamation Plan?
- Will the buried drums be unearthed and sampled?

4.1.5 Subsidence

- How would hydrology be affected by future occurrences of subsidence?
- How will future occurrences of mining-related subsidence be addressed?
- A bond commensurate with the possibility of further incidents and their mitigation should be considered.
- Closure of the mine will result in a small underground lake; further faulting in the area may produce blowouts or rupture the mine plugs, causing discharges of polluted water to surface waters.

4.1.6 Other

Other issues and concerns identified during the scoping process that do not concern the other primary disciplines include:

- Request for Forest Service to require an Environmental Impact Statement, as an Environmental Assessment is not adequate to cover Genesis' proposal.
- A cash bond should be required to cover all obvious and potential problems that could be associated with reclamation and perpetual maintenance, including mandatory review at ten-year intervals.
- Request for copies of the Reclamation Plan to be placed at Heron Library, Noxon High School Library, and Libby Library.
- No reclamation other than tree planting should occur – the mine should be left open to be used as an educational resource and a tourist attraction.
- Comment stating concern regarding blowing dust from the tailings impoundment area.
- Releasing water down pipelines to tailings impoundment provides water needed for revegetation.
- Discharge of mine water to tailings impoundment would provide incentives for wise development and re-use of tailings area.
- The Northern Lights Dam at the mouth of Lake Creek traps sediment; it has been dredged twice, but no analysis of sediments has been conducted to determine extent of contamination from Troy Mine operations.
- The milling process uses iron, but occurrences of 'iron flock' have been attributed to natural processes; fate and transport of iron should be studied as carefully as copper.
- The amount of topsoil currently stockpiled is unsubstantiated, as some of the stockpile was used in creation of toe ponds, which hide violations of the Clean Water Act.

4.2 Issues Considered But Not Recommended For Further Evaluation

The scoping process has provided opportunities for the public to present concerns and issues to be considered during the EA process. Most of the technical issues raised will be carried forward for further consideration in the EA process. However, several issues were not recommended for further evaluation by consensus during the IDT meetings. These issues will be discussed and summarized in the EA, but no additional analysis is anticipated for these issues which are described below.

In December 2002 the Cabinet Resource Group (CRG) initiated a lawsuit against ASARCO and Sterling Mining Company (now Revett Silver Company) that alleged that barrels of hazardous waste were buried within the tailings impoundment during ASARCO's operation of the Troy Mine. CRG also claimed that barrels containing solvents and waste oil and grease were buried in the tailings impoundment. The allegations were never substantiated, and in July 2006, U.S. District Court Judge Donald Molloy dismissed the case without prejudice.

ASARCO has admitted to burying the drums but has also stated that the drums did not contain contaminants. The DEQ has reviewed water quality data and determined that CRG claims were not substantiated. The water quality analysis to be conducted for this EA will provide additional data to determine whether there is any merit to the claim of contamination in the buried drums.

Blowing dust from the impoundment area would only be a concern during operation of the mine. Once the mine has closed and ceases to place additional lifts of tailings in the impoundment, the area will be covered with topsoil and seeded with native plant species. The revegetated areas will be irrigated during the first growing season to ensure plant growth. Establishment of vegetation will control erosion and blowing dust.

The suggestion to leave the Troy Mine open as a tourist attraction does not meet the purpose and need for the reclamation project. Public health and safety, water quality concerns, and habitat for wildlife and threatened and endangered species would not be improved by leaving the Troy Mine open.

Subsidence has been addressed in previous technical reports. Previous occurrences of subsidence have been dealt with operationally, and a bond is in place for any future occurrences. The EA will discuss the potential environmental consequences of future occurrences, but the issue would not be included in the Reclamation Plan.

The purpose of the environmental assessment process is to determine whether significant impacts are likely to occur. The preparation of an EA for the Troy Mine Revised Reclamation Plan does not preclude the preparation of an Environmental Impact Statement (EIS). The EA will assist the agencies in identifying any potential impacts that should be considered and analyzed in more detail. Should significant impacts be identified during the EA process, an EIS will be prepared to assess the impacts, identify mitigation measures, and involve the public in resolution of adverse effects.

4.3 Issues to be Considered in the EA

Issues not dismissed for technical reasons will be addressed during the alternatives evaluation process in the EA. They are:

- Mine Water – Should hydraulic plugs be used to close adits? How much water will be held in the mine when it floods, and where will mine water drain to?
- Groundwater Protection – Will the copper attenuation process continue to function indefinitely? Will groundwater be protected from elevated copper levels, and for how long? Will continued transmission of mine water through the tailings lines to the tailings impoundment/decant ponds cause any environmental impacts after the mine ceases operations?
- Long Term Water Quality Prognosis/Monitoring – what will copper concentrations be in the future? Will copper reach Lake Creek? If so, when? Will the toe ponds continue to drain

to Lake Creek? What are the effects to Ross Creek? How long can/should monitoring realistically be continued?

- Water Treatment - Can mine water be treated in the service adit prior to discharge? Is there any treatment of discharged mine water? Why not replicate the decant ponds at the mill site and remove the seven-plus miles of tailings pipelines? Will an MPDES permit be required and obtained?
- Cover Source Materials - Are the impacts of using the stockpiled topsoil greater than using another source of cover material?
- Impoundment stability - How will continued discharge of mine water affect the stability of the tailings impoundment? How will future placement of additional tailings lifts affect stability?

Section 5

Identification of Data Gaps

Most of the information needed to conduct the NEPA/MEPA evaluation is available from existing sources. The following information will be needed to analyze the issues and prepare the EA:

- An analysis of water treatment alternatives related to 1) mine water discharge through tailings pipelines and 2) hydraulic/non-hydraulic plugging of adits and mine flooding
- An analysis of the sustainability of the copper attenuation process
- Hydrology information, i.e., sampling data for Ross Creek and its tributaries
- A cultural resource survey for any additional borrow areas identified

5.1 Water Treatment Alternatives

Evaluation of the method of mine water discharge and its potential impacts to local hydrology and water quality will require additional studies to acquire the necessary data:

- Further study of alternative water treatment approaches, e.g., water might be treated prior to leaving the mine, allowing discharge not necessarily through the tailings pipelines; or water might be treated after exiting the mine in a series of constructed wetlands near the mill site.
- Analysis of hydraulic versus non-hydraulic plugs - Genesis initially proposed the use of hydraulic plugs in the service and conveyor adits to prevent discharge of adit water to the portal patio. This proposal has been dropped **in favor of non-hydraulic plugs** consisting of concrete and backfill from the portal patios. Water would be piped from the adit (through the plug) to the decant ponds for discharge. Non-hydraulic plugs would provide control of the flow rate of water discharging from the mine, would prevent access to the underground workings, and could potentially result in lower oxygen concentrations in the flooded mine and less oxidation of copper minerals (although reduction of oxygen concentration would be greater with hydraulic plugs).

5.2 Copper Attenuation Process

A phased technical study is proposed to evaluate the copper attenuation process. The attenuation of copper via adsorption within the upper portions of the decant pond is likely to continue for a very long time due to the low loading rate of copper onto the organic matter. The precipitation of copper minerals could continue indefinitely, provided that the geochemical conditions that allow the precipitation to occur are maintained. Should those conditions change and become less effective in the future, an additional attenuation mechanism within the aquifer may become important. An analysis will be conducted of the copper attenuation process and any additional mechanisms that can be used to prevent contamination of groundwater or Lake Creek.

5.3 Surface Water Quality

Surface water quality data will be used to evaluate the quality of water present in drainages. Additional data needs identified during the scoping process include:

- Water quality and flow rates of Ross Creek and its tributaries, which have not been routinely monitored under the operating permit, are needed to identify potential impacts from mining operations and assess mitigation measures, if necessary
- Storm water data for drainage are needed to assess impacts of mine features on quantity and quality of storm water and potential re-routing or catchment options

Data gap gathering activities will commence as the particular data are required and as the necessary funds to gather these data are available. DEQ will request water sampling data from Genesis, Inc. to expand upon the existing water quality database and to determine if additional water quality sampling data are needed.

Section 6

References

Genesis, Inc. 2005. Troy Mine Operating Permit #00093.

Genesis, Inc. 2006. Troy Mine Revised Reclamation Plan.

Appendix A

Press Release

REQUEST FOR COMMENT and NOTICE of PUBLIC MEETING Genesis Incorporated Troy Mine Revised Reclamation Plan

USDA Forest Service, Kootenai National Forest
Three Rivers Ranger District, Troy Ranger Station
Lincoln County, Montana

Genesis, Inc. submitted a Troy Mine Revised Reclamation Plan which outlines their closure adit water management and modified reclamation proposal when the mine ceases operations in approximately six years, according to the latest estimate. The Plan describes their proposal including closure of facilities, revegetation, water management, and maintenance/monitoring. The Kootenai National Forest (KNF), and Montana Department of Environmental Quality (DEQ) are beginning the formal process of environmental analysis, and would like your input on the proposal. Your input is important to identify any concerns with the proposed activity.

Copies of the Troy Mine Revised Reclamation Plan are available for review at the KNF Supervisors Office, 1101 Hwy 2 West, Libby, MT 59923, at the DEQ, Environmental Management Bureau, 1520 East 6th Avenue, Helena, MT 59620. Copies are also available for review at the public library in Troy, MT, and at the Three Rivers Ranger District, 1437 Hwy 2 West, Troy, MT., or on the KNF web page at the address shown below.

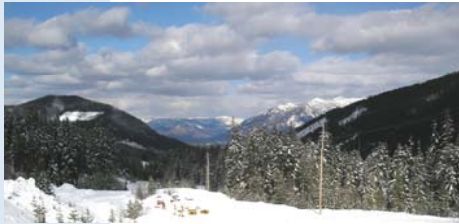
Comments concerning the proposed action must be postmarked by December 28, 2007, to be considered in the draft environmental document. The draft is expected by August, 2008, and the final by February, 2009. Send written comments concerning the Proposed Action to Paul Bradford, Forest Supervisor, Troy Mine Reclamation Project, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923. Send e-mail comments to: comments-northern-kootenai@fs.fed.us. All comments must contain name of commenter, postal service address, and date of comment. A copy on a computer generated CD should accompany all comments over one page in length.

For further information contact the Project Coordinators John McKay, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923, Phone (406) 283-7691, e-mail jmckay@fs.fed.us, or Kathy Johnson, DEQ, PO Box 200901, Helena, MT 59620, Phone (406) 444-1760, e-mail katjohnson@mt.gov or consult the Kootenai National Forest web page at <http://www.fs.fed.us/r1/kootenai/resources/minerals/index.shtml>.

An open house informational and scoping meeting will be held on Tuesday, October 30, 2007, at the Senior Citizens Center, Troy, MT, from 6:30 – 8:30 p.m.

Appendix B

Scoping Document



Troy Mine Revised Reclamation Plan Environmental Assessment

Kootenai National Forest and the Montana Department of Environmental Quality

October 2007

Site Background

The Troy Mine is located in Lincoln County Montana, about 15 miles south of Troy. ASARCO, Inc. originally permitted the copper and silver mine in 1978, commencing full production in the fall of 1982. By 1993, ASARCO stopped production at the mine in the wake of low metals prices. Genesis, Inc. bought the mine during the shutdown period and production resumed in 2005.

In March 2006, Genesis, Inc. submitted their final revised reclamation plan which describes proposed reclamation elements for final mine closure. Closure is anticipated in five to six years, after the ore deposit runs out. The revised plan was needed as a result of new information from studies, data collection, revegetation test plots, prior reclamation plan development reviews, and bonding determinations. Changed conditions (e.g., the toe ponds at the perimeter of the tailings facility dikes) and new regulations (e.g., state water discharge standards that prohibit the direct discharge of produced mine water into Stanley Creek) were also a factor in the need for a revised plan.

DEQ and KNF are reviewing the revised reclamation plan and all of the associated information accumulated during the mine's nearly 30-year life. Through the EA process, they will analyze the impacts of implementing the reclamation plan and all of the developed alternatives.

The Troy Mine has submitted a revised plan for mine reclamation. As required by the **Metal Mine Reclamation Act** (MMRA), an **Environmental Assessment** (EA) must be prepared to ensure the actions proposed are protective of the environment.

This fact sheet provides a brief description of the mine, its history, and the proposed reclamation plan. It also discusses the purpose, scope, and schedule of the EA and describes how the public can provide general information about the project or specific comments on the plan. Sources for copies of the plan and additional information are also provided.

Open House and Public Meeting

The **Kootenai National Forest** (KNF) and the **Montana Department of Environmental Quality** (DEQ) are seeking public comments prior to preparing an EA for the Revised Reclamation Plan for the Troy Mine. Public input is important to identify any concerns with the proposed plan.

Tuesday, October 30, 2007
Senior Citizens Center in Troy

Open house - 6:30 to 7:30 pm
Public meeting - 7:30 to 8:30 pm

The open house is an opportunity for the public to chat one-on-one with technical staff about various areas (such as water quality or mine reclamation) and view maps, posters, etc. At the public meeting, DEQ and KNF will bring everyone up to date with what's going on. People with an interest in the mine are encouraged to attend these events.

DEQ will make reasonable accommodations for persons with disabilities who wish to participate in the open house. If you require accommodation, please contact Kathy Johnson at 406-444-1760.

Everyone is welcome!



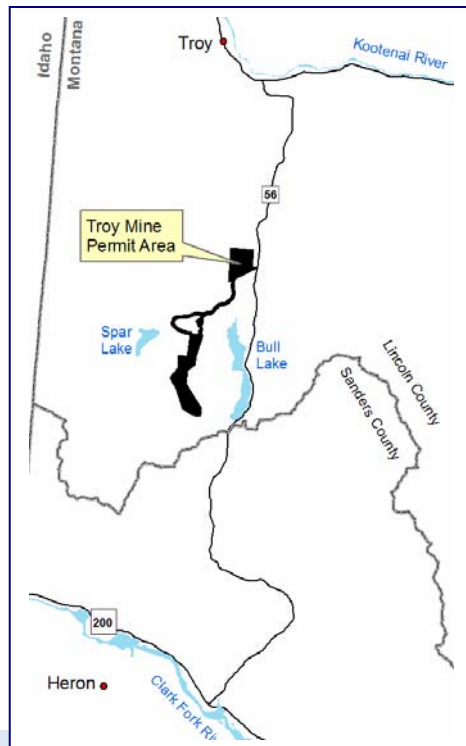
About the Troy Mine

The mine is located on 780 acres of public and private land south of Troy, Montana on US Highway 56 (see map). It consists of an underground mine, mill, and office facilities; tailings and water pipelines; power line; tailings impoundment; and associated support facilities (see pictures at right). The tailings facility and associated disturbances are on private land, and the tailings and water pipelines and the power line are on federal and private land.

Development of the underground copper/silver mine started in 1979 and production of ore concentrate started in 1981. During production, the mine produced approximately 8,500 tons/day of ore. In April 1993, ASARCO suspended mining and milling and placed the Troy Mine on care and maintenance status due to low metal prices. In November 2004, Genesis Inc. reinitiated mining which has focused on development of the East Ore Body portion of the mine.

From 1993 through 2004, the primary activities included:

- Care and maintenance of equipment and facilities
- Reclamation of portions of the tailing facility
- Evaluation of revegetation success
- Testing of methods for tailing reclamation
- Geotechnical monitoring of the tailing facility
- Monitoring of creeks and groundwater
- Initiation of fate and transport studies for mine water
- Preliminary assessment of tunnel-plugging methods
- Evaluation of water treatment methods
- Evaluation of mine flooding during pump shutdown period



What is an EA?

An EA is a detailed study that analyzes the environmental effects of a proposed action and its alternatives. It compares the impacts of the proposed action against those of "no action." DEQ and KNF have committed to preparing an EA to determine the nature and extent of impacts to the environment that may occur with implementation of the proposed revised reclamation plan. An EA provides for a thorough investigation of the potential impacts of a proposed action and also delineates measures that may be implemented to lessen any identified impacts.

The EA has two potential outcomes. If significant impacts to the environment are identified that cannot be avoided or mitigated, an Environmental Impact Statement (EIS) will be prepared. If no significant impacts are identified, a **Finding of No Significant Impact** (FONSI) will be prepared that will state all of the mitigation measures identified, especially those required to make the recommended alternative environmentally acceptable.

Components of the Plan

The original reclamation plan was submitted in the mine permit application in 1978. The revised plan for reclamation submitted by Genesis, Inc. describes their proposals for closure of facilities, revegetation, water management, and maintenance/monitoring.

Modifications to the original reclamation plan include:

- Modification of the adit water management plan
- A reclamation cap for the tailings impoundment
- Identification of new soil borrow areas for the reclamation cap
- Modification of seed mixes
- Incorporation of wetlands and open ponds to be left at closure
- Incorporation of more expansive weed control and weed prevention measures
- Use of the decant ponds to capture storm water runoff from the impoundment
- Use of toe ponds to capture dike runoff and groundwater seepage

Initial Issues of Concern

DEQ and KNF have identified several potential issues of initial concern:

- Air quality – blowing dust from the tailings impoundment area
- Surface and groundwater quality – long-term treatment/disposal of discharge water and potential for groundwater contamination beneath the tailings impoundment
- Wetlands – the creation of wetlands at the tailings impoundment
- Subsidence – reclamation of any future mining-related subsidence
- Tailings impoundment stability
- Effects on fisheries and wildlife (including threatened, endangered, or sensitive species) – potential long-term impacts from water quality degradation and short-term from plan implementation
- Vegetation and noxious weeds
- Cover soil – quality, location, and available volume of sources
- Potential for buried contaminants in tailings impoundment

These issues, and any other substantive issues identified by the public or the agencies' consultants, will be addressed in the EA. **The public is encouraged to come to the open house to discuss issues or to provide comments in writing about the plan.**

The EA Process

- Conduct scoping activities to make stakeholders (the public and local, state, and federal agencies with an interest in the project) aware of the EA
- Research/analyze existing conditions in project area
- Review proposed alternative and, if needed, develop additional alternatives that more fully address the purpose
- Analyze impacts of the alternatives and develop mitigation measures to lessen or avoid them
- Prepare a draft EA and make available to the public and agencies for review and comment
- Hold a public hearing to solicit comment from stakeholders
- Review all comments and apply, where feasible and reasonable
- Provide responses to stakeholders and revise the EA as necessary to incorporate any needed changes
- If no significant impacts are identified, conclude the EA with a FONSI
- If significant impacts are identified that cannot be mitigated, prepare an EIS and a **Record of Decision (ROD)** to document impacts and the decisions made to minimize them

EA Schedule and Deliverables

The EA process began in October 2007 and is anticipated to take approximately 18 months to complete.

Scoping Report.....	December 2007
Draft EA.....	April 2008
Public Hearing.....	September 2008
Comment Period.....	August-Sept. 2008
Review of Comments.....	Sept.-Nov. 2008
Final EA.....	February 2009
FONSI	March 2009

This assumes the EA will be sufficient to analyze the environmental impacts, that no significant impacts will occur, and that an EIS will not be required.

Reclamation Plan Goals

- Re-establish wildlife habitat
- Protect groundwater quality and surface water quality in local creeks
- Protect air quality in surrounding areas
- Provide public access to Federal lands
- Protect public health and safety by removing potential hazards

Do You Want to Learn More?

Copies of the Plan

Copies of the *Troy Mine Revised Reclamation Plan* are available for review at the KNF Supervisors Office, 1101 Highway 2 West, Libby, Montana and at the DEQ Environmental Management Bureau, 1520 East 6th Avenue, Helena, Montana. Copies are also available for review at the public library in Troy and at the Three Rivers Ranger District, 1437 Highway 2 West, Troy, Montana.

Questions or Comments

For further information contact the following Project Coordinators:

- **John McKay**, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923, (406) 283-7691, jmckay@fs.fed.us
- **Kathy Johnson**, DEQ, PO Box 200901, Helena, MT 59620, (406) 444-1760, katjohnson@mt.gov

Information (including an electronic version of the plan) is also available on the KNF and DEQ websites:

- www.fs.fed.us/r1/kootenai/resources/minerals/index.shtml
- www.deq.mt.gov

How to Comment on the Reclamation Plan

Written comments concerning the proposed action must be postmarked by December 28, 2007, to be considered in the draft EA. Comments should be sent to Paul Bradford, Forest Supervisor, Troy Mine Reclamation Project, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923, or e-mailed to: **comments-northern-kootenai@fs.fed.us** with "**Troy Mine Revised Reclamation Plan**" in the subject line.

All comments must contain name of commenter, postal service address, and date of comment. A copy on a computer generated CD should accompany all written comments over one page in length.

Montana Department of Environmental Quality
PO Box 200901
Helena, MT 59620
Attn: Kathy Johnson

Return Service Requested

Can you identify this mystery
mining picture?

(See inside for answer)



Appendix C

Mailing List for Open House/Public Meeting

Company	First Name	Last Name	Address	City	State	Zip
BUREAU OF MINES & GEOLOGY	ROBIN	McCULLOCH	1300 WEST PARK ST RM 200	BUTTE	MT	59701-8998
DAILY INTERLAKE	BILL	SPENCE	727 E IDAHO ST	KALISPELL	MT	59904
MONTANA DNRC	LIBBY UNIT		14096 HWY 37	LIBBY	MT	59923
NORTHWEST MINING ASSOCIATION		220		SPOKANE	WA	99201-0772
WETA	DON	ALLEN	2301 COLONIAL DR	HELENA	MT	59601
SHPO		DEADHEAD	HELENA			
EARTHWORKS	BONNIE	GESTRING	PO Box 8383	MISSOULA	MT	59807
BLM - DILLON FIELD OFFICE	BOB	GUNDERSON	PO BOX 9431	DILLON	MT	59725
BLM - LEWISTOWN FIELD OFFICE	SCOTT	HAIGHT	PO BOX 1160	LEWISTOWN	MT	59457-1160
MEIC	JIM	JENSEN	PO BOX 1184	HELENA	MT	59624
VALLEY COUNTY RD DEPT	RICK	SEILER	PO BOX 1024	GLASGOW	MT	59230
WETA	DON	SERBA	2301 COLONIAL DR STE 2A	HELENA	MT	59601
WETA	DON	SERBA	216 CENTENNIAL ST	MISSOULA	MT	59801
MONTANA HISTORICAL SOCIETY	STAN	WILMOTH, PhD	PO BOX 2001201	HELENA	MT	59620-1201
MONTANA MINING ASSOCIATION			PO BOX 5567	HELENA	MT	59604
MONTANA TECH Library	Document Dept		1300 W Park	BUTTE	MT	59701-8997
NWEST MINING ASSOCIATION			10 N POST SUITE 220	SPOKANE	WA	99201-0772
MONTANA MINING ASSOCIATION			PO BOX 5567	HELENA	MT	59604
GOVERNOR'S OFFICE	CAPT. BLDG RM 204		DEADHEAD			
BLM - LEWISTOWN FIELD OFFICE			PO BOX 1160	LEWISTOWN	MT	59457-1160
BLM-MALTA FIELD OFFICE			HC 65 BOX 5000	MALTA	MT	59538-0047
CHOTEAU COUNTY			PO BOX 459	FORT BENTON	MT	59442
CITY OF HARLOWTON			PO BOX 292	HARLOWTOWN	MT	59036
DNRC			PO BOX 1021	LEWISTOWN	MT	59457-1021
GALLATIN NATIONAL FOREST			PO BOX 130	BOZEMAN	MT	59771
HELENA NATIONAL FOREST			2880 SKYWAY DR	HELENA	MT	59601
L YELLOWSTONE IRR PROJECT			2327 LINCOLN AVE SE	SIDNEY	MT	59270
WHEATLAND COUNTY			PO BOX 1903	HARLOWTOWN	MT	59036
ENRONMENTAL QUALITY COUNCIL	Legislative Environmental Policy Office		PO BOX 201704	HELENA	MT	59620-1704
	CHAS	VINCENT	5957 CHAMPION RD	LIBBY	MT	59923
	AUBYN	CURTISS	PO BOX 216	FORTINE	MT	59918-0216
	RALPH	HEINERT	PO BOX 577	LIBBY	MT	59923-0577
COUNTY COMMISSIONER LINCOLN	JOHN	KONZEN	512 CALIFORNIA AVE	LIBBY	MT	59923
COUNTY COMMISSIONER LINCOLN	MARIANNE	ROOSE	PO BOX 2012	EUREKA	MT	59917
COUNTY COMMISSIONER LINCOLN	rita	WINDOM	512 CALIFORNIA AVE	LIBBY	MT	59923
COUNTY COMMISSIONER SANDERS	CAROL	BROOKER	PO BOX 794	PLAINS	MT	59589
COUNTY COMMISSIONER SANDERS	HAROLD	LAWS	PO BOX 745	THOMPSON FALLS	MT	59873
COUNTY COMMISSIONER SANDERS	JUSTIN	PATTON	129 PATTON RD	LONEPINE	MT	59848
US FISH AND WILDLIFE SERVICE			780 CRESTON HATCHERY RD	KALISPELL	MT	59901
WILDWEST INSTITUTE			PO BOX 7998	MISSOULA	MT	59807
ALLIANCE FOR THE WILD ROCKIES	LIZ	SEDLER	PO BOX 1203	SANDPOINT	ID	83864
CABINET BACK COUNTRY HORSEMEN	ROCKY	SCHAUER	PO BOX 949	LIBBY	MT	59923
CABINET BACK COUNTRY HORSEMEN	LEE	METZGAR	1001 E BROADWAY ST NO 2-206	MISSOULA	MT	59802
MONTANA WILDERNESS ASSOCIATION	CESAR	HERNANDEZ	643 FULKERSON LANE	POLSON	MT	59860
MONTANIANs FOR MULTIPLE USE	STEVE	FUNKE	578 1ST AVE NE	KALISPELL	MT	59901

Company	First Name	Last Name	Address	City	State	Zip
MT CHAPTER AMERICAN FISHERIES SOCIETY	ERIC	REILAND	PO BOX 5186	MISSOULA	MT	59806
MT DEPT OF NATURAL RESOURCES	JOHN	SHOTZBERGER	14096 US HWY 37	LIBBY	MT	59923
MT FISH WILDLIFE AND PARKS	JERRY	BROWN	475 FISH HATCHERY ROAD	LIBBY	MT	59923
UNITED STATES HOUSE OF REPRESENTATIVES	DENNIS	REHBERG	301 E BROADWAY, SUITE 2	MISSOULA	MT	59802
UNITED STATES SENATE	MAX	BAUCUS	8 THIRD STREET EAST	KALISPELL	MT	59901
UNITED STATES SENATE	JON	TESTER	1845 HWY 93 S SUITE 210	KALISPELL	MT	59901
MT STATE REPRESENTATIVE, DIST 13	PAT	INGRAHAM	PO BOX 1151	THOMPSON FALLS	MT	59873
MT STATE SENATOR, DIST 7	JIM	ELLIOT	100 TROUT CREEK RD	TROUT CREEK	MT	59874
ROCK CREEK ALLIANCE	RICK	STERN	210 HIGGINS, SUITE 206	MISSOULA	MT	59802
SOCIETY OF AMERICAN FORESTERS			PO BOX 345	LIBBY	MT	59923
LIBBY MAYOR	TONY	BERGET	PO BOX 1428	LIBBY	MT	59923
TROY MAYOR	JIM	HAMMONS	PO BOX 1030	TROY	MT	59935
US FISH AND WILDLIFE SERVICE	WAYNE	KASWORM	FARM TO MARKET RD	LIBBY	MT	59923
CLARK FORK COALITION	MATT	CLIFFORD	PO BOX 7593	MISSOULA	MT	59807
	HANNAH	HERNANDEZ	71 FOUR CORNERS S	HERON	MT	59844
	JUDY	HUTCHINS	PO BOX 104	HERON	MT	59844
LINCOLN COUNTY COALITION	WAYNE	HIRST	209 MINERAL AVE	LIBBY	MT	59923
YAAK VALLEY FOREST COUNCIL	ROBIN	KING	155 RIVERVIEW DR	TROY	MT	59935
	STEVE	PRIEVE	PO BOX 213	TROY	MT	59935
ENVIRONMENTAL PROTECTION AGENCY	STEVE	POTTS	10 W 15TH STREET SUITE 3200	HELENA	MT	59626
ENVIRONMENTAL PROTECTION AGENCY	BILL	RILEY	1200 6TH AVE	SEATTLE	WA	98101
MONTANA TROUT UNLIMITED	TIM	LINEHAN	472 UPPER FORD ROAD	TROY	MT	59935
NATIONAL WILDLIFE FEDERATION	HANK	FISHER	3342 BUTLER VIEW LN	MISSOULA	MT	59806
SIERRA CLUB	JIM	CURTISS	1318 KHANABAD DR	MISSOULA	MT	59802
USFWS	MARK	WILSON	100 N PARK SUITE 320	HELENA	MT	59601
ARMY CORPS OF ENGINEERS	BOB	NEBEL	215 N 17TH ST	OMAHA	NE	68102-4978
	RAY	REMP	208 N COLORADO AVE	LIBBY	MT	59923
	DASIOS	FOTULA	PO BOX 723	TROY	MT	59935
MONTANA MOUNTAIN VALLEY, LLC			PO BOX 1626	EAGLE RIVER	WI	54521
	RAY	SAMPSON	8952 BULL LAKE RD	TROY	MT	59935
	RON	COTTON	1011 SHINING MTNS TRAIL	TROY	MT	59935
	STEVE	GIBSON	435712 STATE HWY 20	NEWPORT	WA	99156
	JOHN	NORRIS	PO BOX 456	MT BALDY	CA	91759
	ROBERT	JENKINS	1672 SPAR LAKE ROAD # RE	TROY	MT	59935
	DAVID	REID	PO BOX 257	TROY	MT	59935
	RICHARD	GREENE	1365 STAR GAZE CT	COLFAX	CA	95713
ATHENS INVESTMENTS, INC			201 W MAIN STREET STE 201	MISSOULA	MT	59802
	TOM	BAMFORD	PO BOX 971	TROY	MT	59935
	CAROLE	WRIGHT	9504 BULL LAKE RD	TROY	MT	59935
	NEIL	NEWTON	PO BOX 523	SEELEY LAKE	MT	59868
	DON	WEATHERBY	9564 BULL LAKE RD	TROY	MT	59935
	ALLAN	BACON	9566 BULL LAKE RD	TROY	MT	59935
NORTHWEST PROPERTIES, LLC			103 W LINCOLN BLVD	LIBBY	MT	59923

Appendix D

Newspaper Advertisements for Open House/Public Meeting

USDA Forest Service, Kootenai National Forest
Three Rivers Ranger District, Troy Ranger Station
Lincoln County, Montana

The Kootenai National Forest (KNF) and Montana Department of Environmental Quality (DEQ) are beginning the formal process of environmental analysis, and would like your input on the proposal. Your input is important to identify any concerns with the proposed activity.

An open house informational and scoping meeting will be held on Tuesday, October 30, 2007, at the Senior Citizens Center, Troy, MT, from 6:30 - 8:30 p.m.

Comments concerning the proposed action must be postmarked by December 28, 2007, to be considered in the draft environmental document. The draft is expected by August 2008, and the final by February 2009. Send written comments concerning the Proposed Action to Paul Bradford, Forest Supervisor, Troy Mine Reclamation Project, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923. Send e-mail comments to: **comments-northern-kootenai@fs.fed.us**. All comments must contain name of commenter, postal service address, and date of comment. A copy on a computer generated CD should accompany all comments over one page in length.

Genesis, Inc. submitted a Troy Mine Revised Reclamation Plan, which outlines their closure adit water management and modified reclamation proposal when the mine ceases operations in approximately six years, according to the latest estimate. The Plan describes their proposal including closure of facilities, revegetation, water management, and maintenance/monitoring.

Copies of the Troy Mine Revised Reclamation Plan are available for review at the KNF Supervisor's Office, 1101 Hwy 2 West, Libby, MT 59923, at the DEQ, Environmental Management Bureau, 1520 East 6th Avenue, Helena, MT 59620. Copies are also available for review at the public library in Troy, MT, and at the Three Rivers Ranger District, 1437 Hwy 2 West, Troy, MT., or on the KNF web page at the address shown below.

For further information contact the Project Coordinators John McKay, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923, Phone (406) 283-7691, e-mail jmckay@fs.fed.us or Kathy Johnson, DEQ, PO Box 200901, Helena, MT 59620, Phone (406) 444-1760, e-mail katjohnson@mt.gov or consult the Kootenai National Forest web page at <http://www.fs.fed.us/r1/kootenai/resources/minerals/index.shtml>.

Adopt a Family Programs throughout Montana to Giving Program. Fifty \$500 Grants are available. Please send a brief letter describing your Adopt A Family Program, P.O. Box 6000, Butte, Montana 59701, November 9, 2007 for Christmas of 2007

Kootenai National Forest Service
10-24-07

Western News

Libby, MT

by Sandra Gubel

From an historical point of view, there have been several newspapers operating through the years in Sanders County--some with a fair amount of longevity.

In a brief stint, Publisher Adam Aulbach published 13 issues of the *Belknap Sun* in the period around 1884. Belknap at that time was the "getting on" point to depart on the train to the Coeur d'Alene gold fields. The

temporary town that held some 6,000 people awaiting departure on the trains, soon diminished, and Aulbach himself eventually decided to climb on board awaiting a "golden" opportunity.

Nearby Thompson (only later named Thompson Falls), became the spot that people inhabited more in the days afterward--those who remained here, or returned after finding the life of the gold country not to their liking.

The *Western Montanian*, noted Lorraine Dufresne in her book *A Heritage Remembered*, sprung up in 1896, published in Thompson. Maurice Helterline, who painstakingly researched that publication and several others, gleaned births, deaths and marriages, noted he'd checked all issues of the *Montanian* during its publication years of 1894 to 1896. D.A. Hendricks was publisher of the paper, which came out each

Saturday. Subscriptions were \$2.50 per year.

Sanders County Signal was published in the Plains/Hot Springs area from 1906 to 1924. The *Plainsman* (in Plains), published in 1895 and 1896, and again from 1900 to 1992. The *Sanders County Democrat* was published in 1911 and 1912 from Plains, with C.C. Willis as publisher. The *Hot Springs Sentinel* was published from 1950 to 1960. There was also a newspaper called the *Hot Springs Citizen*.

Also publishing news were the newspapers *Camas Hot Springs Exchange* and *Camas Hot Springs Record*.

It is believed that the newspaper for whom Edith Colby worked (see additional historical article), the *Independent-Enterprise* was published for a period of time and was in operation before and during the time that the reporter stood trial for the shooting of Thompson businessman Atwood C. Thomas. Editor of that paper was J. Manire.

Of the newspapers that remain in business, the *Little*

REQUEST FOR COMMENT and NOTICE of PUBLIC MEETING Genesis Incorporated Troy Mine Revised Reclamation Plan

USDA Forest Service, Kootenai National Forest
Three Rivers Ranger District, Troy Ranger Station
Lincoln County, Montana

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Genesis, Inc. submitted a Troy Mine Revised Reclamation Plan, which outlines their closure adit water management and modified reclamation proposal when the mine ceases operations in approximately six years, according to the latest estimate. The Plan describes their proposal including closure of facilities, revegetation, water management, and maintenance/monitoring.

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For further information contact the Project Coordinators John McKay, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923, Phone (406) 283-7691, e-mail jmckay@fs.fed.us, or Kathy Johnson, DEQ, PO Box 200901, Helena, MT 59620, Phone (406) 444-1760, e-mail katjohnson@mt.gov or consult the Kootenai National Forest web page at <http://www.fs.fed.us/r1/kootenai/resources/minerals/index.shtml>.

Published in the Sanders County Ledger October 25, 2007

NOTICE

Notice is hereby given th
conjunction with Town of

SHALL THE TOWN COL
ADDITIONAL 75 MILLS,
ADDITIONAL FUNDING

The polls will be open
p.m., on election day for

Town of

DATED this 17th day of

Published in th

Daily Inter Lake

10/21/07

*that a 237-member steering committee led by Hu approved the final candidate list for the Central Committee after two days of consultations with congress delegates. China Central Television showed Hu and other leaders raising their hands in approval.

the town of Musa Qala in Helmand province, sparking a battle that lasted several hours and involved the use of military aircraft, the coalition said in a statement. Nearly three dozen militants were killed.

The battle was the second in two days near Musa

months after British troops left the town following a contentious peace agreement that handed over security responsibilities to Afghan elders. Musa Qala has been in control of Taliban fighters ever since and is in the heart of the world's largest poppy-growing region.

REQUEST FOR COMMENT and NOTICE of PUBLIC MEETING

Genesis Incorporated Troy Mine Revised Reclamation Plan
USDA Forest Service, Kootenai National Forest
Three Rivers Ranger District, Troy Ranger Station
Lincoln County, Montana

The Kootenai National Forest (KNF) and Montana Department of Environmental Quality (DEQ) are beginning the formal process of environmental analysis, and would like your input on the proposal. Your input is important to identify any concerns with the proposed activity. An open house informational and scoping meeting will be held on Tuesday, October 30, 2007, at the Senior Citizens Center, Troy, MT, from 6:30 – 8:30 p.m.

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Kootenai National Forest - Mail 1

*and receive \$10 off any meal
at The Landing Restaurant*

BONNER COUNTY DAILY BEE

10/24/07



THE LANDING

41 Lakeshore Drive, Sagle • Just across The Long Bridge • 265-2000

4029885-1023

REQUEST FOR COMMENT and NOTICE of PUBLIC MEETING

Genesis Incorporated Troy Mine Revised Reclamation Plan
USDA Forest Service, Kootenai National Forest
Three Rivers Ranger District, Troy Ranger Station
Lincoln County, Montana

The Kootenai National Forest (KNF) and Montana Department of Environmental Quality (DEQ) are beginning the formal process of environmental analysis, and would like your input on the proposal. Your input is important to identify any concerns with the proposed activity. An open house informational and scoping meeting will be held on Tuesday, October 30, 2007, at the Senior Citizens Center, Troy, MT, from 6:30 – 8:30 p.m. Comments concerning the proposed action must be postmarked by December 28, 2007, to be considered in the draft environmental document. The draft is expected by August 2008, and the final by February 2009. Send written comments concerning the Proposed Action to Paul Bradford, Forest Supervisor, Troy Mine Reclamation Project, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923. Send e-mail comments to: <mailto:comments-northern-kootenai@fs.fed.us>. All comments must contain name of commenter, postal service address, and date of comment. A copy on a computer generated CD should accompany all comments over one page in length. Genesis, Inc. submitted a Troy Mine Revised Reclamation Plan, which outlines their closure adit water management and modified reclamation proposal when the mine ceases operations in approximately six years, according to the latest estimate. The Plan describes their proposal including closure of facilities, revegetation, water management, and maintenance/monitoring. Copies of the Troy Mine Revised Reclamation Plan are available for review at the KNF Supervisor's Office, 1101 Hwy 2 West, Libby, MT 59923, at the DEQ, Environmental Management Bureau, 1520 East 6th Avenue, Helena, MT 59620. Copies are also available for review at the public library in Troy, MT, and at the Three Rivers Ranger District, 1437 Hwy 2 West, Troy, MT., or on the KNF web page at the address shown below. For further information contact the Project Coordinators John McKay, Kootenai National Forest, 1101 U.S. Hwy 2 West, Libby, MT 59923, Phone (406) 283-7691, e-mail <mailto:jmckay@fs.fed.us>, or Kathy Johnson, DEQ, PO Box 200901, Helena, MT 59620, Phone (406) 444-1760, e-mail to: katjohnson@mt.gov or consult the Kootenai National Forest web page at <http://www.fs.fed.us/r1/kootenai/resources/minerals/index.shtml>.

4022875-1024

*Kootenai National
Forest*

Appendix E

Sign in Sheets for Open House/Public Meeting

PLEASE SIGN IN

Name: Kevin & Brenda Goe
Mailing Address: P.O. Box 24
City: Troy State: MT Zip Code: 59935
Phone: 295-7214 Email: -

Name: Doug Palmer
Mailing Address: 667 E Bedworth
City: Mission State: MT Zip Code: 59801
Phone: _____ Email: _____

Name: Bruce Clark
Mailing Address: PO Box 479
City: Troy State: MT Zip Code: 59935
Phone: 295-4727 Email: _____

Name: STEVE LLOYD
Mailing Address: Box 874
City: Hot Springs State: MT Zip Code: 59845
Phone: 240-1349 Email: _____

Name: Linda Newstrom
Mailing Address: PO Box 524
City: Troy State: MT Zip Code: 59935
Phone: 406/295 5568 Email: X

Name: Donald W. Boney
Mailing Address: 155 Forest Ave
City: Lt ABY State: MT Zip Code: 59923
Phone: (406) 293-7860 Email: _____

PLEASE SIGN IN

Name: Derek L. Feedback
Mailing Address: PO Box 1282
City: Libby State: MT Zip Code: 59923
Phone: 293-6151 Email: rx@backadit.com

Name: Paul Cocu
Mailing Address: 68 Swanson Lodge Rd
City: Libby State: MT Zip Code: 59935
Phone: 295 9916 Email: _____

Name: Rep. Chas Vincent
Mailing Address: 5957 Champion Rd
City: Libby State: MT Zip Code: 59923
Phone: 293.1575 Email: cvvincent@hotmail.com

Name: Torin Hass
Mailing Address: 1688 Iron Creek Rd
City: Troy State: MT Zip Code: 59935
Phone: ~~293~~(406) 295-9877 Email: TorinHass@yahoo.com

Name: Joshua Peterson
Mailing Address: Box 1150
City: Troy State: MT Zip Code: 59935
Phone: 295 6015 Email: _____

Name: James THill
Mailing Address: P.O. Box 250
City: Troy State: MT Zip Code: 59935
Phone: 295-9061 Email: _____

PLEASE SIGN IN

Name: Judy Hutchins
Mailing Address: P.O. Box 104
City: HERON State: MT Zip Code: 59844
Phone: 406-847-2717 Email: Jhutch7494@aol.com

Name: Jeff Frank
Mailing Address: 851 Lake Creek Road
City: Troy State: MT Zip Code: 59935
Phone: 406-295-5232 Email: frankj@troymine.com

Name: Lloyd Doney
Mailing Address: 40 HALO CT. S.
City: Troy State: MT Zip Code: 59935
Phone: 295-5882 Email: DoneyL@Troymine.com

Name: JANINE PRICE
Mailing Address: PO Box 180
City: Troy State: MT Zip Code: 59935
Phone: 295-5248 Email: _____

Name: Carlene Jensen
Mailing Address: 53 Swanson Rdg rd
City: Troy State: MT Zip Code: 59935
Phone: _____ Email: _____

Name: Clint Jensen
Mailing Address: P.O. Box 71
City: Heron State: Mont Zip Code: 59844
Phone: 406-847-2127 Email: _____

PLEASE SIGN IN

Name: Donald B Davis
Mailing Address: 3945 Bull Lk RD
City: Troy State: mont Zip Code: 59935
Phone: 295-14910 Email: _____

Name: Joe Madaski
Mailing Address: P.O. Box 748
City: Troy State: MT. Zip Code: 59935
Phone: (406) 295-4191 Email: _____

Name: Allen Layer
Mailing Address: PO 284
City: Troy State: MT Zip Code: 59935
Phone: 406-295-9190 Email: _____

Name: Hannah Hernandez
Mailing Address: 71 Four Corners So
City: Heron State: MT Zip Code: 59844
Phone: _____ Email: mtpaloma@hotmail.com

Name: Keith O'Blancs
Mailing Address: 74 Homann Ave
City: Libby State: MT Zip Code: 59923
Phone: (406) 293-9832 Email: _____

Name: Darcy O'Blancs
Mailing Address: 74 Homann Ave.
City: Libby State: MT. Zip Code: 59923
Phone: 406-293-9832 Email: _____

PLEASE SIGN IN

Name: Aubyn Curtis
Mailing Address: Box 212
City: Fontaine State: MT Zip Code: 59918
Phone: 406-882-4448 Email: aubyn.a@interbel.net

Name: Debbie Lyman
Mailing Address: 34 Beaver Peak Rd
City: Heron State: MT Zip Code: 59844
Phone: 8472388 Email: lastlyman@blackfoot.net

Name: Ruby Miller
Mailing Address: PO Box 548
City: Troy State: MT Zip Code: 59935
Phone: 295-4255 Email: millerra@cdm.com

Name: Glenn Knodle
Mailing Address: 184 Elk Meadow Ln
City: Heron State: MT Zip Code: 59844
Phone: _____ Email: _____

Name: Paul D KUKAY
Mailing Address: 343 Warren Road
City: Lobby State: MT Zip Code: 59923
Phone: 406-293-6989 Email: PKUKAY@hotmail.com

Name: Robert Winston
Mailing Address: PO Box 775
City: Troy State: MT Zip Code: 59935
Phone: 406-295-5018 Email: info@kroutfitters.com

PLEASE SIGN IN

Name: Colleen Hinds
Mailing Address: 21 Four Corners Rd S
City: Herm State: MT Zip Code: 59844
Phone: _____ Email: _____

Name: Floyd Beck
Mailing Address: 813 Wisconsin Ave
City: Libby State: MT Zip Code: 59923
Phone: (406) 293-8681 Email: _____

Name: _____
Mailing Address: _____
City: _____ State: _____ Zip Code: _____
Phone: _____ Email: _____

Name: _____
Mailing Address: _____
City: _____ State: _____ Zip Code: _____
Phone: _____ Email: _____

Name: _____
Mailing Address: _____
City: _____ State: _____ Zip Code: _____
Phone: _____ Email: _____

Name: _____
Mailing Address: _____
City: _____ State: _____ Zip Code: _____
Phone: _____ Email: _____

Appendix F

Written Comments Received During the Scoping Process

PB

COMMENT ON TRAY MINE RECLIMATION PROJECT
IN CARE OF PAUL BRADFORD FOREST SUPERVISOR
KOOTENAI N.F. 1101 US. HWY 2 WEST LIBBY MT - 59923

NOV 28 2007

GE943

DEAR SIR: I beleive The Tray Mine Should NOT be Reclaimed, I beleive it Should be Kept AS IS EXCEPT FOR CERTAIN LANDSCAPER WHICH COULD be Reclaimed WITH Tree PLANTS AND GRASS RESTORATION. I beleive This Mine The ONLY ONE of ITS Kind of CAVEREN AND PILLAR ON SILVER AND COPPER MINERALS, Should be Used AS A EDUCATION FACILITY For Children of ELEMENTRY Levels AS well AS Colleges, AND ALSO COULD BE Used AS A GREAT TOURIST ATTRACTION ALSO,

I SEVERAL YEARS AGO WHILE VISITING A OPEN COPPER PIT MINE WHICH WAS NOT OPERATING AT THIS TIME, HAD A CHANCE AS A TOURIST TO TOUR A UNDERGROUND MINE WHICH WAS ALSO IDLE,

A RETIRED MINER WHO HAD WORKED THERE WAS THE TOUR GUIDE WHO EXPLAINED THE HISTORY OF THE MINE. THE TRAY MINE COULD ALSO BE A GREAT TOURIST ATTRACTION IF LEFT MOSTLY AS IS. THANK YOU

HARVEY H. FREDERICKSEN
120 MIDDLE THOMPSON LAKE RD.
LIBBY MT 59923 PH 293-4959

321 Shaugnessy Rd
Libby, MT 59923
26 November 2007

PB
McKay

Paul Bradford
Forest Supervisor
Troy Mine Reclamation Project
Kootenai NF
1101 US Hwy 2 W
Libby, MT 59923



Dear Mr. Bradford,

I am asking you to consider several issues of concern regarding the Troy Mine.
These include:

- air quality (especially blowing dust from tailings impoundment)
- surface and groundwater quality-perpetual treatment of discharge water and potential groundwater contamination underneath the tailings
- subsidence-reclamation of future mining related subsidence
- tailing impoundment stability
- effects n fisheries and wildlife
- buried contaminants in the tailings impoundment (buried barrels)

Goals of a reclamation plan should include protecting the groundwater and surface water quality
protecting public health by removing potential hazards
protecting air quality and reestablishing wildlife habitat.

I urge you to consider these issues of impact to the environment. Thank you

sincerely,
Julie Waters-Barcomb

RE: Troy mine Reclamation

Dear Mr. Bradford

The most important consideration should be a cash bond to cover all obvious and potential problems that could possibly be associated with reclamation and perpetual maintenance of that reclamation with inflation rates calculated in the initial bond amount. This could have a mandatory review at ten-year intervals to assess the needs for the future. This bonding requirement would be particularly appropriate for: 1) tailings ponds integrity including sub- permeability to ground water, dike leakage, toe pond integrity, and toxic materials from containers that become apparent at a later date; 2) treatment of mine water that may change in chemistry and volume at a later date; 3) subsidence that may occur in the future due to climate changes and geologic events.

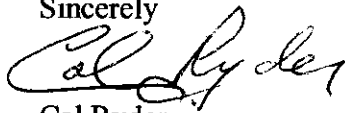
The following concerns identified by the agencies should definitely be addressed and returned to the same or better standards that existed prior to the mine, in a natural self sustaining to perpetuity condition:

1. Air quality in a pristine forest condition through appropriate state-of-the-art restoration of all tailings impoundments.
2. Perpetual monitoring and, if necessary, treatment of ground water and surface water to a purity standard of unimpacted streams of the of the area.
3. Potential for subsidence should be eliminated through any mining technology necessary.
4. The stability and total integrity of all tailings ponds (toe ponds) should be established through testing, analyzing soil conditions and whatever process is necessary to guarantee that total integrity.
5. There should be no negative impacts to fisheries and/or wildlife in the entire area of mine activity, after reclamation.
6. Sub soils, cover soils of the depth and forest soils type, consistent with that forest community, should be established and natural vegetation – grasses, forbs, shrubs and trees of that community restored.
7. All buried containers (many have been located and to the satisfaction of the science, proven to exist), should be located and with any contaminants or other contaminating objects, removed from the tailings and/or any other site in the mine area. These materials or objects should be properly disposed of in accordance with toxic waste disposal.

In addition to the above specific concerns, the entire mine site area should be returned to its natural condition with removal of all objects and materials extrinsic to a natural condition with exception of only, those objects or conditions necessary for access and monitoring the site.

Thank you for the opportunity to participate in this scoping procedure.

Sincerely



Cal Ryder

PO Box 1628

Noxon, MT 59853-1628

406-847-2535

<calrene@blackfoot.net>

My Concerns About the Troy Mine Revised Reclamation Plan

Dear Mr. Bradford,

I wish to express some 'heartfelt thoughts' on this subject at this time of year. As folks ponder "peace on earth and goodwill towards all", the immediate welfare of our surroundings takes on special significance. I am requesting the Forest Service to live up to it's name and serve the forests of our nation by doing everything in it's power to make thorough, conscientious, and far reaching choices.

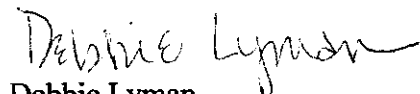
Within the strictures of the law, there are options you may select that can make a huge difference in the outcome of projects. I have no wish to detail 'all of my objections' to the proposed Reclamation Plan, other than to state a few obvious shortcomings in it's overall scope. I believe the monitoring is inadequate in every aspect. I believe the Revett-Troy proposals will require the least output by the company, without the future welfare of the land being of much consequence to them.

The history of this mine and the companies that have operated it does not inspire confidence. After watching this operation for over 25 years, I am left with a bad taste concerning the honesty and transparency of it's actions.

I wish to have the Forest Service require an Environmental Impact Statement for the reclamation plan. I do not believe an Environmental Assessment will be adequate to cover their proposal. When one refers to "in perpetuity", I wonder if that term is understood by the speaker/writer. It means 'FOREVER' and that is a difficult concept to grasp. Monitoring leakages, dust, vegetation, erosion, subsidence etc. " in perpetuity" is impossible.

It is of critical importance to have the Reclamation Plan be held to the highest standards. There is no reason to allow for minimum oversight when the welfare of the land is being decided. I ask that you consider my words. Thank you.

Wishing you a Happy New Year,



Debbie Lyman
34 Beaver Peak Road
Heron, MT 59844
December 26, 2007



lastlyman
<lastlyman@blackfoot.
net>

To: Paul Bradford <comments-northern-kootenai@fs.fed.us>
cc:
Subject: Troy Mine Revised Reclamation Plan

12/26/2007 04:24 PM

Revised Reclamation Plan

Concerns About the Troy Mine

Dear Mr. Bradford,

I will be brief. There is no need to enumerate all the faults and shortcomings in the Revett 'revised reclamation plan'. I will focus on one subject that lies at the core of this 'reclamation'. Water and watersheds are soon to be recognized as the single most important element of our forests. How water is affected will be the primary concern in all land use (and abuse) decision making. We will be spending billions on water conservation and reclamation of damaged watersheds. The idea of perpetual discharge of polluted water will be considered barbaric and stupid. A global paradigm shift, driven by climate change, will make the strategy of disposing of any waste or toxic materials into a watershed a truly criminal act. Clean water will be more valuable than the resources that have been extracted using wasteful and polluting methods. I will remind you that good money is being made mining the tailings piles left over from mining activities abandoned a century ago. Those who look back on our mining practices and our disregard for fresh, clean water will wonder at the incredible waste, the inefficiency and the greed that caused the mess that they will have to clean up.

Please do the right thing and hold Revett to the highest standards for reclamation of their Troy mine. A full EIS. Revett should be held to the highest standard.

Sincerely,

David R. Lyman
34 Beaver Peak Road
Heron, MT 59844
Dec. 26, 2007



Rec'd 11-9-07



Comment Sheet for the Troy Mine Revised Reclamation Plan

The United States Forest Service – Kootenai National Forest (KNF) and the Montana Department of Environmental Quality (DEQ) would appreciate receiving your thoughts and concerns on the Troy Mine Revised Reclamation Plan proposed by Genesis, Inc. An environmental document will be prepared to evaluate the impacts of the revised reclamation plan as well as the impacts of following the No Action alternative. There are several ways to submit your comments on the revised reclamation plan: 1) leave the comment sheet with a study team member at the public meeting; 2) mail your comments to the address shown on the back of this page; or 3) email your comments to comments_northern_kootenai@fs.fed.us – *please include "Troy Mine Revised Reclamation Plan" in the subject line.* All comments must contain name of commenter, postal service address, and date of comment. We would appreciate receiving your comments before **December 28, 2007**. Thank you.

CURRENT PLAN FOR DISCHARGING MINE WATER
ONTO TAILING EMPOUNDMENT IS THE MOST
LOGICAL PROPOSAL FOR FOLLOWING REASONS:

1) SEALING MINE OPENINGS NOT REASONABLY
FEASIBILITY & POSES A FUTURE THREAT OF
RELEASING LARGE VOLUMES OF WATER
SHOULD THERE BE A STRUCTURAL FAILURE.

2) WATER RELEASE DOWN EXISTING PIPELINES
MEETS MT. WATER QUALITY STANDARDS
AND IS VITAL TO FACILITATING THE
REVEGETATION PROCESS... WHERE WOULD
THE WATER COME FROM FOR MAINTAINING
GROWTH IF NOT FROM THIS MOST AVAILABLE
SOURCE? WHY DEplete THE WATER
TABLE WHEN NOT NECESSARY.

3) LOOKING INTO THE FUTURE – THIS AVAILABLE
RESOURCE WOULD PROVIDE INCENTIVES FOR
WISE DEVELOPMENT AND USE OF TAILING AREA.

Name: JANINE PRICE
Mailing Address: PO BOX 180
City: TROY State: MT Zip Code: 59935
Phone: _____ Email: _____



The Cabinet Resource Group

Dec. 27, 2008

Cathy Johnson
Montana DEQ
P.O.B. 200901
Helena, MT 59620-0901

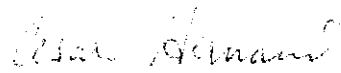
S.O. Paul Bradford
Kootenai National Forest
1101 Highway 2 West
Libby, MT 59923

Dear Cathy Johnson and Paul Bradford,

Attached are the comments of the Cabinet Resource Group concerning scoping for the Troy Unit Revised Reclamation Plan (TRRP). Hard-copy of these comments will be accompanied by approximately 25 pages of supporting documents which are referenced as attachments in the body of CRG's comments. The three attachments (1,2, 3) reference documents in CRG possession, provided by Asarco regarding the Troy Mine CERCLA Lawsuit. A portion of them dealing with field investigation reports and communications between MT/DHES are (or should be) in the possession of DEQ. CRG will provide Bates # copies of said documents which DEQ/KNF may not have upon written request of the agencies. CRG personnel are available for discussion with the agencies and contractor chosen to write the NEPA / MEPA document these comments pertain to.

CRG would like to remind the agencies responsible for this document that DEQ publicly determined the TRRP amounted to a major amendment (Western News 2/28, 3/7, & 3/14/01) under the MMRA Sec. 82-4-342, which requires an EIS. CRG comments raise issues that would be insufficiently addressed in the context of an EA. On behalf of CRG's 175- members, we remind the agencies of their responsibilities under the law and respectfully request that an EIS be conducted for the TRRP.

Thank you.


Cesar Hernandez
Board Member CRG

CRG, Troy Mine Revised Reclamation Plan EA Scoping Comments

1. Topsoil
2. Stability
3. Long Term H2O management plan
4. Barrels
5. MPDES
6. NL Troy Dam and sediment analysis
7. Subsidence
8. Other Discharges
9. Reclamation Bond
10. Catastrophic failure bond

1. (TOPSOIL) The existing Troy Mine Reclamation Plan requires the placement of 18 inches of stockpiled topsoil to cover the tailings impoundment surface upon completion of mining. The Revett / Genesis Troy Revised Reclamation Plan (TRRP) proposes the placement of (18") inches of borrowed subsoil and retention of stockpiled topsoil for toe ponds. This revision is prefaced by the claim that years of experimenting indicate that borrowed subsoil is a better cover medium. Unsaid and unsubstantiated are measurements of the currently existing topsoil stockpile and the true purpose of leaving the stockpiled topsoil in place. For numerous years (especially during operation) the phreatic water level in the tailings impoundment has exceeded the height of the starter dam. This situation was warned against at the earliest inception of the Troy Project as posing the possibility of catastrophic failure for the tailings impoundment. This hazard has been mitigated by the presence of nearly 1,000,000 cubic yards of topsoil at the toe of the tailings impoundment.

Significant amounts of the stockpiled topsoil have also been incorporated into the currently existing toe ponds that were not originally permitted with the facility. These toe ponds contain water that leaks from the tailings impoundment. Removing the toe ponds would expose conditions at the Troy Mine tailings impoundment that violate the Clean Water Act. The illegal toe ponds cover the fact the tailings impoundment was not constructed as permitted (with an internal drainage system). Such a system would have required an MPDES permit, containing permit requirements Asarco realized they would be incapable of meeting.

The toe ponds leak directly into Lake Creek despite the best efforts of the mining company and State regulators to avoid this inconvenient truth and despite their best band-aid efforts to hide this discrepancy. Both entities rely on dilution of the discharge as the solution to this problem.

2. (STABILITY) The stability of the Troy Mine tailings impoundment is tied to its dewatering. The current reclamation plan would move towards accomplishing this objective. The TRRP proposes to perpetually deliver mine water to the tailings impoundment as a method of treating mine water that does not meet surface water criteria. This proposal is little more than a financial decision of the mine owners to avoid more costly treatment options and fails to take into

account the long-term stability requirements of the tailings impoundment. The Reichsmuth (1/31/84) Report was denigrated and dismissed by Asarco and the regulatory agencies. However, sinkholes that appeared in 1988 and 91 as well as piping and recent (March 23 and April 14, 2007) field reports indicate instability features inherent in the facility. These features need to be vigorously and independently investigated.

3a. (H2O MANAGEMENT) Long-term water management of discharges from the mine and related facilities is the most important outstanding reclamation feature. Diffuse discharges from the mine itself towards Ross Creek have begun to show up. Ross Creek is the major tributary feeding Bull Lake. Mapping, monitoring and either eliminating or remediating these discharges are a long-term challenge that needs to be thoroughly addressed.

Discharges from the tailings impoundment are equally important because of its proximity to Lake Creek and for reasons outlined at #2 above. Another issue that must be addressed is that of buried material within the tailings impoundment. Attached is documentation indicating the possibility if not probability that hazardous materials (industrial solvents) were disposed of in the tailings impoundment. The attachments contain information gleaned from discovery documents provided by Asarco regarding a CERCLA lawsuit initiated by the Cabinet Resource Group, which was eventually dismissed without prejudice when Asarco went into bankruptcy. The bills of lading indicate that copious quantities (thousands of gallons) of industrial solvents were used at the facility from 1980-90, while a paper trail of their disposal (as required by EPA) is completely lacking. Documentation and the accounts of some former employees at the facility indicate that illegal and regular dumping of the material occurred under cover of darkness at the tailings impoundment. While Asarco initially denied any dumping of material at the facility, subsequent geophysical investigations located several anomalies consistent with the allegations that barrels are buried in the tailings impoundment. Asarco eventually admitted to burying barrels, but insists the material inside is benign flocculant. The adages; "fool me once shame on you, fool me twice shame on me" and "trust but verify" should prominently guide the actions of regulatory agencies here.

Just as it has taken over 25 years for subsidence at the mine to demonstrate itself a viable issue, so to the issue of buried barrels may take similar time. Buried barrels will deteriorate over time as they rust. The release of hazardous materials is a critica issue that DEQ cannot overlook. It's importance increases as development on private lands surrounding the tailings impoundment increases. Prudence dictates caution, as long-term public health concerns may well have to be dealt with long after the mine has closed, and the ability to hold its operators liable possible. The State of Montana and its Department of Environmental Quality need only look as far as Libby, MT and its Tremolite asbestos crisis, that was ignored for decades to see the effects of what burying its head in the sand over an issue can produce.

3b. (ATTENUATION) The RRP Fate and Transport study makes a great deal concerning the attenuation of copper : Pg. 24 *"all data collected to date indicates that copper is very immobile in the groundwater system surrounding the decant pond. Copper in mine water that is disposed in the decant pond becomes primarily bound in shallow soils/sediment in the*

bottom of the decant pond and very little copper is transported beyond a few tens or hundreds of feet in the groundwater system.” Continued deposition of mine water to the tailings impoundment for treatment purposes would require long-term maintenance of at least one, if not all three, of the pipelines that service the tailings impoundment. Similarly if such is the case, then why aren't the tailings impoundment decant ponds replicated at the mine mill site, which would not require the maintenance of the more than seven miles of buried pipeline from mine to tailings impoundment? See F&T Study pages 30-31 for more information.

4a. (BARRELS) On March 19, 1996 Tom Reid (DEQ) received a Citizen Complaint from former Asarco Troy Mine employee Jim Meyer regarding a number of concerns that took place over the course of his employment (DEQ/ WQD Individual Activity Report 3/19/96). Mr. Meyers was employed at the Troy unit in excess of 10 years. Among his concerns was an allegation that Asarco routinely disposed of 55 gallon barrels (some containing hazardous wastes) in the tailings impoundment. The disposal of barrels under cover of darkness was corroborated by other Asarco employees. A map of the depositions accompanies the Individual Activity Report of 3/19/96. The Citizen Complaint was turned over to DEQ's Waste Management Division –Hard Rock Bureau; which did nothing.

The Cabinet Resource Group filed a CERCLA Law Suit against Asarco on 12/03. Subsequent geophysical investigations by Eco-Tech Geophysical (Missoula) as a part of discovery in the lawsuit identified several anomalies consistent with buried barrels. Asarco initially denied the allegations but admitted a year later that 17 barrels had been disposed of in the tailings impoundment. Asarco claimed the content of the barrels as benign flocculant. Physical copies of the Eco-Tech report and a plan to excavate the anomalies were turned over to Montana DEQ. DEQ was supposed to evaluate the proposed excavation plan and make a determination as to its feasibility and permitting. It did nothing! Asarco's eventual Bankruptcy filing caused the CERCLA case to be dismissed without prejudice. The issue of the buried barrels remains a liability that DEQ and the State of Montana ignore at their peril.

4b. (ATTACHMENTS) As a result of its CERCLA Lawsuit the Cabinet Resource Group was provided some 3000 pages of discovery documents. Those documents consisted of Bills of Lading and Purchase Orders as well as numerous other documents relating to the use and disposal of all industrial fuels and lubricants capable of generating hazardous waste. CRG broke these documents into three categories.

- 1) An inventory of all chemical and lubricant materials that were delivered to the mine in 55 gallon barrels.
- 2) An inventory of industrial solvents use.
3. A list of solvents and other documents possibly indicating their disposition.

From these documents it has been determined that during the 1st decade of Troy Mine operation Asarco relied on local mine personnel to dispose of all mine wastes. Beginning in 1991 (Dave Young tenure) Asarco employed the services of Safety-Kleen to handle the disposal of the mines lubricant waste stream capable of generating hazardous waste.

A synopsis of these documents is provided and some definitive conclusions can be reached.

- a. The Troy mine used generous amounts of lubricants / solvents that contained hazardous substances (1.1.1. Trichlorethylene, 1.1.1 Trichloroethane / Bates # 5446, 5341-43, etc.).
- b. Asarco filed for and received from the State of Montana (DHES / Hazardous Waste Bureau) a Hazardous Waste Generator Permit.
- c. MT/DHES conducted some field inspections concerning use and handling practices of Troy mine generated haz-wastes (Bates # 5263-5410, 5309, 5330-31, etc). Concerns were raised in the field investigation reports regarding the disposal of hazardous materials.
- d. Approximate use of material amounts containing hazardous wastes can be determined (Bates # 5446, 5330-31, 5344-45, etc.).
- e. Only one Hazardous Waste Shipment Manifest (dated 2/18/93, Bates #5263-5410) was produced as a result of discovery.
- f. Documents (Bates # 5297, 5317, 5328, etc.) providing clues that hazardous wastes were disposed of in manner and location violating permit and law.

Actual copies (Bates reg.) of these documents are in possession of CRG and are available for inspection by regulatory agencies. They are also part of the public record in the CERCLA case that was filed.

5. (MPDES) The Troy Mine will discharge into perpetuity. The quantity and quality of that discharge is a responsibility of the operators and a duty they are charged to manage through the MPDES / NPDES process. The regulatory agencies have been entirely deficient in this respect and are only now beginning to acknowledge and move in the direction that the facilities of the mine should have had one.

6. (NORTHERN LIGHTS TROY DAM) The Troy Mine and tailings impoundment has been a source of various discharges since its inception. From tailings line ruptures, storage pad snowplowing, through airborne particulate and piping from the tails impoundment, material from its operations have found their way into the ecosphere. The NL dam at the mouth of Lake Creek acts as a natural sediment trap in the watershed and contains a diary of accumulations in its sediments. The NL dam has been dredged twice since the inception of the Troy Mine yet no analysis of its sediments has been conducted. This is an issue that needs to be addressed. The issue of a catastrophic failure of the Troy mine tailings impoundment on the operation of the NL dam is also an issue that has been overlooked or dismissed.

7. (SUBSIDENCE) Two incidents (2004, 2005) of subsidence reaching the surface at the Troy mine should be regarded as canaries of the coal mine. Several incidents of rock-fall within the mine, one of which recently (2007) resulted in the death of a miner are indicative of the existing poor rock conditions within the mine. Spalling of support pillars is another indicator that stability within the mine may not be what is alleged and needs to be independently assessed. This is critically important given the possibility that removal of some of support pillars is being considered by Revett. Such a situation in a coal mine in Utah resulted in the death of at least six miners and three rescue personnel. Once the mine is closed there will be little opportunity to assess conditions within the mine. A bond commensurate with the possibility of further subsidence incidents and their mitigation for purposes of public safety needs to be considered.

This may be critically important as closure of the mine will result in a small lake forming underground. Rock fall or faulting in the area of the lake may produce pressures that blowout or rupture mine plugs. Subsidence or rock-fall within the mine may also be incidental to the creation of discreet hydrologic discharges to surface waters of the state. A plan and bonding sufficient to cover these possibilities needs to be considered.

8. (DISCHARGES) The issue of iron flock emanating throughout the bluffs between the tailings impoundment and Lake Creek has been blown off as "naturally occurring" throughout the drainage. It is well known the milling process at the Troy Mine utilizes tens of thousands of tons of iron / manganese grind balls. The RRP Fate and Transport Study discusses the use of x-ray spectra and electron microscope technology in the identification of copper-bearing phases important to attenuation. Repeated requests have been made to the regulatory agencies regarding the identification of the iron flock leaching from the Troy tailings impoundment. Identifying the source of the iron flock as native material or coming from the milling process is a relevant issue here. Elevated levels of nitrates, potassium xanthate and some copper have been identified outside of the tailings impoundment adjacent to Lake Creek. It is only reasonable to assume the probability this iron flock comes from the mine milling process. The only way such a definitive identification is possible is if it is investigated as thoroughly as the fate of the copper in the F&T Study.

9. (RECLAMATION BOND) The purpose of a Reclamation Bond is to dismantle the mine facilities after mining has ended and to protect the environment and public from the costs of problems associated with the mining activities. Potential problems that are not addressed and / or thoroughly investigated become public problems when they are dismissed and very serious problems after-the-fact if they have not been bonded for (remember Zortman-Landusky for reference). When viable evidence is provided that an issue should be of concern, it is the public duty and responsibility of the regulatory agencies (in this case DEQ / USFS) to objectively look at the issue and demand viable answers from the entities they are entrusted by the public to issue permits to.

For seven years regulatory agencies (DEQ and the KNF / USFS) have dallied in requiring establishment of a viable long-term Reclamation Bond for the Troy Project Mine.

10. (CATASTROPHIC FAILURE BOND) The Troy Unit Revised Reclamation Plan (Dec. 2000) prepared for Genesis Inc. contains the meeting notes of 10/12/2000 in Appendix I. Under "General Issues" is the quotation: "Forest Service needs a letter of intent to reopen. Their issues are not as problematic as DEQ, since DEQ will eventually own the mine and the tailings." These notes were taken by Len Murray of Klohn Crippen Co. a member of the Louis Berger Group. It can be assumed that Mr. Murray is a person of competence and that he accurately recorded the conversations of the meeting. At least there are no public documents contesting or requesting changes to that statement. The reference he records regarding "eventual ownership of the mine and tailings," is in all probability a reference to the fact that Montana DEQ will eventually own final responsibility and liability through its permitting authority for what becomes of permit #00093 once reclamation has occurred.

Notes: from TRRP

1. (Pg. 3-6) • Impoundment site topsoil stockpiles totaling approximately 560,000 cubic yards of salvaged topsoil located on about 44 acres along the western perimeter of the impoundment dike and acting as a dike containing the toe ponds.
2. (Pg. 4-16) Tailing facility cover-soil volume calculations are included in Table 4-1. Approximately 766,600 cubic yards of cover-soil materials will be required for tailing facility reclamation.
3. (Pg. 7-4) Evidence from mine operations indicates that the SOB and EOB have greater potential to dissipate water than the NOB.
4. (Pg. 7-6) Although the static groundwater levels will fluctuate seasonally, as experienced during previous shutdown of the mine, the water level in the mine is expected to reach equilibrium at approximately the 4248 elevation in the SOB and at the 4225 elevation in the NOB and EOB as shown on Exhibit D. The NOB mine pool will flow to the EOB when it reaches the 4186 elevation.
5. (Pg. 13 Fate & Transport Study) Some soil/sediment samples (DP-1 3 3/4 “ and DP-1 4”) are enriched to the extent that they contain nearly twice the copper content as ore from the Troy mine. Total copper concentrations in decant pond sediments suggest that copper is attenuated in surficial soil/sediments in the decant pond.

Atch: ①

Troy Mine Barrel Inventory

Invoice #	Bates #	# Barrel dep.	#Barrel dep. rtn.
<u>1979</u>			
6552	011592	4	
<hr/>			
<u>1980</u>			
25495	011456	3	
25861	011492	10	
25588	011497	4	
25589	011498	2	
25586	011503	2	
25764	011505	4	
26578	011509	5	
26437	011510	5	
26412	011512	4	
26411	011513	4	
3570	011514		4
3577	011515		1
26382	011524	4	
3416	011528		6
26349	011529	10	
3806	011530		7*
3384	011536		4
26329	011538	10**	
26331	011539	10	
3352	011543		5
26277	011544	2	
2160	011546		9
25745	011547	5	
25750	011548	5	
25662	011552	5	
25659	011553	5	
25631	011554	5	
	011561	5***	
25981	011582	2	
25904	011565	2	
25905	011568	10	
2320	011570		8****
	011570		1****
2260	011571		1
7752	011575	10*****	
25841	011576	4	
Totals		131	46

* Page appears to have 2 barrel return invoices one on top of other. Cannot make out invoice or bates #, but appears to show 10 barrel return. ** Same superimposition as previous notation (*) only for 10 barrel deposit. *** Same superimposition as previous. Same superimposition as previous explanation. ***** Only notation of junk barrels (\$2.00 apiece).

1981			
26601	011348	10	
2103	011349		?*
0189	011350	12	8
0169	011354	1	
0167	011357		9
3702	011359		3
3652	011360		6
928	011361	3	3
937	011363	5	
0247	011364	5	5
0246	011365	5	5
506	011379	2	
893	011381**		
0295	011386	10	5
0299	011387	10	1
747	011389	10	18
0266	011397	10	5
1095	011398	10	21
illegible	011402	5	1
1059	011403	5	1
041	011407	6	
075	011408	10	
074	011408	7	
076	011409		9
8669	011410	6	
8669	011411	100***	
071	011418	10	
063	011421		6
illegible	011428		2
082	011440	10	
6967	011441		8
084	011443	4	
083	011444	6	
080	011448	10	
Totals		252	116

?*Illegible.

** First invoice (dated 7/20/81) that mentions "barrel exchange." Lists 2 barrels but does not indicate in or out.

*** Largest known barrel deposit.

1982

1589	011201	2	
115	011206		10
111	011206		8
096	011207		11
108	011207	10	
098	011209	6	
093	011209	10	
094	011210	10	
099	011210	3	
100	011214	12	
128	011218	1	
143	011219	2	
142	011219	2	
141	011220	2	
116	011220	8	
117	011221	10	
122	011221	10	
124	011223	10	
121	011223	2	
119	011224	4	
132	011224		10
103	011225	6	
129	011225		7
123	011226		5
125	011227	6	
137	011229		10
140	011229		9
135	011230		10
133	011231	10	
134	011231	10	
138	011232	3	
151	011236	10	
155	011237	3	
205	011239	1	
201	011239	4	
206	011240	6	
175	011241		9
170	011241		10
163	011245		10
150	011245		2
179	011246		11

157	011246		10
221	011249		10
227	011253		8
136	011254	10	
139	011255	7	
146	011255	4	
214	011256	10	
212	011257	10	
148	011258	10	
228	011259	10	
145	011260	7	
218	011262	10	
223	011263	6	
189	011264		5
229	011264		10
235	011265		11
232	011267	10	
161	011274	10	
164	011275	6	
169	011276	4	
174	011278	10	
193	011279		10
167	011279	2	
216	011286		11
197	011287	10	
207	011287	1	
211	011290		11
182	011291	1	
178	011291	10	
181	011292	11	
199	011293		10
196	011293		6
185	011295		7
204	011295		11
188	011296		6
236	011297	9	
242	011298	8	
243	011299		6
2195	011307	4	
246	011313	11	
Illelgible	011320	10	
253	011323		10
250	011325	10	
252	011330	9	
532153	011346	15	
Totals		407	267

1983

532188	011065	2	
532197	011066	10	
532189	011067	7	
532216	011080	2	
532212	011083	1	
503331	011084	1	
00000	011093	7	
503401	011104	2	
503402	011114	12	
503336	011117	5	
192	011128	1	
6974	011129		8
50340*	011133	3	
50340*	011134	5	
50339	011140	9	
01302	011159	3	
503412	011189	20	
503409	011190	2	
Totals		99	8

* Two Texaco invoices same number w/different # of barrel deposits?

1984

013012	010950	2	
5300	010964		3
503547	010981	20	
078277	011006	22	
6157	011060	5	
Totals		49	3

1985

6702	010721	4	
078278	010723	10	
078268	010724	10	
11-05-?	010728	20	
078276	010729	10	
075875	010806	2	
Total		56	

1986

Note: This is the year that most invoices are missing. Instead there is a Purchase Order (continued) that is a type of synopsis of total products used.

From this synopsis it can be calculated that 71 / 120lb. grease drums were received, as well as 14 / 400 lb. grease drums for a total of 85 drums. Another 20 drums of transmission fluid, 45 drums of gear lubricant, 10 drums of EP-2 grease and another 40 drums of EP-220 oil were received. Altogether this amounts to **285 drums**.

Note: After 1986 barrel deposits and or returns appear to disappear from petroleum product invoices.

OF SPECIAL NOTE: Petroleum products were not the only thing that came in 55 gallon drums. If you look at the Troy discovery documents listing that I sent much earlier you find the following.

10. 5196-5222 Purchase Order sheets 300,000 lbs. Potassium Amyl Xanthate, pellet form packed 341 lbs. in 55 gal. drums. 7/91-12/93. **Sheet # 5207, item #4 discusses \$20.00 drum deposit.** Sheet # 5217 itemizes sale of scrap metal from tailings area @ \$25 / ton. Other scrap iron sales on other sheets.

The Potassium Amyl Xanthate (a frother) used at the mill came in approximately 870 barrels.

Some barrel deposit invoices (few in number & quantity) then begin to show up in 1990.

* 1981-86 — 554 BARRELS UNACCOUNTED FOR.

Atch (2)

Troy Story solvents use discovery doc. findings 7/9/04

Date / Year	Purchase Invoice	^{BATES} Legal file #	Quantity
1979			
11/12/79	6006	011590	110 gal.
8-17-79	4157	011608	57 gal.
Total			167 gal.

1980*			
1-04-80	7308	011471	55 gal.
1-05-80	5109	011475	110 gal.
2-28-80	6955	011500	110 gal.
12-15-80	004	011511	220 gal.
6-10-80	7988	011558	110 gal.
6-05-80	7970	011559	110 gal.
5-19-80	7903	011563	220 gal.
Total			935 gal.

* Extensive and detailed itemization of barrel deposits and returns. #7752 / 011575 even lists "10 whole junk barrels."

1981*			
7-28-81	037	011374	110 gal.
6-24-81	023	011383	220 gal.
9-10-81	053	011422	220 gal.
9-15-81	8001	011424	220 gal.
Total			770 gal.

* Extensive and detailed itemization of barrel deposits and returns. #8669 / 011411 for 100 drums. Deposit of \$2000

** Jan.-May and Oct.-Dec. invoices appear to be missing.

1982*			
1-21-82	142	011219	110 gal.**
3-03-82	1694	011254	220 gal.
5-03-82	201	011239	220 gal.
3-11-82	146	011255	110 gal.
6-07-82	223	011263	330 gal.
3-22-82	164	011275	330 gal.
Total			1310 gal.

* Indications barrels being returned for credit. See invoices 111 & 115 and others.

Invoice #098 / 12-31-82 indicates solvent barrel return for credit.

**** Amoco Stoddard solvent.**

Note: Period month invoices July – Dec. appear to be missing

1983*

1-07-83	3154	011074	500 gal.
6-08-83	3794	011124	142 gal.
7-06-83	3912	011150	1056 gal.
11-02-83	4704	011164	525 gal.
Total			2223 gal.

* Invoice 6974 only invoice showing / indicating barrel (8) return for deposit.

Note: Believe some invoices for period months Feb.- May and Aug. – Oct. missing.

1984*

* Approximately ½ invoices for 1984 photocopied at 200% causing it to be mostly impossible to tell what the product purchase is. Legible ones listed below.

2-01-84	4872	010919	210 gal.
3-08-84	5014	010943	478 gal.
7-09-84	5574	010993	474 gal.
10-05-84	6241	011018	520 gal.
11-20-84**	6544	011036	300 gal.
Total			1982 gal.

**** Only invoice with (Hyd.) behind Solvent. Reference should be checked.**

Note: Believe some invoices for time April – June may be missing.

1985*

* Approximately ½ invoices for 1985 photocopied at 200% causing it to be mostly impossible to tell what the product purchase is. Legible ones listed below.

7-01-85	7545	010818	300 gal.
8-06-85	7732	010819	480 gal.
10-02-85	7995	010845	623 gal.
11-08-85	8165	010869	418 gal.
12-05-85	9692	010902	410 gal.
Total			2231 gal.

1986

***1986 file unlike others years had only 25 invoices. However, in the 1987 file one sheet was found that appeared to be a 1986 purchase order / all materials only solvent listed below.**

7-22-86	26H-0488B	010513	3600 gal.
Total			3600 gal.

1987

1-29-87	11158	010378	720 gal.
3-30-87	11479	010448	625 gal.
4-21-87	11557	010442	175 gal.
8-03-87	12216	010607	500 gal.
10-13-87	12741	010551	500 gal.
11-18-87	13130	010620	123 gal.
11-30-87	13215	010628	377 gal.
Total			3020 gal.

1988*

2-08-88	13775	008646	500 gal.
3-30-88	S206114	010187	540 gal.
5-02-88	S258856	010167	500 gal.
9-02-88	S3445830	0102433	400 gal.
11-4-88	S347078	010348	470 gal.
12-6-88	S34314?	010332	400 gal.
Total			2810 gal.

*** May be missing some invoices period Jan. - Feb. & June - July.**

1989

1-30-89	S347679	008692	380 gal.
3-02-89	S347951	008784	520 gal.
6-02-89	S346541	008862	500 gal.
7-10-89	15683	008909	15 cases*
7-24-89	S348901	008923	500 gal.
10-04-89	S34883	008983	500 gal.
Total			2400 gal.

*** Only invoice on solvents by the case and with the designation that solvent is bio-degradable.**

Note: Plenty of used oil return invoices.

1990			
2-12-90	S459067	009077	500 gal.
5-03-90	S459236	009135	500 gal.
3-29-90	S458589	009158	410 gal.
7-24-90	S459372	009224	497 gal.
9-24-90	S460182	009344	500 gal.
11-14-90	L002438	009361	516 gal.*
12-28-90	126079	009858	**
Total			2923 gal.

*Notation on invoice reads: "Returned solvent to bulk plant."

Note: 1) Some drum return deposits and exchanges noticed.

2) Legal file # 009323 - 331 Safety-Kleen setup and product delivery sheets for all divisions of mine ops. (drill, main, ballmill, gas, lube room, secondary crusher). None of these noted in previous years, believe it indicates beginning use of Safety-Kleen services.

** Can not determine gal. or lbs. by sheet.

1991			
12-31-91	7566581	009516	1258 lbs.
1-25-91	449546	009836	1249 lbs.
2-09-91	797286	009903	946 lbs.
3-19-91	142929	009912	946 lbs.
3-21-91	L002805	009920	55 gal.**
4-18-91	490337	009894	946 lbs.
5-15-91	832912	009941	946 lbs.
6-13-91	182273	010068	1032 lbs.
7-17-91	532192	009928	1077 lbs. ***
8-13-91	974743	010054	1261 lbs.
9-11-91	324667	010033	1032 lbs.
10-10-91	78585	010010	1249 lbs.
11-05-91	032781	010124	1175 lbs.
12-03-91	388075	010091	1077 lbs.
Total			14,194 lbs.

* Can not determine amount of gal. to Safety-Kleen

**Not a Safety-Kleen ticket. City Service / Libby, MT.

*** Safety-Kleen ticket copy reprint offset. Paid Jul. 22, probably July ticket.

1992

1-29-92	101127	009570	1163 lbs.
1-29-92	101126	009571	72 lbs.
2-26-92	457973	009582	1077 lbs.
3-25-92	813-250	009598	991 lbs.
4-22-92	225898	009619	991 lbs.
4-22-92	225897	009620	72 lbs.
5-20-92	588356	009637	344 lbs.
5-20-92	M28869	009638	688 lbs.
6-17-92	953027	009675	1118 lbs.
7-15-92	317948	009729	1118 lbs.
7-15-92	317947	009730	72 lbs.
8-13-92	682651	009728	1290 lbs.
9-09-92	046194	009711	1118 lbs.
10-07-92	412064	009774	72 lbs.
10-07-92	412065	009775	1290 lbs.
11-02-92	777754	009806	1290 lbs.
12-02-92	142812	009784	1290 lbs.
Total			14,194 lbs.

*Note: 1 gal. Safety Kleen = 6.8 lb. Costs ran about \$1.00 / lb.

* Safety-Kleen appears to be solvents from clean-up terminals???

Note: In reviewing these purchase orders it appears that during the use of Safety-Kleen services no solvents from other sources are being purchased or show up in records. However, the services of Safety-Kleen appear to be rather expensive compared to earlier purchase and use of solvents at mine.

1993

1-27-93	807586	009687	230 gal.*
3-24-93	46813	009480	192 gal.*
2-24-93	139239	009494	193 gal.*
Total			385 gal.

1997

4-23-97	1451	009449	1 drum*
			6 drums**

* sorbents????

** tank bottoms

1999

9-28-99	934009	009440	30 gal.
2-18-99	841659	009441	30 gal.
Total			60 gal.

2000

12-22-00	00014866701	009404	30 gal.
8-31-00	00013676781	009410	30 gal.*
5-16-00	00012496466	009412	30 gal.*
1-27-00	0011292014	009415	30 gal.
Total			120 gal.

* Safety-Kleen also attached appears to be waste disposal form for Petroleum
Naphtha.????

2001

4-16-01	0017097842	009422	30 gal.*
8-06-01	0018239091	009419	30 gal.*
Total			60 gal.

* Safety-Kleen also attached appears to be waste disposal form for Petroleum
Naphtha.????

2002

10-02-02	0021711374	009395	30 gal.
6-08-02	0020893939	009399	30 gal.*
Total			60 gal.

* Safety-Kleen also attached appears to be waste disposal form for Petroleum
Naphtha.????

2003

12-19-03	0025003097	009383	30 gal.*
8-29-03	0024182332	009390	30 gal.
01-21-03	0022567153	009391	30 gal.
5-13-03	0023376874	009392	30 gal.
Total			120 gal.

NOTES:

1. From 1979 – 82 there are detailed and extensive records of deposits for barrels and returns of barrels for deposits. This radically changes from 1983 on when only one (1) barrel deposit is noted. After 1983 there are no invoices or documentation that indicate deposits are being placed on barrels or barrels are being returned for deposit.
2. Use of solvents doubles around 1982, doubles again in 1983 and appears to peak and level out in 1986-7 at perhaps 3000 gallons.
3. Safety-Kleen corporation comes on the scene about Dec. 1990, prior to this date invoices appear to indicate Asarco is handling purchase, use and disposal of solvents at Troy mine.

Troy Lawsuit Discovery Papers Report

BATES	Doc. #	Subject	Remarks
	1. 5000-5002	Requisition sheets	Only 3 sheets (chemicals) might be useful to see all of them.
	2. 5003-5029	Purchase Order sheets	Some from 1988 - 91
	3. 5030-5058	Purchase Order sheets	Safety-Kleen Corp. for purchase and servicing of (3) solvent units used in cleaning mechanic shop parts. 1990-93. Monthly procedure that cost between \$800-1500.
	4. 5059-5061	Purchase Order sheets	Appears to be for lab equipment etc.
	5. 5063-5109	Purchase Order sheets	Diesel, hydraulic oils, antifreeze and grease, 7/90 - 7/91. P.O.S. #5096 for 1-55 Gal. Barrel of safety solvent, no I.D of what type.
	6. 5110-5113	Purchase Order sheets	Dowfroth 2500 & Orzana 4/90 - 1/91.
	7. 5114-5120	Purchase Order sheets	Drill grease 7/90 - 5/91
	8. 5121-5173	Purchase Order sheets	Diesel fuel, oil, transmission fluid and other lubricants. Lubricant comes in 120 lb. barrels.
	9. 5174-5195	Requisition sheets	Antifreeze, diesel, motor oil and used oil pickup. Assorted dates 1993-97. None from before shutdown in 1993.
	10. 5196-5222	Purchase Order sheets	300,000 lbs. Potassium Amyl Xanthate, pellet form packed 341 lbs. in 55 gal. drums. 7/91-12/93. Sheet # 5207, item #4 discusses \$20.00 drum deposit. Sheet # 5217 itemizes sale of scrap metal from tailings area @ \$25 / ton. Other scrap iron sales on other sheets.
	10.1 5213-14	Purchase Order dated 1/5/92	Safety-Kleen concerning mineral spirits solvents and maintaining 12 machine site wash tanks. Service: solvent replacement, removal, cleaning tanks, waste removal.
	11. 5223-5228	Halogenated solvent registration, # 1118	MT Dept. Health & Environmental Services. Registration / reporting form signed by Terry Erskine 10/3/89. Doug Lorentzen listed as Maintenance Supervisor to whom questions may be directed concerning solvent use at facility. Indicates use of at least 440 gallons per year of halogenated solvents, w/dischARGE to drywell, septic tank, lagoon or other industrial sewer.

*Note: MDEQ publication "The Small Business Handbook for Managing Hazardous Waste (2/02) on page 9 states: "Small generators may not dispose of hazardous wastes on their property **without** a permit issued by the MDEQ."*

12.5229-5262 Of note: #5244 Letter 11/14/96 about 500 lbs. used grease stored at mill from 4/93 shutdown. Disposed of on 3/94.

#5247, **Material Safety Data Sheet about Surtac Plus (Trichlorethylene), CAS 799-01-6, black paste with solvent odor.** *This wording precisely validates Jim Meyer's description of the contents he claims was in barrels buried at the Troy tailings impoundment.*

13. 5263-5410 Of note : #5295 Field Investigation Report 2/18/93 about Hazardous Waste Manifest for shipment of 2,225 lbs. of D001/D008 wastes shipped to Romic Chemical Co., Palo Alto, CA. **Shipment of special solvent used to clean ball mill.**

#5297 Memorandum 2/19/93 Re: inspection by Bob Reinke, Solid & Hazardous Waste Bureau. Item. # 4 concerning underground storage of barrels.

#5307 Letter 9/12/91 from DHES to L. Erickson / Asarco about requirement of large and small hazardous waste generators to maintain log books or similar documentation. Want to make sure we ask for this log-book.

#5309 Field Investigation Report, daated 10-28-91, Robt. Reinke. Of note: Discussion on floor drains in last para. Last paragraph, incomplete sentence, "Floor drains do not???????? Following page is missing.

#5317 Memo to file from L.A. Erickson (Unit Manager) dated 8/23/90 concerning inspection by Robert Reinke (MDEQ). (B) **"Solvent tanks / disposal in 55 gal. barrels."**

#5317 Memo to File from L..A Erickson, R.e. Montana Solid & Hazardous Waste Bureau concerning Robt. Reinke inspection of 8-22-90. Item (I)(B) "Solvent tanks / disposal into 55 gal. barrels ? Moore Oil hauls away?. Jan. - Aug. used 1600 gal. 2500 gal./yr.(1990)."

#5328 Letter from Robert Reinke (DHES) 10/1/90 to L. Erickson. 1. Asarco Troy generating @12,940 lbs. of stoddard solvent per year. Asarco must justify haz-waste exemption or submit notification as small generator. 2. Asarco four drum of stoddard solvent sent to Moore Oil of Libby that qualified as hazardous waste. Moore Oil not allowed to handle such material. "Stoddard solvent is a hazardous waste." **What was eventual disposition of 4 barrels of stoddard?**

#5330-31 Field Investigation Report (R. Reinke / DHES) 8/22/90 discussing waste practices. Mine generates 2000 gal. stoddard product per year. **6 / 55 gal. barrels (used for stoddard) stored behind shop. Lee McKinney explanation of waste management for assay lab.** Spent acids discharged to float machine.

✓ #5341-43 Solvent consumption sheets. Great explanation: "Trichlorethane – Chloroethane sp. gr.=1.320, consumption: 440 gal. in 1987 = 37 gal. / mo.= 404 lb./mo. Disposal; evaporates; " ? sludge goes into sewage treatment facility then to?"

#5344-45 Document dated 7/6/88. Reagents used at Troy Unit regarding Toxic Substances Control Act. #7. Chloroethane (1,1,1, Trichlorethylene, usage: 450 gallons / yr. #8 Varasol, solvent / degreaser, usage: 3600 gal. / yr. #9 Stench (Ethanol trichchloroflourmethane, usage 22 lbs. / yr.

#5346 Letter Asarco 7/19/79 Toxic Substances Control Act sheet for various NW mining divisions of Asarco. Couer Unit usage of 1600 lbs. sodium cyanide / yr., Galena Unit 2250 lbs. sodium cyanide / yr., Leadville Unit 10,000 lbs. sodium cyanide / yr.

5350 Material Safety Data Sheet from Dow Chemical concerning Chloroethene VG @ Solvent. Warnings: "Keep out of water supplies. Dumping into sewers, on the ground, or into any body of water is strongly discouraged, and may be illegal."

14. 5411-5634

#5411 Memorandum 7/9/97 Joe Luchini / Asarco; concerning Hazardous Waste disposal at Libby landfill and disposal of vehicle oil filters. Not to be done in one large amount at one time. *Which begs the question of how they were being disposed of at the mine while it was in operation?*

#5414-32 Small Generator Annual Haz-waste Report Form to MDHES for 1993. #5415 speaks to Safety-Kleen issue and Petroleum Naptha waste.

#5431 Part Two (waste description sheet) of 1993 haz-waste Report Form to MDHES. Other chemicals, Trichloroethylene, Toulene, Sodium Hydroxide. Trichloroethylene waste amount generated 155 lbs. (8 lbs. / gal.) and amount shipped off-site, 155 lbs. does not square with (Item 13, #5344-45 which indicates 450 gallons of Trichloroethylene usage annually. Explanation and reconciling needed!

#5433-46 Small Generator Annual Haz-waste Report Form to MDHES for 1992.

#5434 Waste description sheet: Waste Petroleum Naptha Combustible Liquid amount generated 14,886 and amount shipped Off-site 14,886 is considerable and worthy of note!

#5446 Part Two (waste description sheet) reporting on Tetrachloroethylene, Toulene, Sodium Hydroxide. Tetrachloroethylene waste amount generated 155 lbs. (8 lbs. / gal.) and amount shipped off-site, 155 lbs. does not square with (Item 13, #5344-45 which indicates 440 gallons of Trichloroethylene usage annually. Explanation and reconciling needed!

#5470-84 Small Generator Annual Haz-waste Report Form to MDHES for 1991.

#5472 Waste description sheet: Waste Petroleum Naptha Combustible Liquid, 14796.

No other haz-waste reported, No Part Two waste description sheet for Trichloroethylene, Toulene, Sodium Hydroxide. Hard to believe use of such chemicals at mine would have changed so radically. Explanation and reconciling needed

#5485-93 Small Generator Annual Haz-waste Report Form to MDHES for 1990.

#5487 Waste description sheet. Waste Petroleum Naptha Combustible Liquid, 5209.
No other haz-waste reported, No Part Two waste description sheet for Trichloroethylene, Toulene, Sodium Hydroxide. Hard to believe use of such at mine would have changed so radically. Explanation and reconciling needed

#5545- Invoice & Report for Professional Services / Maxim Technologies. "Chemical analysis of soil for Benzene and TCLP Metals. Invoice is addressed Ken. O'Toole, Northern Lights, Inc., P.O.B. 310, Sandpoint, ID 83864. Appears to be related to work done at Troy mine.

#5576-5634 Asarco letter dated 7/3/96 Environmental Practices Survey for the IRRC Environmental Profile on Asarco's Troy Unit. "A possible omission to the supporting detail for the IRRC profile is an environmental lawsuit filed against the Asarco Troy Unit (EPA ID. No. MTDO96199989). The accusation states water is being discharged from the tailings impoundment area without an MPDES permit."

#5583 Para. 6. Discussion on severe water pollution at Squirrel & Kerber Creek Mine near Bonanza, CO. "The wastewater, which resulted from the mine contamination, contains 75 times the safe level of iron, and may cost up to \$10 million to cleanup.(RMN7/26/94)

#5619 40CFR Ch.1 (7-1-91 Edition) Subpart D – Record-keeping and Reporting
262.44 (a) A generator must keep a copy of each manifest signed in accordance with 262.23(a) for three years or until he receives a signed copy from the designated facility which received the waste.

15. 5663-5754

#5663-66 Asarco letter 3/17/82 Dean Knight to R.Thorvilson DHES to the effect that "Asarco's Troy Unit is a mining operation excluded from regulation by the Bevill Amendment." Clarification of Status form. X on "The following temporary exemptions of ARM 16.44.304 (and 40CFR 261.4) remove the need for hazardous waste management facility status for my business."

#5667-8 Asarco letter 8/12/80 Jack Bingham to EPA Region VIII. Notification of Hazardous Waste Facility, Troy Unit, Asarco, Inc.

#5742-49 Asarco letter to All Plant Managers dated 7/30/80 concerning Chemical Substance Inventory program initiation. #5744-49 are an example of all substances

/materials used by Wallace Office. Troy mine would / should have one of these and we should ask for it. Lists all substances and suppliers..

16. 5755-5867 Safety-Kleen Corp. service & delivery sheets 7/99 through 12/90.

17. 5868-6008

#5922 Attention Jack Tuholske: American Cyanamid Invoice (6/29/84) for 13 tons Aluminum Sulfate. Important for Rock Creek MPDES permit that dropped Aluminum monitoring criteria.

#5924 City Service Co. / Libby, MT. Invoice for bulk solvent @ 474 gallons.

#5944 Attention Jack Tuholske: Univar Co. Invoice (8/6/84) for 4000 lb Aluminum sulfate Gran. Same as 5922.

#5954 City Service Co. / Libby, MT. Invoice (10/9/84) for 4 barrels Unoba A2 grease. I believe this is one of the lithium greases used.

#5957 City Service Co. / Libby, MT. Invoice (10/5/84) 520 lbs. or gal. Solvent Rel.

#5969 City Service Co. / Libby, MT. Invoice (11/20/84) 300 lbs. or gal. bulk Solvent (Hyd).

#5993 City Service Co. / Libby, MT. Invoice (12/21/84) Gear lube 2,460 lbs.

#5995 Attention Jack Tuholske: Van Waters & Rogers / Univar. Invoice (12/21/84) 912 - 50 lb. bags of Aluminum Sulfate. Same as #5922

#6002 Attention Jack Tuholske: Van Waters & Rogers / Univar Invoice (10/31/84) 25-100 lb. bags of Aluminum Sulfate Gran. Same as #5922.

#6005 Nalco Chemical Co. Invoice (10/25/84) 7-55 gal. barrels Nalco 3DD758 Liquid D8D. Only Invoice found signed for by Lee McKinney.

18. 6009-6129

#6048 Asarco memo concerning EPA inspection by Charles Frank Stagsdell of 5/14/81 Stan Stoker (electrical foreman) to Jack Bingham concerning EPA inspection of electrical equipment that may contain PCB's. *Interesting persecution complex letter for the times.*

#6116-8 Hazardous Waste Manifest 8/19/94 transport of PCB's from Troy mine, App. 260 gallons and 5 transformer carcasses.

#6124-29 Six forms of some sort indicating testing of PCB contaminated transformers. Notation at bottom of each: "Letter from Continental No PCB's 9-18-80. Tested 4/25/86 56-360 PPM Contaminated." Note #6124-29 indicates 6 contaminated transformers and #6116-8 is invoice for removal of 5 carcasses.

19. 6130-6362

#6130-38 Asarco PCB distribution list. #6138 (4) "Drained and decontaminated carcasses from PCB transformers and soil, rags and other debris (at any conc. PCB's) can be landfilled. Does this include landfilled within a tailings impoundment?

#6157 Asarco letter 6/20/95 D. Young - Mike Balboni (USFS) concerning 5 PCB transformers at Troy mine.

#6189 Bid from Continental Transformer Services Inc.(5/20/94) to remove PCB contaminated transformers. Retrofill / \$23,370. Disposal / \$18,065.00.

#6191 Transformer status & location sheet.

#6205-6 Asarco letter 6/9/94 New bids to remove PCB contaminated transformers. Half as expensive.

#6282 Low concentration PCB spill clean-up certification. Clean-up done with Triclorethane.

#6283 EPA Inspection report 7/25/90 on leaking transformers. Sheet indicates seven (7) leaking transformers. See #6116-8 and #6124-9 for discrepancies.

#6295 Manifest Number 510068. PCB debris drum content from Coeur, Galena and Troy operations.

#6297 Letter (6/6/91) from Ensco to Asarco about PCB debris pick-up and cost / 1-55 gal. bbl. \$1200.

#6302 Asarco Letter 7/27/90 to EPA concerning low-concentration PCB oil stain on two transformer & clean-up.

20. #6363-6414 NA

21. 6415-6844 #6443-4 Industrial Hygiene & Toxicology Data Sheet. Trade name: Super Permalube Grease, important components: Lithium thickener.

#6445-6 Industrial Hygiene & Toxicology Data Sheet. Trade name: Amoco Lithium Multi-Purpose Grease, important components: Lithium thickener.

#6513-16 Copper concentrate: attention to Jim Meyer's symptoms.

#6544 J.T. Baker Chemical Co. / Barium nitrate / oxidizing agent. Material Safety Data Sheet, Attention: Summit Envirosolutions.

#6552-55 J.T. Baker Chemical Co. / Barium Chloride, Dihydrate. Material Safety Data

Sheet, Attention: Summit Envirosolutions.

#6640-43 J.T. Baker Chemical Co. / Sodium Cyanide, Material Safety Data Sheet.

#6672-74 Cyanamid / Material Safety Data sheet: Superfloc 355 Flocculant. Waste Disposal: "Disposal must be made in accordance with applicable government regulations."

22. 6853-7186 #6853 Material Safety Data Sheet. Product name: Chlorothene ® SM Solvent.

#6942 Material Safety Data Sheet. Product name: EutecTrode 2101. Ingredients:
Lithium Chloride 1-5%.

#6960-1 Material Safety Data Sheet. Chemical name; Sodium Thiocyanate. Trade name: Sodium Sulfocyanate.

#7025-26 Fiske Brothers Refining Co. / Material Safety Data Sheet. Product name: Lubriplate Gear Shield. Formula: Lithium soap, mineral oil, and additives.

23. 7187-7717

#7244-45 Material Safety Data Sheet. Product name: Hi-Tri ® Solvent / Ingredient:
Trichloroethylene.

#7247-8 Material Safety Data Sheet. Product name: Cool It. Formula: 1,1,1,
Trichloroethane.

#7330-33 Material Safety Data Sheet. Product name: Sodium Hydroxide. Waste disposal method: PPG recommends disposal of neutralized material in an approved hazardous waste management facility.

#7410-11 Transco NW. Inc. Product name: SC2000 / rubber adhesive. Formula: rubber compound kept liquid with trichloroethylene.

#7414 Shell Oil Co. Material Safety Data Sheet., Product name: Shell Alvanla ® Grease 2. Hazardous ingredients: Lithium soaps.

#7455-58 Thermacote Welco Co., Material Safety Data Sheet, Product name: Welco 1620 Anti-Spatter, Hazardous mixtures: *Specifically, Trichloroethane III.

#7459-61 Technichem Corp., Material Safety Data Sheet (MSDS), Product name: Heavy Duty Hot Tank Powder, Hazardous mixtures: Sodium Hydroxide 70-90%, Sodium Metasilicate 10-20%.

#7464 Thatcher Chemical Company, Material Safety Data Sheet, Product: Aluminum Sulfate Solution. Attention: Jack Tuholske, RCM MPDES permit.

#7470-73 Tip-Top Stahlgruber, MSDS, Product name: SC-2000 Cement, Chemical Characteristic: Trichlorethylene.

#7485-87 Texaco Inc. MSDS, Product name: 01922 Molytex EP 2, Chemical / Common Name: Lithium 12-hydroxystearate.

#7491-93 Texaco Inc., MSDS, Product name: 921 Multifax EP1, Components: Lithium Hydroxide CAS# (1310-65-2) 0.8%.

#7494-97 Texaco Inc., MSDS, Product name: 995 Multifax EP2, Components: Lithium hydroxide 4-10.99% CAS # (1310-65-2)

#7498 Texaco Inc. Dear Customer letter 2/10/87 to Asarco, P.O.B. 868, Troy, MT 59935 "Enclosed are Industrial Hygiene, Toxicology, and Material Safety Data Sheets for Texaco products you recently purchased."

#7515-17 Texaco Inc., MSDS, Product name: 01970 Crater 2X Fluid / Gear lubricant, Precautionary label: Contains 1,1,1,-Trichloroethane.

#7518-20 Texaco Inc., MSDS, Product name: 1970 Crater 2X Fluid / Asphaltic lubricant, Composition: 1,1,1, Trichloroethane 18.5%.

#7521-23 Texaco Inc. MSDS, Product name: 958 Marfax Multi-purpose 2 (grease), Composition: Lithium Hydroxide 0.78% CAS#(1310-65-2)

#7524-26 Texaco Inc. MSDS, Product name: 901 Multifax EP 0 (grease), Composition: Lithium Hydroxide 0.6% CAS#(1310-65-2)

#7561-63 Texaco Inc. MSDS, Product name: 00958 Marfax Multi Purpose 2 (grease), Chemical / Common Name: Lithium 12-hydrosstearate 4-10.99%

#7589-91 Texaco Inc. MSDS, Product name: 1939 Premiiium RB Grease, Composition: Lithium 12-hydroxystearte 15-20%.

#7668-71 Vulcan Chemicals MSDS, Product name: Trichloroethylene 100%

#7688-90 The Whitmore Manufacturing Co. MSDS, Product name: Surtac Plus / black paste w/solvent odor, Trichlorwthylene, CAS (79-01-6)

24. 7718-39 NA

25. 7740-7771

#7746 Used grease and oil pickup sheet 1/9/1991- 11/5/94. Other miscellaneous similar sheets. No data sheets prior to 1991.

26. 7772-7803 Water Pollution Prevention, Spill Prevention and Countermeasure Plan (SPCC) for Asarco Troy Unit.

#7773 Memorandum Brian Hansen (MFG) to Dave Young (Asarco) "I have reviewed SPCC plan..... *The plan indicates that gasoline, diesel fuel, oil, used oil, cleaning solvent and calcium hypochlorite are used on site. I was only able to find discussions on how the diesel fuel and used oil are handled and stored. It would be helpful to briefly discuss how the other substances are handled and stored. * The sections on Inspections and Records and Personnel, training and spill prevention procedures [corresponding to 40 CFR Part 112.7(e)(8)(i) and (e)(10)(iii), page 5 of the plan] should note that records are kept concerning the inspections and training."

It might be beneficial to ask for these documents (Inspections & Records) and maybe even Personnel!

27. 7804-7874 Reagent Consumption Use, Grinding Media and Power Use 1984-1993.
Notes for Jack Tuholske: Aluminum Sulfate use begun in milling process in Dec. 1984. During a period of 7 months in 1985 Aluminum Sulfate use averaged 9 ton per month. Also grind ball size & use varies considerably. ALL indications of some experimentation. July of 1986 work with Aluminum for water clarity completed. Still experimenting w/ milling process. Frother mix replacement.

Last three months of 1987 experimentation w/grinding medium (ball size).

1988, 89 & 90 Reagent Consumption Use, Grinding Media and Power Use MISSING???
See e-mail of 2/5/04 to Jim Kuipers, cc: Bechtold & Rossbach.

1991 Lot of play in the flocculant field.

1992 Not a lot to tell.

1993 Only 3 months of real operation.

28. 7875-7887 Flotation tailing report sheets. ????

29. 7888-7907 Distribution letter concerning Federal Rail Administration / DOT inquiries as to Asarco shipments of Hazardous Waste via rail, primarily heavy metals and contamination.

30. 7908-8636 Filler junk. NA

Recommendations for FURTHER DISCOVERY REQUEST and other questions.

A. See # 9. Should ask for "Requisition sheets" for all oils, grease and lubricants from 1981-93.

B. See #10. Purchase of Potassium Amyl Xanthate. Quantity use confirmed by Reagent Consumption Use sheets indicate at least 900+ / 55gal. barrels came into plant with this product alone. If there was a \$20 barrel deposit as indicated then Asarco was ponying up \$18,000 in deposit money per year on this reagent alone. This is enough deposit money to want to recover. There would have to be some methodology for tracking the return of all these barrels to recover the deposit. **DISCOVERY.**

C. See #11. Question; are the halogenated solvents described Trichloroethylene?

D. See #13 / 5307 Hazardous Waste logbook described in this document. Request copy through **DISCOVERY.**

E. See #13 / #5330-31 Field Investigation Report (R. Reinke / DHES) 8/22/90 discussing waste practices. Mine generates 2000 gal. stoddard product per year. **Lee McKinney explanation of waste management for assay lab.** Spent acids discharged to float machine. *This possibly corroborates some of Greg's testimony about lab waste reporting to tailings impoundment via tails thickener.*

#5341-43 Solvent consumption sheets. Great explanation: "Trichlorethane – Chlorethane sp. gr.=1.320, consumption: 440 gal. in 1987 = 37 gal. / mo.= 404 lb./mo. Disposal; evaporates; ? sludge goes into sewage treatment facility then to?"
See 11 above.

F. See #14 / #5431 Part Two (waste description sheet) of 1993 haz-waste Report Form to MDHES. Other chemicals, Trichloroethylene, Toulene, Sodium Hydroxide. **Trichloroethylene waste amount generated 155 lbs. (8 lbs. / gal.) and amount shipped off-site, 155 lbs. does not square with (Item 13, #5344-45 which indicates 450 gallons of Trichloroethylene usage annually. Explanation and reconciling needed!**
DISCOVERY.

G. See #14 / #5470-84 Small Generator Annual Haz-waste Report Form to MDHES for 1991.

#5472 Waste description sheet: Waste Petroleum Naptha Combustible Liquid, 14796.
No other haz-waste reported, No Part Two waste description sheet for Trichloroethylene, Toulene, Sodium Hydroxide. Hard to believe use of such chemicals at mine would have changed so radically. Explanation and reconciling needed.

#5485-93 Small Generator Annual Haz-waste Report Form to MDHES for 1990.

#5487 Waste description sheet. Waste Petroleum Naptha Combustible Liquid, 5209.
No other haz-waste reported, No Part Two waste description sheet for Trichloroethylene, Toulene, Sodium Hydroxide. Hard to believe use of such at mine would have changed so radically. Explanation and reconciling needed

DISCOVERY : We want to know if Trichloroethylene was purchased and used in these years. Relate to A & D above that might answer this question. If it was purchased and

used as in previous years then at least 155 gallons should show up as having gone to hazardous waste site as per F above.

H. See #15 / #5742-49 Asarco letter to All Plant Managers dated 7/30/80 concerning Chemical Substance Inventory program initiation. #5744-49 is an example of all substances /materials used by Wallace Office. Troy mine would / should have one of these and we should ask for it. Lists all substances and suppliers. DISCOVERY

I. See #18 / #6124-29 Six forms of some sort indicating testing of PCB contaminated transformers.

Notation at bottom of each: "Letter from Continental No PCB's 9-18-80. Tested 4/25/86 56-360 PPM Contaminated." Note #6124-29 indicates 6 contaminated transformers and #6116-8 is invoice for removal of 5 carcasses. This could relate to unsolicited EPA brochure sent to CRG from EPA Region VIII, Denver.

J. See #25 / 25. 7740-7771

#7746 Used grease and oil pickup sheet 1/9/1991- 11/5/94. Other miscellaneous similar sheets. No data sheets prior to 1991.

Relate back to # 13 / #5328 Letter from DHES 10/1/90 to L. Erickson / Asarco four drums of stoddard solvent sent to Moore Oil of Libby that qualified as hazardous waste. Moore Oil not allowed to handle such material. Possibility that Moore Oil of Libby was regularly picking up stuff they shouldn't have been.

Appendix B

Mine Water Balance Analysis



50 West 14th Street, Suite 200
Helena, Montana 59601
tel: 406 441-1400
fax: 406 449-7725

Memorandum

To: Emily Corsi, Herb Rolfes, DEQ

From: Bill Bucher, Kim Chase - CDM

Date: April 12, 2010

Subject: Mine Water Balance Analysis - Troy Mine

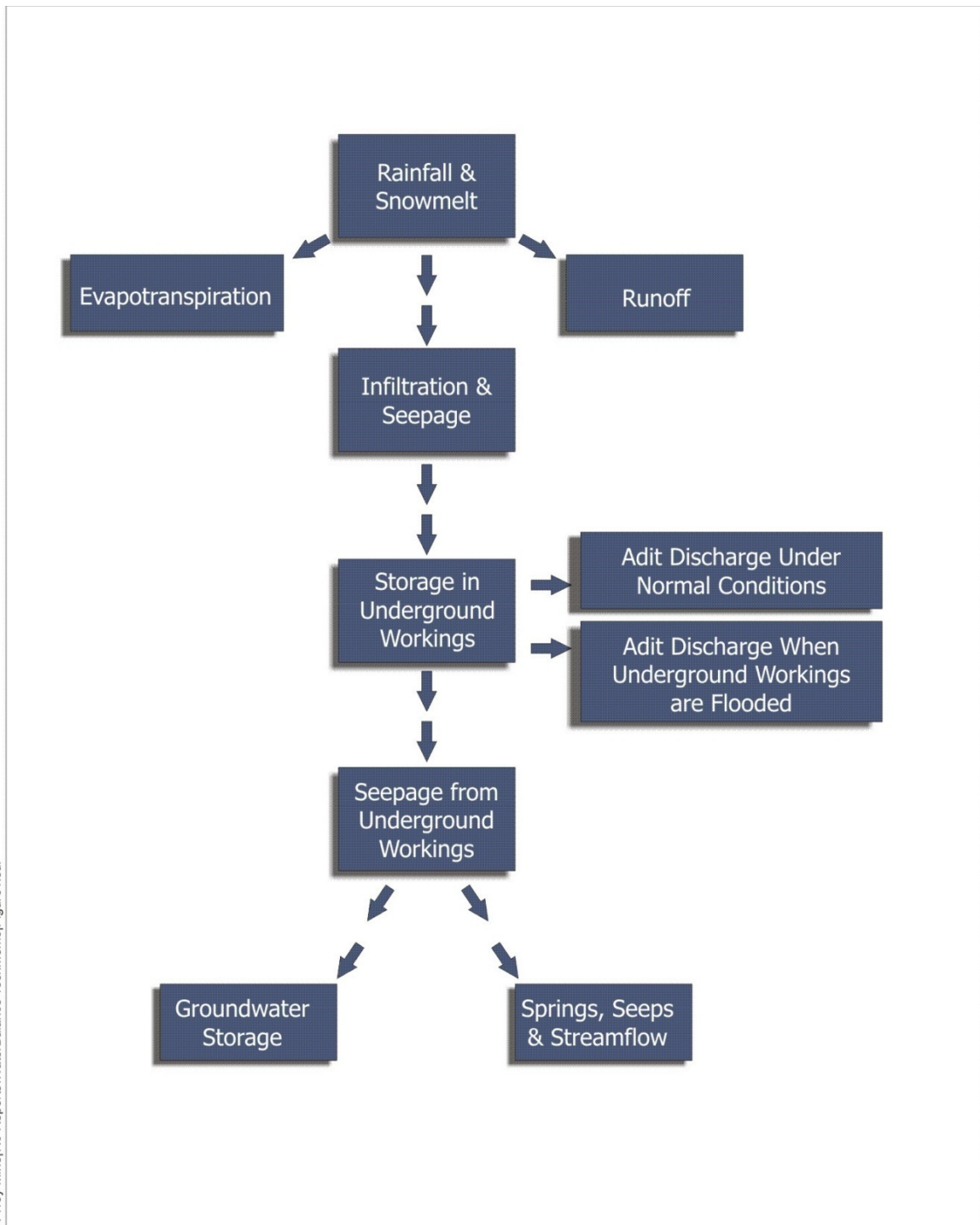
CDM has been retained by the Montana Department of Environmental Quality (DEQ) to develop a water balance model for the Troy Mine to predict expected discharge from the mine workings, if any, after mine closure. The purpose of this memorandum is to present the results of our analysis of existing data that we used to predict whether the Troy Mine will fill with water after pumping ceases to a level that will allow discharge of mine water to the surface. If the mine workings are predicted to discharge, how long it will take to do so and how much discharge may be expected will also be predicted. A review of the existing data and the conceptual water balance are available in a Technical Memorandum prepared by CDM dated January 16, 2009.

The Troy Mine is an underground copper and silver mine located south of the town of Troy in Lincoln County, MT. ASARCO began operating the mine in 1982 and halted production in 1993 due to low metals prices. The mine was sold to Revett Silver Company and production resumed under Genesis, a subsidiary of Revett, in 2005. In 2006, Genesis submitted a Revised Reclamation Plan, for which an Environmental Assessment (EA) must be completed. In support of the EA, this Memorandum assesses the possibility of mine flooding discharging to the surface.

Conceptual Water Balance Model

The conceptual water balance is shown in Figure 1 and is discussed in detail in the January 16, 2009 Technical Memorandum. The water that enters the mine all originates as rainfall or snowmelt above or near the underground workings. This precipitation, minus the portion lost to evapotranspiration, overland or subsurface runoff, enters the groundwater system adjacent to the mine workings and potentially seeps through the mine walls or enters fractures that connect to the mine. The water received by the underground workings may be held as storage or it may seep out through fractures or porous formations. If the underground workings are flooded, water may discharge from the service adit, the lowest

Figure 1. Conceptual Model of Mine Water Balance



potential connection to the surface. Under normal conditions, there is some discharge from the service adit due to water entering from the adit walls and back. It is only the water entering or leaving the mine workings that is of concern in this analysis of whether the or not the mine workings will flood to an elevation which will result in discharge from the mine workings via the service adit.

The conceptual model of mine flooding developed in this study is based on the following assumptions about groundwater in the vicinity of the mine:

- The geology of the rock around the mine workings is fractured metamorphic rock, which allows significant flow through the fractures.
- The aquifers adjacent to the mine workings are local, unconfined aquifers which have the potential to fluctuate rapidly in response to precipitation and runoff patterns as well as mine dewatering activities.
- If pumping ceases and water levels recover, there is a potential for water to discharge through the service adit at a known elevation (4225 ft.).

Figure 2 is a conceptual drawing of a flooded mine that will discharge through an adit. In this scenario, water levels have recovered to an elevation sufficiently high to permit overflow through a discharge point, in this case the service adit of the Troy Mine. Figure 3 shows an alternate scenario in which the recovered water level remains below the overflow point and the mine does not discharge. Determining which scenario occurs in the future after pumping has ceased depends on the recovered elevation of the local aquifer. This is essentially a groundwater analysis problem, which requires an understanding of the local geology and aquifers that we do not have. Therefore, the ability to analyze this problem completely is compromised, and we are limited to the storage data collected by the mine operators to interpret the potential behavior of the local systems under future conditions.

Methods

Because no quantitative information is available about the groundwater conditions surrounding the underground workings, it is not possible to directly calculate the mine water's effect on the surrounding water table. Additionally, there is no way to calculate the magnitude of flows into the underground workings or natural seepage out of the workings. However, it is possible to calculate the difference between inflow and outflow from the change in storage over a given time period. The basic method employed here is the analysis of mine inflow volumes as determined from the pumping rates and storage changes measured by the mine operators through the years. The inflow to the mine workings can be defined as the pumping rate plus the change in storage:

Figure 2. Conceptual Model of Mine Discharging through Adit

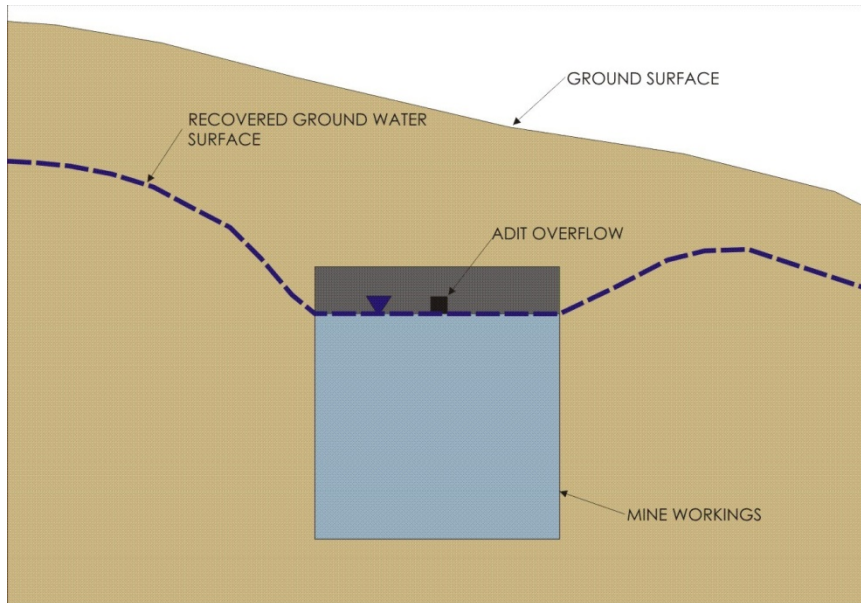
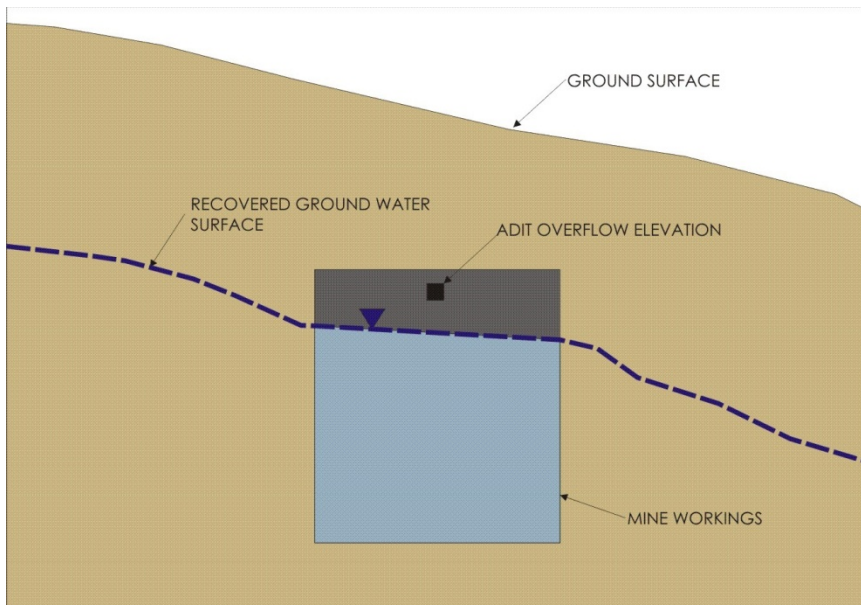


Figure 3. Conceptual Model of Mine without Discharge to Surface



$$I = Q_p + \Delta S$$

Where:

I is the mine inflow for the period
 Q_p is the pumping rate for the period
 ΔS is the change in mine storage for the period.

It is possible that portions of the mine discharge to the surrounding formations where the local water table is below the mine water elevation. This potential is highest when the mine pool is at a high elevation and the local aquifers are depressed after a dry period such as normally occurs in late summer. However, records indicate that this mine generally receives inflow in excess of any outflow, and in this analysis we will use the term mine inflow to mean the net difference between inflow and outflow from the mine workings.

Relationships for inflow were established based on the hypothesis that inflow to the underground workings would be dependent on rain and snowmelt amounts and on water levels in the underground workings. Rain and snowmelt, acting through the groundwater pathway, are believed to be the primary source of water input to the underground workings. When groundwater levels are higher, it follows that the head difference between water in the underground workings and the surrounding groundwater is greater and, therefore, inflow to the underground workings will be greater. Therefore, inflow is proposed to be a function of both recent rain and snowmelt and the elevation of the local groundwater surface:

$$I = f(R, H)$$

Where R is a rainfall and snowmelt and H is a measure of head difference between the mine pool and the local aquifer. The sum of rainfall plus snowmelt is a logical measure of water that has potential to enter the mine. To show that the inflow to the underground workings is related to rainfall and snowmelt, inflow was plotted for each year alongside rainfall and snowmelt.

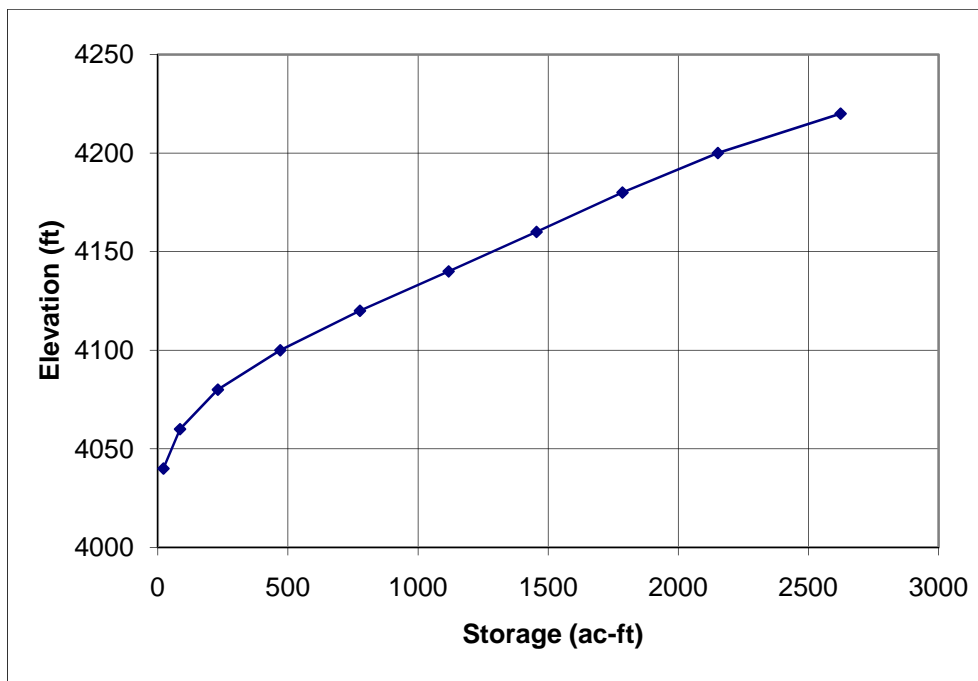
Rain and snowmelt was plotted against inflow over several different time-scales to determine the most appropriate time scale for analysis. This resulted in a method for predicting the inflow to the mine for any precipitation/runoff condition. For the annual time scale, an average value of total rain plus snowmelt was used as the input variable for the years 2004 through 2008 giving a set of four different inflows for these years. Then the actual value of inflow for each year was compared to the value predicted by the linear correlation. This predicted inflow was then plotted against the average water elevation in the underground workings during each year. A second linear correlation was then calculated for inflow as a function of water elevation and was used to predict the expected effect of water level on the change in storage.

With this information, it is possible to estimate the potential for surface discharge from the underground workings and the expected outflow. Further calculations can predict the time until discharge with an assumed initial mine pool elevation and typical precipitation conditions.

Analysis and Results

Our analysis of mine water levels relies on the elevation-storage curve developed by Genesis, Inc. (2004) which is shown in Figure 4 after conversion to acre-feet (ac-ft). This relationship is based on mine surveys of the extent of workings at sequential elevations in the mine with allowance for remaining pillars.

Figure 4. Water Elevation vs. Storage Relation for Troy Mine

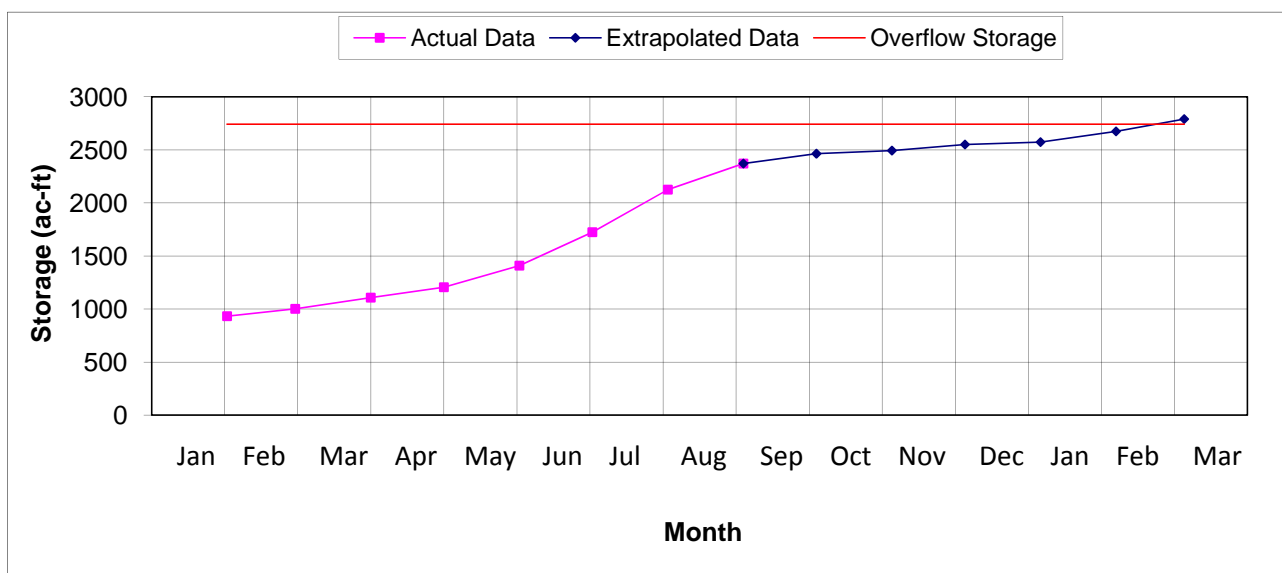


Potential for Mine Flooding

The discharge of mine water from a flooded mine appears to be a real possibility based on data collected in water year 2003 when mine pumping did not occur. Figure 5 plots cumulative mine water storage from January 2003 through August 2003 (end of month values). At the end of August 2003, the storage volume was 2371 ac-ft, which corresponds to an elevation of 4209 ft, the highest recorded water level in this portion of the mine. This level is just 16 feet short of the adit overflow elevation of 4225 ft. In September pumping resumed, the storage volume did not exceed 2371 ac-ft, and pumping continued at a higher rate into 2004, decreasing the storage volume. If the pumping volumes are added to the cumulative volume at the

end of August 2003, it is seen that the storage volume could have continued to increase, potentially exceeding the storage volume of 2,742 ac-ft (corresponding to a discharge elevation of 4225 feet) in February 2004. Thus, if pumping had not resumed, the mine may have started discharging in a relatively short period of time. On the other hand, it is expected that inflow rates decrease with higher water elevations. Therefore, it is also possible that the inflows would have decreased with increasing water elevation sufficiently that the mine would not have discharged. In the following analysis, an attempt is made to determine which of these scenarios is most likely based on analysis of available data.

Figure 5. Measured and Extrapolated Cumulative Mine Water Storage, 2003 - 2004.



Relationship between Inflow and Precipitation Events

As described in the methods section, the relationship for inflow was separated into a function dependent on rain and snowmelt history and a function dependent on water elevation. The first step was to develop the relationship between inflow and precipitation and runoff. This was initiated by examining the temporal sequence of rain and snowmelt at a local SNOTEL site with the inflow to the mine.

Daily precipitation and snow-water equivalent data were gathered from Natural Resources Conservation Service SNOwpack TELemetry (SNOTEL) site 932, Poorman Creek (Natural Resources Conservation Service, 2008). The SNOTEL site is located approximately 14.7 miles to the southeast of Troy Mine. It is located at an elevation of 5100 feet, near the middle of the range of surface elevations directly overlying the mine (4600 feet to 5580 feet). Data for accu-

culated precipitation are available for water years 1999 to 2008 and snow-water equivalent data are available from water year 1969 to 2008. Because of its relative proximity to the mine and its similar elevation, the Poorman Creek SNOTEL site should adequately represent precipitation conditions in the mine area.

The analysis was conducted on monthly, seasonal and annual time scales with special attention paid to lag between surface events and the inflow. The graphs located in Appendix A (Rain+Snow, Storage vs Time) show that the monthly change in storage within the underground workings follows the amount of rainfall and snowmelt very well. However, the lag time between the peak of rainfall plus snowmelt to the peak change in storage is inconsistent. The amount of time that the change in storage peak follows the rainfall plus snowmelt peak varies from zero to two months. In attempting to correlate rainfall and snowmelt with change in storage plus pumped volume during periods of pumping, several time scales were used. The correlation on a monthly basis was weak, with an r-squared of 0.3. When the change in storage was lagged one month behind the rainfall and snowmelt, the correlation weakened even further. The data was then reorganized by season: winter being January, February, March; spring being April, May, June and so on. This improved the correlation to an r-squared of 0.61. On an annual basis, the r-squared improved to 0.73. After removing 2007 from the data, the correlation improved further to an r-squared of 0.87. Because of the poorer correlations for the shorter time scales, the annual time scale was adopted for this analysis.

Figure 6 shows the relationship between mine inflow and rain plus snowmelt on an annual basis for water years 2004 through 2006 and 2008. The data for 2007 is excluded from this analysis because it is suspect for reasons explained later in this memorandum.

Based on the period of record from 1999 to 2008, the average annual rain plus snowmelt for the Poorman Creek SNOTEL site was 79.5 inches. Using the correlation relation established above, this amount of precipitation predicts an average annual inflow into the mine of 1418 ac-ft. To eliminate the effects of rainfall and snowmelt from the inflow relation, predicted inflows were calculated using the relationship in Figure 6 and then subtracted from the actual inflows measured in the corresponding years. These residual amounts were then added to the average annual inflow of 1418 ac-ft resulting in a set of adjusted inflows that are independent of rainfall and precipitation effects to the extent that the data allow. Table 1 summarizes this calculation.

Figure 6. Correlation between Mine Inflow and Rain plus Snowmelt for Water Years 2004 – 2006 and 2008

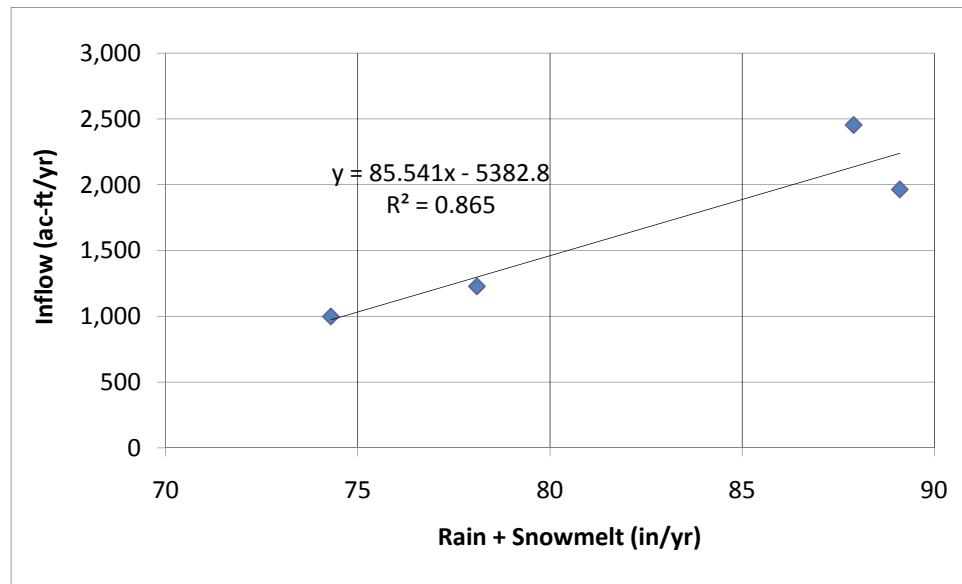


Table 1. Variables and Results of Adjusted Inflow Calculation.

Year	Average Elev (ft)	Measured Inflow (ac-ft)	Rain+snow (in/yr)	Predicted Inflow (ac-ft)	Residual (ac-ft)	Adjusted Inflow (ac-ft)	Q (cfs)
2008	4082.8	2454	87.9	2130	324.2	1736.4	2.4
2006	4169.9	1228	78.1	1291	-63.5	1348.8	1.9
2005	4189.8	999	74.3	966	32.6	1444.9	2.0
2004	4189.9	1965	89.1	2233	-267.5	1144.7	1.6

Relationship between Inflow and Water Elevation

Cumulative mine inflows for the years that the pool elevation was being controlled by pumping are plotted in Figure 7. These curves show that mine flow is generally relatively small and consistent through the fall and winter months and increases greatly, but by varying amounts, in the spring before tapering to lower rates in the summer. As shown in Figure 8, which shows storage levels in the mine throughout this period, stored water was rather high in 2004 through 2006, averaging 1854 ac-ft. In 2007, the water level was drawn down and remained relatively low through 2008. Figure 7 shows clearly the effects of maintaining a low storage level in 2008: the annual inflow for this year was the highest in the pumping period although annual rain plus snowmelt was higher in 2004.

Figure 7 also shows that mine inflow was negative during the winter in 2007. This suggests that water was flowing out of the mine for this period by means other than pumping. Inspection of Figure 8 shows that the mine was being rapidly drawn down through this period. That water would not be flowing into the mine during a period of rapid drawdown is physically difficult to explain and it is suspected that either the storage data or the pumping rates are in error during this period. Therefore, data from water year 2007 have not been used in this analysis.

With a relation between rainfall plus snowmelt and inflow established on an annual basis, the residual information can be used to develop information on the relation between water elevation and inflow. Conceptually the mine workings can be thought of as a large diameter well

Figure 7. Cumulative Inflow for Water Years 2004 through 2008

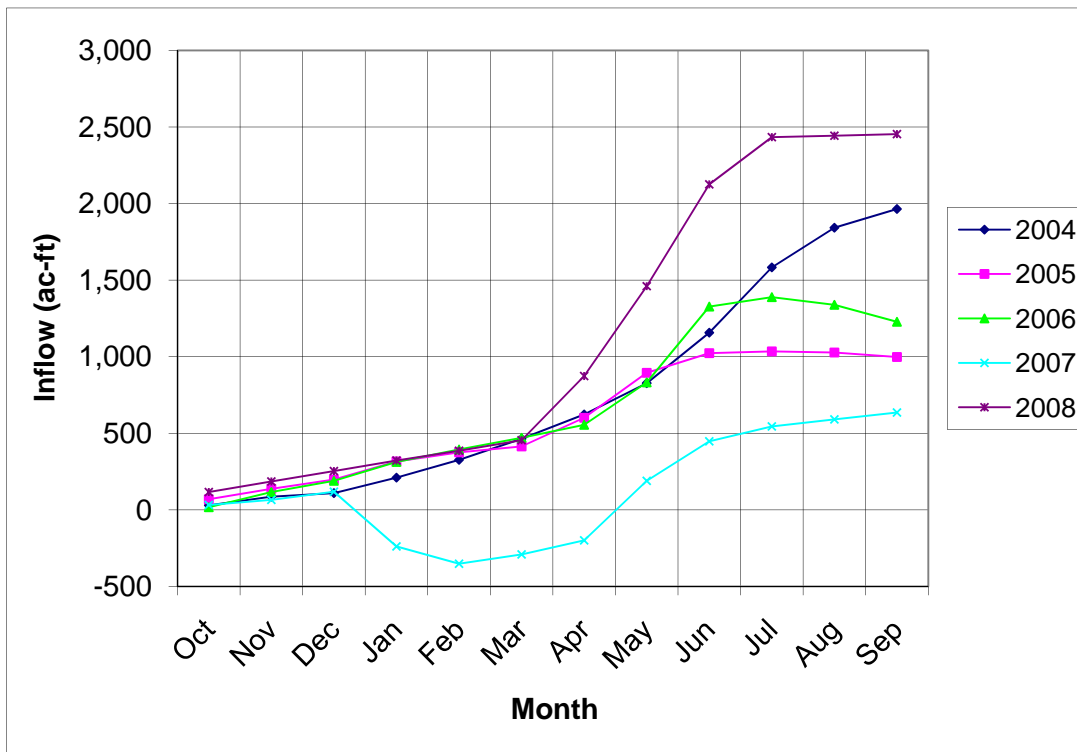
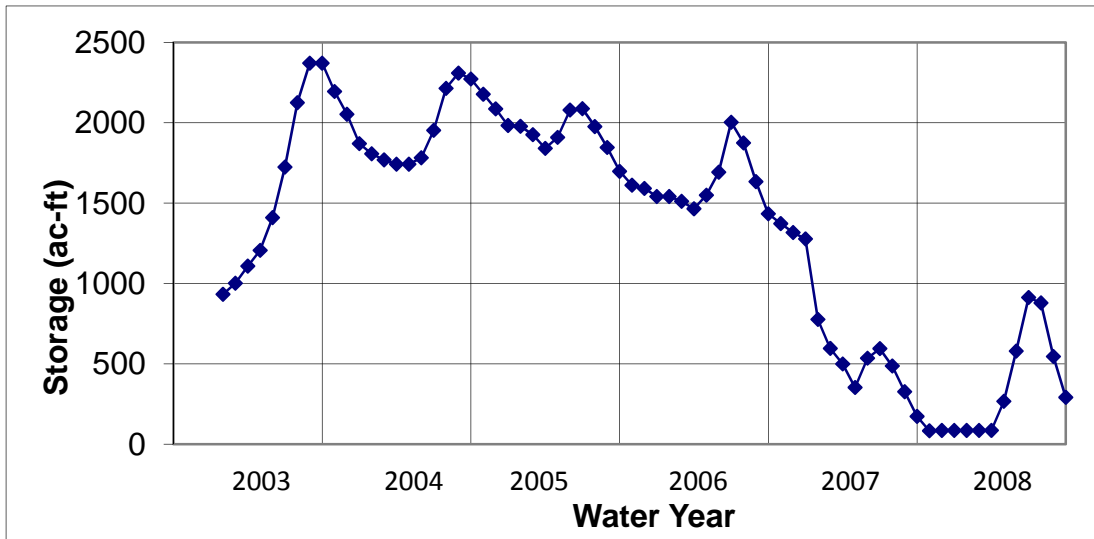


Figure 8. Mine Water Storage 2003 through 2008.



located within an unconfined aquifer. Because the time period for pumping in this analysis is a year, which is relatively long in most pumping applications, the relation of inflow and water elevation should approximate a Theis curve (Freeze and Cherry, 1979). The equation for the Theis curve is:

$$h_0 - h = \frac{Q}{4\pi T} W(u)$$

where

$W(u)$ is the well function

$$u = \frac{r^2 s_y}{4Tt}$$

h_0 is the initial height of the water table

h is the height of the water table at distance r and time t

Q is the flow into the well

r is the radial distance from the well

S_y is the specific yield for an unconfined aquifer

T is the transmissivity of the aquifer

t is time since initiation of pumping

This linear equation relates water elevation (h) and inflow (Q). Plotting water elevation versus inflow for the four years of data presented in Table 1 and fitting a linear regression line through the points results in the graph shown in Figure 9. The r-squared for this correlation is 0.75. Extrapolating the line to the point where $h = h_0$ results in a predicted initial value (or recovered value) for h_0 of 4413 ft., well above the adit overflow elevation of 4225 ft. This suggests the water level in the mine will rise to the adit discharge level when pumping ceases.

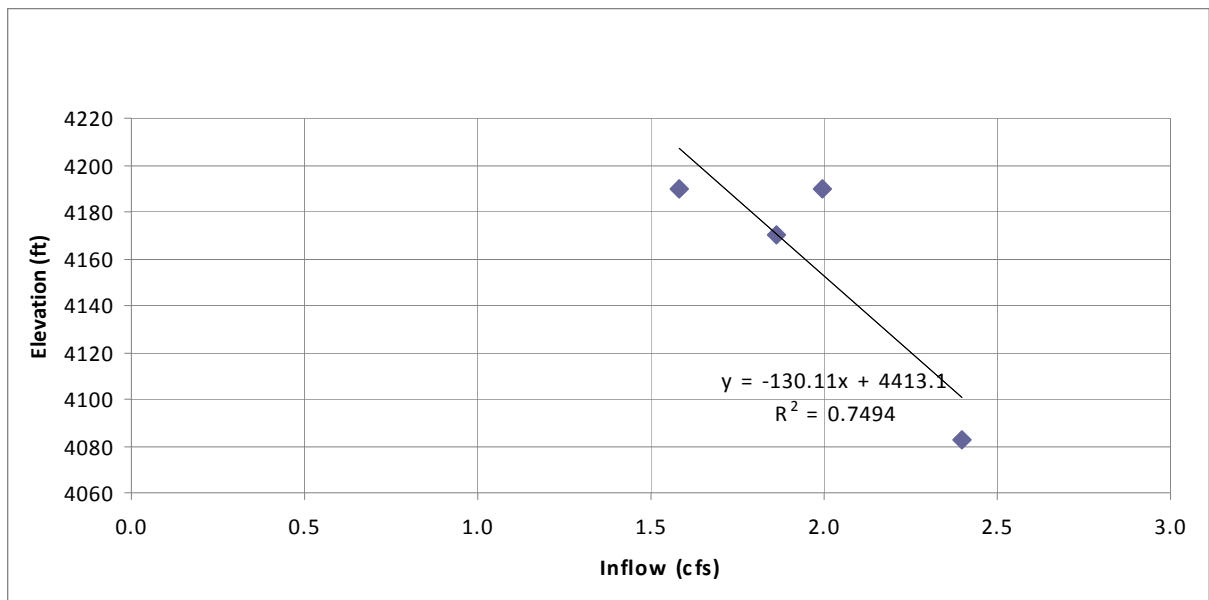
As a check on the validity of the Theis equation in this application, it is shown that a consistent set of aquifer parameters can be derived from the slope of the linear regression (negative 130). If we assume that the typical radius at the edge of the mine workings is related to the area of workings at elevation 4220 ft, the radius for computational purposes is about 800 ft. We can also use a specific yield of 0.05 for fractured crystalline rock (the range is typically zero to 0.10, Freeze and Cherry, 1979), we can calculate from the Theis equation that the formation transmissivity is about 4×10^{-4} ft²/s. If we assume that the thickness of the affected aquifer is about 400 ft., this corresponds to a hydraulic conductivity of about 3×10^{-5} cm/s, which is mid-range in the expected hydraulic conductivities of fractured metamorphic rocks (Freeze and Cherry, 1979). Thus, a reasonable set of hydraulic parameters appears to be consistent with the slope calculated from the data. Note that this is not an attempt to derive actual hydraulic parameters; this calculation is merely a check that it is physically probable that the calculated slope can occur.

Timing and Quantity of Overflow

The relations derived in the previous sections can be used to estimate the time frame in which the mine workings will fill to the overflow point as well as the potential discharge volume. The average precipitation/runoff condition of 79.5 inches per year is assumed and the water elevation versus inflow relation is applied to calculate when the mine will fill to the overflow point. Then the average inflow at that elevation is used to calculate the average annual discharge from the adit due to the mine workings overflow. An additional assumption is that the inflow is distributed through the year approximately as it was in 2006, a year with close to average rain plus snowmelt.

If the mine starts completely pumped down to elevation 4020 in October and then the pumps are shut down, the calculations indicate that the mine will fill under average rain plus snowmelt conditions sometime in June of the following water year, 21 months later. Thereafter, the average rate of discharge to the service adit is anticipated to be about 1.5 cfs or 1,050 ac-ft per

Figure 9. Relation of Water Elevation to Adjusted Mine Inflow.



year. It is possible that the discharge will cease in August and September because of the lowered local water table in these months.

Annual and Seasonal Variations in Adit Discharge

Given the great variations from year to year and within any year in inflow to the mine, it is expected that discharges from the overflowing adit will vary both annually and seasonally. Given, the limited available data, estimates of natural variation in mine discharge rates cannot be made precisely and cannot accurately assess discharge rates during years in which precipitation is substantially higher or lower than during the years for which mine outflow data are available. The additional data necessary to refine this analysis can only be collected during unusually high or low precipitation years, and it cannot reasonably be expected that such opportunities will occur during the time frame in which this environmental analysis must be completed. Other methods of mine discharge estimation would require that data (ground water elevations, spring locations and flow rates, etc) have been collected prior to mining. Such data do not exist and cannot be obtained.

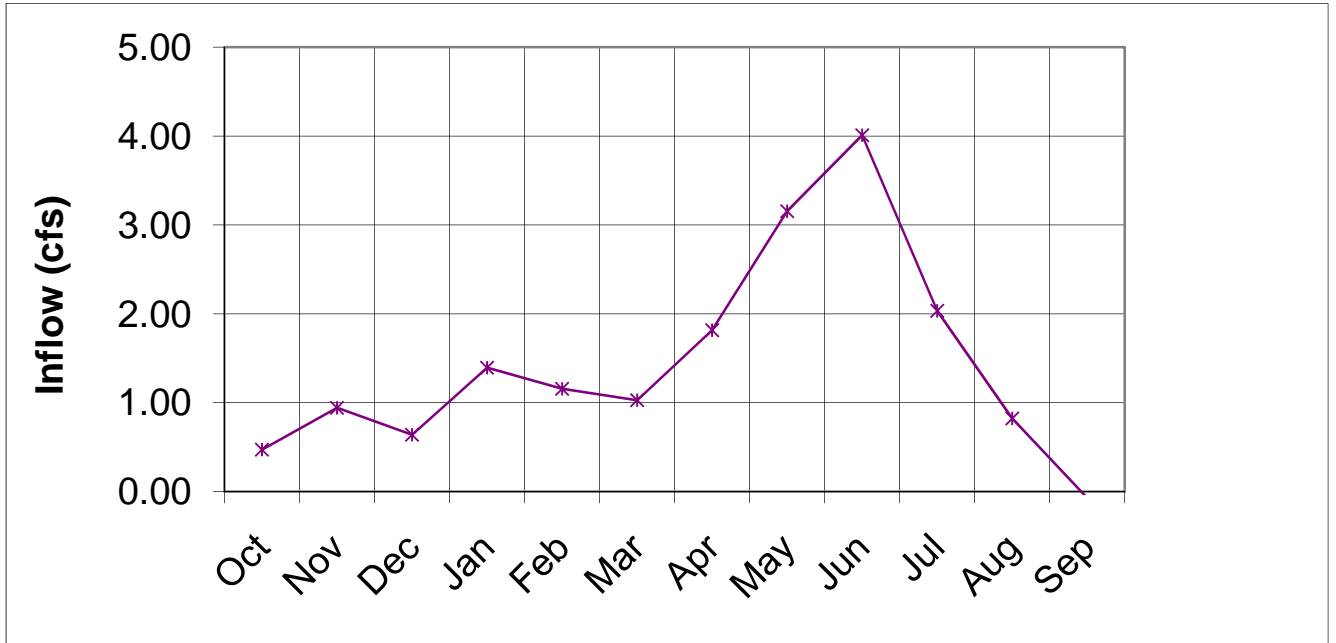
The relationship of water elevation to mine inflow (with no mine discharge) was determined based on four years of storage records. Within these years, 2004 was the wettest year (89.1 inches rain plus snowmelt) and 2005 was the driest (74.3 inches of rain plus snowmelt).

Comparing the calculated annual inflow for these years compared to the mean annual inflow shows that inflows were 39% higher than average in 2004 and 30% lower than average in 2005. If these ratios are assumed to apply to the inflow that would occur at the discharge elevation, the range of annual discharges varies from 1.0 to 2.0 cfs.

These two years, however, are not extreme precipitation years at the Poorman Creek SNOTEL site. In 2002 rain plus snow was about 104.1 inches and 2001 had 45.4 inches at this site, much greater and lesser amounts than were observed in 2004 and 2005. If we attempt to convert these more extreme amounts to inflows using the relationship derived for years 2004 through 2006 plus 2008, considerable extrapolation of the relation is required. In fact, the inflow predicted for 2001 is negative, indicating the inapplicability of this relation during drier years. Without storage-elevation data for more extreme years, it is not possible to predict with any certainty the behavior of adit discharge in extreme precipitation years. However, using the maximum rain plus snow value in 2002 (104.1 inches) from the ten-year period of record at Poorman Creek, an estimated annual adit discharge of 3.6 cfs is calculated. For lack of better data, this figure may approximate a 10-year return interval annual discharge.

Determining seasonal variations in predicted adit discharge is also problematic. As discussed previously, the attempts to develop robust, predictive models of monthly or seasonal flow were not successful. However, an estimation of the seasonal variation in predicted discharge for average conditions can be developed from the high storage level records of 2004 to 2006. Using the mean monthly values for these three years, and proportioning them by the ratio of the predicted annual average flow (1.45 cfs) to the annual flows for this period (1.93 cfs), the monthly flows presented in Figure 10 were calculated. The peak monthly flow is 4.0 cfs in June and the minimum is zero in September, when no discharge is expected. Inspection of the mine inflow data for individual months from 2004 to 2006 and correcting to average annual discharge from the mine adit indicates that a maximum discharge of 6.25 cfs could occur at the adit. In this 36 month period of record, four months would have had no discharge (always the August-September period). However, 2006, when the peak monthly discharge is predicted, was not an extremely high precipitation year so larger discharges than 6.25 cfs are possible.

Figure 10. Average expected monthly mine discharges based on 2004 to 2006 mine inflow data.



References

Freeze, R. A. and J. A. Cherry, 1979. Groundwater, Sections 2.3, 2.5 and 8.3. Prentice-Hall, Englewood Cliffs, N.J.

Genesis, Inc., 2004. Mine Flooding Report, Troy Mine.

Appendix C

Mine Water Management Analysis



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Technical Memorandum

*To: Emily Corsi, Herb Rolfes - DEQ
Bobbie Lacklen, John McKay - KNF*

From: Bill Bucher, Kim Chase - CDM

Date: December 3, 2010

Subject: Mine Water Management Analysis - Troy Mine

CDM has been retained by the Montana Department of Environmental Quality (DEQ) to analyze proposed mine water management alternatives for the Troy Mine. The purpose of this memorandum is to present the results of analysis of water management alternatives to be considered in the Draft Environmental Assessment (EA) for closure of the Troy Mine. This memorandum only addresses the physical management of water; water chemistry issues are being addressed separately.

The Troy Mine is an underground copper and silver mine located south of the town of Troy in Lincoln County, MT. ASARCO began operating the mine in 1982 and halted production in 1993 due to low metals prices. The mine was sold to Revett Silver Company and production resumed under Genesis, a subsidiary of Revett, in 2005. In 2006, Genesis submitted a Revised Reclamation Plan in support of this EA. This memorandum analyzes water management alternatives presented in the EA.

It has previously been shown in the *Mine Water Balance Analysis – Troy Mine* Technical Memorandum (CDM, 2010) that the Troy Mine is expected to discharge mine water after closure. The three closure alternatives to be evaluated in the draft EA consider different water management scenarios:

- Alternative 1 is the original closure proposal from the 1978 reclamation Plan. It would close the adits with non-hydraulic plugs, which would allow the mine water to discharge from the portals. Mine water would then infiltrate to the groundwater system. The glacial till near the mill site probably is not sufficiently permeable to allow infiltration of the expected quantity of water, and there is likely a substantial risk that saturation of this material could trigger landslides and slumping of material into the creek, causing further pollution. In addition there would be a direct discharge to Stanley Creek from the adit that

would not meet State water quality standards and would not be authorized. Therefore, this alternative is not being considered further in this water management analysis.

- Alternative 2 is the revised reclamation plan of March 2006 proposed by Genesis. In this alternative, the conveyor and service adits would be closed with non-hydraulic plugs, and the mine water would be carried by the existing tailings and reclaim water lines to the decant ponds. At the ponds, the water would infiltrate and evaporate, and metals would be adequately attenuated by the underlying soils to preserve ground and surface water quality in the area.
- Alternative 3 is the agency mitigated alternative. In this alternative, the both the Service and Conveyor adits would be closed with partial backfill only and the mine water would be captured inside the entrances and transported to the decant ponds for treatment and disposal. The existing tailings lines would be removed and replaced with a new, buried line with at least 6.9 cubic feet per second (cfs) capacity. If there is insufficient capacity in the new transport line or a leak or other upset, part or all of the flow would be diverted automatically to the reclaim water line that is presently in place. Sensors installed in the line would detect these conditions and Supervisory Control and Data Acquisition (SCADA) instrumentation would automatically open and close appropriate valves and alert maintenance personnel. The decant ponds would be bermed to prevent the entry of storm water. They will be constructed about 10 feet deep and lined with gravel to prevent the growth of aquatic vegetation.

The analysis of water management for a discharging mine in Alternatives 2 and 3 is the main focus of this memorandum.

Management of Mine Discharge

Mine water discharging from the service and conveyor adits at Troy Mine will be transported through a system of pipelines to the existing decant ponds in the tailings pond area under Alternatives 2 and 3. The locations of these facilities are shown in Figure 1. The question to be answered through a water balance analysis is whether the decant ponds will eventually fill and overtop or if infiltration and evaporation are sufficient to maintain the pond level below this elevation. The following information is needed to develop a water balance for the decant ponds at the tailings impoundment:

- Discharge from the mine adits
- Pond Area
- Pond volume and stage-volume relationship
- Evaporation and precipitation information near the pond

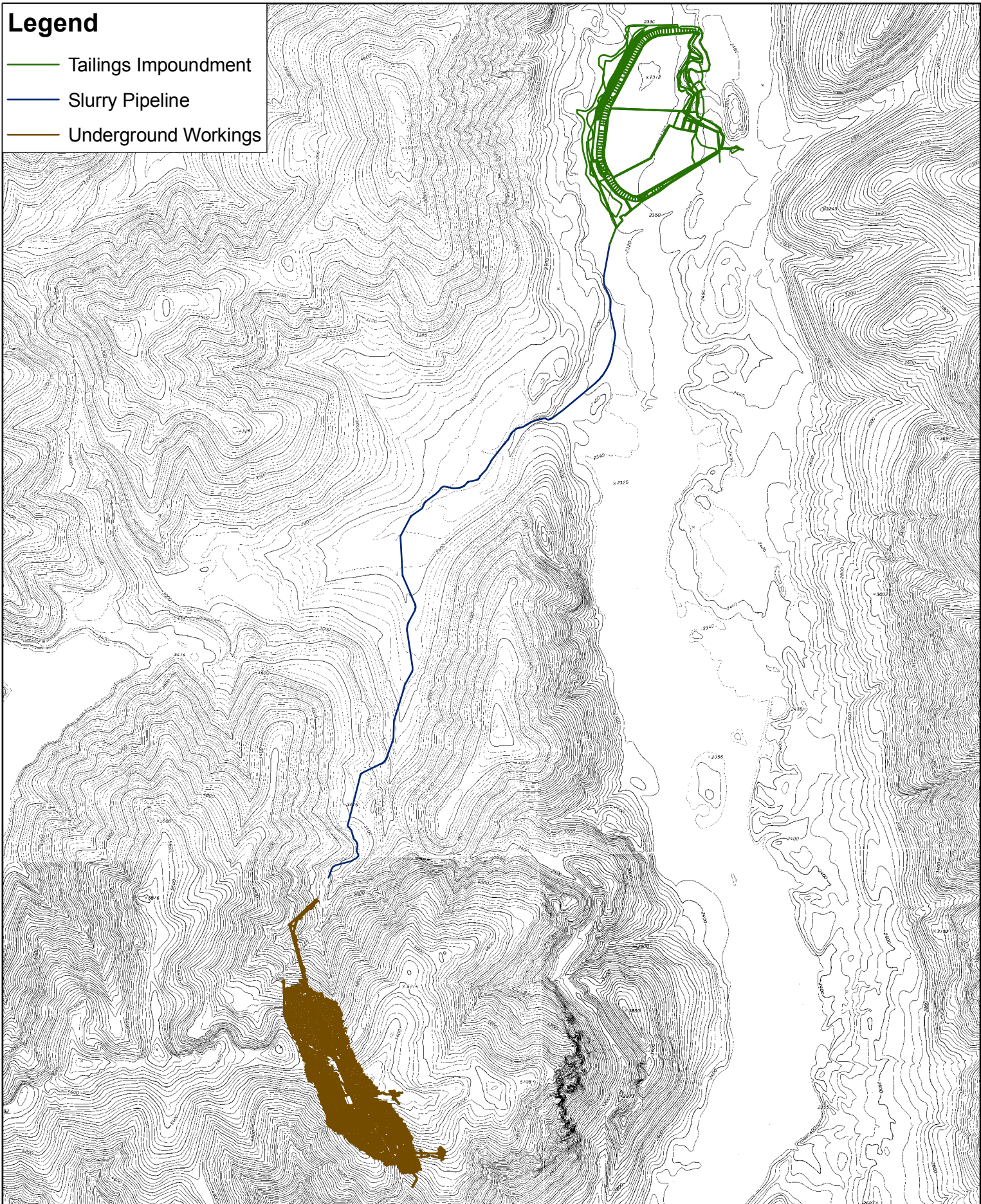


Figure 1: Location of
Underground Workings, Pipeline,
and Tailings Impoundment

0 0.2 0.4 0.8 1.2
Miles



■ Infiltration rate below the pond

For a detailed water balance analysis, a ten-year period with monthly time steps was selected to provide sufficient time to allow a general balance of inputs and outputs and to capture extreme conditions that would correspond roughly to a ten-year recurrence interval. The selected time period was the period of record for the Poorman Creek SNOTEL site, where the available record runs from 1999 through 2008. The Poorman Creek period of record is important to this analysis because the estimated mine discharges are tied to precipitation records at Poorman Creek in the mine water balance memorandum (CDM 2010). Monthly values for the mine discharge and climatic data were developed for the period 1999 -2009 as input to the pond water balance model.

Mine Discharge Data

In the *Mine Water Balance Analysis – Troy Mine* Technical Memorandum, CDM developed an annual relationship between rain plus snowmelt and mine discharge. This relationship demonstrated that the mine would generally discharge water after closure and quantified the annual amount of discharge. Although no robust predictive model of monthly discharge could be developed, evaluation of mine inflows on a monthly basis during a period of high mine water level resulted in an approximate distribution of expected mine discharge through a typical year. To produce the monthly data set, the predictive model was used to calculate estimated annual mine discharge based on precipitation records at the Poorman Creek SNOTEL. The average monthly distribution of mine discharge was then estimated by adding the typical distribution to the difference between the annual distribution for a particular year and the average annual discharge:

$$Q_{Myrx} = Q_{Ayrx} - Q_{Aave} + Q_{Mave}$$

Where: Q_{Myrx} is the calculated monthly discharge for year x

Q_{Ayrx} is the predicted annual flow for year x

Q_{Aave} is the predicted average annual flow

Q_{Mave} is the typical monthly discharge.

Table 1 shows the calculation of the predicted annual discharge based on precipitation at the Poorman Creek SNOTEL site and Table 2 shows the estimated monthly mine discharges for the period of record 1999-2008. The equation used to calculate the monthly discharges occasionally results in negative discharges; these discharges are set to zero.

Table 1. Predicted Annual Inflow and Discharge (Q_{Ayrx}) Based on Inflow Analysis Relations

Year	Rain+ Snowmelt (in)	Predicted Inflow (ac-ft)	Predicted Inflow (cfs)	Adjusted Inflow (cfs)*	Estimated Discharge (cfs)
1999	76.0	1118	1.54	1.54	1.14
2000	84.4	1837	2.54	2.54	1.87
2001	45.4	-1499	-2.07	0.76	0.56
2002	104.1	3522	4.86	4.86	3.59
2003	69.4	554	0.76	0.76	0.56
2004	89.1	2239	3.09	3.09	2.28
2005	74.3	973	1.34	1.34	0.99
2006	78.1	1298	1.79	1.79	1.32
2007	79.0	1375	1.90	1.90	1.40
2008	87.9	2136	2.95	2.95	2.18

*Because the inflow prediction equation results in a negative mine inflow for 2001, the mine inflow is conservatively set to the next lowest inflow (2003).

Table 2. Estimated Monthly Discharges at Mine Adit Elevation 4225 ft.

Month	Estimated Monthly Discharge (cfs)									
	1999	2000	2001	2002	2003	2004	2005	2006	2007	2008
Oct	0.17	0.90	0.00	2.62	0.00	1.31	0.02	0.35	0.43	1.21
Nov	0.64	1.37	0.06	3.09	0.06	1.78	0.49	0.82	0.90	1.68
Dec	0.33	1.07	0.00	2.79	0.00	1.48	0.19	0.52	0.60	1.37
Jan	1.09	1.82	0.51	3.54	0.51	2.23	0.94	1.27	1.35	2.13
Feb	0.85	1.58	0.28	3.30	0.28	1.99	0.70	1.03	1.11	1.89
Mar	0.72	1.46	0.15	3.17	0.15	1.87	0.58	0.91	0.99	1.76
Apr	1.51	2.24	0.94	3.96	0.94	2.65	1.36	1.69	1.77	2.55
May	2.85	3.58	2.27	5.30	2.27	3.99	2.70	3.03	3.11	3.89
Jun	3.70	4.44	3.13	6.16	3.13	4.85	3.56	3.89	3.97	4.74
Jul	1.73	2.46	1.15	4.18	1.15	2.87	1.58	1.91	1.99	2.77
Aug	0.52	1.25	0.00	2.97	0.00	1.66	0.37	0.70	0.78	1.56
Sep	0.00	0.36	0.00	2.07	0.00	0.77	0.00	0.00	0.00	0.66

The estimated monthly discharges presented in Table 2 do not account for the water that originates in that portion of the service adit below the high point of 4,225 ft. or a corresponding high point in the conveyor adit. These adit discharges have been measured semi-annually by the mine operator from July 2005 to February 2008 using weirs. The flows vary relatively little with the average February flow of the two adits being 0.67 cfs and the average July flow being 0.74 cfs. The 0.67 cfs flow was applied to the months of August through March, and the

0.74 cfs flow was applied to the months of April through July, resulting in the total estimated mine discharges shown in Table 3.

Table 3. Estimated Monthly Total Mine Discharges (QM_{yrx})

Month	Estimated Total Mine Discharge (cfs)									
	1999	2000	2001	2002	2003	2004	2005	2006	2007	2008
Oct	0.84	1.57	0.67	3.29	0.67	1.98	0.69	1.02	1.10	1.88
Nov	1.31	2.04	0.73	3.76	0.73	2.45	1.16	1.49	1.57	2.35
Dec	1.00	1.74	0.67	3.46	0.67	2.15	0.86	1.19	1.27	2.04
Jan	1.76	2.49	1.18	4.21	1.18	2.90	1.61	1.94	2.02	2.80
Feb	1.52	2.25	0.95	3.97	0.95	2.66	1.37	1.70	1.78	2.56
Mar	1.39	2.13	0.82	3.84	0.82	2.54	1.25	1.58	1.66	2.43
Apr	2.25	2.98	1.68	4.70	1.68	3.39	2.10	2.43	2.51	3.29
May	3.59	4.32	3.01	6.04	3.01	4.73	3.44	3.77	3.85	4.63
Jun	4.44	5.18	3.87	6.90	3.87	5.59	4.30	4.63	4.71	5.48
Jul	2.47	3.20	1.89	4.92	1.89	3.61	2.32	2.65	2.73	3.51
Aug	1.19	1.92	0.67	3.64	0.67	2.33	1.04	1.37	1.45	2.23
Sep	0.67	1.03	0.67	2.74	0.67	1.44	0.67	0.67	0.67	1.33

Climatic Data at the Tailings Pond

Precipitation data were taken from the Troy, MT weather station 248390 (Western Regional Climate Center 2009). This station is located 9.5 miles north of the tailings facility at an elevation of 1929 ft. It has 48 years of record. The elevation of the tailings ponds is about 2,400 feet, somewhat higher than Troy, so increased precipitation might be expected at this site. However, without local information, the Troy weather station data cannot be reliably adjusted to the site.

Five years of pan evaporation data, 1963-1967, are available from the tailings facility (Genesis, Inc. 2008). To estimate the evaporation amounts for the period 1999-2008, the record was extended using data from Station 4328 at Hungry Horse Dam near Columbia Falls, MT. This station is located at an elevation of 3,160 feet. Fifty-eight years of pan evaporation data, 1948 to 2007, are available from the National Climatic Data Center. The data from the tailings facility were compared to the evaporation values at Hungry Horse Dam and a correction factor was created. Before using the corrected data, a coefficient of 0.75 was applied to the data to adjust for the increased energy and therefore increased evaporation experienced in an evaporation pan as opposed to a lake or pond (Haan et al, 1994).

Average annual precipitation at the Troy, MT weather station is 24.5 inches for the period 1960-2009. November and December are the wettest months and July and August are the driest. Average annual adjusted evaporation is 15.7 inches. The greatest evaporation is generally experienced in August.

In most years, precipitation will exceed evaporation. Therefore, if the area of disposal ponds is increased, any accelerated loss of water due to evaporation will more than be compensated for by increased precipitation, and larger pond areas would not equate to greater disposal of water.

Estimation of Infiltration Rate

There are several sources of information on infiltration rates in the vicinity of the tailings ponds including estimates of seepage from the ponds as well as measurements of the hydraulic conductivity of the underlying materials in the vicinity of the ponds. Seepage was estimated in the Operating Plan (Asarco, 1976) to range between 292 and 803 gallons per minute (gpm), whereas the Draft Environmental Impact Statement (USFS-DSL, 1978) estimated 1,170 gpm. The Tailing Impoundment Design and Operation report (Pfahl, 1989) estimated seepage at 845 gpm based on a simplified water balance. Unfortunately, insufficient information is provided to determine the area over which seepage calculations occurred. Without this information, it is not possible to calculate infiltration on a unit area basis. In Alternatives 2 and 3, the current decant pond would be used to dispose of the mine discharge. This may be a reasonable solution if the bottom of the decant pond is not sealed by slimes. Alternatively, a new pond could be excavated directly in native materials and used to dispose of the discharge water. An estimate of infiltration rate based on hydraulic properties of the native materials is appropriate in the analysis of long-term water disposal options.

The Interim Report of Findings Hydrologic and Hydrogeologic Assessment prepared by Summit et al (1996) offers the most comprehensive data for the materials underlying the tailings pond area. This report includes logs from ten test borings, seven monitoring wells and seven sand-points in the vicinity of the tailings ponds. Data from a hydraulic conductivity test and other information related to infiltration rates are presented. The Summit report measured a hydraulic conductivity of 8.4 feet per day using a well in the middle of the tailings pond and one just below the pond berm. The boring logs from these wells and other wells in the area indicated the unit measured was primarily fine sand. This hydraulic conductivity value is within the expected range for sandy materials.

In November of 1995, ASARCO personnel decreased the flow rate to the decant pond for seven days. During this time, Summit reported a 4.5 foot drop in water levels in the decant pond. The report does not state the amount that the flow rate was decreased or whether flow ceased altogether, so it is not possible to calculate an exact infiltration rate from this information. However, a 4.5 foot drop in water level over the area of the decant pond for a seven-day period equates to a rate of infiltration of 0.64 feet per day and provides a potential pond infiltration rate.

Further data on the infiltration rate at the decant ponds was obtained in 2001 during a tracer test conducted by Hydrometrics (2001). During a period of low mine water discharge, all discharge to the decant ponds was suspended and a 24-day salt tracer test was conducted.

During this period, the ponds dropped at a typical rate of about 0.4 feet per day, somewhat lower than the rate determined in the 1995 study. The infiltration rates used in our analyses vary from 8.4 feet per day for good infiltration conditions to a minimum of 0.4 feet per day.

Water Balance at the Tailings Pond

Precipitation on the surface of the tailings and mine water delivered to the decant ponds are the major input values for a water balance calculation at the tailings facility. The primary outputs are seepage and evaporation from the decant ponds. The general equation used to determine the change (Δ) in storage of water at the decant ponds or tailings impoundment is as follows:

$$\Delta \text{Storage} = \text{Precipitation} + \text{Mine Discharge} - (\text{Evaporation} + \text{Seepage})$$

Because information on the size of the proposed disposal facilities is not included in the proposed Alternatives, the procedure followed in this analysis was to set the size of the proposed disposal facility based on the other inputs to the model. For simplicity, it was assumed that an infiltration pond of a certain area had a capacity of 10 feet of water and vertical sides. This approximate design would need to be revised based on site conditions and other factors during implementation of the pond disposal alternative. In the model input, the pond area was varied to determine at what size the pond would overflow at some time under the 10-year climatic record.

An initial model run was undertaken with the assumption that the hydraulic conductivity of the underlying materials at the site was 8.4 feet/day, a value derived from aquifer tests by Summit et al (1996). A review of the well logs in the vicinity of the east edge of the tailings pond indicates that ground water would be located about 50 feet below the pond bottom. This information was inserted into Darcy's equation to translate the hydraulic conductivity into an infiltration rate:

$$v = k i$$

where v = the Darcy velocity
 k = hydraulic conductivity (ft/day) and
 i = the gradient.

To calculate the gradient, the following terms are defined:

h_A = the height of the pond bottom above the aquifer in feet, and
 h_W = the depth of water in the pond in feet.

For vertical flow, the gradient, i , is the total head ($h_A + h_W$) divided by the length of the media through which it passes (h_A) or, for 50 foot separation from the aquifer:

$$i = (h_A + h_w) / h_A$$

$$i = (50 + h_w) / 50$$

Referring to the spreadsheet model "Water Management" and work sheet "1.6-acre Pond" on the attached compact disc, for a pond with a maximum depth of 10 feet, the infiltration rate will vary from 8.4 to 10.1 ft/ft² - day. Because of the depth to the aquifer is large, the amount of water in a pond with a maximum 10-foot depth does not have a great effect on the magnitude of the infiltration rate. Running the model with various-sized ponds and the precipitation rates for the period of record shows that a 1.6-acre pond that is 10 feet deep would not overtop. Note that the model indicates the pond only fills with water during the wettest portion of 2002, which was the wettest year in the period of record. In part, this is because the infiltration rate modeled (at least 403 acre-feet for a 30-day month) is sufficiently high to accommodate the mine discharge rate, which only averaged 292 acre-feet. Neither precipitation nor evaporation exceed one acre-foot per month, and are almost inconsequential in this calculation. Only when mine inflows exceed those of June 2002 does the model predict the 1.6-acre pond would fill.

It is likely that the high infiltration rate used in the preceding calculation will not be realistic, at least in the long term, due to siltation or vegetation accumulating in the pond. Although pond maintenance should minimize these effects, it is still probable that some decrease in performance will be experienced over time. Therefore, a spreadsheet calculation was also developed to size a pond for the minimum infiltration rate of 0.4 ft per day reported in 2001 by Hydrometrics. Referring to the spreadsheet model "Water Management" and worksheet "23-acre Pond" on the attached compact disc, at this lower infiltration rate, a 23-acre pond is required to prevent overtopping a 10-foot deep pond. For this pond configuration, the model indicates the pond would hold water during early summer in four of the wet years, 2000, 2002, 2004 and 2008. Although the effects of precipitation and evaporation are somewhat more important with a larger surface area pond, they are still generally less than 10 acre-feet per month with minor effects on the calculation.

Implementation

The calculations presented in the previous section suggest that disposal of water through use of a pond is feasible given the transmissive quality of the underlying soils. Even a 23-acre pond would be quite feasible to construct, although it would have significant maintenance requirements. Because the existing 4-acre decant pond system has successfully infiltrated excess mine water during operations, it is expected that this same facility can continue to handle mine discharge water after mine closure. Normally these ponds infiltrate about 2,500 gpm (5.6 cfs) and have infiltrated 3,000 gpm (6.7 cfs), which is nearly the maximum design flow assumed in this memorandum (6.9 cfs) (personal communication, Doug Parker, Hydrometrics, June 16, 2010). If the existing ponds are maintained with a depth of about 10 feet, are gravel lined to increase infiltration, and vegetation growth is controlled, the infiltration rate

should be maintained. The infiltration ponds would need a berm constructed around them to prevent storm water from flowing into them.

Summary

This memorandum presents an analysis of the expected feasibility of mine water disposal in the vicinity of the tailings ponds. Based on previous work, the mine will discharge an average of 2.1 cfs from its adits after pumping ceases and water reaches the overflow elevation of the adits. Flows are expected to range from 0.67 cfs to 6.90 cfs with the largest discharges generally expected in the months of May and June in relatively wet years.

A spreadsheet water balance model was developed for the water disposal pond and used to evaluate 10 years of discharge and climatic data including precipitation, evaporation, and infiltration. The infiltration rates modeled were based on an aquifer test performed at the tailings pond and an observation of infiltration rates at the decant pond. It is not known how representative these data are of the soil conditions in the area. A range of pond sizes, depths, and infiltration rates was modeled using actual discharge plus precipitation rates. The various pond designs were sufficient to contain all expected inflows during the 10-year evaluation period. An important conclusion of this analysis is that the pond size is almost entirely dependent on the magnitude of the mine water discharge and infiltration rate of the subsurface, and relatively unaffected by precipitation and evaporation.

Water disposal ponds sized from 1.6 to 23 acres are feasible to build and maintain. The current decant ponds, which are about 4 acres in size fall within the modeled range of sizes. These ponds have successfully infiltrated the mine discharge for the period of operation of the mine and, with proper maintenance, should continue to infiltrate the discharges expected post closure. To prevent plugging of the ponds and ensure they are not overwhelmed with storm water runoff, berms should be constructed around the ponds to exclude entry of storm water runoff from the surrounding area.

References

CDM, 2010. Mine Water Balance Analysis - Troy Mine. Technical Memorandum to Emily Corsi and Herb Rolfes, Montana Department of Environmental Quality dated April 12, 2010.

Haan, C.T., B.J. Barfield, J.C. Hayes. 1994. Design Hydrology and Sedimentology for Small Catchments. Academic Press. San Diego, CA.

Western Regional Climate Center. Troy Montana Climate Data. 2 Sept 2009.
<<http://www.wrcc.dri.edu/cgi-bin/cliMAIN.pl?mttroy>>

Genesis, Inc. 2008. Troy Mine Precipitation and Evaporation Data. Unpublished Spreadsheet.

Genesis, Inc. 2006. Troy Mine Revised Reclamation Plan. Troy, MT.

Hydrometrics, Inc., 2001. Mine Water Plume Location and Identification, Phase 1 results and Phase 2, & 3 Workplan, Sterling Troy Mine. Prepared for Genesis Inc./Sterling Mining Company, October.

Summit Envirosolutions and McCulley, Frick, and Gilman, Inc. 1996. Interim Report of Findings Hydrologic and Hydrogeologic Assessment, ASARCO Troy Mine. Troy, Montana.

Appendix D
Troy Mine Metals Attenuation Study-
Secondary Processes

Department of Environmental Quality

Troy Mine

Copper Attenuation Study – Secondary Processes

July 8, 2010

Final

Contents

Section 1 Introduction

Section 2 Modeling Methodology

2.1	Water Analyses Internal Consistency Evaluation.....	2
2.2	PHREEQC Thermodynamic Database	5
2.3	Oxidation-Reduction Potential (ORP)	5

Section 3 Geochemical Modeling Results

3.1	Mineral Saturation States.....	7
3.2	Mixing Scenarios.....	10

Section 4 Bench-Scale Tests

4.1	Sample Collection.....	11
4.2	Jar Test Procedures	12
4.3	Results	13

Section 5 Conclusions

Section 6 References

Attachment 1 Analytical Reports

Attachment 2 Photographs

Section 1

Introduction

The fate of the copper, antimony, lead, uranium, arsenic, cadmium, and other metals within the waters discharged to the decant pond is of great interest, as one of the alternatives is to discharge the mine water to the decant pond and to allow natural attenuation mechanisms to remove the metals from solution. Currently (during active operation of the mine), Genesis is using the decant ponds to settle out suspended sediment from the tailings slurry which is piped from the mill. The decant water is then pumped to the mill circuit via a return line. The system is currently semi-closed in that some, but not all of the water is recirculated. As the decant ponds are not lined, some of the decant pond water seeps through the bottom of the pond into the underlying aquifer.

In 2001 Hydrometrics conducted a tracer study in which sodium chloride (salt) was added to the decant ponds (Hydrometrics, 2001). The migration of the salt into the underlying and downgradient groundwater was then monitored by measuring the conductivity within the existing wells. The report concluded that while the deep wells and decant ponds appear to be hydraulically connected, the response time is long, suggesting that the decant water is transported predominantly within shallow sand and gravels above the zone of the then existing wells. Shallow monitoring wells were installed as part of the study.

While the decant pond water appeared to be traveling downgradient of the ponds within the shallow alluvial system, the copper concentrations in the shallow wells were either below the laboratory reporting limit or were significantly lower than in the decant pond water, suggesting that the copper was being removed via some geochemical process. In 2004, a study was conducted in order to determine the fate and transport of copper beneath the decant ponds (Land and Water, 2004). The soil beneath the decant ponds was analyzed for total copper as well as electron microprobe and sequential extraction analyses. The results indicated that copper was attenuated within the upper foot of soil via the precipitation of secondary copper phases (carbonates, silicates, and oxides) as well as adsorption onto organic matter.

Because it is not certain whether the attenuation mechanism(s) occurring within the sediments below the decant pond would have a finite duration or would perpetually precipitate copper within or immediately beneath the pond, the agencies determined that testing should be conducted to evaluate secondary attenuation processes which may occur if copper migrates beyond the decant pond and mixes with downgradient groundwater. The investigation was expanded to include other metals in addition to copper.

The evaluation consisted of two parts;

- 1) Initial geochemical modeling using analyses of the groundwater downgradient of the decant ponds, decant pond water, and mine water to determine if iron hydroxide is predicted to precipitate. Mixing of the decant pond water with the groundwater simulates the current situation during active mining, where mill water would mix with groundwater. In order to simulate a second, post closure scenario, mine water was mixed with groundwater.
- 2) Jar testing to determine the fraction of metals removed during precipitation of iron hydroxide.

Section 2

Modeling Methodology

A geochemical model, PHREEQC (Parkhurst and Appelo, 1999), was used to predict the geochemical reactions that will occur in response to mixing mine water (decant water) with the native groundwater downgradient of the toe ponds. PHREEQC is a thermodynamic equilibrium program designed to model chemical speciation in aqueous solutions, determine the saturation states of solutions with minerals and gases, and predict the results of various reactions, such as dissolution of minerals and oxidation.

The modeling shows which phases or minerals are predicted to become supersaturated (if any) as a result of mixing of the waters. Minerals that are supersaturated would eventually precipitate (assuming the mineral forms at low temperature) as equilibrium is established.

2.1 Water Analyses Internal Consistency Evaluation

The analyses were checked for internal consistency using both charge balance and mass balance relationships. Any solution containing a mixture of positively and negatively charged ions must have an overall charge balance of zero (the sum of the negative ions must equal the sum of the positive ions). If an analysis reported by a lab does not have a calculated charge balance near zero, then it indicates either that one or more errors have occurred in the analysis and/or that ions are present that were not analyzed for.

By evaluating both the mass balance and charge balance, conclusions can be drawn about the accuracy and completeness of the analyses. The possible mass balance and charge balance combinations and the corresponding interpretations are shown in **Table 2-1**.

The charge balance was calculated from the lab analyses as follows:

$$(\Sigma(\text{Cations} \times \text{charge}) - \Sigma(\text{Anions} \times \text{charge})) / (\Sigma(\text{Cations} \times \text{charge}) + \Sigma(\text{Anions} \times \text{charge})) \times 100\% \quad (2-2)$$

Where “cations” refers to the molar concentration of positively-charged ions (moles/L) and “anions” to the molar concentration of negatively-charged ions.

Mass balance is the comparison of the total mass of constituents in solution measured by an analysis called “total dissolved solids” (TDS-Measured) compared to a summation of all the constituents measured in each individual analysis (TDS-Calculated). TDS is determined by evaporating a measured volume of the solution to dryness and weighing the residual salts. Again, if the measured TDS is different from the calculated TDS, either some parameter is missing or there is an error in the analysis.

The mass balance was calculated using the following relationship:

$$\text{TDS-Calculated} / \text{TDS-Measured} \quad (2-3)$$

TDS was calculated by summing the concentrations of all species in mg/L. Adjustments were made in cases where the species that would be formed upon evaporation were in a different form than that provided by the laboratory. For instance, silicon reported as “Si” (atomic mass = 28.09 g/mole) was converted to “SiO₂” (atomic mass = 60.09 g/mole) using the factor 2.14 (60.09g/mole / 28.09 g/mole = 2.14). In addition, the bicarbonate concentration was multiplied by a factor of 0.49 to account for loss of carbon dioxide gas during evaporation.

Table 2-1 – Interpretation of Charge and Mass Balance Results

TDS-Calculated/ TDS-Measured	Charge Balance	Interpretation ¹
>1	Positive	Cations are over-reported
>1	Negative	Anions are over-reported
<1	Negative	Cations are under-reported and/or one or more important cations were not analyzed
<1	Positive	Anions are under-reported and/or one or more important anions were not analyzed

1. Note that the interpretation represents the least complex explanation. In some cases, multiple problems with an analysis may have caused the inconsistencies.

The acceptability criteria for the internal consistency checks are reported in **Table 2-2**.

Table 2-2 Chemical Analysis Acceptability Criteria

Test	Criteria
Charge Balance	+/-10%
Calculated TDS/Measured TDS	0.9-1.1

The wells used to represent the downgradient groundwater were selected by looking at the chemistry, particularly the iron content within the water (see Section 3). A summary of the waters used in the modeling is provided in **Table 2-3**

Table 2-3 – Summary of Water Samples Used for PHREEQC Geochemical Modeling – EPA Column Experiment.

Sample ID	Date Sampled	Notes
Decant 1	5/14/09	Decant pond
MW-01-15	5/13/09 6/08/09 ¹	Well along southern edge of the decant pond. Screened at a depth of approximately 40 feet below ground surface (bgs).
MW 95-4 ²	5/12/09	Well located just north of the southernmost toe pond (in between ponds). Screened at a depth of approximately 15 feet bgs.
MW 95-5 ²	5/12/09	Located downgradient of the toe ponds and screened at a depth of approximately 10 feet bgs.

1. Resampled for dissolved copper due to a suspected laboratory error.
2. Completed at or near the groundwater table within a sand and gravel zone.

The internal consistency checks of the data along with the chemical data used in the modeling are shown in **Table 2-4**.

Table 2-4 – Summary of analytical data Used in the PHREEQC Modeling (mg/L dissolved unless noted otherwise)

Parameter	Decant 1	MW-01-15	MW 95-4	MW 95-5
Date Sampled	5/14/09	6/08/09	5/12/09	5/12/09
Temperature (°C)	9.5	14.1	8	8.8
pH-Lab	7.6	7.5	6.5	5.9
pH-Field	7.45	7.3	7.3	6.04
Total Alkalinity (as CaCO ₃)	82	85	77	31
Chloride	4	4	1	<1
Sulfate (as SO ₄)	33	41	111	3
Sulfide	0.04	<0.04	0.22	<0.04
Ammonia (as N)	6.0	3.64	3.72	<0.05
Nitrite + Nitrate (as N)	17.6	14.5	0.02	0.09
Phosphorous (as P)	0.134	0.025	0.014	0.016
Arsenic	0.002	<0.001	0.002	<0.001
Antimony	0.046	0.029	<0.003	<0.003
Barium	0.097	0.107	0.114	0.016
Cadmium	0.00089	0.00136	0.00093	0.00089
Calcium	17	23	29	6
Copper	0.028	0.009 ²	<0.001	0.002
Iron	0.06	<0.05	7.84	0.18
Lead	0.0026	<0.0005	<0.0005	<0.0005
Magnesium	4	5	8	1
Manganese	0.554	0.467	2.43	0.457
Potassium	27	23	17	5
Silica	8.5	11.4	46.4	60.8
Sodium	31	32	13	5
Uranium	0.0029	0.0018	<0.0003	<0.0003
Conductivity (µmhos/cm)	427	405	445	82
TDS Measured	305	247	292	97
TDS Calculated	259	259	286	100
Charge Balance (%) ¹	-0.44	1.59	-1.93	2.71
TDS-C/TDS-M	0.85	1.05	0.98	1.03

1. Calculated using PHREEQC.
2. Re-sampled due to an apparent lab error.

The results indicate that all of the waters meet the chemical analysis acceptability criteria with the exception of Decant 1, which has a mass balance which is out of range. As the charge balance is very good, the error could simply be an over-reporting of the TDS analysis, which would have no impact on the modeling. There may also be a neutral species that was not included within the analysis, but this would have minimal impact on the modeling

2.2 PHREEQC Thermodynamic Database

The PHREEQC database is provided as a starting point and often new phases need to be added to address specific modeling requirements. The thermodynamic database used by PHREEQC was supplemented to include several additional iron phases. The phases that were added and the thermodynamic data used are shown in **Table 2-5**.

Table 2-5 - Summary of Thermodynamic Data Added to the PHREEQC Database

Phase	Reaction	Log Ksp	ΔH_r° (kcal)
Troilite	$\text{Fe}^{+2}(\text{aq}) + \text{S}^{-2}(\text{aq}) \rightarrow \text{FeS (troilite)}$	19.1	-10.9
Pyrite	$\text{Fe}^{+2}(\text{aq}) + \text{S}^0(\text{aq}) + \text{S}^{-2}(\text{aq}) \rightarrow \text{FeS}_2(\text{pyrite})$	29.3	-27.6
Marcasite	$\text{Fe}^{+2}(\text{aq}) + \text{S}^0(\text{aq}) + \text{S}^{-2}(\text{aq}) \rightarrow \text{FeS}_2(\text{marcasite})$	28.6	-26.6
Mackinawite	$\text{Fe}^{+2}(\text{aq}) + \text{S}^{-2}(\text{aq}) \rightarrow \text{FeS (mackinawite)}$	17.6	-8.5
Mackinawite (alkaline) ¹	$\text{Fe}^{+2}(\text{aq}) + \text{HS}^{-}(\text{aq}) \rightarrow \text{FeS (mackinawite)} + \text{H}^{+}$	-3.5	-
Greigite	$\text{Fe}^{+2}(\text{aq}) + 2 \text{Fe}^{+3}(\text{aq}) + 4 \text{S}^{-2}(\text{aq}) \rightarrow \text{Fe}_3\text{S}_4(\text{greigite})$	90.4	-57.4
Magnetite	$3\text{Fe}^{+2} + 4\text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + 8\text{H}^{+} + 2\text{e}^{-}$	-45.8	-
Lepidocrocite	$\text{Fe}^{+2} + 2\text{H}_2\text{O} \rightarrow \gamma\text{-FeOOH} + 3\text{H}^{+} + \text{e}^{-}$	-16.7	-

1. "alkaline" refers to the solubility of the mackinawite phase which forms at pH>7 (Rickard, 2006)

More common phases, such as hematite were already included in the PHREEQC database and are not included in **Table 2-5**. Pyrite was in the PHREEQC database, but only the solubility constant (log Ksp) was provided. The log Ksp was recalculated, along with the enthalpy (ΔH_r°) to make sure the data were consistent.

2.3 Oxidation-Reduction Potential (ORP)

The oxidation-reduction potential (ORP) of the solution (also referred to as "Eh") is one of the most important parameters for geochemical modeling, but is also one of the most difficult to accurately measure.

Eh was calculated several ways for potential use in the modeling, including the following:

- From the platinum electrode measurements collected in the field
- From the sulfate/sulfide redox couple
- From the nitrite/ammonia redox couple

ORP Calculation Methods

Oxidation reduction potential (ORP) is used by the PHREEQC model to determine the relative proportions of the parameters that have more than one oxidation state in solution. For example iron can exist in solution as ferric iron (Fe^{+3}) or as ferrous iron (Fe^{+2}). The ORP can be directly measured using a platinum electrode, which has several problems, including interferences by sediment, certain ions, or dissolved organic matter. Lindberg and Runnels (1984) recommend evaluating the redox couples within a solution, such as sulfate/sulfide, ferrous iron/ferric iron, or nitrate/nitrite to determine the ORP and if the system is in redox equilibrium. Often, the Eh measured by the platinum electrode is a mixed potential resulting from multiple redox couples. The Eh can be calculated for each redox couple and compared. The system can be considered in redox equilibrium only when the Eh determined from each couple is the same. A summary of Eh values calculated using various methods along with dissolved oxygen (DO) values for comparison is presented in **Table 2-6**. The ferric/ferrous couple was not used due to the very low aqueous concentrations of ferric iron at low Eh and near-neutral pH. Obtaining ferric iron by difference (subtracting ferrous iron from total iron) could not be performed due to the low concentration of dissolved ferric iron in these systems (dissolved ferrous and total iron are essentially identical) and the error in the analytical techniques. Dissolved oxygen (DO) concentrations are listed in the table for comparison.

Table 2-6 Summary of Eh Values Calculated Using Different Methods

Parameter	Decant 1	MW-01-15	MW 95-4	MW 95-5		
Eh _{Pt Electrode} (mV)	397	351	272	397		
Eh _{SO4-2/S-2} (mV)	-200	>198 ²	-183	>-110 ²		
Eh _{NO2-/NH4+} (mV) ¹	347	349	336	>452 ³		
DO (mg/L)	8.90	4.35	0.18	1.77		

1. The assumption was made that nitrate+nitrite = nitrite
2. Using the detection limit of 0.04 mg/L for sulfide, so the Eh can be considered a minimum.
3. Using the detection limit of 0.05 mg/L for ammonia, so the Eh can be considered a minimum

Eh_{Pt electrode} = The Eh as determined using the platinum electrode

Eh_{SO4-2/S-2} = The Eh calculated from the sulfate/sulfide couple

Eh_{NO2-/NH4+} = The Eh calculated from the nitrite/ammonia couple

A comparison of the Eh values indicates that in general, the waters are not in redox equilibrium. For instance, well MW 95-4 contains both significant sulfide (a reduced form) and nitrite or nitrate (an oxidized form), which would not occur together under equilibrium conditions. The decant water contains significant concentrations of both ammonia and oxygen. Given enough time, the ammonia will be oxidized by the DO, producing nitrite and/or nitrate and water.

Section 3

Geochemical Modeling Results

3.1 Mineral Saturation States

The PHREEQC model uses a term called the saturation index (SI) to quantify the degree of saturation of a mineral. SI is defined as follows:

$$SI = \text{Log} (IAP/K_{sp}) \quad (3-1)$$

Where IAP is the ion activity product and K_{sp} is the solubility product constant for the phase in question. For phases at saturation, $IAP=K_{sp}$ and $SI = 0$. A negative SI indicates that the phase is unsaturated ($IAP<K_{sp}$) while a positive SI ($IAP>K_{sp}$) indicates the phase is supersaturated. In practice, a range of 0 ± 0.5 SI units is considered saturated due to uncertainties in analytical and thermodynamic data (Hem, 1970). In the present study, a value of 0 ± 1.0 SI units was used to allow for additional uncertainties in the redox measurements and Eh calculations.

The mineral saturation states for the unmixed waters are provided in **Table 3-1**.

Table 3-1 – Summary of Saturation Indices (SIs) Determined Using PHREEQC

Phase	MW 95-4	MW 95-5	MW-01-15	Decant 1
Barite (BaSO ₄)	0.85	-1.30	0.35	0.33
Poorly Ordered Mackinawite (FeS)	0.68	-2.93	-2.94	-3.21
Alkaline Mackinawite (FeS)	0.68	-	-3.13	-3.26
Amorphous Fe(OH) ₃	-	0.90	1.76¹	2.48²
Siderite (FeCO ₃)	0.62	-2.51	-3.01	-3.35
Vivianite (Fe ₃ (PO ₄) ₂ 8H ₂ O)	-0.31	-8.94	-10.96	-10.32
Rhodochrosite (MnCO ₃)	0.42	-1.78	-0.13	0.01
Amorphous Silica (SiO ₂)	-0.25	-0.14	-0.92	-1.00

1. Supersaturation likely due to using half the reporting limit for the iron concentration (0.025 mg/L)

2. Supersaturation likely due to uncertainties in the reported iron concentration (0.06 mg/L) so close to the reporting limit of 0.05 mg/L

Bold indicates supersaturated phases

Shaded cells represent phases at equilibrium ($SI=0\pm1$ SI unit)

Minerals that are at saturation ($SI = 0 \pm 1$) are in equilibrium and will not precipitate or dissolve unless the geochemical conditions are changed. Supersaturated solutions ($SI > 1$) are predicted to precipitate, while undersaturated phases ($SI < 1$) are predicted to dissolve (if present).

In addition to the minerals shown above, other minerals which were not included in the thermodynamic database are possible. For example, there is field evidence for the presence of the mineral delafossite (CuFeO₂), based on the work conducted by Land and Water (2004).

Ferrous Iron Minerals

Ferrous iron minerals have the capacity to coprecipitate copper. The high pH mackinawite phase investigated by Rickard (2006) which has been named “alkaline mackinawite” for the purposes of this report was only included for sites which had pH values greater than 7 (MW 95-4, MW-01-15, and Decant 1).

Well MW 95-4 is at saturation with respect to the iron monosulfide mackinawite, as well as siderite, and vivianite, all of which are ferrous iron (Fe^{+2}) minerals. The other waters, being more oxidized are undersaturated with respect to the ferrous iron minerals.

Amorphous Silica

Amorphous silica is at saturation for all of the waters analyzed. However, silica does not typically incorporate copper into precipitates except as inclusions of other phases, such as iron oxyhydroxides or sulfides.

Barite

Barite is at saturation in all of the waters except for well MW 95-5. Barite does not typically coprecipitate metals such as copper.

Rhodochrosite

Rhodochrosite is at saturation in all of the waters except for MW 95-5. Rhodochrosite likely does not incorporate copper into the crystal structure of the mineral to a great extent as divalent copper and divalent manganese have significantly different ionic radii (71 picometers [pm] for divalent copper vs 80 pm for divalent manganese). However, as the ionic radii are within 30% of each other limited substitution of copper for manganese in the mineral structure may be possible according to “Goldschmidt’s Rules.”

Amorphous $\text{Fe}(\text{OH})_3$

Amorphous $\text{Fe}(\text{OH})_3$ is calculated to be at saturation in well MW 95-5 and is supersaturated for well MW-01-15 and Decant 1. The supersaturation of MW-01-15 is likely due to the use of half the reporting limit for the iron concentration. For Decant 1, the iron concentration may be in error, as the value reported is close to the reporting limit, where instrument “noise” is usually significant. Instrument noise is the fluctuation in laboratory results as the limitations of the analytical method are approached. In this case, the error is minimal, as supersaturation or saturation of amorphous $\text{Fe}(\text{OH})_3$ are both indications that the mineral is present in the system already or will precipitate given enough time.

Another explanation for the supersaturation of amorphous $\text{Fe}(\text{OH})_3$ is that fine particulate iron is being included in the dissolved analysis. Studies have shown that very fine grained particulate matter (less than 0.45 microns in diameter) can pass through the filters. During processing (digestion) of the sample the fine particles dissolve and are included as part of the dissolved fraction of iron.

Comparison of Historic Iron Concentrations

As the presence of iron in the groundwater is important for the removal mechanism being investigated, a review of the historic iron concentrations within the groundwater was conducted. In May 2009, well MW 95-5, had surprisingly low iron compared to the March 2001 analysis where a concentration of 36 mg/L was measured. **Figure 1** shows that both wells MW 95-4 and MW 95-5 have had large variations in dissolved iron concentrations since 1995. The decrease in iron concentration in MW-95-5 from 36 mg/L to 0.08 mg/L in April 2003 and 0.18 mg/L in May 2009 was likely due to oxidation of the iron by decant water, infiltration, or dust control irrigation. It appears that the downgradient groundwater at times is influenced by oxygenated water of unknown origin that may be seasonal in nature or the result of changes in decant water influx (i.e. Mine Flooding study from September 2002 through February 2004) or the water balance of the groundwater system.

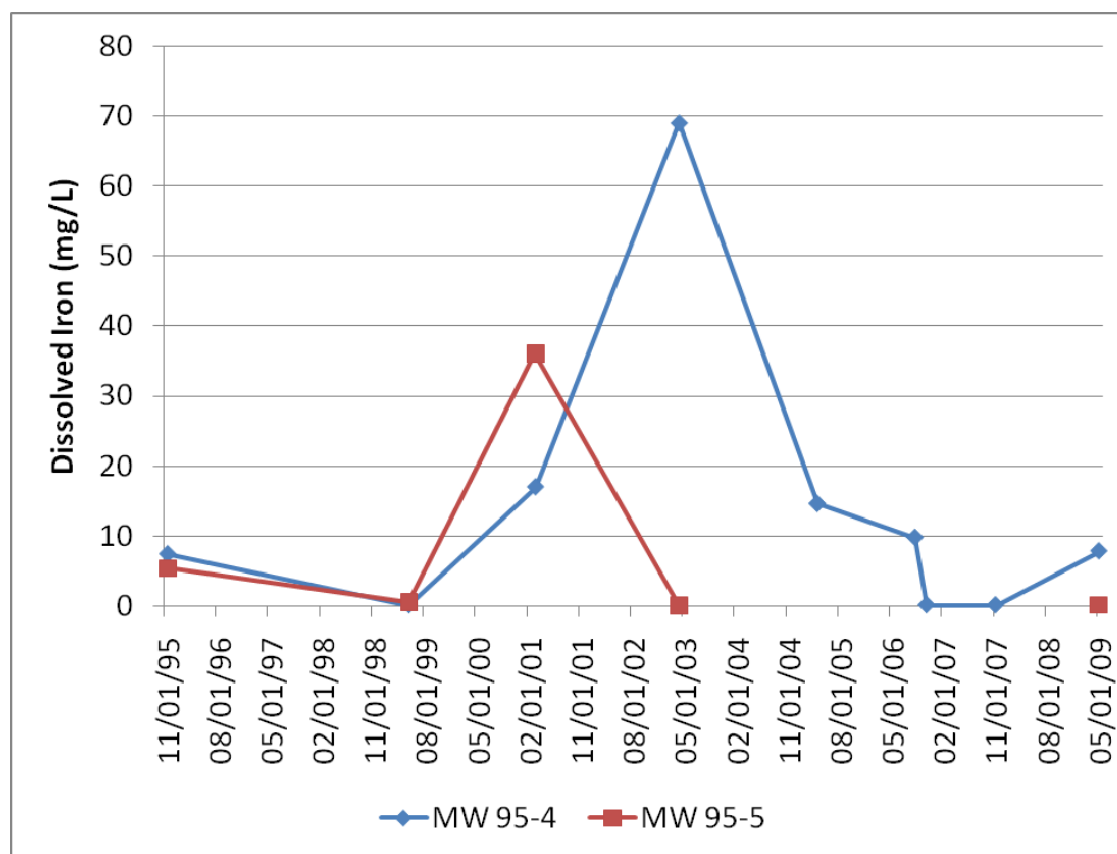


Figure 1 – Variations in dissolved iron concentrations with time for wells MW 95-4 and MW 95-5.

Iron concentrations within wells MW-1, MW-2, and MW-3 have also varied significantly, as shown in **Figure 2**. The inconsistency in the iron concentrations within the downgradient groundwaters may indicate that the secondary removal process is more effective at some times than others due possibly to seasonal processes.

3.2 Mixing Scenarios

Mixing scenarios were modeled using the waters from well MW-01-15 and well MW 95-4 in equal proportions. The decant water was not modeled, as it has the same measured Eh value and an even higher DO concentration than MW-01-15. The mixing ratio was set at 50% for each water as this is the optimal degree of mixing where both iron and DO concentrations are high. At lower fractions of MW 95-4 the iron concentration becomes lower, while at greater fractions the DO (and Eh) becomes lower. The water from well MW-01-15 is beneath the decant ponds and is the water that will eventually travel downgradient to mix with the natural groundwater.

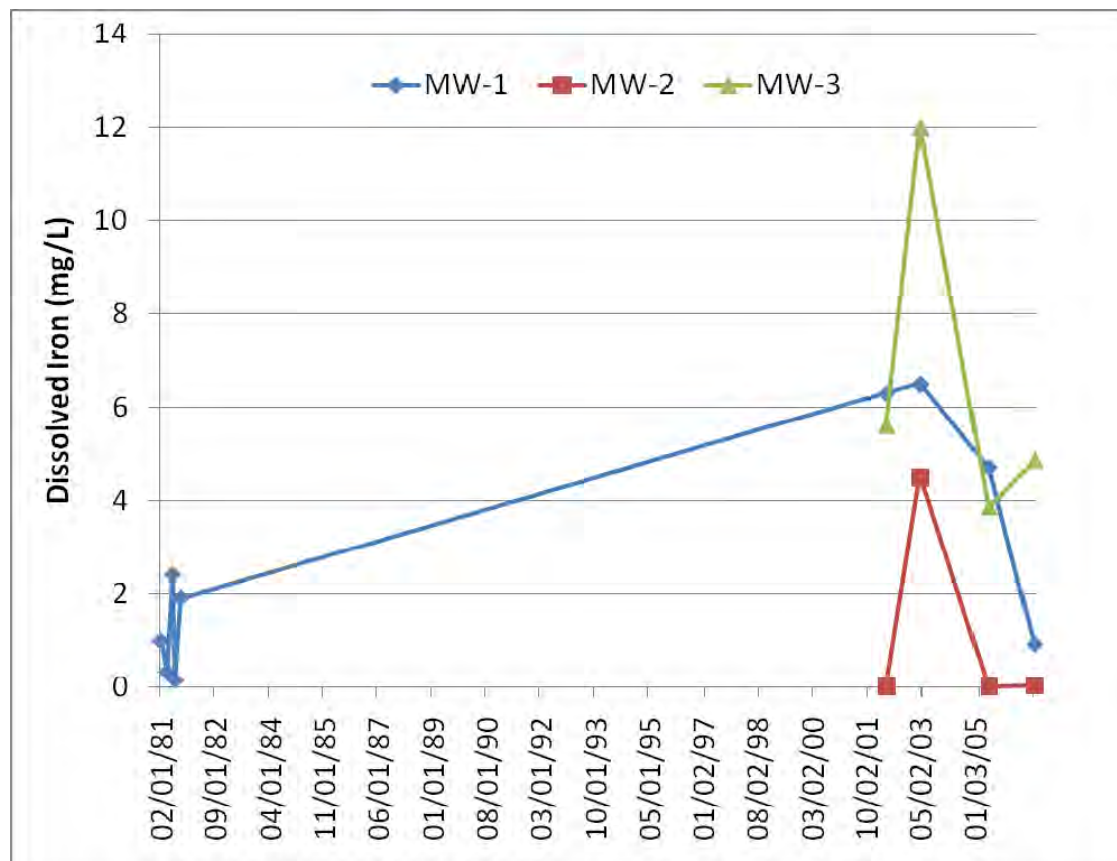


Figure 2 – Variations in dissolved iron concentrations with time for wells MW-1, MW-2 and MW-3

The results of the mixing modeling using three different assumptions to set the Eh of the mixture is shown in **Table 3-2**.

Table 3-2 – Summary of the 50:50 Mixing Simulation Results Using PHREEQC

Scenario	Iron Precipitated as $\text{Fe}(\text{OH})_3$	Percent of Iron Precipitated
MW 95-4 and MW-01-15, Assume all of the DO in the mixture reacts with the iron	3.90 mg/L	100%
MW 95-4 and MW-01-15, Assume that the Eh attained by the mixture is the same as for well MW 95-5 (397 mV)	3.90 mg/L	100%
MW 95-4 and MW-01-15, Assume that the final Eh is the average of the two waters mixed (83 mV)	3.87 mg/L	98.3%

The Eh of the mixture had to be set within PHREEQC, as the model does not accurately predict the final Eh of the mixture. The three scenarios were as follows:

- Assume all of the DO in the mixture reacts with the iron – The DO provided by the decant water is allowed to react with any ferrous iron present in MW 95-4. While this makes sense from an equilibrium standpoint, oxygen does not usually attain equilibrium within surface waters.
- Assume that the Eh attained by the mixture is the same as for well MW 95-5 (397 mV) – As it is likely that well MW 95-5 has already undergone mixing, the use of the “final” Eh of a real site water is reasonable. As the mine or decant water is similar to an oxygenated infiltration water in terms of the redox properties, this is a reasonable assumption.
- Assume that the final Eh is the average of the two waters – While this scenario is the least likely of the three, it was included in order to show the result of a lower final Eh. In reality, the Eh of mixed waters is rarely linear as this method implies.

The results show that the mixing of the decant-influenced water (MW-01-15) and the natural groundwater (MW 95-4) results in the precipitation of Fe(OH)₃. Should copper from the decant or mine water ever reach the area of MW 95-4, the precipitation of the amorphous ferric hydroxide would likely result in coprecipitation of copper as follows:



Where “x” is the fraction of iron in the precipitated phase (Other divalent [charge of +2] metals would be removed in a similar fashion). However, the value of x is difficult to predict and determination of the relative fractions of iron and copper in the coprecipitate requires bench-scale testing.

Section 4

Bench-Scale Tests

4.1 Sample Collection

To determine the relative fractions of iron and copper in the coprecipitate, samples were collected for bench-scale testing. Three locations were selected at the Troy Mine site for sampling. To represent post-mining water quality, water was collected underground in the mine behind the Seven East dam. This water was chosen as an appropriate closure mine water surrogate because it was collected in an area of little recent blasting activity evidenced by its relatively low ammonia content (from explosives residues). Mine water would be discharged at closure to the decant pond, which was selected as the second sampling location. The third sample was ground water collected from location MW 95-4, at the farthest downgradient edge of the tailings impoundment.

On December 16, 2009, after meeting with mine personnel and reviewing safety procedures, sampling personnel mobilized to well MW95-4 for sampling. A static water level was collected and then MW 95-4 was purged of approximately 5 well volumes using a disposable bailer. Field water quality readings were collected throughout the purging using a portable water quality meter. After purging the well, a sample was analyzed using field test equipment to check for ferrous iron concentrations in the well water. Results showed approximately 5 mg/L of ferrous iron; this was above the minimum requirement of 2 mg/L, so analytical and jar test samples were collected. Samples collected for laboratory analysis were immediately preserved in the field. The preserved samples would be laboratory analyzed to determine the initial concentrations of metals. An additional volume was collected for jar testing in unpreserved jugs.

The decant pond was sampled next. Due to the cold temperatures and icy conditions, access to the pond was limited to the pump dock. A hole was chipped through the ice to gain access to the pond water. Field readings were collected and then analytical and jar test samples were collected, as before.

Mine water was then collected next from behind the underground dam. Field readings, analytical samples, and jar test samples were collected, similar to the previous locations.

Field notes and photographs were collected during the sampling process, and are attached.

4.2 Jar Test Procedures

Once the sampling was completed, jar tests were performed using the unpreserved samples. A control plus mixtures were selected for testing, as follows:

- Control - 1000 mL MW 95-4 water
- Jar 1 - 250 mL of MW 95-4 water with 750 mL decant water (25% A+ 75% B)
- Jar 2 - 500 mL of MW 95-4 water with 500 mL decant water (50% A+ 50% B)
- Jar 3 - 750 mL of MW 95-4 water with 250 mL decant water (75% A+ 25% B)
- Jar 4 - 250 mL of MW 95-4 water with 750 mL mine water (25% A+ 75% C)
- Jar 5 - 500 mL of MW 95-4 water with 500 mL mine water (50% A+ 50% C)
- Jar 6 - 750 mL of MW 95-4 water with 250 mL mine water (75% A+ 25% C)
- Jar 7 - 500 mL of MW 95-4 water with 500 mL mine water + 20mL 25% hydrogen peroxide [H₂O₂] (75% A+ 25% C + H₂O₂)

The appropriate volumes of each water sampled were added and mixed in each test jar, and then the jars were allowed to react for duration of 20 days (from December 16, 2009 to January 5, 2010).

The mixed samples were allowed to react at room temperature within the uncovered sample jars. After the reaction period, a portion of each test jar was filtered, preserved, and submitted to the lab. Field parameters were again collected from the test jars. The remaining volume from each jar was labeled and archived.

All test jars were photographed prior to and after filtration.

4.3 Results

The results of the jar testing are summarized in **Table 4-1**, while the full analytical reports are provided as Attachment 1.

Table 4-1 - Summary of Jar Test Results

Sample	Description	Antimony (mg/L)	Arsenic (mg/L)	Cadmium (mg/L)	Copper (mg/L)	Iron (mg/L)	Manganese (mg/L)	Uranium (mg/L)	Lead (mg/L)
MW95-4	Field Sample	<0.001	<0.003	<0.00008	0.006	15.5	5.34	0.0005	<0.0005
Decant Pond	Field Sample	0.068	<0.003	<0.00008	0.032	<0.05	0.862	0.0033	0.0008
Mine Water	Field Sample	0.015	<0.003	<0.00008	0.114	<0.05	0.105	0.0011	0.0010
Jar 1	25% MW95-4 75% Decant	0.046	<0.001	0.0001	0.007	0.008	2.17	0.0017	<0.0005
Jar 2	50% MW95-4 50% Decant	0.026	<0.001	0.0001	0.005	0.009	3.49	0.0003	<0.0005
Jar 3	75% MW95-4 25% Decant	0.010	<0.001	0.0002	0.003	0.008	4.86	0.0004	<0.0005
Jar 4	25% MW95-4 75% Mine	0.010	<0.001	<0.0001	0.012	<0.005	1.52	0.0007	<0.0005
Jar 5	50% MW95-4 50% Mine	0.006	<0.001	0.0001	0.005	0.007	2.74	0.0004	<0.0005
Jar 6	75% MW95-4 25% Mine	0.002	<0.001	0.0002	0.003	<0.005	4.47	<0.0003	<0.0005
Jar 7	50% MW95-4 50% Mine w/ H ₂ O ₂	0.007	<0.001	<0.0001	0.002	0.005	2.79	<0.0003	<0.0005
Control	100% MW95-4	<0.001	<0.001	0.0002	0.002	0.009	6.03	<0.0003	<0.0005

The differences in the reporting limit for arsenic were due to the fact that the three initial samples were analyzed using the standard reporting limit used by the laboratory. For all subsequent samples low level arsenic analyses were requested. However, the higher initial reporting limit was not a problem, as all subsequent analyses, including the control were below the low level reporting limit.

Reductions in concentrations of iron, copper, and to a lesser extent, antimony and uranium were observed as a result of mixing mine or decant water with the natural groundwater. Arsenic (as discussed above) and cadmium were below the analytical reporting limit in all three source waters. However, cadmium was detected in several of the mixed samples. As it is impossible to have a concentration in a mixture greater than either of the two component waters, it suggests a laboratory error, which is to be expected at levels so close to the reporting limit.

Due to the effect of dilution when two waters are mixed, the percent removals were calculated from the concentrations that would have resulted as a result of mixing alone (without coprecipitation or adsorption). The results are shown in **Table 4-2**.

Table 4-2 - Metal Removal Efficiency (Factoring in Dilution)

Sample	Antimony - Dilution (mg/L)	Percent Antimony Removal	Copper - Dilution (mg/L)	Percent Copper Removal	Iron - Dilution (mg/L)	Percent Iron Removal	Manganese - Dilution (mg/L)	Percent Manganese Removal	Uranium - Dilution (mg/L)	Percent Uranium Removal
Jar 1	0.051	10.5%	0.0255	72.5%	3.9	99.8%	2.0	-9.5%	0.0026	34.6%
Jar 2	0.035	25.2%	0.019	73.7%	7.8	99.9%	3.1	-12.5%	0.0019	84.2%
Jar 3	0.018	44.8%	0.0125	76.0%	11.6	99.9%	4.2	-15.2%	0.0012	66.7%
Jar 4	0.012	14.0%	0.087	86.2%	3.9	99.9%	1.4	-7.5%	0.0010	26.3%
Jar 5	0.008	27.3%	0.06	91.7%	7.8	99.9%	2.7	-0.6%	0.0008	50.0%
Jar 6	0.005	59.0%	0.033	90.9%	11.6	100.0%	4.0	-10.9%	0.0007	76.9%
Jar 7	0.008	15.2%	0.06	96.7%	7.8	99.9%	2.7	-2.5%	0.0	81.3%
Control	N/A	N/A	N/A	N/A	N/A	N/A	N/A	N/A	N/A	N/A

Greater than 99% of the iron was removed during the 20 day duration of the jar tests, indicating that sufficient time was given to oxidize the ferrous iron to ferric iron and result in precipitation. The red-brown color of the jar tests indicated that iron oxyhydroxides had precipitated (see Attachment 2 for photographs). However, the manganese did not oxidize, even in the sample to which hydrogen peroxide was added (jar 7).

In general, copper was coprecipitated with the iron to a greater extent than antimony or uranium. However, the water of concern for the post-closure scenario would be mine water which had a significantly lower antimony (0.015 mg/L for the mine water vs. 0.068 for the current decant water) and uranium (0.0033 mg/L vs. 0.0011 mg/L) concentrations. An estimated 73-76% of the copper was removed from the decant/groundwater mixtures, with the removal efficiency increasing with increasing fractions of groundwater.

A similar trend was observed for the mine water/groundwater mixtures, although the efficiency was higher, ranging from 92-98% removal. Greater removal efficiencies may be expected at mixtures containing groundwater fractions higher than 75%. This higher groundwater fraction would be expected at the leading edge of any decant water plume containing copper that could hypothetically migrate from the tailings ponds in the future.

Section 5

Conclusions

The results of the modeling and bench-scale testing show that should the initial removal mechanisms documented by Land and Water (2004) to occur beneath the decant ponds become less effective over time, there are additional existing

geochemical mechanisms capable of removing at least 73-98% of the copper, 11-59% of the antimony, and 35-84% of the uranium from groundwater.

Section 6

References

Hem, 1970. *Study and Interpretation of the Chemical Characteristics of Natural Water*. 2nd ed. U.S. Geological Survey Water-Supply Paper 1473.

Hydrometrics. 2001. *Mine Water Plume Location and Identification Phase 1 Results and Phase 2 & 3 Workplan*. Prepared for Genesis Inc./Sterling Mining Company. October 2001.

Land and Water Consulting, Inc. 2004. *Assessment of Fate and Transport of Copper in Decant Pond System – Troy Mine, Lincoln County, Montana*. Prepared for Revett Mining Company. January 2004.

Langmuir, D. 1971. Eh-pH determination. In R.E. Carver Ed. *Procedures in Sedimentary Petrology*. Wiley, New York. p. 597-635.

Lindberg, R.D. and D.D. Runnells. 1984. Ground water redox reactions: An analysis of equilibrium state applied to Eh measurements and geochemical modeling. *Science*. v. 225. p. 925-927.

Parkhurst, D.L. and Appelo, C.A.J. 1999. User's Guide to PHREEQC (Version 2)- A Computer Program for Speciation, Batch-Reaction, One-Dimensional Transport, and Inverse Geochemical Calculations. USGS Water-Resources Investigations Report 99-4259.

Rickard, D. 2006. The solubility of FeS. *Geochimica Cosmochimica Acta*, v. 70 p. 5779-5789.

Attachment 1

Analytical Reports



ANALYTICAL SUMMARY REPORT

January 28, 2010

Paul Kukay

Genesis Inc.

7000 Asarco Rd

Troy, MT 59935

Workorder No.: H10010040

Project Name: Troy Mine Jar Testing

Energy Laboratories Inc received the following 8 samples for Genesis Inc. on 1/5/2010 for analysis.

Sample ID	Client Sample ID	Collect Date	Receive Date	Matrix	Test
H10010040-001	Control - 010410	01/04/10 14:56	01/05/10	Aqueous	Metals by ICP/ICPMS, Dissolved
H10010040-002	Jar1 - 010410	01/04/10 15:02	01/05/10	Aqueous	Same As Above
H10010040-003	Jar2 - 010410	01/04/10 15:10	01/05/10	Aqueous	Same As Above
H10010040-004	Jar3 - 010410	01/04/10 15:16	01/05/10	Aqueous	Same As Above
H10010040-005	Jar4 - 010410	01/04/10 15:25	01/05/10	Aqueous	Same As Above
H10010040-006	Jar5 - 010410	01/04/10 15:30	01/05/10	Aqueous	Same As Above
H10010040-007	Jar6 - 010410	01/04/10 15:35	01/05/10	Aqueous	Same As Above
H10010040-008	Jar7 - 010410	01/04/10 15:40	01/05/10	Aqueous	Same As Above

BRANCH LABORATORY LOCATIONS

eli-b - Energy Laboratories, Inc. - Billings, MT, EPA # MT00005
eli-c - Energy Laboratories, Inc. - Casper, WY, EPA# WY00002
eli-g - Energy Laboratories, Inc. - Gillette, WY, EPA# WY00006
eli-h - Energy Laboratories, Inc. - Helena, MT, EPA# MT00945
eli-r - Energy Laboratories, Inc. - Rapid City, SD, EPA# SD00012
eli-t - Energy Laboratories, Inc. - College Station, TX, EPA# TX01520

SUBCONTRACTING ANALYSIS

Subcontracting of sample analyses to an outside laboratory may be required. If so, ENERGY LABORATORIES, INC. will utilize its branch laboratories or qualified contract laboratories for this service. Any such laboratories are indicated within the Laboratory Analytical Report.

SAMPLE TEMPERATURE COMPLIANCE: 4°C (±2°C)

Temperature of samples received may not be considered properly preserved by accepted standards. Samples that are hand delivered immediately after collection shall be considered acceptable if there is evidence that the chilling process has begun.

ELI appreciates the opportunity to provide you with this analytical service. For additional information, including certifications, and analytical services visit our web page www.energylab.com.

Report Approved By: _____



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine Jar Testing
Lab ID: H10010040-001
Client Sample ID Control - 010410

Report Date: 01/28/10
Collection Date: 01/04/10 14:56
DateReceived: 01/05/10
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Antimony	ND	mg/L		0.001		E200.8	01/08/10 23:35 / dck
Arsenic	ND	mg/L		0.001		E200.8	01/08/10 23:35 / dck
Cadmium	0.0002	mg/L		0.0001		E200.8	01/08/10 23:35 / dck
Copper	0.002	mg/L		0.001		E200.8	01/08/10 23:35 / dck
Iron	0.009	mg/L		0.005		E200.8	01/12/10 01:28 / dck
Manganese	6.03	mg/L		0.005		E200.8	01/09/10 20:32 / dck

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine Jar Testing
Lab ID: H10010040-002
Client Sample ID Jar1 - 010410

Report Date: 01/28/10
Collection Date: 01/04/10 15:02
DateReceived: 01/05/10
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Antimony	0.046	mg/L		0.001		E200.8	01/08/10 23:42 / dck
Arsenic	ND	mg/L		0.001		E200.8	01/08/10 23:42 / dck
Cadmium	0.0001	mg/L		0.0001		E200.8	01/08/10 23:42 / dck
Copper	0.007	mg/L		0.001		E200.8	01/08/10 23:42 / dck
Iron	0.008	mg/L		0.005		E200.8	01/12/10 01:48 / dck
Manganese	2.17	mg/L		0.005		E200.8	01/08/10 23:42 / dck

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine Jar Testing
Lab ID: H10010040-003
Client Sample ID Jar2 - 010410

Report Date: 01/28/10
Collection Date: 01/04/10 15:10
DateReceived: 01/05/10
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Antimony	0.026	mg/L		0.001		E200.8	01/08/10 23:48 / dck
Arsenic	ND	mg/L		0.001		E200.8	01/08/10 23:48 / dck
Cadmium	0.0001	mg/L		0.0001		E200.8	01/08/10 23:48 / dck
Copper	0.005	mg/L		0.001		E200.8	01/08/10 23:48 / dck
Iron	0.009	mg/L		0.005		E200.8	01/12/10 02:36 / dck
Manganese	3.49	mg/L		0.005		E200.8	01/08/10 23:48 / dck

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine Jar Testing
Lab ID: H10010040-004
Client Sample ID Jar3 - 010410

Report Date: 01/28/10
Collection Date: 01/04/10 15:16
DateReceived: 01/05/10
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Antimony	0.010	mg/L		0.001		E200.8	01/09/10 00:43 / dck
Arsenic	ND	mg/L		0.001		E200.8	01/09/10 00:43 / dck
Cadmium	0.0002	mg/L		0.0001		E200.8	01/09/10 00:43 / dck
Copper	0.003	mg/L		0.001		E200.8	01/09/10 00:43 / dck
Iron	0.008	mg/L		0.005		E200.8	01/09/10 00:43 / dck
Manganese	4.86	mg/L		0.005		E200.8	01/09/10 21:39 / dck

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine Jar Testing
Lab ID: H10010040-005
Client Sample ID Jar4 - 010410

Report Date: 01/28/10
Collection Date: 01/04/10 15:25
DateReceived: 01/05/10
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Antimony	0.010	mg/L		0.001		E200.8	01/09/10 00:50 / dck
Arsenic	ND	mg/L		0.001		E200.8	01/09/10 00:50 / dck
Cadmium	ND	mg/L		0.0001		E200.8	01/09/10 00:50 / dck
Copper	0.012	mg/L		0.001		E200.8	01/09/10 00:50 / dck
Iron	ND	mg/L		0.005		E200.8	01/09/10 00:50 / dck
Manganese	1.52	mg/L		0.005		E200.8	01/09/10 00:50 / dck

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine Jar Testing
Lab ID: H10010040-006
Client Sample ID Jar5 - 010410

Report Date: 01/28/10
Collection Date: 01/04/10 15:30
DateReceived: 01/05/10
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Antimony	0.006	mg/L		0.001		E200.8	01/09/10 00:56 / dck
Arsenic	ND	mg/L		0.001		E200.8	01/09/10 00:56 / dck
Cadmium	0.0001	mg/L		0.0001		E200.8	01/09/10 00:56 / dck
Copper	0.005	mg/L		0.001		E200.8	01/09/10 00:56 / dck
Iron	0.007	mg/L		0.005		E200.8	01/09/10 00:56 / dck
Manganese	2.74	mg/L		0.005		E200.8	01/09/10 00:56 / dck

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine Jar Testing
Lab ID: H10010040-007
Client Sample ID Jar6 - 010410

Report Date: 01/28/10
Collection Date: 01/04/10 15:35
DateReceived: 01/05/10
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Antimony	0.002	mg/L		0.001		E200.8	01/09/10 01:03 / dck
Arsenic	ND	mg/L		0.001		E200.8	01/09/10 01:03 / dck
Cadmium	0.0002	mg/L		0.0001		E200.8	01/09/10 01:03 / dck
Copper	0.003	mg/L		0.001		E200.8	01/09/10 01:03 / dck
Iron	ND	mg/L		0.005		E200.8	01/09/10 01:03 / dck
Manganese	4.47	mg/L		0.005		E200.8	01/09/10 01:03 / dck

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine Jar Testing
Lab ID: H10010040-008
Client Sample ID Jar7 - 010410

Report Date: 01/28/10
Collection Date: 01/04/10 15:40
DateReceived: 01/05/10
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Antimony	0.007	mg/L		0.001		E200.8	01/09/10 01:10 / dck
Arsenic	ND	mg/L		0.001		E200.8	01/09/10 01:10 / dck
Cadmium	ND	mg/L		0.0001		E200.8	01/09/10 01:10 / dck
Copper	0.002	mg/L		0.001		E200.8	01/09/10 01:10 / dck
Iron	0.005	mg/L		0.005		E200.8	01/09/10 01:10 / dck
Manganese	2.79	mg/L		0.005		E200.8	01/09/10 01:10 / dck

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.
Project: Troy Mine Jar Testing

Report Date: 01/28/10
Work Order: H10010040

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8								Analytical Run: ICPMS204-B_100108A		
Sample ID: QCS-090602A,090609		6	Initial Calibration Verification Standard						01/08/10 13:32	
Antimony		0.0499	mg/L	0.050	100	90	110			
Arsenic		0.0497	mg/L	0.0050	99	90	110			
Cadmium		0.0258	mg/L	0.0010	103	90	110			
Copper		0.0507	mg/L	0.010	101	90	110			
Iron		0.261	mg/L	0.030	102	90	110			
Manganese		0.247	mg/L	0.010	99	90	110			
Sample ID: ICSA-090423A		6	Interference Check Sample A						01/08/10 13:39	
Antimony		0.000189	mg/L	0.050						
Arsenic		0.000151	mg/L	0.0050						
Cadmium		0.000421	mg/L	0.0010						
Copper		0.00118	mg/L	0.010						
Iron		99.6	mg/L	0.030	100	70	130			
Manganese		0.00223	mg/L	0.010						
Sample ID: ICSAB-090423A,09010		6	Interference Check Sample AB						01/08/10 13:46	
Antimony		0.000153	mg/L	0.050		0	0			
Arsenic		0.0105	mg/L	0.0050	105	70	130			
Cadmium		0.0101	mg/L	0.0010	101	70	130			
Copper		0.0208	mg/L	0.010	104	70	130			
Iron		96.9	mg/L	0.030	97	70	130			
Manganese		0.0220	mg/L	0.010	110	70	130			
Sample ID: QCS-090602A,090609		6	Initial Calibration Verification Standard						01/08/10 22:07	
Antimony		0.0498	mg/L	0.050	100	90	110			
Arsenic		0.0498	mg/L	0.0050	100	90	110			
Cadmium		0.0263	mg/L	0.0010	105	90	110			
Copper		0.0503	mg/L	0.010	101	90	110			
Iron		0.264	mg/L	0.030	104	90	110			
Manganese		0.251	mg/L	0.010	101	90	110			
Sample ID: ICSA-090423A		6	Interference Check Sample A						01/08/10 22:14	
Antimony		0.000180	mg/L	0.050						
Arsenic		0.000103	mg/L	0.0050						
Cadmium		0.000326	mg/L	0.0010						
Copper		0.00123	mg/L	0.010						
Iron		98.6	mg/L	0.030	99	70	130			
Manganese		0.00219	mg/L	0.010						
Sample ID: ICSAB-090423A,09010		6	Interference Check Sample AB						01/08/10 22:20	
Antimony		0.000156	mg/L	0.050		0	0			
Arsenic		0.0104	mg/L	0.0050	104	70	130			
Cadmium		0.0101	mg/L	0.0010	101	70	130			
Copper		0.0207	mg/L	0.010	104	70	130			
Iron		96.3	mg/L	0.030	96	70	130			

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.
Project: Troy Mine Jar Testing

Report Date: 01/28/10
Work Order: H10010040

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8								Analytical Run: ICPMS204-B_100108A		
Sample ID: ICSAB-090423A,09010		6	Interference Check Sample AB					01/08/10 22:20		
Manganese		0.0210	mg/L	0.010	105	70	130			
Sample ID: QCS-090602A,090609		6	Initial Calibration Verification Standard					01/09/10 05:14		
Antimony		0.0493	mg/L	0.050	99	90	110			
Arsenic		0.0494	mg/L	0.0050	99	90	110			
Cadmium		0.0259	mg/L	0.0010	104	90	110			
Copper		0.0504	mg/L	0.010	101	90	110			
Iron		0.254	mg/L	0.030	99	90	110			
Manganese		0.256	mg/L	0.010	103	90	110			
Sample ID: ICSA-090423A		6	Interference Check Sample A					01/09/10 05:21		
Antimony		0.000209	mg/L	0.050						
Arsenic		0.000148	mg/L	0.0050						
Cadmium		0.000313	mg/L	0.0010						
Copper		0.00130	mg/L	0.010						
Iron		97.9	mg/L	0.030	98	70	130			
Manganese		0.00207	mg/L	0.010						
Sample ID: ICSAB-090423A,09010		6	Interference Check Sample AB					01/09/10 05:27		
Antimony		0.000150	mg/L	0.050		0	0			
Arsenic		0.0102	mg/L	0.0050	102	70	130			
Cadmium		0.00988	mg/L	0.0010	99	70	130			
Copper		0.0208	mg/L	0.010	104	70	130			
Iron		98.9	mg/L	0.030	99	70	130			
Manganese		0.0217	mg/L	0.010	108	70	130			
Method: E200.8								Batch: R59545		
Sample ID: LRB		6	Method Blank			Run: ICPMS204-B_100108A			01/08/10 14:46	
Antimony		ND	mg/L	3E-05						
Arsenic		ND	mg/L	2E-05						
Cadmium		ND	mg/L	1E-05						
Copper		ND	mg/L	3E-05						
Iron		ND	mg/L	6E-05						
Manganese		3E-05	mg/L	2E-05						
Sample ID: LFB		6	Laboratory Fortified Blank			Run: ICPMS204-B_100108A			01/08/10 14:53	
Antimony		0.0511	mg/L	0.050	102	85	115			
Arsenic		0.0497	mg/L	0.0050	99	85	115			
Cadmium		0.0495	mg/L	0.0010	99	85	115			
Copper		0.0485	mg/L	0.010	97	85	115			
Iron		4.91	mg/L	0.030	98	85	115			
Manganese		0.0497	mg/L	0.010	99	85	115			
Sample ID: H10010040-003AMS		6	Sample Matrix Spike			Run: ICPMS204-B_100108A			01/08/10 23:55	
Antimony		0.0751	mg/L	0.0050	98	70	130			

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.
Project: Troy Mine Jar Testing

Report Date: 01/28/10
Work Order: H10010040

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8										Batch: R59545
Sample ID: H10010040-003AMS 6 Sample Matrix Spike										Run: ICPMS204-B_100108A 01/08/10 23:55
Arsenic		0.0505	mg/L	0.0050	100	70	130			
Cadmium		0.0484	mg/L	0.0010	97	70	130			
Copper		0.0525	mg/L	0.010	95	70	130			
Iron		4.81	mg/L	0.030	96	70	130			
Manganese		3.40	mg/L	0.010		70	130			A
Sample ID: H10010041-001CMS 6 Sample Matrix Spike										Run: ICPMS204-B_100108A 01/09/10 01:57
Antimony		0.0507	mg/L	0.0050	101	70	130			
Arsenic		0.0501	mg/L	0.0050	100	70	130			
Cadmium		0.0498	mg/L	0.0010	96	70	130			
Copper		0.0467	mg/L	0.010	93	70	130			
Iron		4.84	mg/L	0.030	97	70	130			
Manganese		17.2	mg/L	0.010		70	130			A
Sample ID: H10010041-001CMSD 6 Sample Matrix Spike Duplicate										Run: ICPMS204-B_100108A 01/09/10 02:04
Antimony		0.0498	mg/L	0.0050	100	70	130	1.7	20	
Arsenic		0.0505	mg/L	0.0050	100	70	130	0.6	20	
Cadmium		0.0495	mg/L	0.0010	95	70	130	0.6	20	
Copper		0.0472	mg/L	0.010	94	70	130	1.1	20	
Iron		4.89	mg/L	0.030	98	70	130	1.1	20	
Manganese		17.3	mg/L	0.010		70	130	0.6	20	A

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

A - The analyte level was greater than four times the spike level. In accordance with the method % recovery is not calculated.



QA/QC Summary Report

Client: Genesis Inc.
Project: Troy Mine Jar Testing

Report Date: 01/28/10
Work Order: H10010040

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8								Analytical Run: ICPMS204-B_100109B		
Sample ID: QCS-090602A,090609	Initial Calibration Verification Standard									
Manganese		0.248	mg/L	0.010	99	90	110			01/09/10 18:57
Sample ID: ICSA-090423A	Interference Check Sample A									
Manganese		0.00225	mg/L	0.010						01/09/10 19:03
Sample ID: ICSAB-090423A,09010	Interference Check Sample AB									
Manganese		0.0224	mg/L	0.010	112	70	130			01/09/10 19:10
Sample ID: QCS-090602A,090609	Initial Calibration Verification Standard									
Manganese		0.251	mg/L	0.010	101	90	110			01/10/10 00:14
Method: E200.8								Batch: R59556		
Sample ID: LRB	Method Blank									
Manganese		4E-05	mg/L	2E-05			Run: ICPMS204-B_100109B			01/09/10 20:11
Sample ID: LFB	Laboratory Fortified Blank									
Manganese		0.0510	mg/L	0.010	102	85	115			01/09/10 20:18
Sample ID: H10010040-001AMS	Sample Matrix Spike									
Manganese		5.96	mg/L	0.010		70	130			01/09/10 20:39 A
Sample ID: H10010040-001AMSD	Sample Matrix Spike Duplicate									
Manganese		5.96	mg/L	0.010		70	130	0		01/09/10 20:45 20 A
Sample ID: H10010041-001CMS	Sample Matrix Spike									
Manganese		16.1	mg/L	0.010		70	130			01/09/10 21:53 A
Sample ID: H10010041-001CMSD	Sample Matrix Spike Duplicate									
Manganese		14.2	mg/L	0.010		70	130	13		01/09/10 21:59 20 A

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

A - The analyte level was greater than four times the spike level. In accordance with the method % recovery is not calculated.



QA/QC Summary Report

Client: Genesis Inc.
Project: Troy Mine Jar Testing

Report Date: 01/28/10
Work Order: H10010040

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8		Analytical Run: ICPMS204-B_100111A								
Sample ID: QCS-090602A,090609		Initial Calibration Verification Standard								01/11/10 17:17
Iron		0.260	mg/L	0.030	102	90	110			
Sample ID: ICSA-090423A		Interference Check Sample A								01/11/10 17:24
Iron		101	mg/L	0.030	101	70	130			
Sample ID: ICSAB-090423A,09010		Interference Check Sample AB								01/11/10 17:30
Iron		101	mg/L	0.030	101	70	130			
Sample ID: QCS-090602A,090609		Initial Calibration Verification Standard								01/12/10 05:25
Iron		0.255	mg/L	0.030	100	90	110			
Sample ID: ICSA-090423A		Interference Check Sample A								01/12/10 05:32
Iron		101	mg/L	0.030	101	70	130			
Sample ID: ICSAB-090423A,09010		Interference Check Sample AB								01/12/10 05:38
Iron		99.9	mg/L	0.030	100	70	130			
Method: E200.8		Batch: R59601								
Sample ID: LRB		Method Blank								01/12/10 01:01
Iron		0.0004	mg/L	6E-05						
Sample ID: LFB		Laboratory Fortified Blank								01/12/10 01:08
Iron		4.85	mg/L	0.030	97	85	115			
Sample ID: H10010040-003AMS		Sample Matrix Spike								01/12/10 02:43
Iron		4.86	mg/L	0.030	97	70	130			
Sample ID: H10010040-003AMSD		Sample Matrix Spike Duplicate								01/12/10 02:50
Iron		4.85	mg/L	0.030	97	70	130	0.1	20	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



Workorder Receipt Checklist



H10010040

Login completed by: Tracy L. Lorash

Date and Time Received: 1/5/2010 1:26 PM

Reviewed by: BL2000\wjohanson

Received by: TLL

Reviewed Date: 1/27/2010 12:39:03 PM

Carrier name: Hand Del

Shipping container/cooler in good condition?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	Not Present <input type="checkbox"/>
Custody seals intact on shipping container/cooler?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	Not Present <input type="checkbox"/>
Custody seals intact on sample bottles?	Yes <input type="checkbox"/>	No <input type="checkbox"/>	Not Present <input checked="" type="checkbox"/>
Chain of custody present?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Chain of custody signed when relinquished and received?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Chain of custody agrees with sample labels?	Yes <input type="checkbox"/>	No <input checked="" type="checkbox"/>	
Samples in proper container/bottle?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Sample containers intact?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Sufficient sample volume for indicated test?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
All samples received within holding time?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Container/Temp Blank temperature:	4.4°C On Ice		
Water - VOA vials have zero headspace?	Yes <input type="checkbox"/>	No <input type="checkbox"/>	No VOA vials submitted <input checked="" type="checkbox"/>
Water - pH acceptable upon receipt?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	Not Applicable <input type="checkbox"/>

Contact and Corrective Action Comments:

Jar 3 has 010409 as part of sample ID and as the date. Contacted Dustin and it is supposed to read 010410. TI



Chain of Custody and Analytical Request Record

Page 1 of 1

PLEASE PRINT- Provide as much information as possible.

Company Name: CAMP DRESSER MCKEE / CDM		Project Name, PWS, Permit, Etc. TROY MINE JAR TESTING		Sample Origin State: MT		EPA/State Compliance: Yes <input type="checkbox"/> No <input checked="" type="checkbox"/>	
Report Mail Address: 50 W 14TH ST, #200 HELENA, MT 59601		Contact Name: DUSTIN KLEMPER		Phone/Fax: (406) 441-1405		Email: KLEMPER.DT@CDM.COM	
Invoice Address: P.O. Box 1660 Troy, MT 59935		Invoice Contact & Phone: PAUL KUKAY (406) 295-5882 x221		Purchase Order: _____		Quote/Bottle Order: _____	
Special Report/Formats - ELI must be notified prior to sample submittal for the following: <input type="checkbox"/> DW <input type="checkbox"/> GSA <input type="checkbox"/> POTW/MWTP <input type="checkbox"/> State: _____ <input type="checkbox"/> Other: _____ <input type="checkbox"/> A2LA <input type="checkbox"/> EDD/EDT (Electronic Data) Format: _____ <input type="checkbox"/> LEVEL IV <input type="checkbox"/> NELAC		ANALYSIS REQUESTED DISSOLVED METALS LOW LEVEL SEE ATTACHED Normal Turnaround (TAT)		Contact ELI prior to RUSH sample submittal for charges and scheduling - See Instruction Page R U S H		Shipped by: Hand Del. Cooler ID(s): _____ Receipt Temp: 46.7 °C On Job: Yes <input checked="" type="checkbox"/> No <input type="checkbox"/> Custody Seal: YN Bottles/Coolers: YN Intact: YN Signature Match: YN	
SAMPLE IDENTIFICATION (Name, Location, Interval, etc.)		Collection Date	Collection Time	MATRIX	Number of Containers Air Water Soils/Solids Vegetation Bioassay Other	LABORATORY USE ONLY	
1 CONTROL-010410	01/04/10	15:56	L	X		ALL SAMPLES FIELD FILTERED & PRESERVED BY SAMPLER	
2 JAR1-010410		15:02		X			
3 JAR2-010410		15:10		X			
4 JAR3-010410		15:16		X			
5 JAR4-010410		15:25		X			
6 JAR5-010410		15:30		X			
7 JAR6-010410		15:35		X			
8 JAR7-010410		15:40		X			
9							
10							
* LOW LEVEL DISSOLVED METALS = COPPER, ANTIMONY, IRON, MANGANESE, CADMIUM, ARSENIC							
Custody Record MUST be Signed		Relinquished by (print): Dustin Klempel	Date/Time: 01/05/09 13:26	Signature: <i>[Signature]</i>	Received by (print): Troyland	Date/Time: 1/5/10 13:26	Signature: <i>[Signature]</i>
Sample Disposal: _____		Return to Client: _____		Lab Disposal: _____	Received by Laboratory: _____		

In certain circumstances, samples submitted to Energy Laboratories, Inc. may be subcontracted to other certified laboratories in order to complete the analysis requested. This serves as notice of this possibility. All sub-contract data will be clearly notated on your analytical report. Visit our web site at www.energylab.com for additional information, downloadable fee schedule, forms, and links.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine
Lab ID: H09120170-001
Client Sample ID MW 95-4

Report Date: 01/07/10
Collection Date: 12/16/09 10:20
Date Received: 12/17/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	6.4	s.u.		0.1		A4500-H B	12/17/09 17:50 / hm
Conductivity	801	umhos/cm		1		A2510 B	12/17/09 17:47 / hm
Solids, Total Dissolved TDS @ 180 C	561	mg/L		10		A2540 C	12/18/09 09:09 / JG
INORGANICS							
Alkalinity, Total as CaCO3	80	mg/L		1		A2320 B	12/17/09 20:25 / hm
Bicarbonate as HCO3	97	mg/L		1		A2320 B	12/17/09 20:25 / hm
Carbonate as CO3	ND	mg/L		1		A2320 B	12/17/09 20:25 / hm
Chloride	1	mg/L		1		E300.0	12/17/09 22:13 / hm
Sulfate	320	mg/L		1		E300.0	12/17/09 22:13 / hm
Sodium Adsorption Ratio (SAR)	0.63	unitless		0.01		Calculation	01/07/10 10:10 / wjj
Fluoride	ND	mg/L		0.1	4	A4500-F C	12/18/09 13:31 / JG
Sulfide	ND	mg/L		0.04		A4500 S-D	12/21/09 11:30 / eli-b
Hardness as CaCO3	232	mg/L		1		A2340 B	12/24/09 10:32 / sld
AGGREGATE ORGANICS							
Organic Carbon, Total (TOC)	3.2	mg/L		0.5		A5310 C	12/21/09 15:40 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	5.7	mg/L	D	0.5		E350.1	12/21/09 16:24 / stp
Nitrogen, Nitrate+Nitrite as N	ND	mg/L		0.01		E353.2	12/18/09 14:55 / stp
Phosphorus, Orthophosphate as P	ND	mg/L		0.001		E365.1	12/18/09 17:01 / stp
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.8	12/17/09 19:00 / dck
Antimony	ND	mg/L		0.003		E200.8	12/17/09 19:00 / dck
Arsenic	ND	mg/L		0.003		E200.8	12/17/09 19:00 / dck
Barium	0.255	mg/L		0.005		E200.8	12/17/09 19:00 / dck
Beryllium	ND	mg/L		0.001		E200.8	12/17/09 19:00 / dck
Cadmium	ND	mg/L		0.00008		E200.8	12/17/09 19:00 / dck
Calcium	62	mg/L		1		E200.7	12/17/09 15:30 / sld
Chromium	ND	mg/L		0.001		E200.8	12/17/09 19:00 / dck
Copper	0.006	mg/L		0.001		E200.8	12/17/09 19:00 / dck
Iron	15.5	mg/L		0.05		E200.7	12/17/09 15:30 / sld
Lead	ND	mg/L		0.0005		E200.8	12/17/09 19:00 / dck
Magnesium	19	mg/L		1		E200.7	12/17/09 15:30 / sld
Manganese	5.34	mg/L		0.005		E200.8	12/18/09 13:06 / dck
Mercury	ND	mg/L		0.00001		E200.8	12/17/09 19:00 / dck
Nickel	ND	mg/L		0.01		E200.8	12/17/09 19:00 / dck
Potassium	25	mg/L		1		E200.7	12/17/09 15:30 / sld
Selenium	ND	mg/L		0.001		E200.8	12/17/09 19:00 / dck
Silica	43.1	mg/L		0.2		E200.8	12/18/09 13:06 / dck

**Report
Definitions:**

RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



ENERGY LABORATORIES, INC. * 3161 E Lyndale (59604) * PO Box 5688 * Helena, MT 59601
Toll Free 877.472.0711 * 406.442.0711 * FAX 406.442.0712 * helena@enerylab.com

LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine
Lab ID: H09120170-001
Client Sample ID MW 95-4

Report Date: 01/07/10
Collection Date: 12/16/09 10:20
Date Received: 12/17/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Silver	ND	mg/L		0.0005		E200.8	12/17/09 19:00 / dck
Sodium	22	mg/L		1		E200.7	12/23/09 10:25 / sld
Thallium	ND	mg/L		0.0002		E200.8	12/17/09 19:00 / dck
Uranium	0.0005	mg/L		0.0003		E200.8	12/17/09 19:00 / dck
Zinc	0.01	mg/L		0.01		E200.8	12/17/09 19:00 / dck
METALS, TOTAL							
Iron	17.4	mg/L		0.03		E200.7	12/22/09 12:41 / sld

Report RL - Analyte reporting limit.
Definitions: QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine
Lab ID: H09120170-002
Client Sample ID Decant Pond

Report Date: 01/07/10
Collection Date: 12/16/09 11:10
Date Received: 12/17/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	8.1	s.u.		0.1		A4500-H B	12/17/09 17:51 / hm
Conductivity	685	umhos/cm		1		A2510 B	12/17/09 17:49 / hm
Solids, Total Dissolved TDS @ 180 C	436	mg/L		10		A2540 C	12/18/09 09:09 / JG
INORGANICS							
Alkalinity, Total as CaCO ₃	100	mg/L		1		A2320 B	12/17/09 20:31 / hm
Bicarbonate as HCO ₃	120	mg/L		1		A2320 B	12/17/09 20:31 / hm
Carbonate as CO ₃	ND	mg/L		1		A2320 B	12/17/09 20:31 / hm
Chloride	6	mg/L		1		E300.0	12/17/09 22:29 / hm
Sulfate	68	mg/L		1		E300.0	12/17/09 22:29 / hm
Sodium Adsorption Ratio (SAR)	2.55	unitless		0.01		Calculation	01/07/10 10:10 / wjj
Fluoride	0.3	mg/L		0.1	4	A4500-F C	12/18/09 13:31 / JG
Sulfide	ND	mg/L		0.04		A4500 S-D	12/21/09 11:30 / eli-b
Hardness as CaCO ₃	78	mg/L		1		A2340 B	12/18/09 12:19 / abb
AGGREGATE ORGANICS							
Organic Carbon, Total (TOC)	1.2	mg/L		0.5		A5310 C	12/21/09 15:51 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	11.2	mg/L	D	0.5		E350.1	12/21/09 16:26 / stp
Nitrogen, Nitrate+Nitrite as N	37.6	mg/L	D	0.5		E353.2	12/18/09 15:01 / stp
Phosphorus, Orthophosphate as P	0.070	mg/L		0.001		E365.1	12/18/09 17:02 / stp
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.8	12/17/09 19:07 / dck
Antimony	0.068	mg/L		0.003		E200.8	12/17/09 19:07 / dck
Arsenic	ND	mg/L		0.003		E200.8	12/17/09 19:07 / dck
Barium	0.091	mg/L		0.005		E200.8	12/17/09 19:07 / dck
Beryllium	ND	mg/L		0.001		E200.8	12/17/09 19:07 / dck
Cadmium	ND	mg/L		0.00008		E200.8	12/17/09 19:07 / dck
Calcium	22	mg/L		1		E200.7	12/17/09 15:39 / sld
Chromium	ND	mg/L		0.001		E200.8	12/17/09 19:07 / dck
Copper	0.032	mg/L		0.001		E200.8	12/17/09 19:07 / dck
Iron	ND	mg/L		0.05		E200.7	12/17/09 15:39 / sld
Lead	0.0008	mg/L		0.0005		E200.8	12/17/09 19:07 / dck
Magnesium	6	mg/L		1		E200.7	12/17/09 15:39 / sld
Manganese	0.862	mg/L		0.005		E200.8	12/17/09 19:07 / dck
Mercury	ND	mg/L		0.00001		E200.8	12/17/09 19:07 / dck
Nickel	ND	mg/L		0.01		E200.8	12/17/09 19:07 / dck
Potassium	32	mg/L		1		E200.7	12/17/09 15:39 / sld
Selenium	ND	mg/L		0.001		E200.8	12/17/09 19:07 / dck
Silica	9.7	mg/L		0.2		E200.8	12/17/09 19:07 / dck

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine
Lab ID: H09120170-002
Client Sample ID Decant Pond

Report Date: 01/07/10
Collection Date: 12/16/09 11:10
Date Received: 12/17/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Silver	ND	mg/L		0.0005		E200.8	12/17/09 19:07 / dck
Sodium	52	mg/L		1		E200.7	12/17/09 15:39 / sld
Thallium	ND	mg/L		0.0002		E200.8	12/17/09 19:07 / dck
Uranium	0.0033	mg/L		0.0003		E200.8	12/17/09 19:07 / dck
Zinc	ND	mg/L		0.01		E200.8	12/17/09 19:07 / dck
METALS, TOTAL							
Iron	0.45	mg/L		0.03		E200.7	12/22/09 12:44 / sld

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine
Lab ID: H09120170-003
Client Sample ID Mine Water

Report Date: 01/07/10
Collection Date: 12/16/09 13:10
Date Received: 12/17/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.5	s.u.		0.1		A4500-H B	12/17/09 17:52 / hm
Conductivity	288	umhos/cm		1		A2510 B	12/17/09 17:50 / hm
Solids, Total Dissolved TDS @ 180 C	184	mg/L		10		A2540 C	12/18/09 09:10 / JG
INORGANICS							
Alkalinity, Total as CaCO3	70	mg/L		1		A2320 B	12/17/09 20:52 / hm
Bicarbonate as HCO3	85	mg/L		1		A2320 B	12/17/09 20:52 / hm
Carbonate as CO3	ND	mg/L		1		A2320 B	12/17/09 20:52 / hm
Chloride	ND	mg/L		1		E300.0	12/17/09 22:46 / hm
Sulfate	42	mg/L		1		E300.0	12/17/09 22:46 / hm
Sodium Adsorption Ratio (SAR)	0.11	unitless		0.01		Calculation	01/07/10 10:10 / wjj
Fluoride	ND	mg/L		0.1	4	A4500-F C	12/18/09 13:31 / JG
Sulfide	ND	mg/L		0.04		A4500 S-D	12/21/09 11:30 / eli-b
Hardness as CaCO3	114	mg/L		1		A2340 B	12/24/09 10:32 / sld
AGGREGATE ORGANICS							
Organic Carbon, Total (TOC)	ND	mg/L		0.5		A5310 C	12/21/09 16:00 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	0.39	mg/L		0.05		E350.1	12/21/09 16:00 / stp
Nitrogen, Nitrate+Nitrite as N	7.4	mg/L	D	0.1		E353.2	12/18/09 15:03 / stp
Phosphorus, Orthophosphate as P	0.003	mg/L		0.001		E365.1	12/18/09 17:03 / stp
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.8	12/17/09 19:14 / dck
Antimony	0.015	mg/L		0.003		E200.8	12/17/09 19:14 / dck
Arsenic	ND	mg/L		0.003		E200.8	12/17/09 19:14 / dck
Barium	0.047	mg/L		0.005		E200.8	12/17/09 19:14 / dck
Beryllium	ND	mg/L		0.001		E200.8	12/17/09 19:14 / dck
Cadmium	ND	mg/L		0.00008		E200.8	12/17/09 19:14 / dck
Calcium	30	mg/L		1		E200.7	12/17/09 15:42 / sld
Chromium	ND	mg/L		0.001		E200.8	12/17/09 19:14 / dck
Copper	0.114	mg/L		0.001		E200.8	12/17/09 19:14 / dck
Iron	ND	mg/L		0.05		E200.7	12/17/09 15:42 / sld
Lead	0.0010	mg/L		0.0005		E200.8	12/17/09 19:14 / dck
Magnesium	10	mg/L		1		E200.7	12/17/09 15:42 / sld
Manganese	0.105	mg/L		0.005		E200.8	12/17/09 19:14 / dck
Mercury	ND	mg/L		0.00001		E200.8	12/17/09 19:14 / dck
Nickel	ND	mg/L		0.01		E200.8	12/17/09 19:14 / dck
Potassium	ND	mg/L		1		E200.7	12/17/09 15:42 / sld
Selenium	ND	mg/L		0.001		E200.8	12/17/09 19:14 / dck
Silica	5.8	mg/L		0.2		E200.8	12/17/09 19:14 / dck

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



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LABORATORY ANALYTICAL REPORT

Client: Genesis Inc.
Project: Troy Mine
Lab ID: H09120170-003
Client Sample ID Mine Water

Report Date: 01/07/10
Collection Date: 12/16/09 13:10
Date Received: 12/17/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Silver	ND	mg/L		0.0005		E200.8	12/17/09 19:14 / dck
Sodium	3	mg/L		1		E200.7	12/23/09 10:28 / sld
Thallium	ND	mg/L		0.0002		E200.8	12/17/09 19:14 / dck
Uranium	0.0011	mg/L		0.0003		E200.8	12/17/09 19:14 / dck
Zinc	0.02	mg/L		0.01		E200.8	12/17/09 19:14 / dck
METALS, TOTAL							
Iron	ND	mg/L		0.03		E200.7	12/22/09 12:47 / sld

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: A2510 B								Analytical Run: COND_091217A		
Sample ID: ICV1_091217A	Initial Calibration Verification Standard									12/17/09 17:45
Conductivity	716	umhos/cm	1.0	100	90	110				
Method: A2510 B								Batch: 091217A-COND-PROBE-W		
Sample ID: LCS1_091217A	Laboratory Control Sample				Run: COND_091217A				12/17/09 17:46	
Conductivity	1400	umhos/cm	1.0	99	90	110				
Sample ID: H09120186-001ADUP	Sample Duplicate				Run: COND_091217A				12/17/09 17:54	
Conductivity	520	umhos/cm	1.0		0.1	10				
Method: A2540 C								Batch: 7607		
Sample ID: MB-7607	Method Blank				Run: ACCU-124 (14410200)_091218				12/18/09 09:13	
Solids, Total Dissolved TDS @ 180 C	6	mg/L	1							
Sample ID: LCS-7607	Laboratory Control Sample				Run: ACCU-124 (14410200)_091218				12/18/09 09:07	
Solids, Total Dissolved TDS @ 180 C	1010	mg/L	10	100	90	110				
Sample ID: H09120186-002AMS	Sample Matrix Spike				Run: ACCU-124 (14410200)_091218				12/18/09 09:10	
Solids, Total Dissolved TDS @ 180 C	2140	mg/L	10	99	80	120				
Sample ID: H09120186-002AMSD	Sample Matrix Spike Duplicate				Run: ACCU-124 (14410200)_091218				12/18/09 09:11	
Solids, Total Dissolved TDS @ 180 C	2140	mg/L	10	99	80	120	0	10		
Method: A4500-H B								Batch: 091217A-PH-W		
Sample ID: LCS1_091217A	Laboratory Control Sample				Run: PH2_091217A				12/17/09 17:44	
pH	7.06	s.u.	0.10	101	99	101				
Method: A2320 B								Batch: R59017		
Sample ID: MBLK	Method Blank				Run: MAN-TECH_091217A				12/17/09 17:18	
Alkalinity, Total as CaCO3	ND	mg/L	0.9							
Sample ID: LCS	Laboratory Control Sample				Run: MAN-TECH_091217A				12/17/09 17:25	
Alkalinity, Total as CaCO3	570	mg/L	4.0	95	90	110				
Sample ID: H09120170-002A MS	Sample Matrix Spike				Run: MAN-TECH_091217A				12/17/09 20:38	
Alkalinity, Total as CaCO3	730	mg/L	4.0	104	90	110				
Sample ID: H09120170-002A MSD	Sample Matrix Spike Duplicate				Run: MAN-TECH_091217A				12/17/09 20:46	
Alkalinity, Total as CaCO3	710	mg/L	4.0	102	90	110	1.6	20		

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: A4500 S-D								Analytical Run: SUB-B140740		
Sample ID: ICV1_091221A	Initial Calibration Verification Standard									12/21/09 11:30
Sulfide		0.216	mg/L	0.040	106	90	110			
Method: A4500 S-D								Batch: B_091221A-SULFIDE-MB-W		
Sample ID: B09121305-002BMS	Sample Matrix Spike									12/21/09 11:30
Sulfide		9.14	mg/L	1.0	90	70	130			
Sample ID: B09121305-002BMDS	Sample Matrix Spike Duplicate									12/21/09 11:30
Sulfide		8.99	mg/L	1.0	89	70	130	1.6	30	
Sample ID: MBLK1_091221A	Method Blank									12/21/09 11:30
Sulfide		ND	mg/L	0.02						
Sample ID: LFB1_091221A	Laboratory Fortified Blank									12/21/09 11:30
Sulfide		0.220	mg/L	0.040	109	70	130			
Method: A4500-F C								Analytical Run: PH_091218A		
Sample ID: ICV1_091218A	Initial Calibration Verification Standard									12/18/09 13:31
Fluoride		0.746	mg/L	0.10	99	90	110			
Method: A4500-F C								Batch: 091218A-F-ISE-W		
Sample ID: MBLK1_091218A	Method Blank									12/18/09 13:31
Fluoride		ND	mg/L	0.03						
Sample ID: LFB1_091218A	Laboratory Fortified Blank									12/18/09 13:31
Fluoride		0.477	mg/L	0.10	95	90	110			
Sample ID: LCS1_091218A	Laboratory Control Sample									12/18/09 13:31
Fluoride		0.746	mg/L	0.10	96	90	110			
Sample ID: H09120170-001ADUP	Sample Duplicate									12/18/09 13:31
Fluoride		0.0398	mg/L	0.10					20	
Sample ID: H09120170-002AMS	Sample Matrix Spike									12/18/09 13:31
Fluoride		0.753	mg/L	0.10	96	85	115			
Sample ID: H09120170-002AMSD	Sample Matrix Spike Duplicate									12/18/09 13:31
Fluoride		0.745	mg/L	0.10	94	85	115	1.1	20	

Qualifiers:

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QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E300.0								Analytical Run: IC101-H_091218A		
Sample ID: CCV	2	Continuing Calibration Verification Standard								12/17/09 20:34
Chloride		23	mg/L	1.0	92	90	110			
Sulfate		50	mg/L	1.0	101	90	110			
Sample ID: ICV	2	Initial Calibration Verification Standard								12/17/09 20:51
Chloride		2.4	mg/L	1.0	97	90	110			
Sulfate		9.4	mg/L	1.0	94	90	110			
Method: E300.0								Batch: R59015		
Sample ID: LFB	2	Laboratory Fortified Blank								12/17/09 21:07
Chloride		24	mg/L	1.0	98	90	110			
Sulfate		53	mg/L	1.0	106	90	110			
Sample ID: MBLK	2	Method Blank								12/17/09 21:24
Chloride		ND	mg/L	0.05						
Sulfate		ND	mg/L	0.1						
Sample ID: H09120175-001A MS	2	Sample Matrix Spike								12/17/09 23:19
Chloride		26	mg/L	1.0	96	90	110			
Sulfate		61	mg/L	1.0	106	90	110			
Sample ID: H09120175-001A MSD	2	Sample Matrix Spike Duplicate								12/17/09 23:35
Chloride		26	mg/L	1.0	96	90	110	0.2	20	
Sulfate		60	mg/L	1.0	105	90	110	1.1	20	
Method: A5310 C								Analytical Run: SUB-C127879		
Sample ID: ICV-5085		Initial Calibration Verification Standard								12/21/09 09:15
Organic Carbon, Total (TOC)		10.4	mg/L	0.50	104	90	110			
Method: A5310 C								Batch: C_R127879		
Sample ID: MBLK		Method Blank								12/21/09 09:40
Organic Carbon, Total (TOC)		0.06	mg/L	0.04						
Sample ID: LCS-5085		Laboratory Control Sample								12/21/09 14:59
Organic Carbon, Total (TOC)		10.3	mg/L	0.50	103	90	110			
Sample ID: H09120170-003E		Sample Matrix Spike								12/21/09 17:21
Organic Carbon, Total (TOC)		5.14	mg/L	0.50	99	85	115			
Sample ID: H09120170-003E		Sample Matrix Spike Duplicate								12/21/09 17:31
Organic Carbon, Total (TOC)		5.36	mg/L	0.50	103	85	115	4.1	10	

Qualifiers:

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QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E350.1								Analytical Run: NUTRIENTS_091221A		
Sample ID: ICV-1		Initial Calibration Verification Standard								12/21/09 15:16
Nitrogen, Ammonia as N		1.01	mg/L	0.10	101	90	110			
Sample ID: ICB		Initial Calibration Blank, Instrument Blank								12/21/09 15:24
Nitrogen, Ammonia as N		-0.0100	mg/L	0.10		0	0			
Sample ID: CCV-19		Continuing Calibration Verification Standard								12/21/09 15:52
Nitrogen, Ammonia as N		0.530	mg/L	0.10	106	90	110			
Sample ID: CCV-32		Continuing Calibration Verification Standard								12/21/09 16:18
Nitrogen, Ammonia as N		0.530	mg/L	0.10	106	90	110			
Method: E350.1								Batch: A2009-12-21_5_NH3_01		
Sample ID: LCS-2		Laboratory Control Sample								12/21/09 15:18
Nitrogen, Ammonia as N		8.19	mg/L	0.19	103	90	110			
Sample ID: LFB-3		Laboratory Fortified Blank								12/21/09 15:20
Nitrogen, Ammonia as N		1.00	mg/L	0.10	100	90	110			
Sample ID: MBLK-6		Method Blank								12/21/09 15:26
Nitrogen, Ammonia as N		ND	mg/L	0.02						
Sample ID: H09120170-003CMS		Sample Matrix Spike								12/21/09 16:02
Nitrogen, Ammonia as N		1.43	mg/L	0.10	104	90	110			
Sample ID: H09120170-003CMSD		Sample Matrix Spike Duplicate								12/21/09 16:04
Nitrogen, Ammonia as N		1.40	mg/L	0.10	101	90	110	2.1	20	
Sample ID: H09120168-001BDUP		Sample Duplicate								12/21/09 16:22
Nitrogen, Ammonia as N		5.19	mg/L	0.47				0	20	

Qualifiers:

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ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E353.2								Analytical Run: NUTRIENTS_091218A		
Sample ID: ICV-1		Initial Calibration Verification Standard								12/18/09 14:33
Nitrogen, Nitrate+Nitrite as N		1.01	mg/L	0.050	101	90	110			
Sample ID: CCV-4		Continuing Calibration Verification Standard								12/18/09 14:39
Nitrogen, Nitrate+Nitrite as N		0.520	mg/L	0.050	104	90	110			
Sample ID: ICB		Initial Calibration Blank, Instrument Blank								12/18/09 14:43
Nitrogen, Nitrate+Nitrite as N		-0.0100	mg/L	0.050		0	0			
Method: E353.2								Batch: A2009-12-18_5_NO3_01		
Sample ID: LCS-2		Laboratory Control Sample				Run: NUTRIENTS_091218A			12/18/09 14:35	
Nitrogen, Nitrate+Nitrite as N		18.9	mg/L	0.095	103	90	110			
Sample ID: LFB-3		Laboratory Fortified Blank				Run: NUTRIENTS_091218A			12/18/09 14:37	
Nitrogen, Nitrate+Nitrite as N		1.03	mg/L	0.050	103	90	110			
Sample ID: MBLK-7		Method Blank				Run: NUTRIENTS_091218A			12/18/09 14:45	
Nitrogen, Nitrate+Nitrite as N		ND	mg/L	0.010						
Sample ID: H09120170-001CMS		Sample Matrix Spike				Run: NUTRIENTS_091218A			12/18/09 14:57	
Nitrogen, Nitrate+Nitrite as N		0.970	mg/L	0.050	97	90	110			
Sample ID: H09120170-001CMSD		Sample Matrix Spike Duplicate				Run: NUTRIENTS_091218A			12/18/09 14:59	
Nitrogen, Nitrate+Nitrite as N		0.970	mg/L	0.050	97	90	110	0	20	

Qualifiers:

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ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual	
Method: E365.1								Analytical Run: FIA202-HE_091218B			
Sample ID: ICV	Initial Calibration Verification Standard										12/18/09 16:54
Phosphorus, Orthophosphate as P		0.249	mg/L	0.010	100	90	110				
Sample ID: CCV	Continuing Calibration Verification Standard										12/18/09 16:57
Phosphorus, Orthophosphate as P		0.252	mg/L	0.010	101	90	110				
Sample ID: ICB	Initial Calibration Blank, Instrument Blank										12/18/09 16:59
Phosphorus, Orthophosphate as P		0.00103	mg/L	0.010		0	0				
Method: E365.1										Batch: R59073	
Sample ID: LCS	Laboratory Control Sample					Run: FIA202-HE_091218B			12/18/09 16:55		
Phosphorus, Orthophosphate as P		3.11	mg/L	0.010	101	90	110				
Sample ID: LFB	Laboratory Fortified Blank					Run: FIA202-HE_091218B			12/18/09 16:56		
Phosphorus, Orthophosphate as P		0.206	mg/L	0.010	103	90	110				
Sample ID: MBLK	Method Blank					Run: FIA202-HE_091218B			12/18/09 17:00		
Phosphorus, Orthophosphate as P		ND	mg/L	0.001							
Sample ID: H09120170-003AMS	Sample Matrix Spike					Run: FIA202-HE_091218B			12/18/09 17:04		
Phosphorus, Orthophosphate as P		0.196	mg/L	0.010	97	90	110				
Sample ID: H09120170-003AMSD	Sample Matrix Spike Duplicate					Run: FIA202-HE_091218B			12/18/09 17:05		
Phosphorus, Orthophosphate as P		0.196	mg/L	0.010	97	90	110	0.1	20		
Sample ID: H09120198-002ADUP	Sample Duplicate					Run: FIA202-HE_091218B			12/18/09 17:08		
Phosphorus, Orthophosphate as P		0.0306	mg/L	0.010				4.6	20		
Method: E200.7										Batch: 7638	
Sample ID: MB-7638	Method Blank					Run: ICP1-HE_091222A			12/22/09 12:35		
Iron		ND	mg/L	0.008							
Sample ID: LCS-7638	Laboratory Control Sample					Run: ICP1-HE_091222A			12/22/09 12:38		
Iron		2.46	mg/L	0.030	98	85	115				
Sample ID: H09120185-001AMS5	Sample Matrix Spike					Run: ICP1-HE_091222A			12/22/09 12:59		
Iron		2.36	mg/L	0.030	93	70	130				
Sample ID: H09120185-001AMSD5	Sample Matrix Spike Duplicate					Run: ICP1-HE_091222A			12/22/09 13:02		
Iron		2.57	mg/L	0.030	101	70	130	8.4	20		

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7										Analytical Run: ICP1-HE_091217B
Sample ID: ICV	5	Initial Calibration Verification Standard								12/17/09 14:22
Calcium		41.9	mg/L	1.0	105	95	105			
Iron		4.12	mg/L	0.030	103	95	105			
Magnesium		41.8	mg/L	1.0	104	95	105			
Potassium		38.6	mg/L	1.0	97	95	105			
Sodium		40.5	mg/L	1.0	101	95	105			
Sample ID: CCV-1	5	Continuing Calibration Verification Standard								12/17/09 14:30
Calcium		26.0	mg/L	1.0	104	95	105			
Iron		2.59	mg/L	0.030	103	95	105			
Magnesium		25.4	mg/L	1.0	101	95	105			
Potassium		24.3	mg/L	1.0	97	95	105			
Sodium		25.6	mg/L	1.0	102	95	105			
Sample ID: ICSA	5	Interference Check Sample A								12/17/09 14:44
Calcium		556	mg/L	1.0	111	80	120			
Iron		199	mg/L	0.030	99	80	120			
Magnesium		572	mg/L	1.0	114	80	120			
Potassium		0.0146	mg/L	1.0		0	0			
Sodium		0.0725	mg/L	1.0		0	0			
Sample ID: ICSAB	5	Interference Check Sample AB								12/17/09 14:47
Calcium		478	mg/L	1.0	96	80	120			
Iron		186	mg/L	0.030	93	80	120			
Magnesium		528	mg/L	1.0	106	80	120			
Potassium		20.3	mg/L	1.0	102	80	120			
Sodium		21.9	mg/L	1.0	109	80	120			
Sample ID: CCV	5	Continuing Calibration Verification Standard								12/17/09 14:51
Calcium		27.1	mg/L	1.0	108	90	110			
Iron		2.70	mg/L	0.030	108	90	110			
Magnesium		26.3	mg/L	1.0	105	90	110			
Potassium		24.2	mg/L	1.0	97	90	110			
Sodium		25.1	mg/L	1.0	100	90	110			
Sample ID: CCV	5	Continuing Calibration Verification Standard								12/17/09 15:33
Calcium		22.7	mg/L	1.0	91	90	110			
Iron		2.39	mg/L	0.030	95	90	110			
Magnesium		22.9	mg/L	1.0	92	90	110			
Potassium		24.0	mg/L	1.0	96	90	110			
Sodium		25.3	mg/L	1.0	101	90	110			
Sample ID: ICSA	5	Interference Check Sample A								12/17/09 15:54
Calcium		474	mg/L	1.0	95	80	120			
Iron		180	mg/L	0.030	90	80	120			
Magnesium		513	mg/L	1.0	103	80	120			

Qualifiers:

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QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7								Analytical Run: ICP1-HE_091217B		
Sample ID: ICSA	5	Interference Check Sample A								12/17/09 15:54
Potassium		-0.0107	mg/L	1.0		0	0			
Sodium		0.0732	mg/L	1.0		0	0			
Sample ID: ICSAB	5	Interference Check Sample AB								12/17/09 15:57
Calcium		464	mg/L	1.0	93	80	120			
Iron		194	mg/L	0.030	97	80	120			
Magnesium		487	mg/L	1.0	97	80	120			
Potassium		19.9	mg/L	1.0	99	80	120			
Sodium		20.3	mg/L	1.0	101	80	120			
Method: E200.7								Batch: R59005		
Sample ID: MBLK	5	Method Blank								Run: ICP1-HE_091217B 12/17/09 15:02
Calcium		ND	mg/L	0.1						
Iron		0.002	mg/L	0.002						
Magnesium		0.06	mg/L	0.02						
Potassium		ND	mg/L	0.04						
Sodium		ND	mg/L	0.1						
Sample ID: LFB	5	Laboratory Fortified Blank								Run: ICP1-HE_091217B 12/17/09 15:05
Calcium		8.98	mg/L	1.0	90	85	115			
Iron		0.916	mg/L	0.030	91	85	115			
Magnesium		9.14	mg/L	1.0	91	85	115			
Potassium		9.76	mg/L	1.0	98	85	115			
Sodium		9.16	mg/L	1.0	92	85	115			
Sample ID: LCS	5	Laboratory Control Sample								Run: ICP1-HE_091217B 12/17/09 15:08
Calcium		20.2	mg/L	1.0	101	90	110			
Iron		2.04	mg/L	0.030	102	90	110			
Magnesium		20.5	mg/L	1.0	102	90	110			
Potassium		19.7	mg/L	1.0	99	90	110			
Sodium		21.1	mg/L	1.0	106	90	110			
Sample ID: H09120169-001BMS2	5	Sample Matrix Spike								Run: ICP1-HE_091217B 12/17/09 15:24
Calcium		43.8	mg/L	1.0	97	70	130			
Iron		1.91	mg/L	0.030	96	70	130			
Magnesium		66.2	mg/L	1.0	97	70	130			
Potassium		20.0	mg/L	1.0	95	70	130			
Sodium		44.4	mg/L	1.0	103	70	130			
Sample ID: H09120169-001BMSD2	5	Sample Matrix Spike Duplicate								Run: ICP1-HE_091217B 12/17/09 15:27
Calcium		44.7	mg/L	1.0	101	70	130	1.9	20	
Iron		1.93	mg/L	0.030	96	70	130	0.8	20	
Magnesium		64.4	mg/L	1.0	88	70	130	2.7	20	
Potassium		20.1	mg/L	1.0	95	70	130	0.7	20	
Sodium		42.3	mg/L	1.0	93	70	130	4.7	20	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.
Project: Troy Mine

Report Date: 01/07/10
Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7										Analytical Run: ICP1-HE_091222A
Sample ID: ICV		Initial Calibration Verification Standard								12/22/09 10:31
Iron		3.88	mg/L	0.030	97	95	105			
Sample ID: CCV-1		Continuing Calibration Verification Standard								12/22/09 10:39
Iron		2.53	mg/L	0.030	101	95	105			
Sample ID: ICSA		Interference Check Sample A								12/22/09 10:49
Iron		196	mg/L	0.030	98	80	120			
Sample ID: ICSAB		Interference Check Sample AB								12/22/09 10:52
Iron		181	mg/L	0.030	91	80	120			
Sample ID: CCV		Continuing Calibration Verification Standard								12/22/09 12:16
Iron		2.44	mg/L	0.030	98	90	110			
Sample ID: ICSA		Interference Check Sample A								12/22/09 13:58
Iron		180	mg/L	0.030	90	80	120			
Sample ID: ICSAB		Interference Check Sample AB								12/22/09 14:01
Iron		180	mg/L	0.030	90	80	120			
Method: E200.7										Batch: R59144
Sample ID: MBLK		Method Blank								12/22/09 11:09
Iron		0.005	mg/L	0.002						
Sample ID: LFB		Laboratory Fortified Blank								12/22/09 11:12
Iron		1.00	mg/L	0.030	100	85	115			
Sample ID: LCS		Laboratory Control Sample								12/22/09 11:15
Iron		1.96	mg/L	0.030	98	90	110			
Sample ID: H09120198-006EMS2		Sample Matrix Spike								12/22/09 12:22
Iron		1.92	mg/L	0.030	96	70	130			
Sample ID: H09120198-006EMSD2		Sample Matrix Spike Duplicate								12/22/09 12:26
Iron		1.91	mg/L	0.030	95	70	130	0.7	20	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7								Analytical Run: ICP1-HE_091223A		
Sample ID: ICV	Initial Calibration Verification Standard									12/23/09 09:30
Sodium		41.0	mg/L	1.0	103	95	105			
Sample ID: CCV-1	Continuing Calibration Verification Standard									12/23/09 09:38
Sodium		25.6	mg/L	1.0	102	95	105			
Sample ID: ICSA	Interference Check Sample A									12/23/09 09:52
Sodium		0.0690	mg/L	1.0		0	0			
Sample ID: ICSAB	Interference Check Sample AB									12/23/09 09:55
Sodium		21.8	mg/L	1.0	109	80	120			
Sample ID: CCV	Continuing Calibration Verification Standard									12/23/09 09:59
Sodium		26.1	mg/L	1.0	104	90	110			
Sample ID: ICSA	Interference Check Sample A									12/23/09 12:37
Sodium		0.0589	mg/L	1.0		0	0			
Sample ID: ICSAB	Interference Check Sample AB									12/23/09 12:40
Sodium		21.8	mg/L	1.0	109	80	120			
Sample ID: ICSA	Interference Check Sample A									12/23/09 13:57
Sodium		0.0611	mg/L	1.0		0	0			
Sample ID: ICSAB	Interference Check Sample AB									12/23/09 14:01
Sodium		21.3	mg/L	1.0	107	80	120			
Sample ID: ICSA	Interference Check Sample A									12/23/09 20:02
Sodium		0.0665	mg/L	1.0		0	0			
Sample ID: ICSAB	Interference Check Sample AB									12/23/09 20:06
Sodium		20.3	mg/L	1.0	102	80	120			
Method: E200.7								Batch: R59176		
Sample ID: MBLK	Method Blank									12/23/09 10:06
Sodium		ND	mg/L	0.1						
Sample ID: LFB	Laboratory Fortified Blank									12/23/09 10:09
Sodium		10.2	mg/L	1.0	102	85	115			
Sample ID: LCS	Laboratory Control Sample									12/23/09 10:12
Sodium		20.9	mg/L	1.0	105	90	110			
Sample ID: H09120170-003BMS2	Sample Matrix Spike									12/23/09 10:31
Sodium		23.3	mg/L	1.0	103	70	130			
Sample ID: H09120170-003BMDS2	Sample Matrix Spike Duplicate									12/23/09 10:34
Sodium		23.5	mg/L	1.0	104	70	130	0.9	20	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



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QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7										Batch: R59176
Sample ID: LFB		Laboratory Fortified Blank					Run: ICP1-HE_091223A			12/23/09 11:44
Sodium		10.2	mg/L	1.0	102	85	115			

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.
Project: Troy Mine

Report Date: 01/07/10
Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8						Analytical Run: ICPMS204-B_091217A				
Sample ID: QCS-090602A,090609 18 Initial Calibration Verification Standard						12/17/09 16:14				
Aluminum		0.256	mg/L	0.10	102	90	110			
Antimony		0.0503	mg/L	0.050	101	90	110			
Arsenic		0.0503	mg/L	0.0050	101	90	110			
Barium		0.0504	mg/L	0.10	101	90	110			
Beryllium		0.0257	mg/L	0.0010	103	90	110			
Cadmium		0.0261	mg/L	0.0010	104	90	110			
Chromium		0.0499	mg/L	0.010	100	90	110			
Copper		0.0511	mg/L	0.010	102	90	110			
Lead		0.0509	mg/L	0.010	102	90	110			
Manganese		0.254	mg/L	0.010	102	90	110			
Mercury		0.00197	mg/L	0.0010	99	90	110			
Nickel		0.0498	mg/L	0.010	100	90	110			
Selenium		0.0497	mg/L	0.0050	99	90	110			
Silicon		0.500	mg/L	0.10	100	90	110			
Silver		0.0255	mg/L	0.0050	102	90	110			
Thallium		0.0510	mg/L	0.10	102	90	110			
Uranium		0.0199	mg/L	0.0010	100	90	110			
Zinc		0.0511	mg/L	0.010	102	90	110			
Sample ID: ICSA-090423A 18 Interference Check Sample A						12/17/09 16:20				
Aluminum		44.4	mg/L	0.10	111	70	130			
Antimony		0.000263	mg/L	0.050						
Arsenic		0.000181	mg/L	0.0050						
Barium		0.000322	mg/L	0.10						
Beryllium		5.00E-06	mg/L	0.0010						
Cadmium		0.000452	mg/L	0.0010						
Chromium		0.000533	mg/L	0.010						
Copper		0.00151	mg/L	0.010						
Lead		7.50E-05	mg/L	0.010						
Manganese		0.00244	mg/L	0.010						
Mercury		3.00E-05	mg/L	0.0010						
Nickel		0.00123	mg/L	0.010						
Selenium		1.20E-05	mg/L	0.0050						
Silicon		0.00272	mg/L	0.10		0	0			
Silver		0.000206	mg/L	0.0050						
Thallium		3.80E-05	mg/L	0.10						
Uranium		1.10E-05	mg/L	0.0010						
Zinc		0.00113	mg/L	0.010						
Sample ID: ICSAB-090423A,09010 18 Interference Check Sample AB						12/17/09 16:27				
Aluminum		44.6	mg/L	0.10	112	70	130			
Antimony		0.000177	mg/L	0.050		0	0			
Arsenic		0.0111	mg/L	0.0050	111	70	130			
Barium		7.80E-05	mg/L	0.10		0	0			

Qualifiers:

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ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.
Project: Troy Mine

Report Date: 01/07/10
Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8								Analytical Run: ICPMS204-B_091217A		
Sample ID: ICSAB-090423A,09010 18 Interference Check Sample AB										12/17/09 16:27
Beryllium		ND	mg/L	0.0010		0	0			
Cadmium		0.0112	mg/L	0.0010	112	70	130			
Chromium		0.0228	mg/L	0.010	114	70	130			
Copper		0.0233	mg/L	0.010	117	70	130			
Lead		6.80E-05	mg/L	0.010		0	0			
Manganese		0.0250	mg/L	0.010	125	70	130			
Mercury		1.90E-05	mg/L	0.0010		0	0			
Nickel		0.0234	mg/L	0.010	117	70	130			
Selenium		0.0109	mg/L	0.0050	109	70	130			
Silicon		0.00297	mg/L	0.10		0	0			
Silver		0.0213	mg/L	0.0050	106	70	130			
Thallium		3.20E-05	mg/L	0.10		0	0			
Uranium		7.00E-06	mg/L	0.0010		0	0			
Zinc		0.0120	mg/L	0.010	120	70	130			
Sample ID: QCS-090602A,090609 18 Initial Calibration Verification Standard										12/17/09 21:27
Aluminum		0.255	mg/L	0.10	102	90	110			
Antimony		0.0508	mg/L	0.050	102	90	110			
Arsenic		0.0502	mg/L	0.0050	100	90	110			
Barium		0.0513	mg/L	0.10	103	90	110			
Beryllium		0.0252	mg/L	0.0010	101	90	110			
Cadmium		0.0266	mg/L	0.0010	106	90	110			
Chromium		0.0510	mg/L	0.010	102	90	110			
Copper		0.0521	mg/L	0.010	104	90	110			
Lead		0.0502	mg/L	0.010	100	90	110			
Manganese		0.252	mg/L	0.010	101	90	110			
Mercury		0.00190	mg/L	0.0010	95	90	110			
Nickel		0.0513	mg/L	0.010	103	90	110			
Selenium		0.0501	mg/L	0.0050	100	90	110			
Silicon		0.490	mg/L	0.10	98	90	110			
Silver		0.0260	mg/L	0.0050	104	90	110			
Thallium		0.0509	mg/L	0.10	102	90	110			
Uranium		0.0199	mg/L	0.0010	99	90	110			
Zinc		0.0511	mg/L	0.010	102	90	110			
Sample ID: ICSA-090423A 18 Interference Check Sample A										12/17/09 21:34
Aluminum		43.8	mg/L	0.10	109	70	130			
Antimony		0.000268	mg/L	0.050						
Arsenic		0.000182	mg/L	0.0050						
Barium		0.000356	mg/L	0.10						
Beryllium		ND	mg/L	0.0010						
Cadmium		0.000439	mg/L	0.0010						
Chromium		0.000544	mg/L	0.010						
Copper		0.00145	mg/L	0.010						

Qualifiers:

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QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8										Analytical Run: ICPMS204-B_091217A
Sample ID: ICSA-090423A	18	Interference Check Sample A								12/17/09 21:34
Lead		8.60E-05	mg/L	0.010						
Manganese		0.00246	mg/L	0.010						
Mercury		3.20E-05	mg/L	0.0010						
Nickel		0.00126	mg/L	0.010						
Selenium		9.80E-05	mg/L	0.0050						
Silicon		0.00267	mg/L	0.10		0	0			
Silver		0.000117	mg/L	0.0050						
Thallium		5.10E-05	mg/L	0.10						
Uranium		8.00E-06	mg/L	0.0010						
Zinc		0.00113	mg/L	0.010						
Sample ID: ICSAB-090423A,09010	18	Interference Check Sample AB								12/17/09 21:41
Aluminum		42.7	mg/L	0.10	107	70	130			
Antimony		0.000182	mg/L	0.050		0	0			
Arsenic		0.0113	mg/L	0.0050	113	70	130			
Barium		7.30E-05	mg/L	0.10		0	0			
Beryllium		2.00E-06	mg/L	0.0010		0	0			
Cadmium		0.0110	mg/L	0.0010	110	70	130			
Chromium		0.0226	mg/L	0.010	113	70	130			
Copper		0.0232	mg/L	0.010	116	70	130			
Lead		7.60E-05	mg/L	0.010		0	0			
Manganese		0.0244	mg/L	0.010	122	70	130			
Mercury		2.10E-05	mg/L	0.0010		0	0			
Nickel		0.0233	mg/L	0.010	116	70	130			
Selenium		0.0109	mg/L	0.0050	109	70	130			
Silicon		0.00276	mg/L	0.10		0	0			
Silver		0.0204	mg/L	0.0050	102	70	130			
Thallium		4.20E-05	mg/L	0.10		0	0			
Uranium		5.00E-06	mg/L	0.0010		0	0			
Zinc		0.0112	mg/L	0.010	112	70	130			
Method: E200.8										Batch: R59022
Sample ID: LRB	18	Method Blank								Run: ICPMS204-B_091217A
Aluminum		0.0001	mg/L	7E-05						12/17/09 17:20
Antimony		ND	mg/L	3E-05						
Arsenic		ND	mg/L	2E-05						
Barium		4E-05	mg/L	3E-05						
Beryllium		ND	mg/L	1E-05						
Cadmium		ND	mg/L	1E-05						
Chromium		6E-05	mg/L	5E-05						
Copper		5E-05	mg/L	3E-05						
Lead		ND	mg/L	5E-06						
Manganese		5E-05	mg/L	2E-05						
Mercury		ND	mg/L	7E-06						

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc.

Report Date: 01/07/10

Project: Troy Mine

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8										Batch: R59022
Sample ID: LRB	18	Method Blank				Run: ICPMS204-B_091217A				12/17/09 17:20
Nickel		ND	mg/L	3E-05						
Selenium		ND	mg/L	3E-05						
Silicon		0.002	mg/L	0.0004						
Silver		8E-05	mg/L	2E-05						
Thallium		ND	mg/L	1E-05						
Uranium		ND	mg/L	1E-05						
Zinc		0.0009	mg/L	0.0001						
Sample ID: LFB	18	Laboratory Fortified Blank				Run: ICPMS204-B_091217A				12/17/09 17:27
Aluminum		0.0504	mg/L	0.10	100	85	115			
Antimony		0.0491	mg/L	0.050	98	85	115			
Arsenic		0.0485	mg/L	0.0050	97	85	115			
Barium		0.0487	mg/L	0.10	97	85	115			
Beryllium		0.0510	mg/L	0.0010	102	85	115			
Cadmium		0.0487	mg/L	0.0010	97	85	115			
Chromium		0.0490	mg/L	0.010	98	85	115			
Copper		0.0485	mg/L	0.010	97	85	115			
Lead		0.0493	mg/L	0.010	99	85	115			
Manganese		0.0499	mg/L	0.010	100	85	115			
Mercury		0.000950	mg/L	0.0010	95	85	115			
Nickel		0.0486	mg/L	0.010	97	85	115			
Selenium		0.0497	mg/L	0.0050	99	85	115			
Silicon		0.202	mg/L	0.10	100	85	115			
Silver		0.0185	mg/L	0.0050	92	85	115			
Thallium		0.0494	mg/L	0.10	99	85	115			
Uranium		0.0503	mg/L	0.0010	101	85	115			
Zinc		0.0497	mg/L	0.010	97	85	115			
Sample ID: H09120173-001AMS	18	Sample Matrix Spike				Run: ICPMS204-B_091217A				12/17/09 19:27
Aluminum		0.0583	mg/L	0.10	103	70	130			
Antimony		0.0510	mg/L	0.0030	101	70	130			
Arsenic		0.0535	mg/L	0.0010	100	70	130			
Barium		0.0780	mg/L	0.10	98	70	130			
Beryllium		0.0523	mg/L	0.0010	105	70	130			
Cadmium		0.0491	mg/L	0.0010	98	70	130			
Chromium		0.0493	mg/L	0.010	97	70	130			
Copper		0.0506	mg/L	0.010	94	70	130			
Lead		0.0543	mg/L	0.0010	100	70	130			
Manganese		0.102	mg/L	0.010	98	70	130			
Mercury		0.000970	mg/L	0.00020	97	70	130			
Nickel		0.0478	mg/L	0.010	95	70	130			
Selenium		0.0532	mg/L	0.0050	106	70	130			
Silicon		11.0	mg/L	0.10		70	130			A
Silver		0.0176	mg/L	0.0050	88	70	130			

Qualifiers:

RL - Analyte reporting limit.

A - The analyte level was greater than four times the spike level. In accordance with the method % recovery is not calculated.

ND - Not detected at the reporting limit.



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QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8										Batch: R59022
Sample ID: H09120173-001AMS	18	Sample Matrix Spike				Run: ICPMS204-B_091217A				12/17/09 19:27
Thallium		0.0502	mg/L	0.0010	100	70	130			
Uranium		0.0743	mg/L	0.0010	102	70	130			
Zinc		0.160	mg/L	0.010	81	70	130			
Sample ID: H09120173-001AMSD	18	Sample Matrix Spike Duplicate				Run: ICPMS204-B_091217A				12/17/09 19:33
Aluminum		0.0569	mg/L	0.10	100	70	130			20
Antimony		0.0507	mg/L	0.0030	101	70	130	0.6		20
Arsenic		0.0546	mg/L	0.0010	102	70	130	2		20
Barium		0.0779	mg/L	0.10	98	70	130			20
Beryllium		0.0522	mg/L	0.0010	104	70	130	0.4		20
Cadmium		0.0486	mg/L	0.0010	97	70	130	1		20
Chromium		0.0497	mg/L	0.010	98	70	130	0.9		20
Copper		0.0509	mg/L	0.010	95	70	130	0.6		20
Lead		0.0539	mg/L	0.0010	100	70	130	0.7		20
Manganese		0.101	mg/L	0.010	96	70	130	0.9		20
Mercury		0.000949	mg/L	0.00020	95	70	130	2.2		20
Nickel		0.0479	mg/L	0.010	95	70	130	0.3		20
Selenium		0.0531	mg/L	0.0050	106	70	130	0.2		20
Silicon		10.8	mg/L	0.10		70	130	2.3		20 A
Silver		0.0177	mg/L	0.0050	88	70	130	0.2		20
Thallium		0.0497	mg/L	0.0010	99	70	130	1.2		20
Uranium		0.0747	mg/L	0.0010	102	70	130	0.5		20
Zinc		0.159	mg/L	0.010	79	70	130	0.6		20

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

A - The analyte level was greater than four times the spike level. In accordance with the method % recovery is not calculated.



QA/QC Summary Report

Client: Genesis Inc.

Project: Troy Mine

Report Date: 01/07/10

Work Order: H09120170

Analyte	Count	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8								Analytical Run: ICPMS204-B_091218A		
Sample ID: QCS-090602A,090609	2	Initial Calibration Verification Standard								12/18/09 11:45
Manganese		0.255	mg/L	0.010	102	90	110			
Silicon		0.480	mg/L	0.10	96	90	110			
Sample ID: ICSA-090423A	2	Interference Check Sample A								12/18/09 11:52
Manganese		0.00257	mg/L	0.010						
Silicon		0.00255	mg/L	0.10		0	0			
Sample ID: ICSAB-090423A,09010	2	Interference Check Sample AB								12/18/09 11:58
Manganese		0.0251	mg/L	0.010	125	70	130			
Silicon		0.00270	mg/L	0.10		0	0			
Method: E200.8								Batch: R59047		
Sample ID: LRB	2	Method Blank								12/18/09 12:39
Manganese		6E-05	mg/L	2E-05						
Silicon		0.002	mg/L	0.0004						
Sample ID: LFB	2	Laboratory Fortified Blank								12/18/09 12:46
Manganese		0.0500	mg/L	0.010	100	85	115			
Silicon		0.200	mg/L	0.10	99	85	115			
Sample ID: H09120175-001BMS	2	Sample Matrix Spike								12/18/09 13:12
Manganese		1.30	mg/L	0.010		70	130			A
Silicon		11.9	mg/L	0.10		70	130			A
Sample ID: H09120175-001BMSD	2	Sample Matrix Spike Duplicate								12/18/09 13:19
Manganese		1.29	mg/L	0.010		70	130	0.4	20	A
Silicon		12.2	mg/L	0.10		70	130	2.3	20	A

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

A - The analyte level was greater than four times the spike level. In accordance with the method % recovery is not calculated.



Energy Laboratories Inc

Workorder Receipt Checklist



H09120170

Genesis Inc Troy Mine

Login completed by: Tracy L. Lorash

Date and Time Received: 12/17/2009 9:48 AM

Reviewed by: BL2000\wjohanson

Received by: rlt

Reviewed Date: 12/21/2009 9:20:08 AM

Carrier name: Hand Del

Shipping container/cooler in good condition?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	Not Present <input type="checkbox"/>
Custody seals intact on shipping container/cooler?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	Not Present <input type="checkbox"/>
Custody seals intact on sample bottles?	Yes <input type="checkbox"/>	No <input type="checkbox"/>	Not Present <input checked="" type="checkbox"/>
Chain of custody present?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Chain of custody signed when relinquished and received?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Chain of custody agrees with sample labels?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Samples in proper container/bottle?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Sample containers intact?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Sufficient sample volume for indicated test?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
All samples received within holding time?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Container/Temp Blank temperature:	1.6°C On Ice		
Water - VOA vials have zero headspace?	Yes <input type="checkbox"/>	No <input type="checkbox"/>	No VOA vials submitted <input checked="" type="checkbox"/>
Water - pH acceptable upon receipt?	Yes <input type="checkbox"/>	No <input checked="" type="checkbox"/>	Not Applicable <input type="checkbox"/>

Contact and Corrective Action Comments:

All dissolved metals/hardness/SAR samples filtered and preserved w/ 1ml. HNO3 in lab. Ferric Iron sample poured off and preserved in lab w/ 1ml. HNO3. TL



Chain of Custody and Analytical Request Record

PLEASE PRINT - Provide as much information as possible.

Company Name: CDM		Project Name, PWS, Permit, Etc. TROY MINE		Sample Origin State: MT		EPA/State Compliance: Yes <input type="checkbox"/> No <input checked="" type="checkbox"/>					
Report Mail Address: 50 W 14TH ST, # 200 HELENA, MT 59607		Contact Name: DUSTIN KEMPPEL		Phone/Fax: (406) 441-1405		Email: KLEMPPELDT@CDM.COM		Sampler: (Please Print) DUSTIN KLEMPPEL			
Invoice Address: TROY MINE PO BOX 1660 TROY, MT 59935		Invoice Contact & Phone: PAUL KUKAY (406) 295-5882		x221 PKUKAY@HOTMAIL.COM		Purchase Order:		Quote/Bottle Order:			
Special Report/Formats - ELI must be notified prior to sample submittal for the following: <input type="checkbox"/> DW <input type="checkbox"/> A2LA <input type="checkbox"/> GSA <input type="checkbox"/> EDD/EDT (Electronic Data) <input type="checkbox"/> POTW/MMWTP Format: <input type="checkbox"/> State: <input type="checkbox"/> LEVEL IV <input type="checkbox"/> Other: <input type="checkbox"/> NELAC		Number of Containers Sample Type: A W S V B O Vegetation Bioassay Other		ANALYSIS REQUESTED				Contact ELI prior to RUSH sample submittal for charges and scheduling - See Instruction Page		Shipped by: Hand	
				SEE ATTACHED				Comments:		Cooler ID(s):	
MATRIX		Collection Date		Collection Time		Normal Turnaround (TAT)		RUSH		Receipt Temp 1.6 °C	
1 MW 95-4		12/16/09		13:20		X X		X X		On Ice: <input checked="" type="checkbox"/> Yes <input type="checkbox"/> No	
2 DECANT POND		12/16/09		11:10		X X		X X		Custody Seal: Y N Bottles/Coolers: B C	
3 MINE WATER		12/16/09		13:10		X X		X X		Intact: Y N Signature Match: Y N	
4										LABORATORY USE ONLY 409120170	
5											
6											
7											
8											
9											
10											
Custody Record MUST be Signed		Relinquished by (print): DUSTIN KEMPPEL		Date/Time: 12/17/09 09:43		Signature: Dd Kk		Received by (print):		Date/Time: Signature:	
Sample Disposal:		Return to Client:		Lab Disposal:		Received by Laboratory: ROXANNE WILKINS		Date/Time: 12.17.09 9:48		Signature: Roxanne Wilkins	

In certain circumstances, samples submitted to Energy Laboratories, Inc. may be subcontracted to other certified laboratories in order to complete the analysis requested. This serves as notice of this possibility. All sub-contract data will be clearly notated on your analytical report. Visit our web site at www.energylab.com for additional information, downloadable fee schedule, forms, and links.

Attachment 2

Photographs



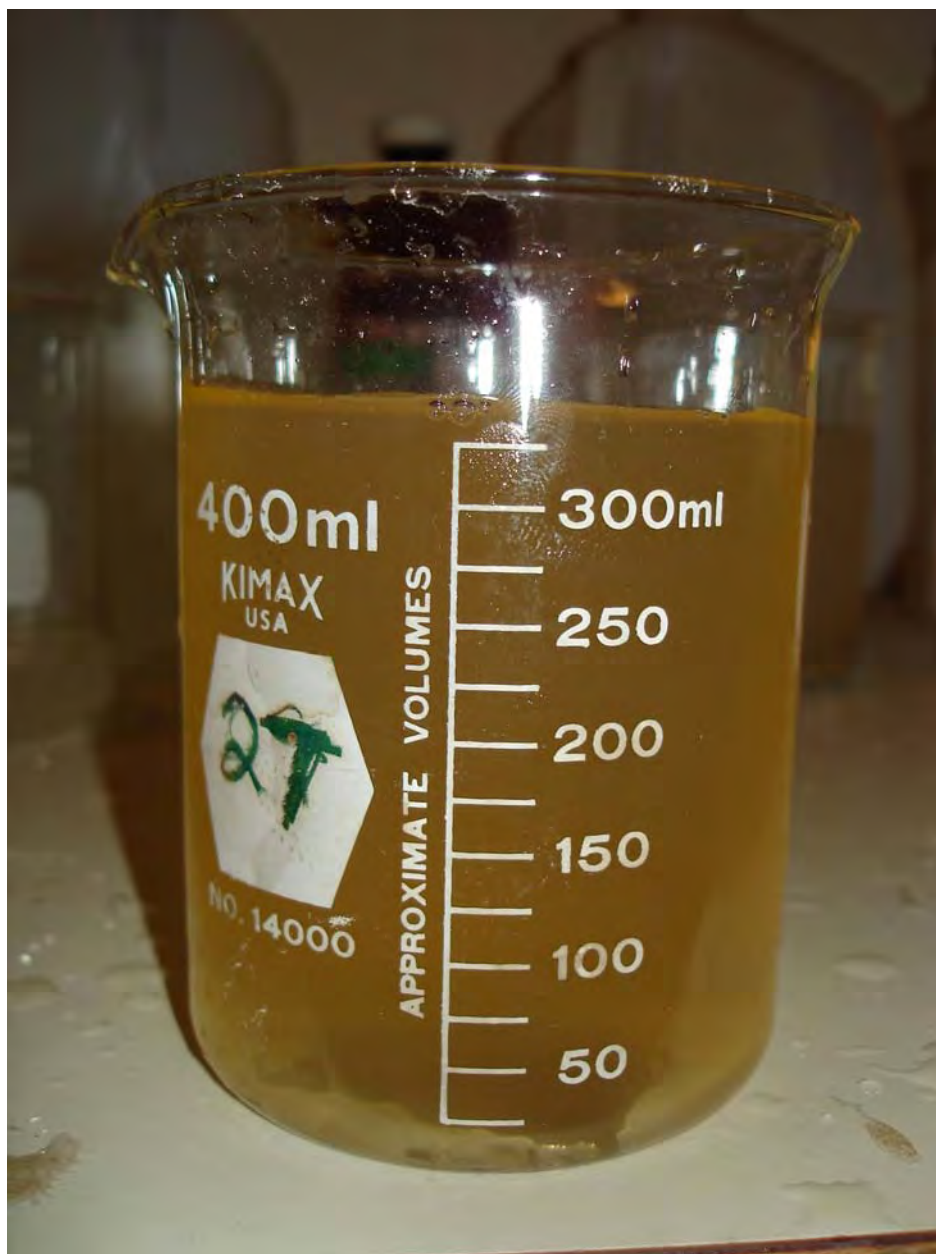
Monitoring Well MW95-4.



Jar Test Sample from MW95-4.



Mixed Jar Sample #5 (mixture of MW95-4 and Mine Water).



Test sample of MW95-4 water with peroxide added – to show presence of iron.



Test Jar #4 after reaction period. Bottle on left is unfiltered. Bottle on right is filtered and preserved.

Appendix E
Troy Mine Mill Site Conceptual
Channel Design



50 West 14th Street, Suite 200
Helena, Montana 59601
tel: 406 441-1400
fax: 406 449-7725

Draft Memorandum

*To: Emily Corsi, Herb Rolfes – DEQ
Bobbie Lacklen, John McKay - KNF*

From: Bill Bucher- CDM

Date: December 3, 2010

Subject: Troy Mine Mill Site Conceptual Channel Design

CDM has prepared a design and cost estimate for a conceptual drainage channel design at the Troy Mine mill site to support the *Draft Environmental Assessment* currently being prepared for final reclamation of the mine. The mill pad and waste rock dump located just below the mine portal are built across a drainage that enters the site from the west. Currently water from this drainage is collected in a detention basin on the upslope side of the millsite and routed beneath the mill site through two culverts to the downslope side of the site where it enters Stanley Creek. The culverts are over 600 feet long and could not be maintained with any assurance in the long term. Under the Agency-Mitigated Alternative for the reclamation plan, these culverts will be plugged and the drainage channel reconstructed over or around the mill pad/waste rock dump.

The purpose of this memorandum is to present a conceptual design for the new channel and provide approximate reclamation costs for construction of the new channel. Maintenance of the new channel is designed to be minimal and should not extend beyond the first few years after construction.

Alternative Considered but not Selected

The most desirable solution for channel design is removal of the fill at the mill site patio and construction of a channel along the original natural alignment at the original grade. This solution would eliminate the difficulty of maintaining a stable channel on the face of the fill and allow for a construction of a naturally functioning channel, which could minimize long-term maintenance costs. However, rebuilding the original channel for the larger drainage will require removing a large portion of the mill site fill. It will also require relocating the access road and buried pipeline that are currently at the mill site.

Two scenarios were investigated to determine the feasibility of establishing a more natural channel gradient across the site. In the first scenario, the existing twin culverts would be

plugged and the surrounding depression filled. The channel would cross the filled depression and follow an approximate 4H:1V grade to the toe of the mill site fill slope at the location of the original channel. For estimation purposes, it is assumed that the bottom width of the cut would be 24 feet to accommodate the new channel, and the side slopes of the cut would rise at 2H:1V. An AutoCAD drawing of the required cut is attached as Figure 1. The volume of cut would be 71,000 cy. The excavated material would be transported to the relatively flat portion of the mill site, where it would be placed in lifts and compacted to form a fill at a slope no greater than 2H:1V.

A second scenario would plug the twin culverts but begin excavation approximately at the top of the culverts. The channel would be excavated at a slightly lesser gradient because it is starting at a lower elevation, and would terminate at the some point in the original channel. This layout results in about 110,000 cy of excavation. Although this channel is presumably near the original channel grade, it may not replicate the original channel because we do not know the alignment of the original channel.

If we assume the excavation can be accomplished for about \$4 per cubic yard, the cost of excavation would be in the range of \$280,000 to \$440,000. This cost would be in addition to costs for building the new channel and revegetating the fresh cut slopes. Because of the greatly increased cost of this alternative compared to the costs of the options discussed below, this alternative was not pursued further.

Selected Alternative

Two options for the selected alternative were considered for evaluation as shown in the attached Figure 2. Both options originate from the end of the natural channel above the current detention basin, which will be backfilled during reclamation. Option 1 routes the drainage in an easterly direction across the top of the reclaimed mill pad/waste rock dump, down the face of the dump and into an energy dissipation basin before discharging to the original channel and Stanley Creek. These options are shown on the attached layout for the mill site.

Option 2 routes the water to the northeast on reclaimed ground until it reaches the north end of the mill pad/waste rock dump where it descends a steep slope. At the base of the steepest portion of the slope, an energy dissipation basin will reduce the velocity of the stream before it is routed through a constructed channel the remaining distance to Stanley Creek.

There is a minor drainage that enters the southern portion of the mill site that would also be routed across the mill patio to the road. Then it would be routed on the uphill (west) side of the road to the larger channel.. This would require shaping a small channel through the reclaimed area and lining it with appropriately sized rock. This design applies to Option 1 or Option 2. These costs are incidental compared to the costs for the main drainage and are not considered further in this memorandum.

Hydrology and Hydraulics

The design storm for permanent drainage channels at mine sites in Montana is the 100-year 24-hour storm. The Forest Service sizes culverts for the 100-year peak flow, which is the event that occurs with a probability of 1% in any given year. The 100-year flow was calculated using the USGS *Methods for Estimating Flood Frequency in Montana Based on Data through Water Year 1998* (Parrett and Johnson, 2004). It was assumed that a burn could occur in the even-aged forest in this tributary drainage so a forest cover factor of 50% was used to represent 50% crown kill. The calculated 100-year flow under this condition is 33 cfs. An attachment provides documentation of the calculation.

Hydraulic calculations were performed using the Bathurst resistance calculation method as described in the *Stormwater Collection Systems Design Handbook* (Mays, 2001, Section 16.5). This iterative method determines the rock size necessary to prevent movement of the rock with a selected factor of safety. A typical industry standard factor of safety of 1.3 was selected for this calculation. This method is specifically meant for steep (greater than 10%) channel slopes where Manning's equation is not valid. Calculations show that a channel with a bottom width of 3 feet and side slopes of 2H:1V will be more than adequate to accommodate the calculated peak flow. The attached sketch shows the typical cross-sections for gentler and steeper portions of the channel.

For the portions of the channel with a slope of less than 10%, a rock size of $d_{50} = 16$ inches was calculated. For the steeper portions of the channels (77% for Option 1 and 50% for Option 2) a $d_{50} = 5.6$ feet is required. The latter gradation will range from a d_{10} of 2.8 feet to a d_{90} of 7 feet. Excel spreadsheet calculations are attached.

Conceptual Design

The channel options were designed as rock-lined channels using guidance from the USACE manual *Hydraulic Design of Flood Control Channels* (ASCE, 1995). The design includes a 9-inch thick filter bed of $d_{50} = 6$ inches beneath the riprap. Below this will be 1-inch minus bedding to protect the liner to be constructed where the channel crosses the mill patio. The liner will prevent infiltration of water into the mill fill. This liner could be constructed of HDPE, geosynthetic clay liner (GCL), or clay. Bedding material may need to be placed below the liner as well to provide a suitable surface for installation.

On the steep portion of the channel, installation of a liner will not be feasible. Riprap should be underlain by rock filter layers that prevent scour of the underlying material. It may be that appropriate sized material is already in place in this portion of the mill site, and it will not be necessary to place filter layers in that case.

The design includes a four foot culvert for passage of the drainage beneath the realigned and reconstructed road. A four foot culvert is oversized for the design flow and slope but is preferred to smaller sizes because of its ability to pass debris more readily.

The energy dissipation basin at the base of the steep slopes is intended to reduce velocities before a milder channel routes the discharge to Stanley Creek. The final design of this basin is recommended to be in accordance with US Department of Transportation (USDOT) guidelines (USDOT, 1975) for energy dissipaters. For purposes of this conceptual design, it is assumed that a basin 30 feet wide and 50 feet long will be more than adequate to dissipate energy from the design storm. The rock size for the basin is assumed to have a d_{50} of 16 inches but this may need to be adjusted during final design.

Although this channel has been designed for cost estimating purposes as a riprap channel, modifications can be made to provide a more natural appearance and function, especially in the gentler reaches. Woody debris can be incorporated between the rocks, rocks can be arranged to create step pools, and shrubs and trees can be planted between rocks on the banks. However, the framework of large rock sizes is still needed throughout the channel to provide stability under high flows. Addition of these items will increase installation time but should not significantly affect material costs relative to the high cost of the rock.

Both options present construction problems with using large rock on steep slopes. The 2H:1V slope of Option 2 is tractable with conventional equipment. Placement of large rock on the angle of repose slope of Option 1 will be more difficult. Equipment tethered from above may be able to shape the channel and place the required large rock. If sufficiently large rock is already present in the face of the dump, it may not be necessary to place rock here, making this option more attractive. However, without incising a channel in the face with equipment there will be no certainty that the discharge will be funneled to the energy dissipation basin.

Cost Estimate

A cost estimate for each option was prepared and these are attached to this memorandum. The largest and most uncertain cost is the cost of providing and placing rock. It is assumed that the $d_{50} = 16$ -inch and $d_{50} = 6$ -inch gradations can be found in the waste rock dump or other nearby areas and a cost of \$30/cy will be sufficient for collection and placement of these materials. However, the $d_{50} = 5$ -foot plus rock may need to be imported from a greater distance. A cost of \$60/cy has been used for this rock assuming it can be found within 30 miles.

Other elements accounted for in the cost estimates are mobilization, erosion control, clearing and grubbing of timbered areas, installation of a culvert, and construction of an access road from the existing access road to construct the energy dissipater. Construction costs include a 15% contingency. Engineering design and oversight costs are estimated at 20% of construction costs. The total cost estimate for Option 1 is about \$238,000 and the estimated total cost of Option 2 is \$290,000. The attached cost sheets present breakdowns of the estimated costs of the two options.

References

- American Society of Civil Engineers, 1995. Hydraulic Design of Flood Control Channels. Technical Engineering and Design Guides as Adapted from the US Army Corps of Engineers, No. 10. New York, N. Y.
- Mays, L., 2001. Stormwater Collection Systems Design Handbook.
- Parrett, C. and D.L.R. Johnson, 2004. Methods for Estimating Flood Frequency in Montana Based on Data through Water Year 1998. US Geological Survey Water-Resources Investigations Report 03-4308, Helena MT. February 2004.
- US Department of Transportation, 1975. Hydraulic Design of Energy Dissipators for Culverts and Cahnnels. Hydraulic Engineering Circular No. 14. June 1975.

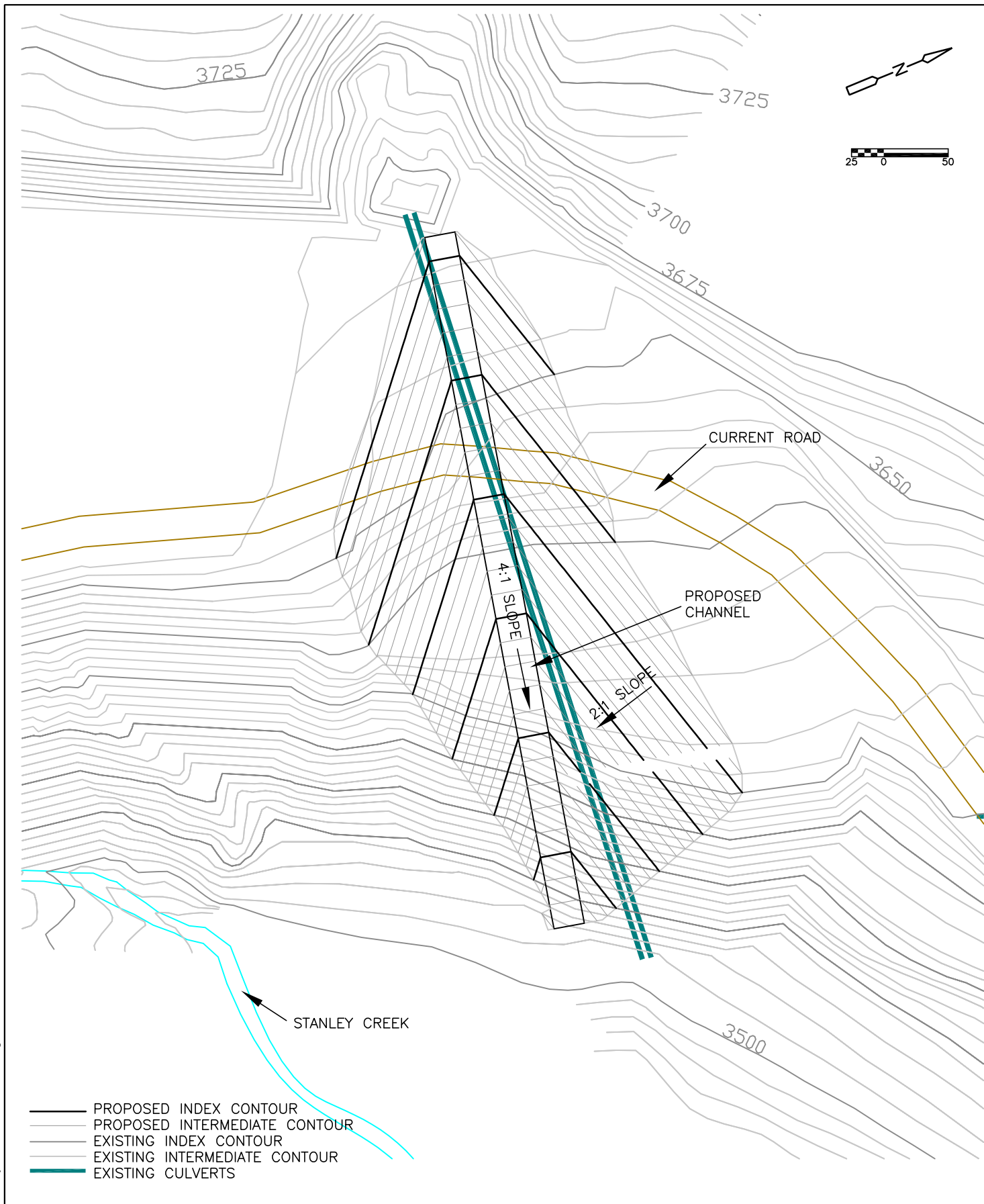


FIGURE 1
Excavation for 4:1 Channel Slope
Troy Mine Waterline Conceptual
Channel Design

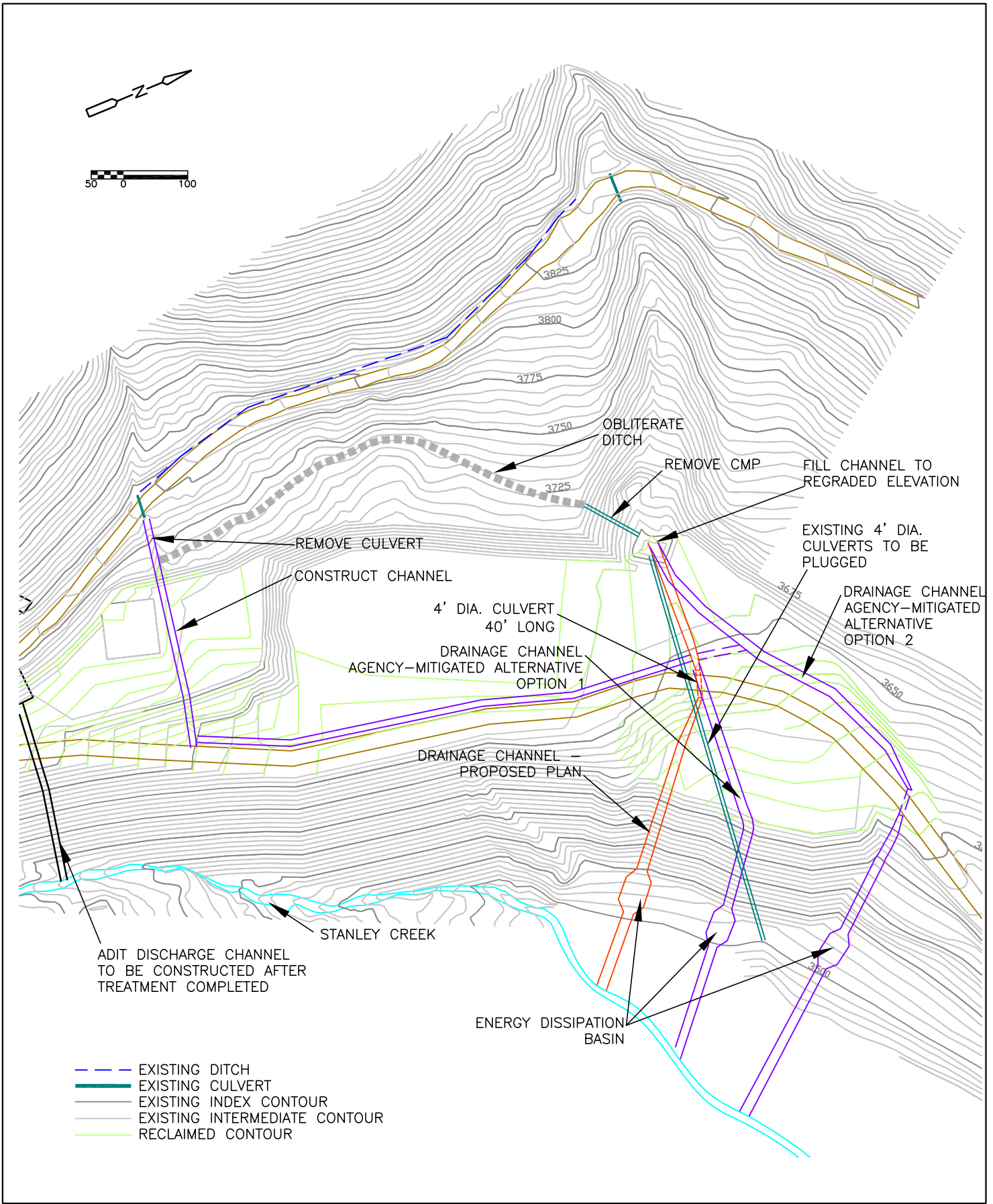


FIGURE 2
Channel Options
Troy Mine Conceptual
Channel Design

Determine peak flow of 100 year event at mill site drainage. Use NR 03-4308 (Parrelle Johnson, 2004) to calculate 100-yr flow.

Drainage Area = 0.179 sq. mi. (USGS topo)
 Annual precip = 24" (Troy average precip)
 Forest cover = 50% Burned condition with 50% crown kill.

Use West Region Equation:

$$Q_{100} = 18.7 A^{0.812} P^{1.06} (F+1)^{-0.664}$$

$$Q_{100} = 18.7 \times (0.179)^{0.812} (24)^{1.06} (51)^{-0.664}$$

$$Q_{100} = 33 \text{ cfs.}$$

Steep Slope Riprap Design

Ref: Stormwater Collection Systems Design Handbook, Larry Mays, 2001, Section 16.5

Troy - Mine, Millpad Channel, 10% slope

Target Flow = 33 cfs

Channel Dimensions

Bottom Width(ft) = 3.28
Bottom Width(m) = 1
Sideslopes = 2 :1(H:V)
Slope (m/m) = 0.1
Length (ft) = 581
Length (m) = 177

Trial Lining

D90 (ft) = 1.64
D90 (m) = 0.5
D50 (ft) = 1.31
D50 (m) = 0.4
D10 (ft) = 0.66
D10 (m) = 0.2

Range of Depth to be Evaluated

Least Depth (m) = 0.1
Maximum Depth (m)= 2
Increment (m) = 0.02

Angles

Side Slope, degrees 26.57
Bed Slope, degrees 5.74
Riprap Angle of
Repose, degrees 43
(Get from Charts)

Factor of Safety 1.3

Hydraulic Calculations

Depth		Area		Wetted Perimeter		Top Width		Hydraulic Radius	
(ft)	(m)	(sf)	(sm)	(ft)	(m)	(ft)	(m)	(ft)	(m)
0.328	0.100	1.292	0.120	4.748	1.447	4.593	1.4	0.272	0.083
0.394	0.120	1.602	0.149	5.042	1.537	4.856	1.48	0.318	0.097
0.459	0.140	1.929	0.179	5.335	1.626	5.118	1.56	0.362	0.110
0.525	0.160	2.273	0.211	5.628	1.716	5.381	1.64	0.404	0.123
0.591	0.180	2.635	0.245	5.922	1.805	5.643	1.72	0.445	0.136
0.656	0.200	3.014	0.280	6.215	1.894	5.906	1.8	0.485	0.148
0.722	0.220	3.410	0.317	6.509	1.984	6.168	1.88	0.524	0.160
0.787	0.240	3.823	0.355	6.802	2.073	6.430	1.96	0.562	0.171
0.853	0.260	4.254	0.395	7.096	2.163	6.693	2.04	0.600	0.183
0.919	0.280	4.702	0.437	7.389	2.252	6.955	2.12	0.636	0.194
0.984	0.300	5.167	0.480	7.683	2.342	7.218	2.2	0.673	0.205
1.050	0.320	5.649	0.525	7.976	2.431	7.480	2.28	0.708	0.216
1.115	0.340	6.148	0.571	8.269	2.521	7.743	2.36	0.744	0.227
1.181	0.360	6.665	0.619	8.563	2.610	8.005	2.44	0.778	0.237
1.247	0.380	7.199	0.669	8.856	2.699	8.268	2.52	0.813	0.248
1.312	0.400	7.750	0.720	9.150	2.789	8.530	2.6	0.847	0.258
1.378	0.420	8.318	0.773	9.443	2.878	8.793	2.68	0.881	0.268
1.444	0.440	8.904	0.827	9.737	2.968	9.055	2.76	0.914	0.279
1.509	0.460	9.507	0.883	10.030	3.057	9.318	2.84	0.948	0.289
1.575	0.480	10.127	0.941	10.324	3.147	9.580	2.92	0.981	0.299
1.640	0.500	10.764	1.000	10.617	3.236	9.843	3	1.014	0.309
1.706	0.520	11.418	1.061	10.910	3.326	10.105	3.08	1.047	0.319
1.772	0.540	12.090	1.123	11.204	3.415	10.367	3.16	1.079	0.329
1.837	0.560	12.779	1.187	11.497	3.504	10.630	3.24	1.111	0.339
1.903	0.580	13.485	1.253	11.791	3.594	10.892	3.32	1.144	0.349
1.969	0.600	14.208	1.320	12.084	3.683	11.155	3.4	1.176	0.358
2.034	0.620	14.949	1.389	12.378	3.773	11.417	3.48	1.208	0.368
2.100	0.640	15.707	1.459	12.671	3.862	11.680	3.56	1.240	0.378
2.165	0.660	16.482	1.531	12.965	3.952	11.942	3.64	1.271	0.387

Note: Gradation calculation (D90 and D10) based on USFS Specs. for Roads and Bridges

Bathurst Resistance Calculations

Depth (ft)	Depth (m)	b	c1	f(FR)	c2	f(REG)	f(CG)	V*	Sum of f (functions) n	V/V*	Velocity (V) m/sec	Froude Number	Correction for slope 0.98			
													Flow, Q m^3/sec	Flow, Q ft^3/sec	To N/m^2	
0.33	0.1	0.180	0.624	1.319 ^{1/4} c1	1.745	3.233	0.602	0.313	2.568	0.082	2.650	0.830	0.838	0.100	3.5	80.9
0.39	0.12	0.199	0.580	1.220 ^{1/4} c1	1.944	3.700	0.582	0.343	2.627	0.082	2.722	0.934	0.861	0.139	4.9	94.5
0.46	0.14	0.215	0.545	1.153 ^{1/4} c1	2.118	4.137	0.565	0.371	2.696	0.082	2.805	1.040	0.887	0.186	6.6	107.6
0.52	0.16	0.231	0.515	1.105 ^{1/4} c1	2.273	4.552	0.550	0.396	2.770	0.081	2.893	1.146	0.915	0.242	8.5	120.2
0.59	0.18	0.244	0.490	1.070 ^{1/4} c1	2.413	4.946	0.538	0.420	2.846	0.080	2.983	1.253	0.943	0.307	10.8	132.4
0.66	0.2	0.256	0.469	1.042 ^{1/4} c1	2.540	5.324	0.527	0.443	2.922	0.079	3.073	1.361	0.972	0.381	13.4	144.3
0.72	0.22	0.268	0.450	1.020 ^{1/4} c1	2.656	5.687	0.517	0.465	2.998	0.078	3.162	1.469	1.000	0.465	16.4	155.9
0.79	0.24	0.278	0.433	1.003 ^{1/4} c1	2.763	6.039	0.508	0.485	3.073	0.077	3.250	1.577	1.028	0.560	19.8	167.2
0.85	0.26	0.288	0.419	0.988 ^{1/4} c1	2.862	6.379	0.499	0.505	3.147	0.076	3.337	1.685	1.055	0.666	23.5	178.4
0.92	0.28	0.297	0.405	0.976 ^{1/4} c1	2.954	6.711	0.491	0.524	3.219	0.075	3.422	1.794	1.082	0.783	27.6	189.3
0.98	0.3	0.306	0.393	0.966 ^{1/4} c1	3.040	7.034	0.484	0.542	3.290	0.075	3.506	1.902	1.109	0.913	32.2	200.1
1.05	0.32	0.314	0.381	0.958 ^{1/4} c1	3.122	7.350	0.477	0.560	3.360	0.074	3.588	2.010	1.135	1.055	37.2	210.7
1.12	0.34	0.321	0.371	0.950 ^{1/4} c1	3.198	7.660	0.471	0.578	3.429	0.073	3.669	2.119	1.160	1.210	42.7	221.2
1.18	0.36	0.328	0.361	0.944 ^{1/4} c1	3.271	7.963	0.465	0.594	3.496	0.072	3.748	2.227	1.185	1.379	48.7	231.6
1.25	0.38	0.335	0.352	0.938 ^{1/4} c1	3.341	8.262	0.459	0.611	3.562	0.071	3.825	2.335	1.210	1.562	55.1	241.8
1.31	0.4	0.342	0.344	0.934 ^{1/4} c1	3.407	8.556	0.454	0.626	3.627	0.070	3.901	2.444	1.234	1.760	62.1	252.0
1.38	0.42	0.348	0.336	0.929 ^{1/4} c1	3.470	8.846	0.449	0.642	3.690	0.069	3.976	2.552	1.257	1.972	69.6	262.1
1.44	0.44	0.354	0.329	0.926 ^{1/4} c1	3.531	9.132	0.444	0.657	3.753	0.069	4.050	2.661	1.281	2.201	77.7	272.1
1.51	0.46	0.360	0.322	0.922 ^{1/4} c1	3.589	9.415	0.439	0.672	3.815	0.068	4.122	2.769	1.303	2.446	86.3	282.0
1.57	0.48	0.366	0.315	0.919 ^{1/4} c1	3.646	9.694	0.435	0.686	3.875	0.067	4.193	2.877	1.326	2.707	95.5	291.8
1.64	0.5	0.371	0.309	0.917 ^{1/4} c1	3.700	9.971	0.430	0.700	3.935	0.067	4.263	2.986	1.348	2.986	105.4	301.6
1.71	0.52	0.376	0.303	0.915 ^{1/4} c1	3.753	10.245	0.426	0.714	3.993	0.066	4.332	3.094	1.370	3.282	115.8	311.4
1.77	0.54	0.381	0.297	0.913 ^{1/4} c1	3.804	10.516	0.422	0.728	4.051	0.065	4.400	3.203	1.391	3.597	126.9	321.0

Maximum Boundary Shear Stresses

Depth (ft)	d (m)	B/R	Kbs	Kss	Fsf	T bed N/m^2	T bank N/m^2	Stresses Allowable?	
								Bed	Bank
0.33	0.1	12.060	1.176	0.923	0.892	95.2	84.9	OK	OK
0.39	0.12	10.327	1.201	0.942	0.895	113.5	101.6	OK	OK
0.46	0.14	9.074	1.223	0.960	0.898	131.6	118.2	OK	OK
0.52	0.16	8.123	1.244	0.977	0.901	149.5	134.7	OK	OK
0.59	0.18	7.373	1.264	0.992	0.904	167.3	151.3	OK	OK
0.66	0.2	6.766	1.282	1.006	0.907	185.0	167.8	OK	OK
0.72	0.22	6.262	1.299	1.020	0.910	202.5	184.3	OK	OK
0.79	0.24	5.837	1.315	1.033	0.913	220.0	200.8	OK	OK
0.85	0.26	5.473	1.331	1.045	0.915	237.4	217.3	OK	OK
0.92	0.28	5.156	1.345	1.056	0.918	254.7	233.8	OK	OK
0.98	0.3	4.878	1.359	1.067	0.920	272.0	250.3	OK	OK
1.05	0.32	4.632	1.373	1.078	0.923	289.2	266.9	OK	
1.12	0.34	4.413	1.386	1.088	0.925	306.5	283.4	OK	
1.18	0.36	4.215	1.398	1.097	0.927	323.7	300.0	OK	
1.25	0.38	4.036	1.410	1.106	0.929	340.9	316.7	OK	
1.31	0.4	3.873	1.421	1.115	0.931	358.1	333.3	OK	
1.38	0.42	3.725	1.432	1.124	0.933	375.2	350.1	OK	
1.44	0.44	3.588	1.442	1.132	0.935	392.4	366.8	OK	
1.51	0.46	3.461	1.453	1.140	0.936	409.6	383.6	OK	
1.57	0.48	3.345	1.463	1.148	0.938	426.8	400.4	OK	
1.64	0.5	3.236	1.472	1.156	0.940	444.0	417.3	OK	
1.71	0.52	3.135	1.481	1.163	0.941	461.3	434.2	OK	
1.77	0.54	3.040	1.490	1.170	0.943	478.5	451.2	OK	

Where C is =		0.785
z	C	
1.5	0.76	
2	0.785	
3	0.85	
4	0.935	
6	0.97	

Permissible Shear Stresses

	D50 = (m)		Tc = (N/m^2)		SF =		Cz =		Cw=		Cr=		Tpbed = N/m^2	
	0.4		600		1.3		0.89		1		0.54		461.54	
Depth (ft)	d	E	Ab	Cl	Cb	Ca	Al (estimated)	Tp_bank N/m^2						
0.33	0.1	6.78	0.15	2.01	1.00	0.42	0	250.8						
0.39	0.12	5.67	0.18	2.01	1.00	0.42	0	250.9						
0.46	0.14	4.87	0.21	2.01	1.00	0.42	0	251.0						
0.52	0.16	4.27	0.23	2.01	1.00	0.42	0	251.2						
0.59	0.18	3.80	0.26	2.01	1.00	0.42	0	251.3						
0.66	0.2	3.43	0.29	2.01	1.00	0.42	0	251.4						
0.72	0.22	3.12	0.32	2.01	1.00	0.42	0	251.5						
0.79	0.24	2.87	0.35	2.01	1.00	0.42	0	251.7						
0.85	0.26	2.65	0.38	2.01	1.00	0.42	0	251.8						
0.92	0.28	2.46	0.41	2.01	1.00	0.42	0	251.9						
0.98	0.3	2.30	0.43	2.02	1.00	0.42	0	252.0						
1.05	0.32	2.16	0.46	2.02	1.00	0.42	0	252.2						
1.12	0.34	2.03	0.49	2.02	1.00	0.42	0	252.3						
1.18	0.36	1.92	0.52	2.02	1.00	0.42	0	252.4						
1.25	0.38	1.82	0.55	2.02	1.00	0.42	0	252.6						
1.31	0.4	1.73	0.58	2.02	1.00	0.42	0	252.7						
1.38	0.42	1.64	0.61	2.02	1.00	0.42	0	252.8						
1.44	0.44	1.57	0.64	2.02	1.00	0.42	0	253.0						
1.51	0.46	1.50	0.67	2.02	1.00	0.42	0	253.1						
1.57	0.48	1.44	0.70	2.02	1.00	0.42	0	253.2						
1.64	0.5	1.38	0.73	2.03	1.00	0.42	0	253.3						
1.71	0.52	1.33	0.75	2.03	1.00	0.42	0	253.5						
1.77	0.54	1.28	0.78	2.03	1.00	0.42	0	253.6						

Steep Slope Riprap Design

Ref: Stormwater Collection Systems Design Handbook, Larry Mays, 2001, Section 16.5

Troy - Mine, Millpad Channel, 77% slope

Target Flow = 33 cfs

Channel Dimensions

Bottom Width(ft) = 2.99
Bottom Width(m) = 0.91
Sideslopes = 2 :1(H:V)
Slope (m/m) = 0.77
Length (ft) = 200
Length (m) = 61

Trial Lining

D90 (ft) = 6.97
D90 (m) = 2.125
D50 (ft) = 5.58
D50 (m) = 1.7
D10 (ft) = 2.79
D10 (m) = 0.85

Range of Depth to be Evaluated

Least Depth (m) = 0.1
Maximum Depth (m)= 2
Increment (m) = 0.02

Angles

Side Slope, degrees 26.57
Bed Slope, degrees 50.35
Riprap Angle of
Repose, degrees 43
(Get from Charts)

Factor of Safety 1.3

Hydraulic Calculations

Depth		Area		Wetted Perimeter		Top Width		Hydraulic Radius	
(ft)	(m)	(sf)	(sm)	(ft)	(m)	(ft)	(m)	(ft)	(m)
0.328	0.100	1.195	0.111	4.453	1.357	4.298	1.31	0.268	0.082
0.394	0.120	1.485	0.138	4.746	1.447	4.560	1.39	0.313	0.095
0.459	0.140	1.793	0.167	5.040	1.536	4.823	1.47	0.356	0.108
0.525	0.160	2.118	0.197	5.333	1.626	5.085	1.55	0.397	0.121
0.591	0.180	2.461	0.229	5.627	1.715	5.348	1.63	0.437	0.133
0.656	0.200	2.820	0.262	5.920	1.804	5.610	1.71	0.476	0.145
0.722	0.220	3.197	0.297	6.213	1.894	5.873	1.79	0.515	0.157
0.787	0.240	3.591	0.334	6.507	1.983	6.135	1.87	0.552	0.168
0.853	0.260	4.002	0.372	6.800	2.073	6.398	1.95	0.588	0.179
0.919	0.280	4.430	0.412	7.094	2.162	6.660	2.03	0.625	0.190
0.984	0.300	4.876	0.453	7.387	2.252	6.923	2.11	0.660	0.201
1.050	0.320	5.339	0.496	7.681	2.341	7.185	2.19	0.695	0.212
1.115	0.340	5.819	0.541	7.974	2.431	7.448	2.27	0.730	0.222
1.181	0.360	6.316	0.587	8.268	2.520	7.710	2.35	0.764	0.233
1.247	0.380	6.831	0.635	8.561	2.609	7.972	2.43	0.798	0.243
1.312	0.400	7.363	0.684	8.855	2.699	8.235	2.51	0.831	0.253
1.378	0.420	7.911	0.735	9.148	2.788	8.497	2.59	0.865	0.264
1.444	0.440	8.478	0.788	9.441	2.878	8.760	2.67	0.898	0.274
1.509	0.460	9.061	0.842	9.735	2.967	9.022	2.75	0.931	0.284
1.575	0.480	9.662	0.898	10.028	3.057	9.285	2.83	0.963	0.294
1.640	0.500	10.280	0.955	10.322	3.146	9.547	2.91	0.996	0.304
1.706	0.520	10.915	1.014	10.615	3.236	9.810	2.99	1.028	0.313
1.772	0.540	11.567	1.075	10.909	3.325	10.072	3.07	1.060	0.323
1.837	0.560	12.236	1.137	11.202	3.414	10.335	3.15	1.092	0.333
1.903	0.580	12.923	1.201	11.496	3.504	10.597	3.23	1.124	0.343
1.969	0.600	13.627	1.266	11.789	3.593	10.860	3.31	1.156	0.352
2.034	0.620	14.348	1.333	12.082	3.683	11.122	3.39	1.188	0.362
2.100	0.640	15.087	1.402	12.376	3.772	11.385	3.47	1.219	0.372
2.165	0.660	15.842	1.472	12.669	3.862	11.647	3.55	1.250	0.381

Note: Gradation calculation (D90 and D10) based on USFS Specs. for Roads and Bridges

Bathurst Resistance Calculations

Depth (ft)	Depth (m)	b	c1	f(FR)	c2	f(REG)	f(CG)	V*	Sum of f (functions) n	V/V*	Velocity (V) m/sec	Froude Number	Correction for slope 0.98			
													Flow, Q m³/sec	Flow, Q ft³/sec	To N/m²	
0.33	0.1	0.109	0.842	2.222 ^{1/4} c1	1.478	1.300	0.740	0.869	2.138	0.098	2.210	1.921	1.940	0.213	7.5	394.2
0.39	0.12	0.120	0.800	1.972 ^{1/4} c1	1.605	1.454	0.726	0.952	2.081	0.104	2.162	2.058	1.897	0.284	10.0	459.8
0.46	0.14	0.130	0.765	1.803 ^{1/4} c1	1.714	1.596	0.713	1.028	2.053	0.107	2.142	2.203	1.880	0.367	13.0	522.7
0.52	0.16	0.138	0.737	1.681 ^{1/4} c1	1.809	1.728	0.703	1.099	2.041	0.110	2.138	2.351	1.877	0.463	16.3	583.5
0.59	0.18	0.146	0.713	1.588 ^{1/4} c1	1.891	1.853	0.693	1.166	2.040	0.112	2.145	2.501	1.882	0.572	20.2	642.4
0.66	0.2	0.153	0.692	1.516 ^{1/4} c1	1.965	1.970	0.685	1.229	2.046	0.113	2.158	2.652	1.894	0.695	24.5	699.8
0.72	0.22	0.160	0.674	1.458 ^{1/4} c1	2.030	2.082	0.677	1.289	2.056	0.114	2.175	2.804	1.909	0.833	29.4	755.8
0.79	0.24	0.166	0.658	1.410 ^{1/4} c1	2.090	2.190	0.670	1.346	2.069	0.115	2.195	2.956	1.926	0.986	34.8	810.7
0.85	0.26	0.172	0.643	1.369 ^{1/4} c1	2.144	2.293	0.664	1.401	2.085	0.115	2.218	3.108	1.946	1.155	40.8	864.5
0.92	0.28	0.177	0.630	1.335 ^{1/4} c1	2.193	2.393	0.658	1.454	2.101	0.115	2.241	3.259	1.966	1.341	47.3	917.5
0.98	0.3	0.182	0.618	1.305 ^{1/4} c1	2.239	2.490	0.652	1.505	2.119	0.115	2.265	3.410	1.988	1.545	54.5	969.6
1.05	0.32	0.187	0.607	1.279 ^{1/4} c1	2.281	2.584	0.647	1.555	2.138	0.115	2.290	3.561	2.010	1.766	62.3	1021.1
1.12	0.34	0.191	0.597	1.256 ^{1/4} c1	2.321	2.676	0.642	1.603	2.157	0.115	2.316	3.711	2.032	2.006	70.8	1072.0
1.18	0.36	0.195	0.588	1.236 ^{1/4} c1	2.358	2.765	0.637	1.649	2.177	0.115	2.341	3.860	2.054	2.265	79.9	1122.3
1.25	0.38	0.199	0.579	1.218 ^{1/4} c1	2.393	2.853	0.632	1.694	2.197	0.115	2.367	4.010	2.077	2.545	89.8	1172.1
1.31	0.4	0.203	0.571	1.202 ^{1/4} c1	2.426	2.939	0.628	1.738	2.217	0.115	2.392	4.158	2.099	2.844	100.4	1221.5
1.38	0.42	0.207	0.563	1.187 ^{1/4} c1	2.457	3.023	0.624	1.781	2.237	0.114	2.418	4.306	2.122	3.165	111.7	1270.5
1.44	0.44	0.210	0.556	1.173 ^{1/4} c1	2.487	3.106	0.620	1.823	2.257	0.114	2.443	4.454	2.144	3.508	123.8	1319.1
1.51	0.46	0.213	0.549	1.160 ^{1/4} c1	2.515	3.187	0.616	1.864	2.278	0.114	2.469	4.601	2.166	3.874	136.7	1367.3
1.57	0.48	0.217	0.542	1.149 ^{1/4} c1	2.543	3.267	0.612	1.904	2.298	0.113	2.494	4.748	2.188	4.262	150.4	1415.3
1.64	0.5	0.220	0.536	1.138 ^{1/4} c1	2.569	3.346	0.608	1.943	2.318	0.113	2.519	4.895	2.210	4.674	165.0	1463.0
1.71	0.52	0.223	0.530	1.129 ^{1/4} c1	2.594	3.424	0.605	1.982	2.338	0.113	2.543	5.041	2.232	5.111	180.4	1510.4
1.77	0.54	0.226	0.524	1.119 ^{1/4} c1	2.618	3.501	0.601	2.020	2.357	0.112	2.568	5.186	2.253	5.573	196.7	1557.7

Maximum Boundary Shear Stresses

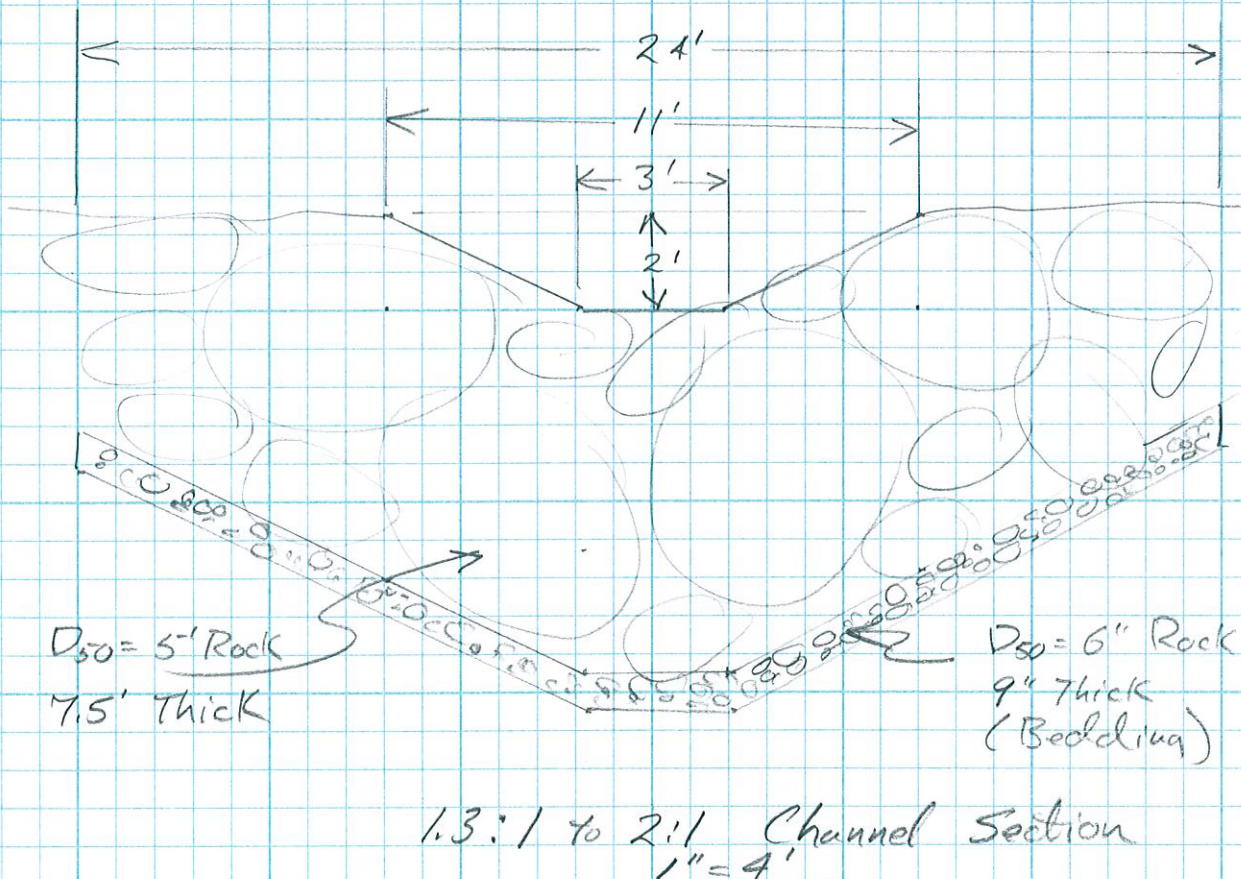
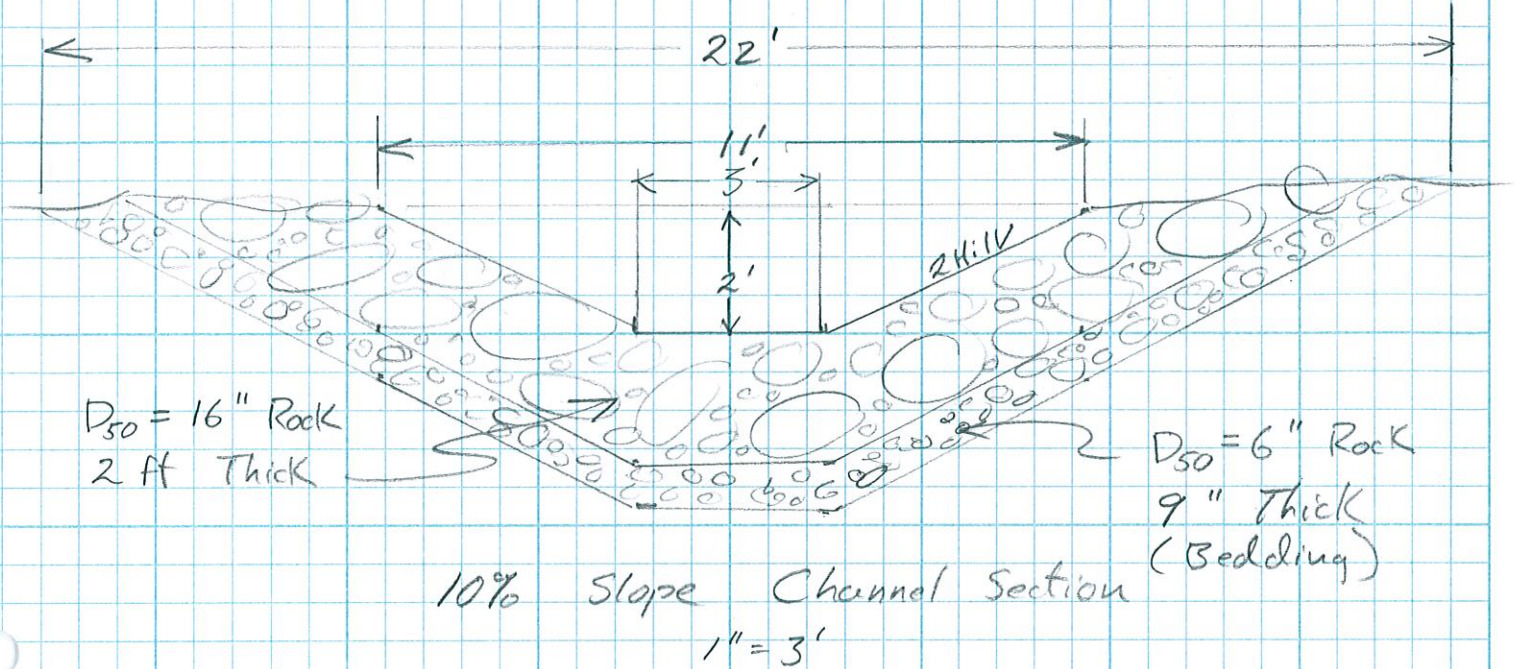
Depth (ft)	d (m)	B/R	Kbs	Kss	Fsf	T bed N/m^2	T bank N/m^2	Stresses Allowable?	
								Bed	Bank
0.33	0.1	11.127	1.188	0.933	0.976	468.4	457.2	OK	OK
0.39	0.12	9.540	1.214	0.953	0.974	558.2	543.7	OK	OK
0.46	0.14	8.390	1.238	0.972	0.973	647.1	629.6	OK	OK
0.52	0.16	7.516	1.260	0.989	0.973	735.1	715.1	OK	OK
0.59	0.18	6.827	1.280	1.005	0.973	822.4	800.3	OK	OK
0.66	0.2	6.267	1.299	1.020	0.974	909.1	885.2	OK	OK
0.72	0.22	5.803	1.317	1.034	0.975	995.3	970.0	OK	OK
0.79	0.24	5.410	1.334	1.047	0.975	1081.1	1054.6	OK	OK
0.85	0.26	5.073	1.350	1.059	0.976	1166.7	1139.2	OK	OK
0.92	0.28	4.780	1.365	1.071	0.978	1252.0	1223.9	OK	OK
0.98	0.3	4.523	1.379	1.082	0.979	1337.1	1308.5	OK	OK
1.05	0.32	4.295	1.393	1.093	0.980	1422.1	1393.3	OK	OK
1.12	0.34	4.091	1.406	1.104	0.981	1507.0	1478.1	OK	OK
1.18	0.36	3.908	1.418	1.113	0.982	1591.9	1563.1	OK	OK
1.25	0.38	3.742	1.431	1.123	0.983	1676.8	1648.2	OK	OK
1.31	0.4	3.591	1.442	1.132	0.984	1761.6	1733.6	OK	OK
1.38	0.42	3.452	1.453	1.141	0.985	1846.5	1819.1	OK	OK
1.44	0.44	3.325	1.464	1.149	0.986	1931.4	1904.7	OK	OK
1.51	0.46	3.208	1.475	1.158	0.987	2016.4	1990.6	OK	OK
1.57	0.48	3.099	1.485	1.166	0.988	2101.5	2076.8	OK	OK
1.64	0.5	2.998	1.495	1.173	0.989	2186.7	2163.1	OK	OK
1.71	0.52	2.904	1.504	1.181	0.990	2271.9	2249.6	OK	OK
1.77	0.54	2.816	1.513	1.188	0.991	2357.3	2336.4	OK	OK

Where C is = 0.785

z	C
1.5	0.76
2	0.785
3	0.85
4	0.935
6	0.97

Permissible Shear Stresses

	D50 = (m)		Tc = (N/m^2)		SF =		Cz =		Cw=		Cr=		Tpbed = N/m^2	
	1.7		2550		1.3		0.89		1		0.54		1961.54	
Depth (ft)	d	E	Ab	Cl	Cb	Ca	Al (estimated)	Tp_bank N/m^2						
0.33	0.1		5.35	0.19	2.01	1.00	0.42	0	1066.5					
0.39	0.12		4.50	0.22	2.01	1.00	0.42	0	1067.2					
0.46	0.14		3.88	0.26	2.01	1.00	0.42	0	1067.9					
0.52	0.16		3.42	0.29	2.01	1.00	0.42	0	1068.5					
0.59	0.18		3.06	0.33	2.01	1.00	0.42	0	1069.2					
0.66	0.2		2.76	0.36	2.01	1.00	0.42	0	1069.8					
0.72	0.22		2.52	0.40	2.01	1.00	0.42	0	1070.5					
0.79	0.24		2.32	0.43	2.02	1.00	0.42	0	1071.1					
0.85	0.26		2.15	0.47	2.02	1.00	0.42	0	1071.8					
0.92	0.28		2.00	0.50	2.02	1.00	0.42	0	1072.4					
0.98	0.3		1.87	0.53	2.02	1.00	0.42	0	1073.1					
1.05	0.32		1.76	0.57	2.02	1.00	0.42	0	1073.8					
1.12	0.34		1.65	0.60	2.02	1.00	0.42	0	1074.4					
1.18	0.36		1.56	0.64	2.02	1.00	0.42	0	1075.1					
1.25	0.38		1.48	0.67	2.02	1.00	0.42	0	1075.7					
1.31	0.4		1.41	0.71	2.02	1.00	0.42	0	1076.4					
1.38	0.42		1.34	0.74	2.03	1.00	0.42	0	1077.1					
1.44	0.44		1.28	0.78	2.03	1.00	0.42	0	1077.7					
1.51	0.46		1.23	0.81	2.03	1.00	0.42	0	1078.4					
1.57	0.48		1.18	0.85	2.03	1.00	0.42	0	1079.1					
1.64	0.5		1.13	0.88	2.03	1.00	0.42	0	1079.8					
1.71	0.52		1.09	0.92	2.03	1.00	0.42	0	1080.4					
1.77	0.54		1.05	0.96	2.03	1.00	0.42	0	1081.1					



**TROY MINE RECLAMATION
MILLSITE CHANNEL COST
December 2, 2010**

OPTION 1

Item	Quantity		Unit Cost		Cost	
Mobilization @ 10%					\$15,681	
Erosion control @ 5%					\$7,467	
Clear and Grub (wooded areas)	1	AC	4000		\$4,000	
Access Road	700	LF	16		\$11,200	USDA FS Cost Estimating Guide for Road Construction
<u>Channel</u>	<u>length (ft)</u>	<u>section (ft²)</u>				
D ₅₀ = 16" Rock	650	30	722	CY	\$30.00	\$21,667
D ₅₀ = 5' Rock	150	180	1000	CY	\$60.00	\$60,000
D ₅₀ = 6" Rock	800	18	533	CY	\$30.00	\$16,000
<u>Energy Dissipation Basin</u>						
D ₅₀ = 16" Rock	30	46	51	CY	\$30.00	\$1,533
D ₅₀ = 6" Rock	30	23	26	CY	\$30.00	\$767
<u>Liner</u>						
HDPE			1200	SY	\$15.00	\$18,000
Bedding			400	CY	\$20.00	\$8,000
<u>Culvert</u>						
4' Dia. CMP			40	LF	\$143.00	\$5,720 Means 33 41 13 40 2200
Excavation, Backfill, Compaction			40	LF	\$61.30	\$2,452 Means 31 23 16 13 0500, 2020 and 31 23 23 23 7000
<u>Construction contingency @ 15%</u>						<u>\$25,873</u>
Total Construction Cost						\$198,359
Engineering @ 20%						<u>\$39,672</u>
Total Project Cost						\$238,031

**TROY MINE RECLAMATION
MILLSITE CHANNEL COST
December 2, 2010**

OPTION 2

Item	Quantity		Unit Cost		Cost	
Mobilization @ 10%					\$19,086	
Erosion control @ 5%					\$9,089	
Clear and Grub (wooded areas)	1	AC	4000		\$4,000	
Access Road	400	LF	16		\$6,400	USDA FS Cost Estimating Guide for Road Construction
<u>Channel</u>	<u>length (ft)</u>	<u>section (ft²)</u>				
D ₅₀ = 16" Rock	870	30	967	CY	\$30.00	\$29,000
D ₅₀ = 5' Rock	200	180	1333	CY	\$60.00	\$80,000
D ₅₀ = 6" Rock	970	18	647	CY	\$30.00	\$19,400
<u>Energy Dissipation Basin</u>						
D ₅₀ = 16" Rock	30	46	51	CY	\$30.00	\$1,533
D ₅₀ = 6" Rock	30	23	26	CY	\$30.00	\$767
<u>Liner</u>						
HDPE			1500	SY	\$15.00	\$22,500
Bedding			500	CY	\$20.00	\$10,000
<u>Culvert</u>						
4' Dia. CMP			40	LF	\$143.00	\$5,720 Means 33 41 13 40 2200
Excavation, Backfill, Compaction			40	LF	\$61.30	\$2,452 Means 31 23 16 13 0500, 2020 and 31 23 23 23 7000
<u>Construction contingency @ 15%</u>						<u>\$31,492</u>
Total Construction Cost						\$241,439
Engineering @ 20%						<u>\$48,288</u>
Total Project Cost						\$289,726

Appendix F

Troy Mine Area Surface Water Evaluation



Technical Memorandum

*To: Emily Corsi, Herb Rolfes – DEQ
Bobbie Lacklen, John McKay – KNF*

From: Kent Whiting, Mark Hills, Bill Bucher – CDM Helena

Date: December 3, 2010

Subject: Troy Mine Area Surface Water Evaluation

The purpose of this memorandum is to evaluate the potential for adverse effects to surface water, seeps and springs in the vicinity of the Troy mine workings after closure of the mine. After closure the groundwater entering the mine will be allowed to flood up to a level of 4,225 feet, where it will be routed through the Service and Conveyor adits and transported by pipeline to the tailings impoundment area. The current decant ponds at this location will be used to infiltrate the water. This water will be treated through attenuation of metals as it infiltrates through the underlying soils as described in *Troy Mine Copper Attenuation Study – Secondary Processes* (CDM, 2010a). Further descriptions of this system are presented in the *Mine Water Management Technical Memorandum* (CDM, 2010b) and the *Conceptual Design of New Water Line for Transport of Mine Water after Mine Closure* (CDM, 2010c).

The primary concern is that as the elevation of water in the mine workings increases, the increased head in the mine workings will force mine water through permeable pathways in the vicinity of the mine. Movement of untreated mine pool water through these pathways has the potential to cause contamination of existing surface water, seeps and springs and create new seeps or springs, which may discharge contaminated water. This evaluation investigates the relationship between surface water, seeps and springs and the mine workings using a hydrogeological approach with a geochemical perspective. In the hydrogeologic evaluation, the surface water expressions in the vicinity of the mine workings are related to the present and future water elevations, and the expected changes are described. In the geochemical evaluation, the present water quality of the surface water, seeps and springs is described and used to determine the potential connection between the surface water, springs and seeps and the flooded mine workings.

Hydrogeologic Evaluation

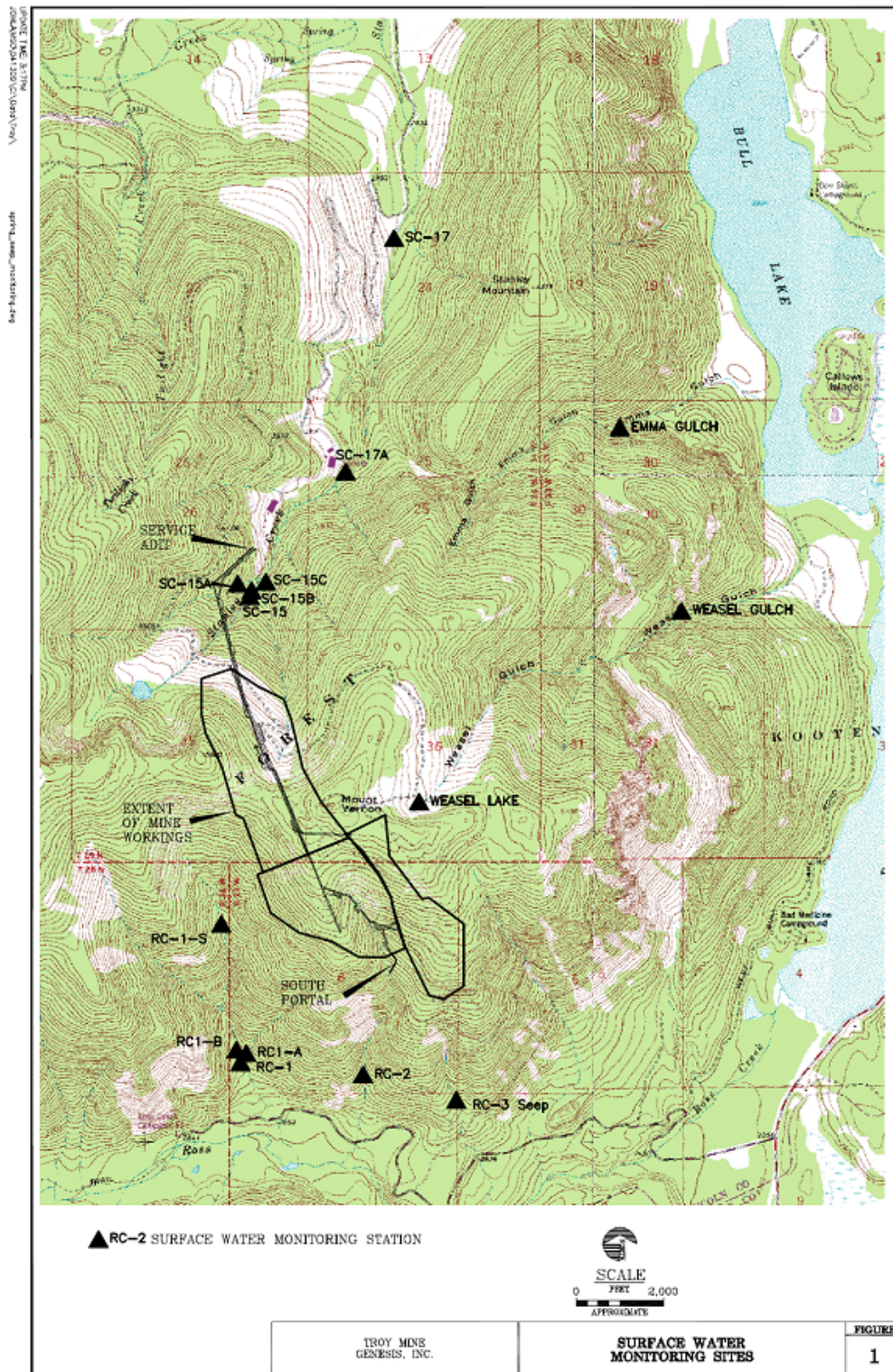
When the mine is closed, dewatering pumps will be shut-off and water will flood the workings. The physical and hydrologic features of the mine workings and terrain in the vicinity of the

Troy Mine were evaluated to assess changes that would occur in the hydrologic system. Data used in the evaluation were limited to the location, elevation, and distribution of identified springs and seeps near the Troy Mine, the current pumped mine water elevation in the mine workings, the expected flooded post-mine water elevation, and the local topography. Limited geologic and quantitative hydrologic data were available for this evaluation. For example, fracture systems are only well mapped within the ore body but not in the surrounding formations, and no wells were available outside the mine footprint to verify static water elevations in the surrounding area.

The general extent of the mine workings and the location of identified surface water sites in the vicinity of the Troy Mine are shown in **Figure 1**. The monitored surface water sites are located in low lying areas within the Stanley Creek drainage north of the mine and the Ross Creek drainage south of the mine. These surface water sites represent known points of groundwater discharge during the period of mine operation and dewatering. Hydrometrics (2009) reported two springs (RC-1-S and RC-3) in the Ross Creek drainage and apparent groundwater-supplied baseflow in two Ross Creek tributaries (RC-1 and RC-2). Hydrometrics also reported that expressions of groundwater exist in Stanley Creek within the defined stream channel. These springs appear to represent groundwater expression at elevations from 3,800 to 3,900 feet. The SC-15 sites (near the mine Service Adit) generally have flow even during late summer and fall. This discharge during the dry portion of the annual water cycle suggests that these springs are connected to an aquifer of significant extent.

To assess the potential impacts on groundwater discharge in the vicinity of the mine after closure, a series of cross-sections were drawn. Sections A-A' through E-E' intersect the mine workings and surface water monitoring sites. The cross-sections are attachments to this document. The plan view of the mine area is shown on each of the drawings to reference the respective sections. During mine operation, the active workings are dewatered to 4,050 feet above mean sea level (amsl). After mine closure, the maximum "flooded" groundwater elevation will be 4,225 feet amsl. The maximum flooded elevation is set by the base of the high point in the service adit that will remain open and serve to drain the mine workings. Following mine closure, elevation head within the mine workings will increase about 175 feet. The projected "flooded" water table (light blue line) relative to the current "pumped" water table (dark blue line) on the sections depicts the relative magnitude of shift that is anticipated for the surface water monitoring sites.

The cross-sections indicate the approximate stratigraphy and elevation of the geologic units. This information is most accurate in the immediate vicinity of the mine workings and is extrapolated in areas more distant from the mine workings. The geologic sections show that current and post-closure water levels are within the Revett formation. The Revett Formation consists of fractured quartzite and groundwater flow is through fractures. At the coarse scale of these cross-sections, fracture flow can be represented similarly to that of a porous medium, and groundwater elevations have been estimated based on this assumption.



Source: Genesis, Inc.

The evaluation of the hydrogeology of the area surrounding the mine workings indicates that the surface water measurement sites are typically far below the elevation of the mine workings. The monitoring sites in Stanley Creek are 250 feet below the elevation of the operational (pumped) water level in the mine, and the monitoring sites in Ross Creek drainage are at least 700 feet below the operational water level. The projected post-closure rise of water level would be relatively small compared to the hydraulic head difference between the operational (pumped) mine water level and most surface water measurement sites (see attachments). Using the principles of groundwater flow, it is expected that this head difference would decrease in an outward direction from the mine to no more than a few tens of feet in the vicinity of the monitoring sites. This would limit the potential for migration of water sources in the vicinity of these sites horizontally and vertically, as well as limit the potential increase inflow to a small percentage of current flow.

It is known that there are perennial surface water flows in the upper reaches of Stanley Creek. One small perennial stream originates above the elevation of the mine workings presumably from a local aquifer. The upper portions of this stream would not be affected by flooding of the mine because it originates above the maximum mine flooding elevation. However, other springs and seeps may become active at elevations below the final mine pool elevation after mine closure. These springs would most likely occur in the interval between the current pumping level of 4,050 feet and the final pool elevation of 4,225 feet, an interval of 175 feet. Whether these springs appear depends largely on whether faults that intercept the ground surface in this interval can transmit sufficient water to result in surface expressions. Although there is limited knowledge of surface faults in the perimeter around the mine workings, it is known that most faults within the mine workings tend to be vertically oriented. They therefore transmit groundwater primarily in the vertical direction, with horizontal flow being limited by their length and a lesser hydraulic gradient. Horizontal joints, fractures, and faults that could transmit water to the surface are not documented. This geological understanding suggests that the potential for development of new springs in the interval of the flooded mine workings is limited.

Geochemical Evaluation

The data available for surface water, springs and seeps analysis was spotty, both in terms of the collection frequency and the parameters analyzed. The data used in the analysis were collected from a number of sources, which were collected as needed for different reasons. A summary of the data includes:

- ASARCO, mostly pre-mining baseline data, 1977-1985 (Genesis Appendix E of March 2008 Annual Report)
- Forest Service Database, 1973-2005

- DEQ Field Inspection Reports, 1991-2009
- Genesis Springs and Seep Data, February 2009
- Miscellaneous ASARCO data, 1979-1985
- Supplemental spring sampling performed by DEQ in August 2010

The surface water data were compared to data for the mine pool, the decant pond, and the mine adit collected mainly by Genesis but also with some data collected by DEQ and CDM. A summary of the data used is provided in Table 1.

Table 1 - Summary of Sample Data for Used for the Geochemical Evaluation.

Diagram	Number of Samples	Date Range
TDS Box Plot	31	1973-2007
Ternary Diagram	13	1975-2010
Dissolved Copper Box Plot	5	2002-2006
Dissolved Iron Box Plot	5	2003-2004

Major Ions

Major ions, such as calcium, magnesium, sulfate, chloride, sodium, and potassium, can be used to evaluate the source of a groundwater. Often, groundwaters derived from distinct aquifers will have distinct major ion chemistry, both in terms of the total dissolved load (i.e. TDS) and in the ratios of one ion to another. A comparison of the TDS concentrations for the mine area surface water compared to other water sources is provided in **Figure 2** below.

The mine area springs and seeps and surface waters have low TDS values compared to the mining related waters (mine void, decant pond, service adit) and the groundwaters in the vicinity of the tailings impoundment (a.k.a. "Lake Creek GW"). Most of the TDS in the mine related waters and tailings impoundment area groundwaters is composed of calcium, magnesium, and bicarbonate, although there is a higher contribution of sodium and potassium in the decant pond water, which may reflect the addition of reagents during the flotation process. The higher TDS and the calcium-magnesium carbonate composition of these waters likely is the result of fairly long residence times of ground water in local rocks, such as dolostone or clastic sediments cemented by the mineral dolomite ($\text{CaMg}(\text{CO}_3)_2$). The surface waters and mine area seeps and springs are not calcium-magnesium bicarbonate waters and are low in dissolved solids, suggesting that the source of these waters is precipitation runoff and/or

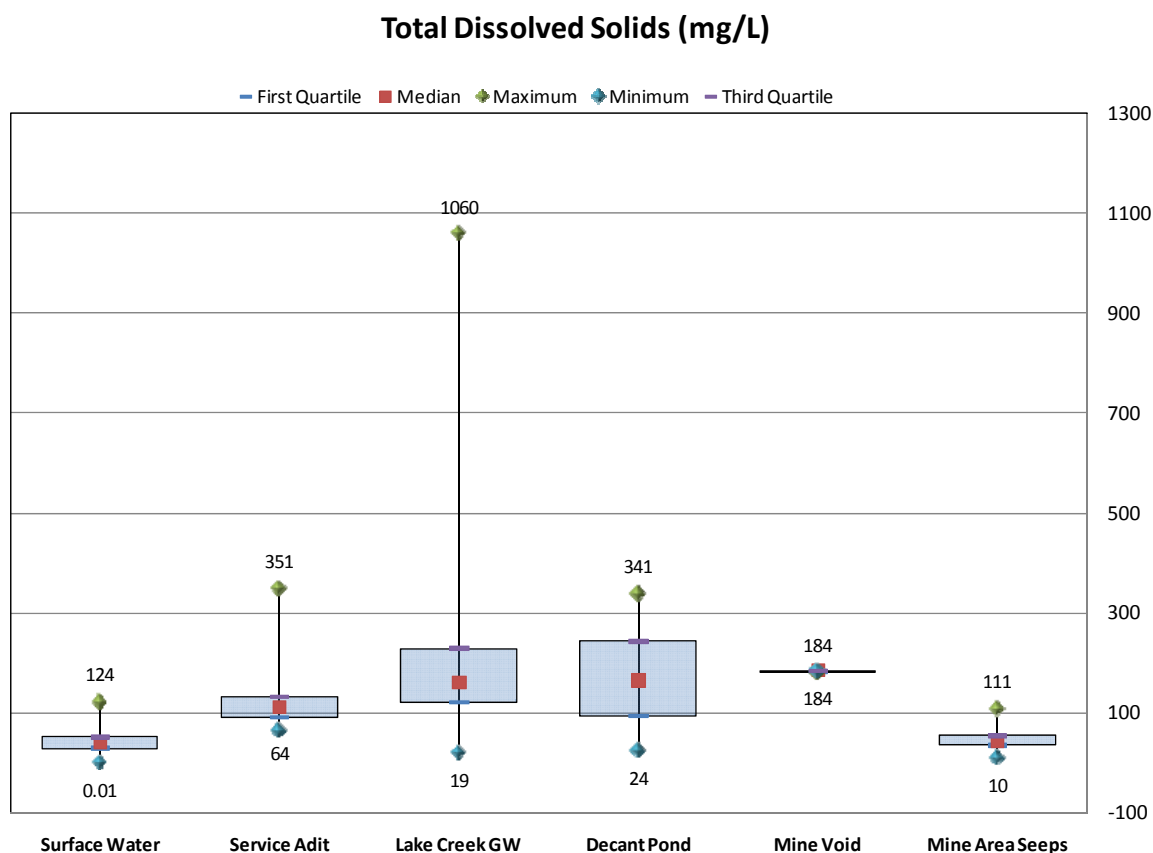


Figure 2 - Comparison of the TDS values for the mine area surface water (seeps) to other water types.

near-surface groundwater which has not had enough residence time or sufficient contact with dolomite-bearing aquifer materials.

Figure 3 is a ternary diagram showing the differences in the relative chemistry between the mine area surface water (as represented by a single sample collected within the Ross Creek drainage) and some of the mine-related waters. The triangular field on the bottom left is used to plot the positively-charged ions (calcium, magnesium, sodium, potassium), and the triangular field on the bottom right is used to plot the negatively-charged ions (chloride, sulfate, bicarbonate). The central diamond-shaped field reflects the whole chemistry composition of the water. In general, similar waters plot close to each other on a ternary diagram, while dissimilar waters plot at a distance from each other. The Ross Creek baseline data indicate a calcium-magnesium sulfate water, while the mine-related waters and springs in the vicinity are calcium-

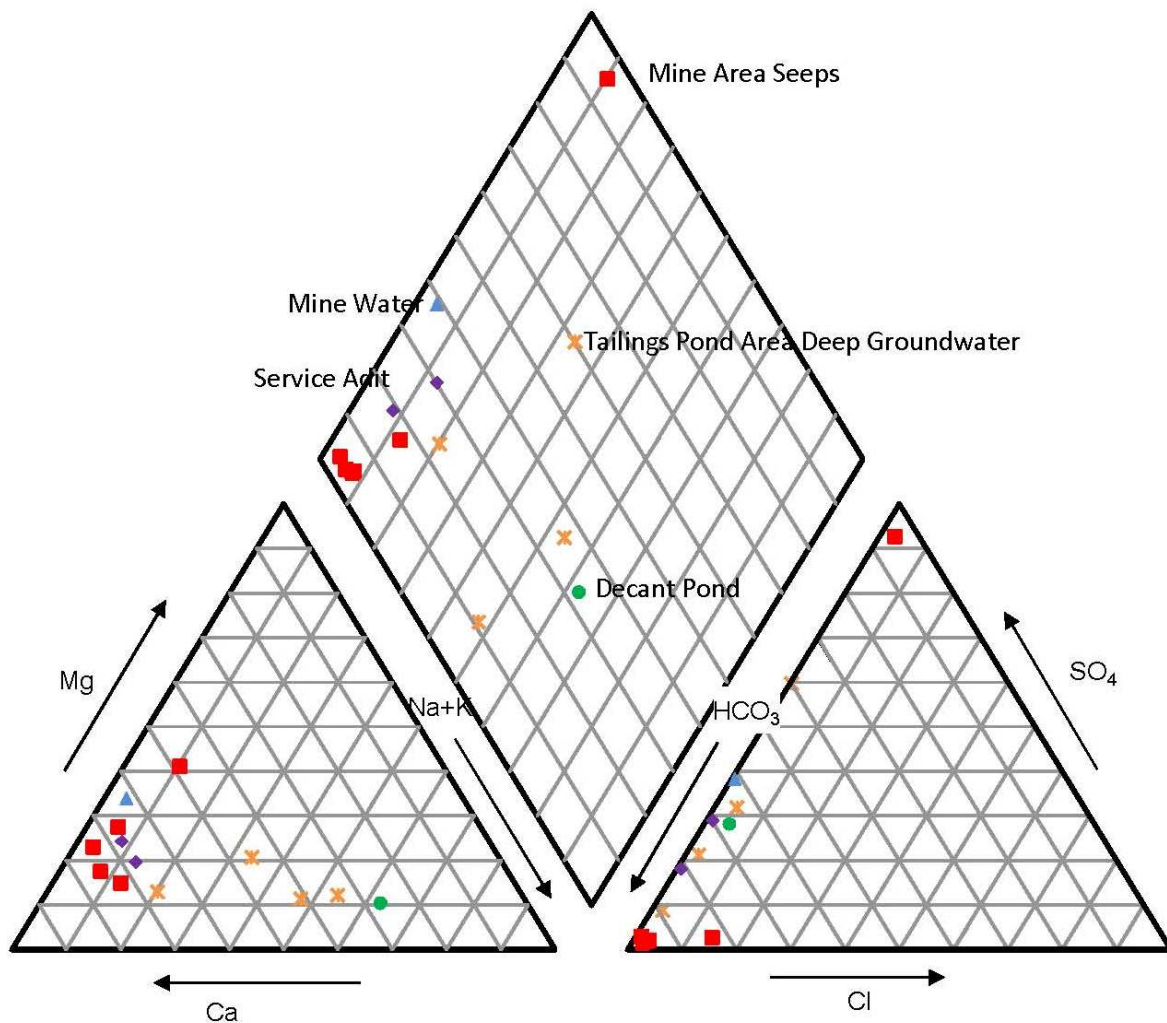


Figure 3 - Ternary diagram showing the ratios of the major ions for each water type.

magnesium carbonate waters with relatively lower magnesium ratios. A summary of the samples used for the diagram is shown in Table 2.

Figure 3 shows that Ross Creek surface water is very different from the mine-related waters. The difference in major ion ratios and TDS between the Ross Creek surface water and the other waters is a strong confirmation that the surface water in Ross Creek tributaries is not derived from the mine void.

Table 2 - Summary of Sample Data Used in Ternary Diagram

Sample ID	Date Collected	Type
Service Adit-D	5/12/09	Adit
Service Adit-P	5/12/09	
Decant Pond	5/14/2009	Decant Pond
Mine Water	12/16/2009	Mine Water
95-4	5/12/2009	Tailings Pond Area Deep Groundwater
95-5	5/12/2009	
95-8	6/12/2009	
MW-01-15	6/8/2009	
SC-15A	8/12/10	Mine Area Surface Water
SC-15C	8/12/10	
Ross Creek #11	8/12/10	
RC-1	8/12/10	
Ross Creek #10	8/12/10	
Ross Cr 74-10	5/20/75	

Metals

The metals data for mine area surface water is limited to dissolved and total recoverable concentrations of arsenic, antimony, copper, lead, iron, manganese, and zinc. Of these, few values above the analytical reporting limit were observed, namely total recoverable and dissolved copper, iron, and manganese, and dissolved zinc. In general, when detections were observed, the concentrations were low, as shown by the comparison of dissolved copper concentrations for each water type (**Figure 4**).

Dissolved zinc concentrations of 0.003 mg/L were reported for Ross Creek drainage sites RC-1 and RC-3 for samples collected in July 2006. Dissolved zinc concentrations within the mine void waters were higher, at 0.020 mg/L for the sample collected in December 2009. However, much of the other zinc data could not be compared to these two detections, as the reporting limit was too high (0.010 mg/L).

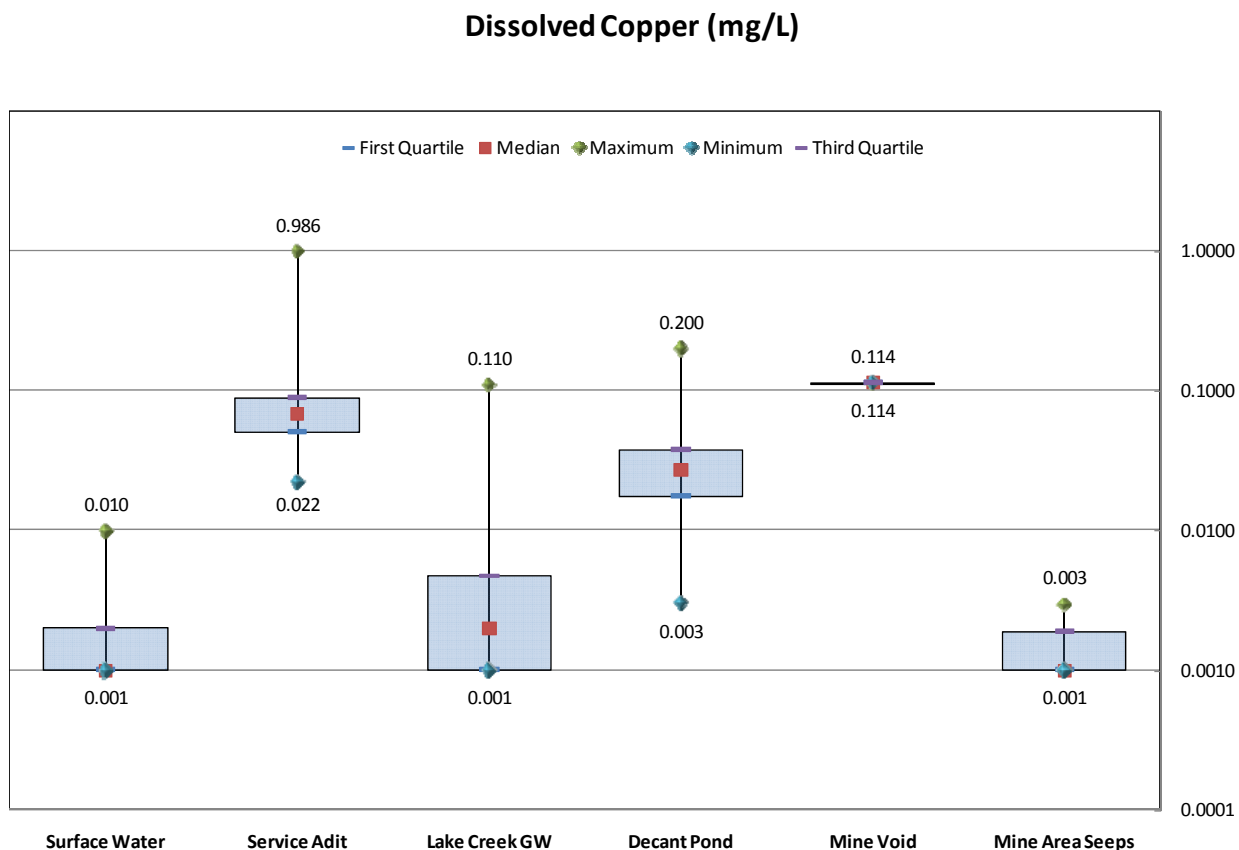


Figure 4 – Comparison of dissolved copper concentrations between the mine area surface water (seeps) and other waters on the site (note the logarithmic scale).

A comparison of the dissolved iron concentrations is shown in **Figure 5** below.

The highest dissolved iron concentrations (up to 69 mg/L) are associated with the deep groundwaters in the vicinity of the tailings impoundment. These waters have variable chemistry, as shown in the wide spread of the data points on the ternary diagram (Figure 3). The geochemical conditions are sometimes reducing, which results in high iron concentrations, while at other times are more oxidizing and have lower iron concentrations (see the *Troy Mine Copper Attenuation Study*, CDM, 2010, for more details). The wells completed in the Lake Creek Valley near the tailings impoundment (“MW” series) have shown erratic iron concentrations over time, which likely reflects a fluctuating redox gradient. The mine area surface waters generally have higher dissolved iron concentrations than the mine related waters, but not as high as the deep groundwaters in the vicinity of Lake Creek and the tailings impoundment.

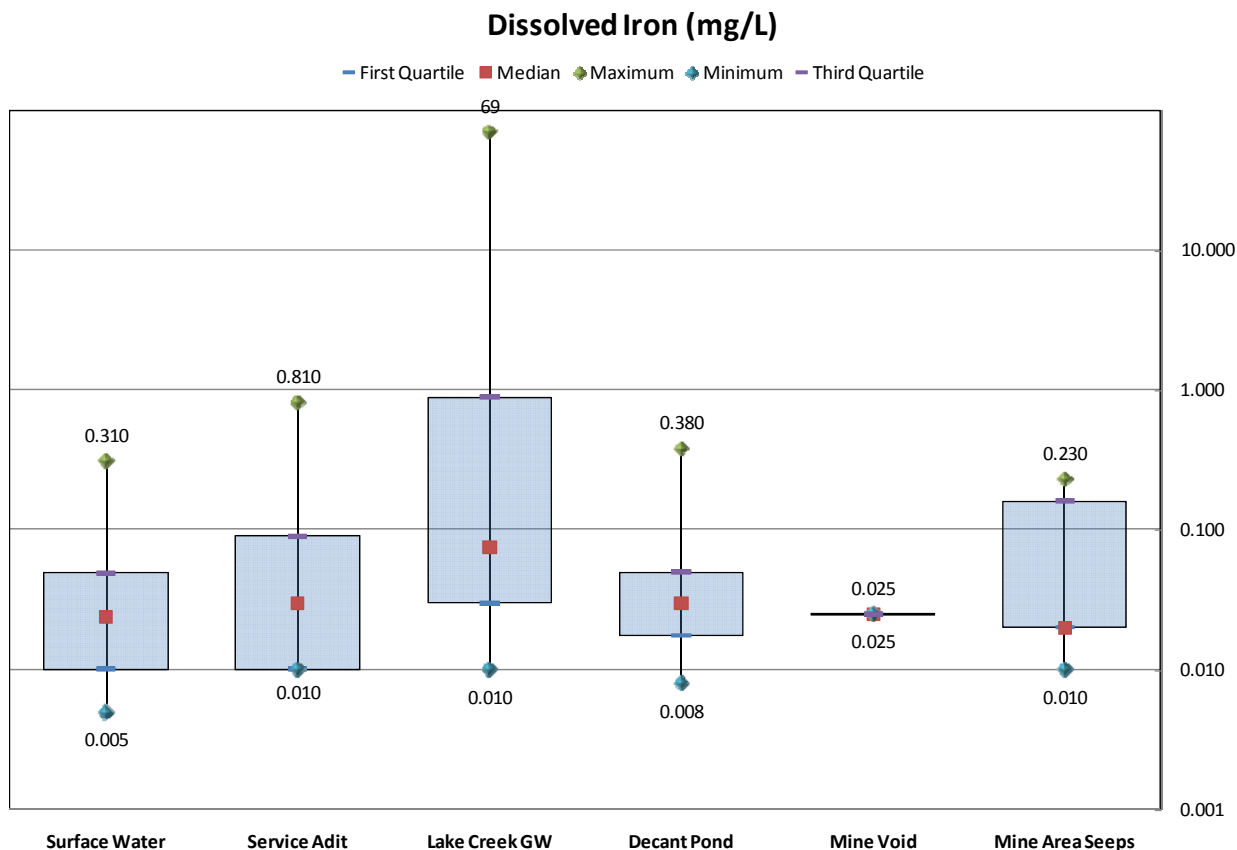
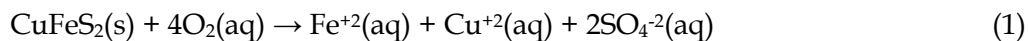


Figure 5 – Comparison of dissolved iron concentrations between the mine area surface water (seeps) and the other waters (note the logarithmic scale).

Post Closure Mine Water Quality

Following closure of the mine, the underground workings will be flooded up to the spillover point at the Service and Conveyor adits. Copper (Cu^{+2}) is released into water flooding the workings from the mineralized wall rocks by oxidation of the copper sulfide minerals present. The reaction requires an oxidant, such as dissolved oxygen (O_2) in order to occur. An example for the oxidation of chalcopyrite (CuFeS_2) is shown in equation 1.



Under flooded conditions, the dissolved oxygen concentrations will be expected to decrease within the flooded portions of the workings, and the rate of oxidation of copper sulfide minerals should decrease as well. Skousen et al. (2006) in a study of several coal mines in the eastern United States over a 35 year period found that after the first few mine volumes of water pass through the system the water becomes much less acid or even net alkaline, and metals concentrations decrease. Although the Troy Mine is not acid producing, a similar beneficial

effect of flooding on metals concentrations is expected. The quality of the mine water is expected to improve or remain constant after closure.

Nitrate and ammonia concentrations would also decrease once blasting is discontinued. Nitrogen compounds in adit water result from incomplete detonations or from poor housekeeping (spilled explosives). Explosives typically contain nitrogen compounds, which are present in the mine water as ammonia, and once oxidized, as nitrate and/or nitrite. Once the initial mine water is flushed from the system (by discharge at the spillover point and transport to the decant ponds), nitrate and ammonia concentrations would decline. In fact, nitrate+nitrite and ammonia concentrations within the disused and partially flooded portions of the mine are significantly lower than for the decant pond water, and to some extent, the adit discharge as shown in Table 3.

Table 3 – Comparison of Nitrate and Ammonia Concentrations for the Active and Inactive Portions of the Mine.

Water	Sample Date	Ammonia Concentration (mg/L)	Nitrate+Nitrite Concentration (mg/L)
Mine Water (disused area)	12/16/09	0.39	7.4
Decant Pond	12/16/09	11.2	37.6
Decant Pond (Decant 1)	5/14/09	6	17.6
Decant Pond (Decant)	5/14/09	6	19.4
Decant Pond	6/9/09	4.4	14.5
Adit Pipe	5/12/09	2.33	7.99

The difference in the concentrations of ammonia and nitrite+nitrate is not as pronounced for the adit discharge as for the decant pond. However, the adit discharge is composed of drainage from both used and disused portions of the mine, and would be expected to have an intermediate concentration between the active and disused water quality. The decant pond water, which is derived from the mill circuit, would be expected to have the highest ammonia and nitrate+nitrite as the crushed ore bears fresh blasting residue.

Table 4 shows data for the decant ponds and the service adit before, during and after the mine shutdown from April 1993 to December 2004. For both the decant ponds and the Service Adit, there is a decrease in nitrate plus nitrate during mine shutdown, indicating that mine water quality improved for this parameter when operations ceased. Again, the effect is more pronounced in decant ponds water than service adit water although the mill was not operating in this period. A similar decrease in nitrite plus nitrate is expected after final closure of the mine.

Table 4 – Comparison of Average Nitrate plus Nitrite Concentrations (mg/L) before, during and after Mine Shutdown.

Water	Pre-Shutdown	Shutdown	Post Shutdown
Decant Pond	17.6	1.72	12.0
Service Adit	10.0	1.43	7.24

Conclusions

The evaluation of the hydrogeology of the area surrounding the mine workings shows that most identified springs are well below the elevation of the present (pumped) water level in the mine. The geochemical evaluation indicates that identified surface waters in the vicinity of the mine are not being impacted by mine water. The expected water level rise after mine closure is relatively small compared to the current head difference between the mine water level and most surface water expressions in the mine area. Using the principles of groundwater flow, it is expected that this head difference will decrease in an outward direction from the mine and become no more than tens of feet in the vicinity of the springs. Thus, the rate of discharge from the existing surface water expressions may increase to some degree due to the increased head, but the location and spatial distribution of the expressions is not expected to change appreciably after mine closure.

The evaluation of the major ions and metals within the mine area surface water and other mine waters reveals that the mine area surface waters are chemically distinct from the deep groundwater and mine related waters but quite similar to area surface waters. The data suggest that the source of mine area surface water is not from the mine void, but is likely derived from precipitation or freshly recharged near-surface aquifers. After the mine is flooded at closure, it is expected that mine water will improve in quality as oxidation is reduced in the flooded portions of the mine void and concentrations of metals and nitrogen compounds decrease. Therefore, any effects of mine water that may reach surface water expressions are expected to be reduced after mine flooding.

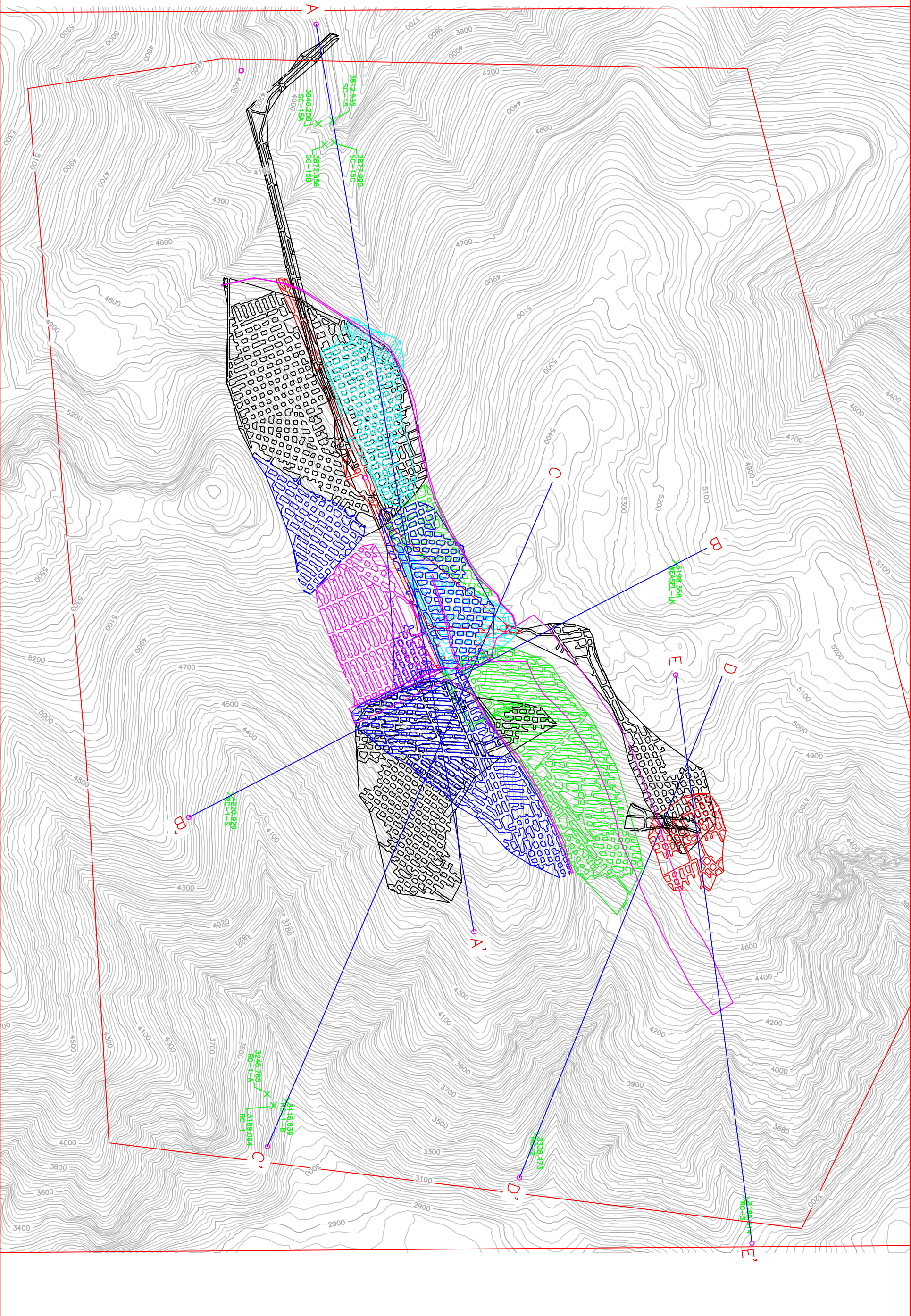
References

- CDM. 2010a. *Troy Mine Copper Attenuation Study – Secondary Processes*. Prepared for DEQ by CDM, Inc. July 2010.
- CDM, 2010b. *Mine Water Management Analysis – Troy Mine*. Technical Memorandum prepared for Montana Department of Environmental Quality and USDA Forest Service, Kootenai National Forest dated December 3, 2010.

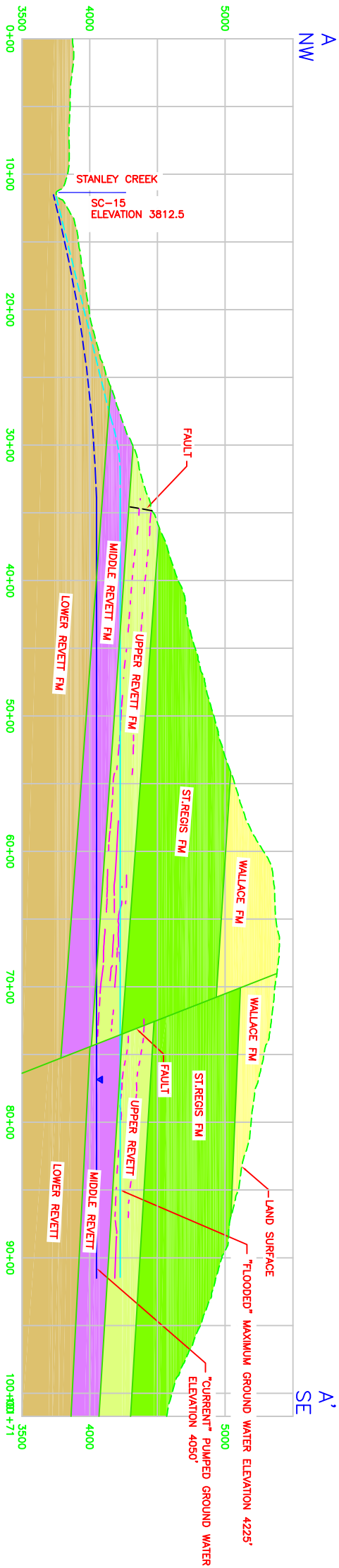
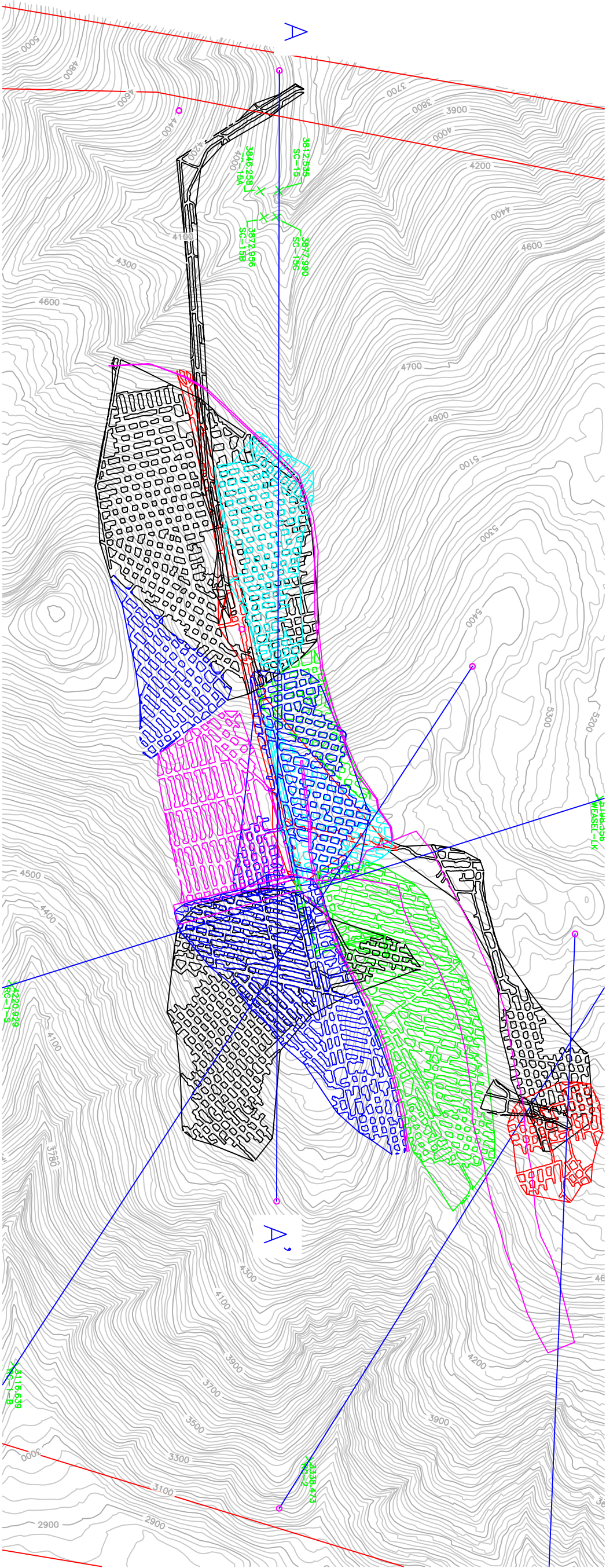
CDM, 2010c. Conceptual Design of New Water Line for Transport of Mine Water after Mine Closure. Technical Memorandum prepared for Montana Department of Environmental Quality and USDA Forest Service, Kootenai National Forest dated December 3, 2010.

Hydrometrics, 2009. Memorandum from Juliann L. Clum to Doug Parker regarding Troy Mine Seep/Spring Flow Data Compilation Procedures. Hydrometrics, Inc., Helena, Montana. April 8, 2009.

Skousen, J., L. McDonald, B. Mack, and J. Demchak, 2006. Water quality from above-drainage underground mines over a 35-year period. *Proceedings of the 7th International Conference on Acid Rock Drainage (ICARD), March 26-30, 2006, St. Louis, MO.*



OVERALL SITE PLAN



- LEGEND**
- LAND SURFACE
 - TOP OF MINE WORKINGS
 - CURRENT PUMPED ELEVATION
 - MAXIMUM FLOODED ELEVATION
 - FAULTS

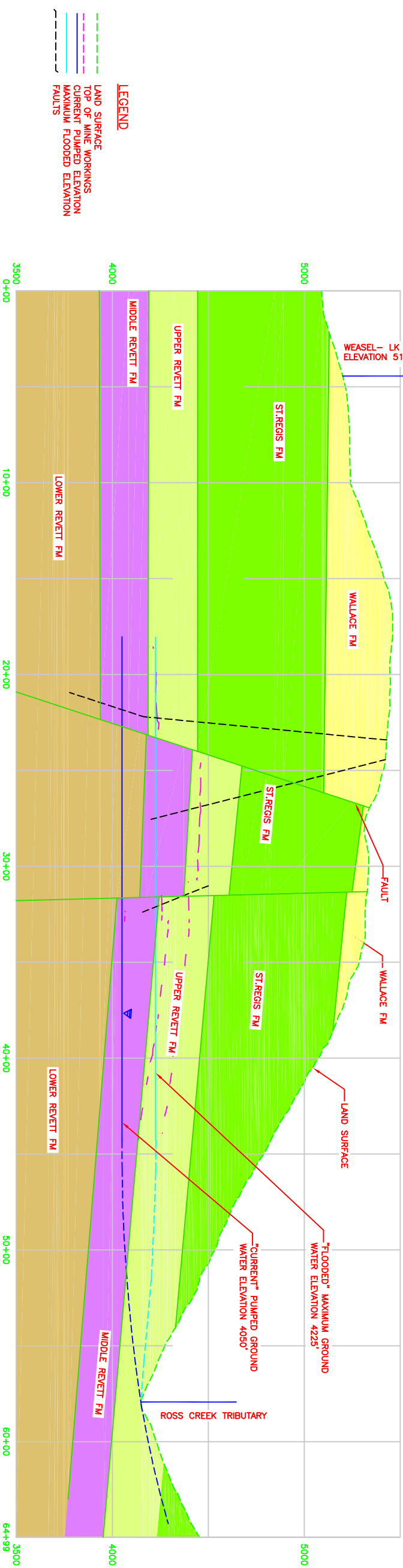
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2. FLOODED GROUNDWATER ELEVATION CONTROLLED BY HIGH POINT OF SERVICE ADIT.





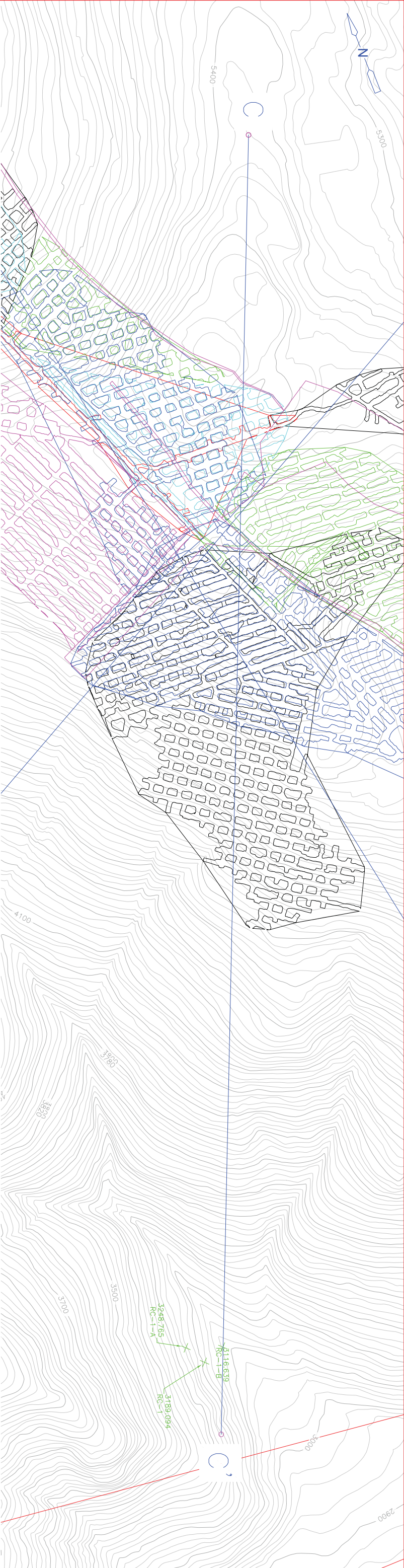
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2. FLOODED GROUNDWATER ELEVATION CONTROLLED BY HIGH POINT OF SERVICE ADIT.

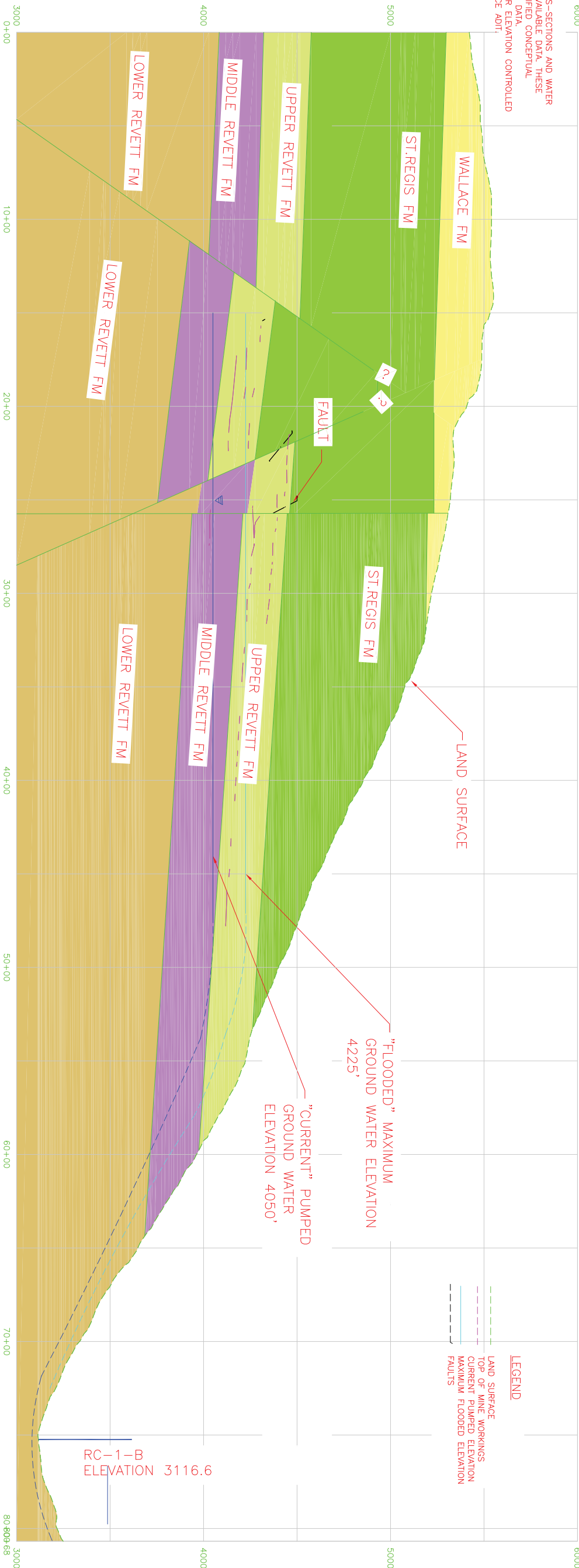




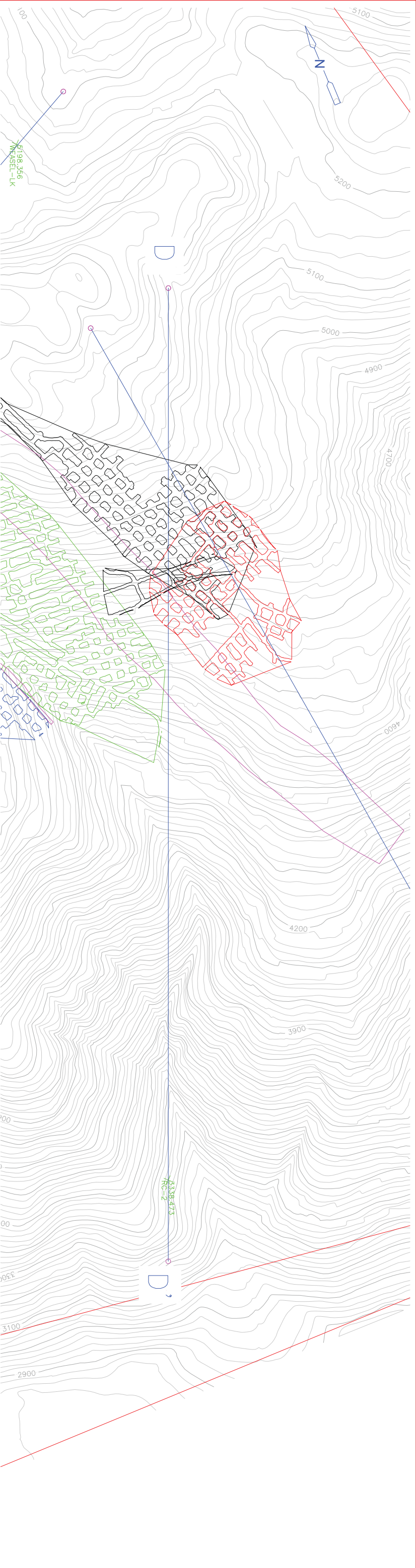
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2. FLOODED GROUNDWATER ELEVATION CONTROLLED BY HIGH POINT OF SERVICE ADIT.

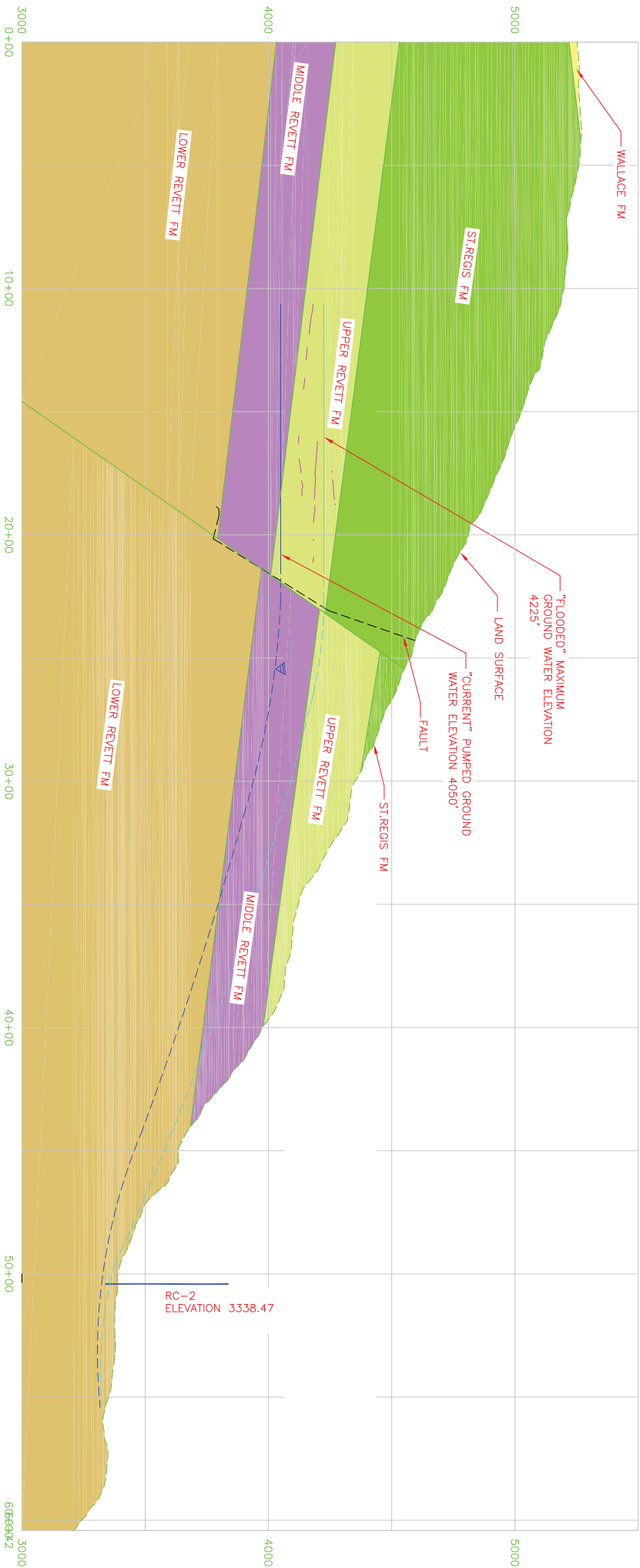


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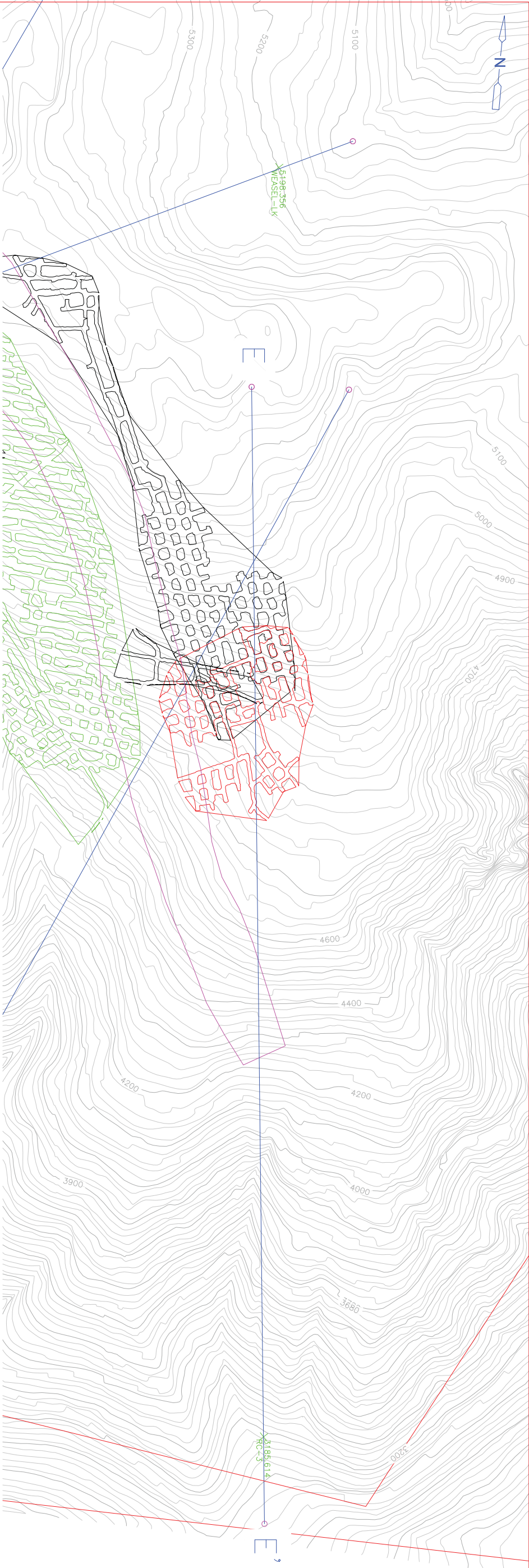


- LEGEND
- LAND SURFACE
 - TOP OF MINE WORKINGS
 - CURRENT PUMPED ELEVATION
 - MAXIMUM FLOODED ELEVATION
 - FAULTS

NOTES:
1. THE GEOLOGICAL CROSS-SECTIONS AND WATER LEVELS ARE BASED ON AVAILABLE DATA. THESE SECTIONS PRESENT SIMPLIFIED CONCEPTUAL INTERPRETATIONS OF THIS DATA.
2. FLOODED GROUNDWATER ELEVATION CONTROLLED BY HIGH POINT OF SERVICE ADIT.



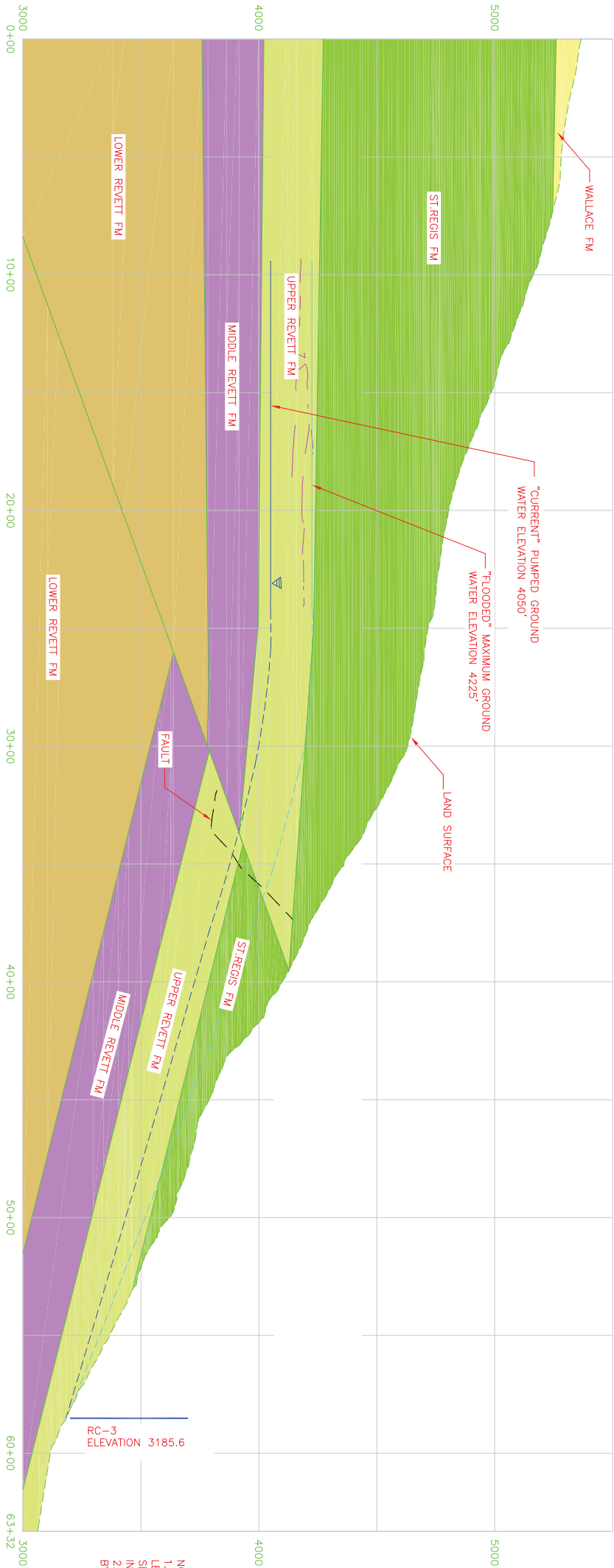
SECTION D



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- LEGEND**
- LAND SURFACE
 - TOP OF MINE WORKINGS
 - CURRENT PUMPED ELEVATION
 - MAXIMUM FLOODED ELEVATION
 - FAULTS



NOTES:
1. THE GEOLOGICAL CROSS-SECTIONS AND WATER LEVELS ARE BASED ON AVAILABLE DATA. THESE SECTIONS PRESENT SIMPLIFIED CONCEPTUAL INTERPRETATIONS OF THIS DATA.
2. FLOODED GROUNDWATER ELEVATION CONTROLLED BY HIGH POINT OF SERVICE ADIT.



SECTION E

Appendix G

**Troy Mine Conceptual Design of New
Water Line for Transport of Mine Water
after Mine Closure**



50 West 14th Street, Suite 200
Helena, Montana 59601
tel: 406 441-1400
fax: 406 449-7725

Technical Memorandum

To: *Emily Corsi, Herb Rolfes – DEQ
Bobbie Lacklen, John McKay - KNF*

From: *Darrel Stordahl, P.E., Bill Bucher, P.E., Terry Cowan, P.E. - CDM*

Date: *December 3, 2010*

Subject: *Conceptual Design of New Water Line for Transport of Mine Water
after Mine Closure – Troy Mine*

CDM has been retained by the Montana Department of Environmental Quality (DEQ) and the Kootenai National Forest (KNF) to prepare a conceptual design and cost estimate for a new pipeline and capture system at the Troy Mine. The pipeline will transport mine pool discharge through the service and conveyor adits to the infiltration pond located at the tailings impoundment via a new, buried pipeline. The purpose of this memorandum is to present the conceptual design and cost estimate to be used in decision making by the agencies.

The Troy Mine is an underground copper and silver mine located south of the town of Troy in Lincoln County, MT. ASARCO began operating the mine in 1982 and halted production in 1993 due to low metals prices. The mine was sold to Revett Silver Company and production resumed under Genesis, a subsidiary of Revett, in 2005. In 2006, Genesis submitted a Revised Reclamation Plan, for which an Environmental Assessment (EA) must be completed. In support of the EA, this memorandum analyzes a closure water management alternative presented in the EA.

Mine Discharge Data

It has been shown in the *Revised Mine Water Management Analysis – Troy Mine* and *Mine Water Balance Analysis – Troy Mine* Technical Memoranda (CDM, 2010a and CDM, 2010b) that the Troy Mine is expected to discharge mine pool water through the service and conveyor adits after closure. The *Revised Mine Water Management Analysis* memorandum estimates the maximum water discharge to be 6.9 cubic feet per second (cfs), and this rate is used in this memorandum for the conceptual design of the new closure water line. The discharge estimation is based on a limited amount of available data. However, it is consistent with observed maximum discharges measured by the mine over the period of operation. This

memorandum includes a backup system for the proposed design that should handle any reasonably foreseeable discharges.

Water Line Conceptual Design and Calculations

The water line was designed to follow the existing tailings/slurry pipes, approximate the land gradient, and have the hydraulic capacity to handle all of the estimated peak discharge of 6.9 cfs under gravity flow (no pressure flow). The land gradient is rather steep near the adit openings and flattens as it approaches the tailings impoundment. The upper quarter of the pipe run has slopes around 12%, the next quarter close to 3.5%, and the bottom half has slopes approximately 1.0% down to 0.3%. With the wide range of slopes, and their corresponding pipe capacities, multiple pipe sizes were used in the design to provide a cost savings. See the *Cost Estimate and Assumptions* portion of this report for a list of pipe sizes, lengths, and costs. The static capacity of the water line is about 250,000 cubic feet when full (see attached calculation sheet). This is the maximum amount of water that could drain from the pipeline should a break occur at the lowest point in the line. This amount could be reduced with the addition of automatically controlled valves along the length of the pipeline, but this design feature is not incorporated in this design.

Smooth HDPE was chosen as the pipe material for this conceptual design. The smooth interior of HDPE provides hydraulic efficiency, has a long service life (estimated at 75 years), is virtually chemically inert, is easy to install, is watertight, and provides an overall cost savings. Other materials such as iron, steel and concrete would not provide the equivalent performance at as low a cost.

FlowMaster software was used in the hydraulic calculations. Pipe lengths and slopes were estimated from *Exhibit A, Troy Mine Facilities Current Status of Genesis' March 2006 Troy Mine Revised Reclamation Plan*. FlowMaster calculations and a topographic map are shown in attached Appendix A.

Systems Control and Data Acquisition

Supervisory Control and Data Acquisition (SCADA) is included as part of the new water line. The SCADA can detect variability of flow in the pipe system and if a leak is present it will alert the proper personnel and divert mine adit discharge into the existing 10-12" tailings reclaim water line.

Leakage or breakage in the nearly six-mile pipe can be determined by using two flow meters, one placed close to the mine adit and the other placed close to the discharge end. The SCADA system would compare the two flow rates, with a time delay allotted to compensate for the length of pipe between the flowmeters. The flow rates should correspond within a small percentage range. If a leak or pipeline break occurs, the flow rate at the discharge would not correspond to the flow rate at the intake. The SCADA system alarm will be

activated, and a valve would reroute the water to another pipe. The existing pipe would be taken out of commission until repairs were made.

The flow meters would be powered using a solar-powered panel-mounted system with battery backup. The flow meter at the infiltration pond will transmit its flow rate to a compact programmable logic controller (PLC) located inside the solar power panel. The PLC will forward the flow rate by Ethernet radio to another compact PLC located in a panel close to the mine portal. The values will be compared, and, if the difference is outside a predetermined range, the PLC will close the pipe valve and reroute flow to the existing reclaim water line. The SCADA system will notify appropriate personnel of an alarm situation.

The existing reclaim water line will be used as a backup system. It is assumed that the existing reclaim water line consists of a 10" steel pipe on the upper, steeper section of the pipe run and a 12" steel pipe on the lower, flatter run. Assuming the pipeline follows the land gradient, it is estimated that the reclaim water line will handle approximately half of the estimated maximum of 6.9 cfs under gravity flow conditions. It is likely that the reclaim water line could handle the full estimated maximum flow if it is retrofitted with appropriate vacuum relief and pressure relief valves.

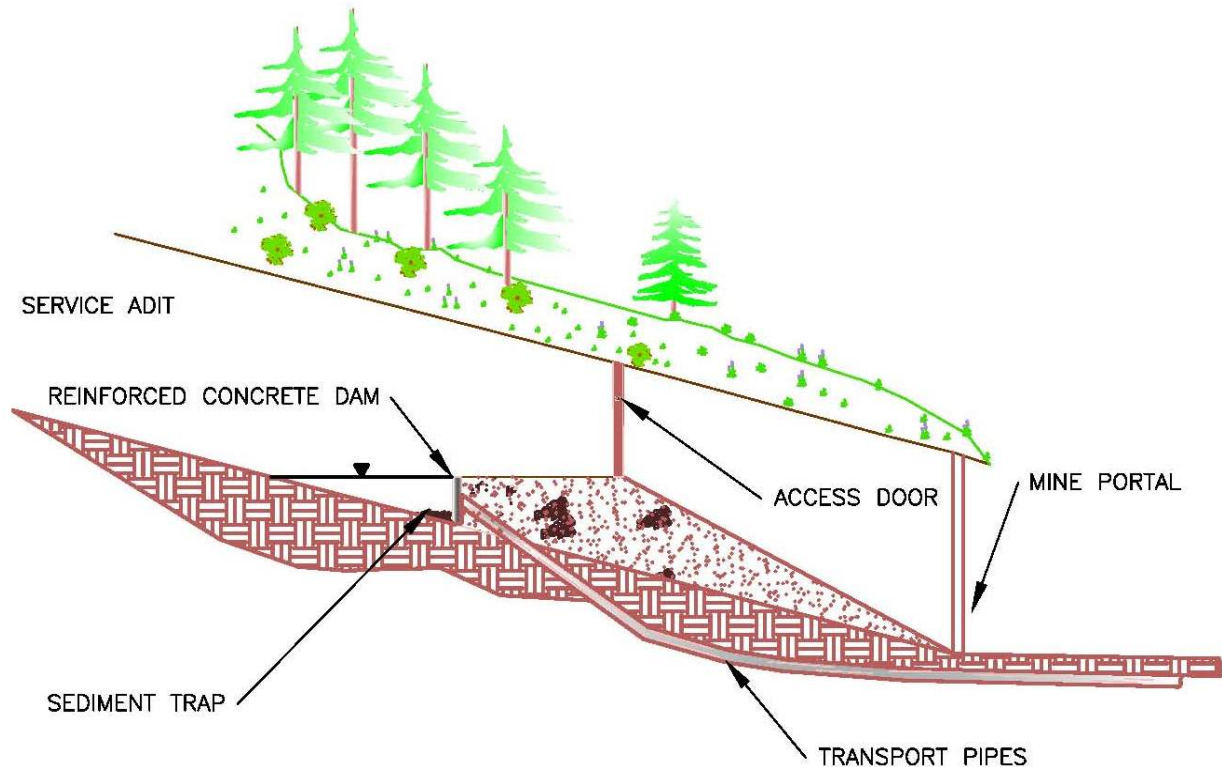
Mine Water Capture

Systems to capture the mine water and deliver it into the transport lines are needed at the service and conveyor adit portals. We propose to install these systems inside the adit to minimize freezing issues. As shown in Figure 1, the capture systems consist of concrete dams about six feet high constructed inside the adits about 50 feet back from the portal entrances. The entrance to the transport pipes would be set about two feet below the top of the dam so sediment transported by the mine discharge would be trapped below the pipe entrances. The 50-foot run of transport pipe between the dams and portals would be backfilled with material from the mine portal pads to bury the transport pipes and provide frost protection at the portals. To maintain access for maintenance of the sediment traps and transport pipe intakes, locking doors would be constructed at the portals. These doors will be sized to accommodate a small backhoe and truck to periodically clean out the sediment traps.

Small reservoirs would be created in the adits behind the concrete dams. The accumulated water could potentially infiltrate into fractures or unconsolidated material in the adit floors and walls and be released to the environment. Fractures and other leakage paths in the area below the elevation of the top of the concrete dam should be grouted to minimize any such releases.

The design has two transport pipes with their intake elevations staggered so the higher one would only come on line when flow into the mine increases. The lower intake would be for

Figure 1. Conceptual design of mine water capture system.

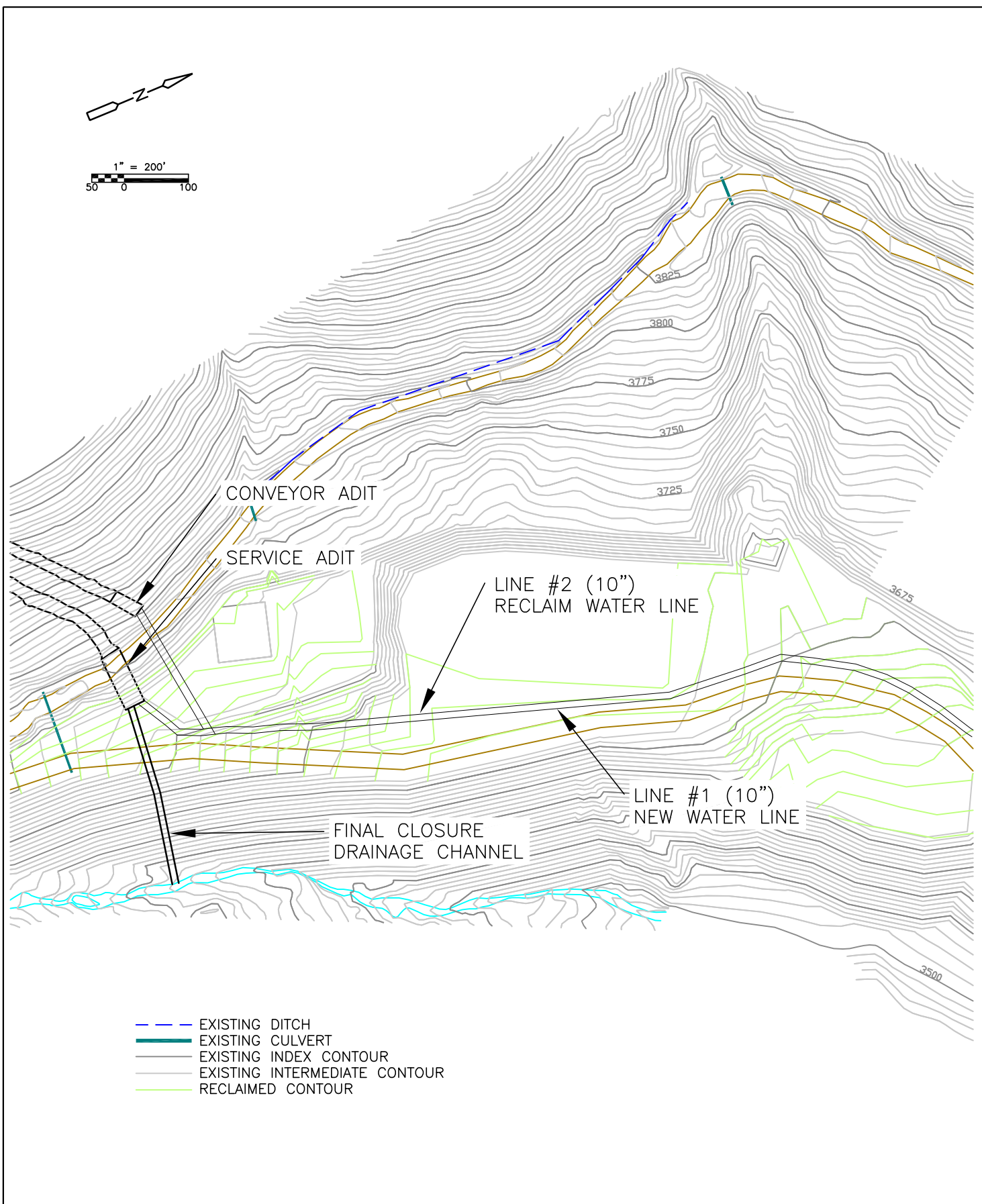


the new transport line described in this memorandum. The higher intake would tie to the existing reclaim water line. As the new transport line reaches capacity, flow would enter the reclaim water line without the need for a valve. With proper maintenance of the system, it is anticipated that any foreseeable discharge could be accommodated by the two lines flowing at full capacity.

In the unlikely event that the capacity of both these lines is exceeded, water would flow over the dam and pass through the rock backfill. At the mine portal it would drain to a constructed channel which would lead to Stanley Creek. This emergency channel would require a MPDES permit. Figure 2 is a conceptual design showing these water lines and the channel to Stanley Creek.

If, at some point in the future, the mine water quality improves to the point that it can be directly discharged to Stanley Creek, the capture system will be altered so the mine discharge

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is routed to a small channel that routes directly to Stanley Creek as shown on Figure 2. The transport lines will then be abandoned.

Cost Estimate and Assumptions

Table 1 is a preliminary cost estimate for construction of the capture system and new transport pipeline. Unit prices for the pipeline used in this cost estimate were taken from several contractors' bids for similar, prior work. The run of pipe would follow the existing tailings/slurry line as shown on *Exhibit A, Troy Mine Facilities Current Status* of Genesis's March 2006 *Troy Mine Revised Reclamation Plan* report (Genesis, 2006). Pipe slopes were estimated from topographic maps with the assumption that pipes will closely follow the land gradient. The pipe diameter was designed to correspond to the anticipated gradient of the pipe run. Pipe capture items were estimated using Means Costworks cost estimating software. The cost of installing buried power to the mine portal area for operation of the valves is included. The total cost estimate of about \$2.9 million includes installation of all components, mobilization, erosion control, 10% engineering, and 15% contingencies. Operation and maintenance costs are not included. Details on the cost estimate are attached in Appendix B.

Table 1. Cost Estimate

Unit Description	Quantity	Unit	Engineer's Estimate	
			Unit Price	Total Price
10" Smooth HDPE	7,708	LF (installed)	\$50	\$385,400
12" Smooth HDPE	5,604	LF (installed)	\$55	\$308,220
15" Smooth HDPE	10,246	LF (installed)	\$65	\$665,990
18" Smooth HDPE	6,763	LF (installed)	\$75	\$507,225
Reinforced Concrete Dam	2	LS	\$8,000	\$16,000
Tunnel Grouting	2	LS	\$9,900	\$19,800
Door construction	2	LS	\$10,000	\$20,000
Adit Backfill	466	CY	\$4	\$1,864
Flowmeters	2	each	\$10,500	\$21,000
Solar Panel	2	each	\$3,600	\$7,200
PLC	2	each	\$2,400	\$4,800
Ethernet Radio	2	each	\$2,100	\$4,200
Buried power	15,000	LF (installed)	\$0.50	\$7,500
Mobilization (10%)	-	-	-	\$196,170
Erosion Control (5%)	-	-	-	\$98,085
Engineering (10%)	-	-	-	\$226,345
Contingency (15%)	-	-	-	\$373,470
Total Estimate:				\$2,863,269

Conclusion

This technical memorandum was prepared to provide a conceptual design of a new, buried water line to transport mine pool discharge through the service and conveyor adits to the infiltration pond located in the tailings impoundment. The design flow of 6.9 cfs is estimated to be the maximum discharge (CDM, 2010a). The new pipe system conceptual design consists of 10", 12", 15", and 18" smooth HDPE pipes operating under gravity flow. The smooth interior of HDPE provides hydraulic efficiency, has a long service life (estimated at 75 years), is virtually chemically inert, is easy to install, is watertight, and provides an overall cost savings as compared to other materials. SCADA would be incorporated to detect leaks in the pipe line, and, if needed, alert personnel and automatically divert flow into the existing water reclamation line while any leak is repaired. The existing reclaim water line will be kept as a backup transport pipe should the capacity of the new pipe be exceeded or the new pipe require repair. Water collection systems have been conceptually designed inside the adit entrances to capture the mine pool water and direct it into the transport pipes. The cost estimate to construct this project is about \$2.8 million.

References

- CDM, 2010a. Mine Water Management Analysis – Troy Mine. Prepared for Montana Department of Environmental Quality and Kootenai National Forest, December 3, 2010.
- CDM, 2010b. Mine Water Balance Analysis - Troy Mine. Technical Memorandum to Emily Corsi and Herb Rolfes, Montana Department of Environmental Quality dated April 12, 2010.
- Genesis, Inc., 2006. Troy Mine Revised Reclamation Plan. March, 2006. Troy, MT.

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Appendix A

FlowMaster Calculations and Topo Map

Worksheet for 10" Smooth HDPE

Project Description

Friction Method Manning Formula
Solve For Full Flow Slope

Input Data

Roughness Coefficient	0.012	
Channel Slope	8.45157	%
Normal Depth	10.00	in
Diameter	10.00	in
Discharge	6.90	ft ³ /s

Results

Channel Slope	8.45157	%
Normal Depth	10.00	in
Flow Area	0.55	ft ²
Wetted Perimeter	2.62	ft
Top Width	0.00	ft
Critical Depth	0.83	ft
Percent Full	100.0	%
Critical Slope	0.07994	ft/ft
Velocity	12.65	ft/s
Velocity Head	2.49	ft
Specific Energy	3.32	ft
Froude Number	0.00	
Maximum Discharge	7.42	ft ³ /s
Discharge Full	6.90	ft ³ /s
Slope Full	0.08452	ft/ft
Flow Type	SubCritical	

GVF Input Data

Downstream Depth	0.00	in
Length	0.00	ft
Number Of Steps	0	

GVF Output Data

Upstream Depth	0.00	in
Profile Description		
Profile Headloss	0.00	ft
Average End Depth Over Rise	0.00	%
Normal Depth Over Rise	100.00	%

Worksheet for 10" Smooth HDPE

GVF Output Data

Downstream Velocity	Infinity	ft/s
Upstream Velocity	Infinity	ft/s
Normal Depth	10.00	in
Critical Depth	0.83	ft
Channel Slope	8.45157	%
Critical Slope	0.07994	ft/ft

Worksheet for 12" Smooth HDPE

Project Description

Friction Method	Manning Formula
Solve For	Full Flow Slope

Input Data

Roughness Coefficient	0.012
Channel Slope	3.19622 %
Normal Depth	12.00 in
Diameter	12.00 in
Discharge	6.90 ft ³ /s

Results

Channel Slope	3.19622 %
Normal Depth	12.00 in
Flow Area	0.79 ft ²
Wetted Perimeter	3.14 ft
Top Width	0.00 ft
Critical Depth	0.97 ft
Percent Full	100.0 %
Critical Slope	0.02830 ft/ft
Velocity	8.79 ft/s
Velocity Head	1.20 ft
Specific Energy	2.20 ft
Froude Number	0.00
Maximum Discharge	7.42 ft ³ /s
Discharge Full	6.90 ft ³ /s
Slope Full	0.03196 ft/ft
Flow Type	SubCritical

GVF Input Data

Downstream Depth	0.00 in
Length	0.00 ft
Number Of Steps	0

GVF Output Data

Upstream Depth	0.00 in
Profile Description	
Profile Headloss	0.00 ft
Average End Depth Over Rise	0.00 %
Normal Depth Over Rise	100.00 %

Worksheet for 12" Smooth HDPE

GVF Output Data

Downstream Velocity	Infinity	ft/s
Upstream Velocity	Infinity	ft/s
Normal Depth	12.00	in
Critical Depth	0.97	ft
Channel Slope	3.19622	%
Critical Slope	0.02830	ft/ft

Worksheet for 15" Smooth HDPE

Project Description

Friction Method	Manning Formula
Solve For	Full Flow Slope

Input Data

Roughness Coefficient	0.012	
Channel Slope	0.97226	%
Normal Depth	15.00	in
Diameter	15.00	in
Discharge	6.90	ft ³ /s

Results

Channel Slope	0.97226	%
Normal Depth	15.00	in
Flow Area	1.23	ft ²
Wetted Perimeter	3.93	ft
Top Width	0.00	ft
Critical Depth	1.05	ft
Percent Full	100.0	%
Critical Slope	0.00927	ft/ft
Velocity	5.62	ft/s
Velocity Head	0.49	ft
Specific Energy	1.74	ft
Froude Number	0.00	
Maximum Discharge	7.42	ft ³ /s
Discharge Full	6.90	ft ³ /s
Slope Full	0.00972	ft/ft
Flow Type	SubCritical	

GVF Input Data

Downstream Depth	0.00	in
Length	0.00	ft
Number Of Steps	0	

GVF Output Data

Upstream Depth	0.00	in
Profile Description		
Profile Headloss	0.00	ft
Average End Depth Over Rise	0.00	%
Normal Depth Over Rise	100.00	%

Worksheet for 15" Smooth HDPE

GVF Output Data

Downstream Velocity	Infinity	ft/s
Upstream Velocity	Infinity	ft/s
Normal Depth	15.00	in
Critical Depth	1.05	ft
Channel Slope	0.97226	%
Critical Slope	0.00927	ft/ft

Worksheet for 18" Smooth HDPE

Project Description

Friction Method	Manning Formula
Solve For	Full Flow Slope

Input Data

Roughness Coefficient	0.012	
Channel Slope	0.36769	%
Normal Depth	18.00	in
Diameter	18.00	in
Discharge	6.90	ft ³ /s

Results

Channel Slope	0.36769	%
Normal Depth	18.00	in
Flow Area	1.77	ft ²
Wetted Perimeter	4.71	ft
Top Width	0.00	ft
Critical Depth	1.02	ft
Percent Full	100.0	%
Critical Slope	0.00571	ft/ft
Velocity	3.90	ft/s
Velocity Head	0.24	ft
Specific Energy	1.74	ft
Froude Number	0.00	
Maximum Discharge	7.42	ft ³ /s
Discharge Full	6.90	ft ³ /s
Slope Full	0.00368	ft/ft
Flow Type	SubCritical	

GVF Input Data

Downstream Depth	0.00	in
Length	0.00	ft
Number Of Steps	0	

GVF Output Data

Upstream Depth	0.00	in
Profile Description		
Profile Headloss	0.00	ft
Average End Depth Over Rise	0.00	%
Normal Depth Over Rise	100.00	%

Worksheet for 18" Smooth HDPE

GVF Output Data

Downstream Velocity	Infinity	ft/s
Upstream Velocity	Infinity	ft/s
Normal Depth	18.00	in
Critical Depth	1.02	ft
Channel Slope	0.36769	%
Critical Slope	0.00571	ft/ft

Worksheet for Existing 10" Steel

Project Description

Friction Method	Manning Formula
Solve For	Full Flow Capacity

Input Data

Roughness Coefficient	0.012
Channel Slope	3.60000 %
Normal Depth	10.00 in
Diameter	10.00 in
Discharge	4.50 ft ³ /s

Results

Discharge	4.50 ft ³ /s
Normal Depth	10.00 in
Flow Area	0.55 ft ²
Wetted Perimeter	2.62 ft
Top Width	0.00 ft
Critical Depth	0.81 ft
Percent Full	100.0 %
Critical Slope	0.03203 ft/ft
Velocity	8.26 ft/s
Velocity Head	1.06 ft
Specific Energy	1.89 ft
Froude Number	0.00
Maximum Discharge	4.84 ft ³ /s
Discharge Full	4.50 ft ³ /s
Slope Full	0.03600 ft/ft
Flow Type	SubCritical

GVF Input Data

Downstream Depth	0.00 in
Length	0.00 ft
Number Of Steps	0

GVF Output Data

Upstream Depth	0.00 in
Profile Description	
Profile Headloss	0.00 ft
Average End Depth Over Rise	0.00 %
Normal Depth Over Rise	100.00 %

Worksheet for Existing 10" Steel

GVF Output Data

Downstream Velocity	Infinity	ft/s
Upstream Velocity	Infinity	ft/s
Normal Depth	10.00	in
Critical Depth	0.81	ft
Channel Slope	3.60000	%
Critical Slope	0.03203	ft/ft

Worksheet for Existing 12" Steel

Project Description

Friction Method	Manning Formula
Solve For	Full Flow Capacity

Input Data

Roughness Coefficient	0.012
Channel Slope	1.03000 %
Normal Depth	12.00 in
Diameter	12.00 in
Discharge	3.92 ft ³ /s

Results

Discharge	3.92 ft ³ /s
Normal Depth	12.00 in
Flow Area	0.79 ft ²
Wetted Perimeter	3.14 ft
Top Width	0.00 ft
Critical Depth	0.84 ft
Percent Full	100.0 %
Critical Slope	0.00988 ft/ft
Velocity	4.99 ft/s
Velocity Head	0.39 ft
Specific Energy	1.39 ft
Froude Number	0.00
Maximum Discharge	4.21 ft ³ /s
Discharge Full	3.92 ft ³ /s
Slope Full	0.01030 ft/ft
Flow Type	SubCritical

GVF Input Data

Downstream Depth	0.00 in
Length	0.00 ft
Number Of Steps	0

GVF Output Data

Upstream Depth	0.00 in
Profile Description	
Profile Headloss	0.00 ft
Average End Depth Over Rise	0.00 %
Normal Depth Over Rise	100.00 %

Worksheet for Existing 12" Steel

GVF Output Data

Downstream Velocity	Infinity	ft/s
Upstream Velocity	Infinity	ft/s
Normal Depth	12.00	in
Critical Depth	0.84	ft
Channel Slope	1.03000	%
Critical Slope	0.00988	ft/ft

TROY MINE - NEW WATER LINE VOLUME

Diam (in)	Length (ft)	Volume (cf)
10	7708	31430
12	5604	32906
15	10246	94004
18	6763	89350
Total:		247690

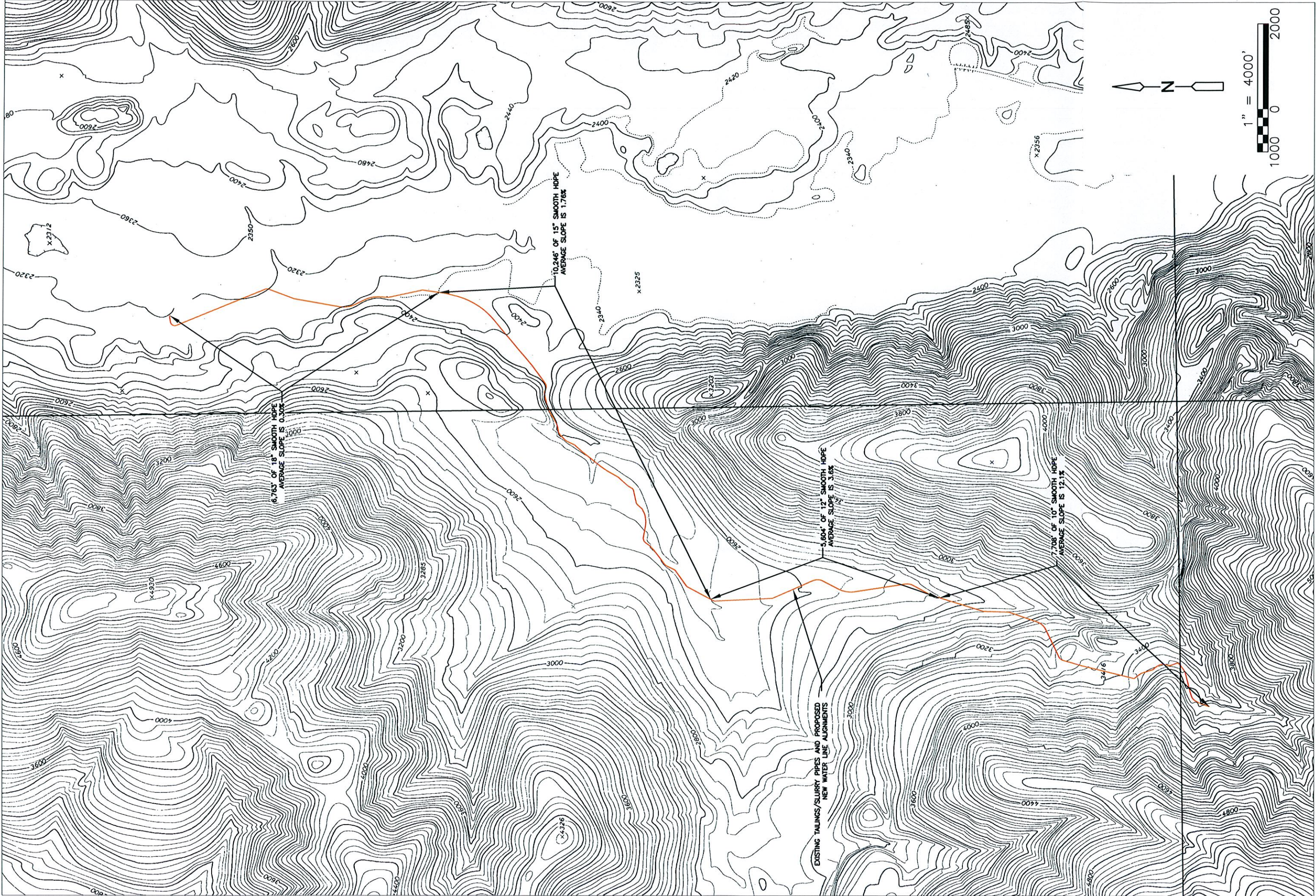


Figure No. 1
Plan View

Appendix B

Cost Estimate

Troy Mine
Mine Pool Water Capture System
Cost Estimate - Single Adit
July 1, 2010

			Engineer's Estimate		
Unit Description	Quantity	Unit	Unit Price	Total Price	Mean's Code
Concrete Dam					
Forms	240	sq ft	\$7.65	\$1,836	03 11 13 85 2000
Rebar	0.2	ton	\$2,126.25	\$425	03 21 10 60 700 + 1000
Concrete	8.9	cy	\$143.00	\$1,273	03 31 05 35 0300
Place Concrete (pumped)	8.9	cy	\$28.00	\$249	
Subtotal - Dam				\$3,783	
Given Location, double cost.				\$8,000	
Tunnel Grouting					
Mobilization				\$4,000	Engineer's Estimate
Grout	100	cu ft	\$59.00	\$5,900	31-73-13-10-0820
Subtotal - Dam				\$9,900	
Door Construction					
	1	ea	\$10,000.00	\$10,000	Engineer's Estimate
Backfill					
	233	cu yd	\$4.00	\$932	Engineer's Estimate

Total Cost: **\$28,832**

Transport pipe cost in pipeline cost estimate.

Appendix H
Mine Water Plume Location and
Identification Phase I Results and Phase 2
& 3 Work Plan

Mine Water Plume Location and Identification
Phase 1 Results and Phase 2 & 3 Workplan
Sterling Troy Mine

Prepared for:

Genesis Inc./Sterling Mining Company
424 S. Sullivan Road
Veradale, WA 99037

Prepared by:

Hydrometrics, Inc.[®]
22 Second Ave West, Suite 1100
Kalispell, MT 59901

October 2001

REVIEW DRAFT

TABLE OF CONTENTS

LIST OF TABLES	ii
LIST OF FIGURES.....	ii
LIST OF APPENDICES	ii
LIST OF ATTACHMENTS.....	ii
1.0 INTRODUCTION.....	1-1
1.1 BACKGROUND.....	1-2
1.2 PHASE 1 SUMMARY.....	1-4
2.0 METHODS.....	2-1
2.1 SAMPLING AND EVALUATION OF MONITORING WELLS.....	2-1
2.2 TRACER TEST.....	2-1
3.0 PHASE 1 RESULTS	3-1
3.1 GROUNDWATER SAMPLING AND EVALUATION	3-1
3.2 TRACER TEST.....	3-1
3.2.1 Chemical Effects of the Tracer on Groundwater	3-1
3.2.2 Decant Pond Infiltration Rate.....	3-2
4.0 PHASE 2 AND 3 WORKPLAN.....	4-1
4.1 CONCEPTUAL MODEL OF HYDROGEOLOGIC SYSTEM.....	4-1
4.2 PHASE 2 MONITORING WELL INSTALLATION AND TESTING.....	4-3
4.3 PHASE 3 GROUNDWATER TRACER TEST.....	4-5
5.0 REFERENCES.....	5-1

REVIEW DRAFT

LIST OF TABLES

TABLE 1 SEPTEMBER 2000 WATER QUALITY DATA.....	3-1
TABLE 2. MARCH 2001 WATER QUALITY DATA (BEFORE TRACER TEST)	3-2
TABLE 3. APRIL 2001 WATER QUALITY DATA (AFTER TRACER ADDITION PERIOD).....	3-3
TABLE 4. IDENTIFIED PERMEABLE ZONES IN THE VICINITY OF THE DECANT PONDS	4-3

LIST OF FIGURES

FIGURE 1 MONITORING WELL LOCATIONS	2-1
FIGURE 2 TYPICAL SHALLOW MONITORING WELL CONSTRUCTION	4-7

REVIEW DRAFT

LIST OF APPENDICES

APPENDIX A. GROUNDWATER MONITORING PROPOSAL

(August 16, 2000 letter from Genesis Inc. to MDEQ)

APPENDIX B. TRACER TEST PROPOSAL, REQUEST AND APPROVAL

(December 4, 2001 letter from Genesis Inc. to MDEQ)

(January 22, 2001 letter from MDEQ to Genesis Inc.)

APPENDIX C. TRACER TEST PROTOCOL REVISION 1

(February 2, 2001 letter from Hydrometrics to MDEQ)

APPENDIX D. TRACER TEST PROTOCOL REVISION 2

(March 6, 2001 letter from Hydrometrics to MDEQ)

APPENDIX E. PIPER DIAGRAMS

APPENDIX F. TRACER TEST RESULTS – SC, Na, CL Trendplots

REVIEW DRAFT

Mine Water Plume Location and Identification

Phase 1 Results and Phase 2 & 3 Workplan
Genesis Inc. Troy Mine

1.0 INTRODUCTION

This report summarizes results of investigations in 2000 and 2001 that were conducted to locate and identify the minewater flowpath from the decant pond at the Troy Mine tailings impoundment. Additionally, this report describes the workplan for continued investigation of the minewater flowpath in 2001 and 2002.

Phase 1 results continue to support the hypothesis that copper in minewater that is discharged to the decant ponds is strongly attenuated (immobilized) in the groundwater system, probably within a short distance of the decant ponds. Groundwater quality data suggests that the groundwater in deeper gravel/sand units near the decant ponds (e.g., units monitored by wells IW-1, MW-95-7, and MW-95-8) is derived primarily from minewater disposed in the decant ponds. This groundwater does not contain elevated copper concentrations (0.003 mg/L or less). However, tracer test results suggest that these gravel units are not the primary groundwater flowpath for decant pond water. Based on Phase 1 results and review of the local geology, it is hypothesized that a shallow gravel unit may provide the primary flowpath for transport of water from the decant ponds. While this shallow unit is believed to transport water primarily laterally away from the decant pond, some slow transport of water downward to the deeper gravels (i.e., leakage through the silts/fine sands that underlie the shallow gravel) is believed to also occur. The slow rate of the vertical transport of water to the deeper gravels is believed to account for the lack of detection of the tracer in the deeper gravels during the 2001 tracer test. Phases 2 and 3 of the evaluation include additional monitoring well drilling targeting the shallow gravel unit and an additional groundwater tracer test.

REVIEW DRAFT

1.1 BACKGROUND

The Troy Mine was developed by Asarco in 1979 and went into production in 1981. The mine was operated by Asarco until 1993 when the mine was placed on temporary shutdown. In 1999, the mine was purchased by Genesis Inc., a subsidiary of Sterling Mining Company. Since 1993, Asarco and Genesis Inc. have disposed of water from the underground workings by discharge to the impoundment in one of three ways:

- Sprinkler irrigation over the impoundment (primarily during the growing season).
- Discharge to the decant pond (also known as the barge pond).
- Discharge to the Section III pond.

The primary chemical of concern in mine water is copper, which is typically present at concentrations of approximately 0.1 mg/L. For comparison, surface water in Lake Creek adjacent to the impoundment typically contains approximately 0.003 mg/L and the chronic aquatic life standard for copper is approximately 0.004 mg/L (assuming hardness of 35 mg/L). In spite of the higher copper concentrations discharged to the impoundment, copper concentrations in Lake Creek and in monitoring wells installed in the impoundment vicinity have not increased measurably as a result of the discharge (*see* Summit, 1999; Parametrix, 1999).

The observation that mine water has been discharged with little or no apparent effect on ambient water quality has led to the preliminary consideration of continued discharge of mine water to the decant pond as a potential long-term method of mine water disposal (see Revised Reclamation Plan, Sterling 2000). In order for impoundment discharge to be considered further, both the physical flow paths and the transport and attenuation mechanisms for copper in the impoundment/groundwater system must be understood. In particular, the flow paths must be identified so that an effective groundwater monitoring system can be installed. Transport and attenuation mechanisms must be understood to determine the long-term fate of constituents in mine water and the effects of long-term discharge on ambient groundwater and stream quality.

REVIEW DRAFT

In September 2000, Hydrometrics prepared a “Minewater Fate and Transport Study Plan” (Hydrometrics, 2000a) to collect information needed to evaluate the fate and transport of mine water that is disposed at the decant pond. An initial obstacle to the study of fate and transport in the system is the absence of any evidence of a copper-bearing mine water discharge plume in the groundwater system. Therefore, the first phase of the fate and transport study is to locate and identify the groundwater flow paths for discharged mine water. Later phases of the study plan will identify probable fate and transport mechanisms, and evaluate potential implications under a long-term discharge scenario.

Plume location and identification is anticipated to consist of three phases as follows:

1. Sampling and evaluation of all available monitoring wells and further evaluation of water quality data.
2. Construction, sampling, and evaluation of additional excavations, borings, and wells to supplement the existing monitoring well network.
3. Tracer testing to verify plume location and flow paths

REVIEW DRAFT

1.2 PHASE 1 SUMMARY

As initially proposed in the study plan, Phase 1 of plume location consisted of sampling and evaluation of all available monitoring wells and further evaluation of water quality data. Based on the water quality evaluation it was thought that one of the existing wells (IW-1, the “irrigation well”) was effectively monitoring the mine water plume. Therefore, a preliminary tracer test was conducted in spring 2001 to test the hypothesis that well IW-1 is located in, and monitors, the minewater plume. Results of this tracer test (summarized in Section 3 of this report) indicate that none of the existing wells is located in the primary flowpath from the pond. Based on these Phase 1 results, it is determined that additional investigations as outlined in Phases 2 and 3 of the Fate and Transport Study are needed. Workplans for Phases 2 and 3 are described in Section 4 of this report.

REVIEW DRAFT

2.0 METHODS

Monitoring and testing plans for Phase 1 work are included in Appendices A through D. The following sections provide an overview of the plans and methods. Locations of wells near the tailings impoundment are shown on Figure 1.

2.1 SAMPLING AND EVALUATION OF MONITORING WELLS

A proposed plan for groundwater monitoring in and near the tailings impoundment was submitted to MDEQ by Genesis Inc. in August 2000. This plan (see Appendix A) proposed sampling of 9 wells near the impoundment (MW-1, MW-2, MW-3, MW-4, MW-95-5, MW-95-7, MW-95-8, IW-1 or the Irrigation well, and the Quonset hut well). All of the wells were sampled on September 7, 2000 except for MW-95-7 which was dry. Wells MW-95-5 and MW-95-8 were sampled by bailing with disposable plastic bailers. The remaining wells were sampled with submersible pumps.

2.2 TRACER TEST

Tracer testing at the decant pond consisted of adding salt (NaCl or sodium chloride) to the mine water flowing to the decant pond to raise the TDS, specific conductivity, sodium, and chloride concentrations of water in the decant pond to levels much higher than ambient groundwater. At the same time, monitoring of groundwater in adjacent wells was conducted in order to identify increases in these chemical parameters in groundwater and thus locate the minewater flowpath in the groundwater system. Methods of conducting the tracer test were originally described in a December 4, 2001 letter from Sterling Mining Co. (Frank Duval) to MDEQ (Abe Horpestad) proposing the test and requesting permission from DEQ to run the test (see Appendix B). As originally proposed, the tracer test was to consist of adding salt to the minewater inflow to the decant pond at the rate of approximately 500 pounds per day for a period of about 18 days. MDEQ granted approval to conduct the tracer test on January 22, 2001 (see Appendix B).

FIGURE 1 MONITORING WELL LOCATIONS

REVIEW DRAFT

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In February 2001, it was recognized that minewater flows were sufficiently low that all minewater flows to the decant pond could be suspended for a period of about 2 weeks. Without minewater inflows to the pond, it was believed that it would be easier to simply add the salt to the pond as a “slug” and eliminate the need for daily salt additions. Moreover, it was believed that slug addition of salt would result in more constant salt concentrations in the pond and overall a better tracer test. The testing protocol was revised in a February 2, 2001 letter to MDEQ (see Appendix C). The revised testing protocol also provided a monitoring schedule and analytical parameter list.

During preparation for the tracer test it was discovered that the system for remote monitoring of some wells in the tailings impoundment area were not functioning, perhaps as the result of snowplowing to access the wells. Therefore, the monitoring schedule for the tracer test was revised to eliminate remote monitoring (see Appendix D).

The tracer test was conducted during the period of March 8 to May 31, 2001. In general, monitoring during the test followed the schedule outlined in Table 1 of Appendix D. A timeline and summary of the tracer test is as follows:

- March 8 – All wells and decant ponds were sampled to collect baseline water quality data for major ions, nutrients, and metals prior to tracer (salt) addition. Decant pond cell #3 and well MW-97-14 were dry and could not be sampled.
- March 12 – Salt addition started at 8 am; 5 tons of salt were added by 9:50 am and 10 tons of salt were added by 11:45 am. Salt was added to the reclaim water pipeline at the mill pond. At 2:30 pm, monitoring of the decant pond specific conductivity was started.
- March 13 – Mine water flows to decant pond were shut off at 5 pm. Decant pond staff gage installed. Decant pond specific conductivity monitoring continued (3 times daily) and daily sampling of wells IW-1, MW-95-3, MW-95-7, and MW-95-8 was initiated.
- March 16 – Decant pond monitoring continued on a daily basis.

REVIEW DRAFT

- March 14 through April 2 – Daily sampling of wells (IW-1, MW-95-3, MW-95-7, and MW-95-8) and decant ponds continued. (March 18 – Decant pond #3 dry and no longer sampled. March 30 – Well MW-95-8 dry and could not be sampled.)
- April 2 – All wells, including MW-1 through MW-4, MW-95-5, and MW-97-12 sampled prior to reduction in monitoring frequency to weekly.
- April 9 – Mine water flow to decant pond resumed; water from section 3 pond allowed to flow to decant pond.
- April 9 through May 31 – Weekly sampling of wells and decant ponds continued. (April 25, decant pond #2 dry and could no longer be sampled. May 31, well MW-95-8 again contained water and was sampled.)
- May 31 – Tracer test terminated.

REVIEW DRAFT

3.0 PHASE 1 RESULTS

Phase 1 results continue to support the hypothesis that copper in minewater that is discharged to the decant ponds is strongly attenuated (immobilized) in the groundwater system, probably within a short distance of the decant ponds. Groundwater quality data suggests that the groundwater in deeper gravel/sand units near the decant ponds (e.g., units monitored by wells IW-1, MW-95-7, and MW-95-8) is derived primarily from minewater disposed in the decant ponds. This groundwater does not contained elevated copper concentrations (0.003 mg/L or less). However, tracer test results suggests that these gravel units are not the primary groundwater flowpath for decant pond water. Based on Phase 1 results and review of the local stratigraphy (see Section 4), it is hypothesized that a shallow gravel unit may provide the primary flowpath for transport of water from the decant ponds. While this shallow unit is believed to transport water primarily laterally away from the decant pond, some slow transport of water downward to the deeper gravels (i.e., leakage through the silts/fine sands that underlie the shallow gravel) is believed to also occur. The slow rate of the vertical transport of water to the deeper gravels is believed to account for the lack of detection of the tracer in the deeper gravels during the relatively short duration of the tracer test.

3.1 GROUNDWATER SAMPLING AND EVALUATION

Groundwater quality data for wells around the decant ponds was collected three times in the past year (September 2000, March 2001, April 2001). Comparisons of groundwater quality data and minewater/decant pond water are provided in Tables 1 through 3 and Piper diagrams in Appendix E. As these comparisons indicate, groundwater in wells near the decant pond (IW-1, MW-95-8, MW-95-7) has an overall chemical signature that is very similar to the decant pond with the exception of the lack of copper in groundwater. This is true not only of major ion chemistry as shown by the Piper diagrams, but also nitrate + nitrite concentration (see Tables 1-3). Of particular note on the Piper diagrams is the seasonal shift in decant water composition and corresponding shifts in groundwater composition between September 2000 and March 2001.

REVIEW DRAFT

REVIEW DRAFT

TABLE 1 SEPTEMBER 2000 WATER QUALITY DATA

	Service Adit Minewater	IW-1	QW-1	MW-4	MW-2	MW-3	MW-1	MW95-8	MW95-5
Alkalinity.	76	72	24	16	156	124	108	170	170
Arsenic	0.003	<0.003	<0.003	<0.003	0.027	0.01	0.007	<0.003	0.003
Calcium	23	26	7	4.3	29	36	33	62	47
Carbonate	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0
Copper	0.056	0.002	0.001	<0.001	<0.001	0.001	<0.001	0.003	0.003
Iron	<0.020	<0.020	0.16	<0.020	<0.020	7.8	7.7	<0.020	30
Hardness	81	86	24	15	109	106	103	184	171
Bicarb.	76	72	24	16	156	124	108	170	170
Potassium	<2	<2.0	<2.0	<2.0	2.7	1.1	1.5	2.1	10
Magnesium	5.5	5	1.7	1	8.7	4	5.3	6.9	13
Manganese	0.024	0.081	0.018	0.04	0.25	0.66	0.57	<0.005	4.8
Sodium	NA	3	<2.0	<2.0	21	5.1	4.9	2.5	7.7
Ammonia	0.1	0.061	<0.050	0.13	0.43	0.39	0.11	0.051	0.21
Nitrate+nitrite	0.69	0.69	<0.050	0.11	0.094	<0.050	0.12	1.2	0.075
Lead	<0.003	<0.003	<0.003	<0.003	<0.003	<0.003	<0.003	<0.003	<0.003
pH	8.3	7.2	6.3	6.6	7.5	6.9	6.9	7.9	6.3
Antimony	0.009	<0.003	<0.003	<0.003	<0.003	<0.003	<0.003	<0.003	<0.003
Silica	NA	9.3	8.4	7.5	11	28	22	11	65
Sulfate	15	19	3.9	2.6	<2.0	<2.0	9.1	15	7.3
TDS	112	115	52	36	197	166	159	235	361
Zinc	<0.010	<0.010	0.32	0.023	<0.010	<0.010	<0.010	0.19	0.35

Note: All units (mg/L) except pH (std. units), alkalinity, carbonate, bicarbonate and hardness (all mg/L as CaCO₃), ammonia and nitrate + nitrite (mg/L as N). All metals are dissolved.

REVIEW DRAFT

TABLE 2. MARCH 2001 WATER QUALITY DATA (BEFORE TRACER TEST)

	MW 95-5	MW 95-4	MW 95-3	MW 97-12	MW-1	MW-2	MW-3	MW-4	MW 95-7	MW 95-8	IW-1	DECANT POND 1	DECANT POND 2
Alkalinity.	154.	22.	89.	61.	114.	21.	113.	168.	71.	81.	75.	74.	76.
Arsenic	0.008	0.004	<0.002	<0.002	0.006	<0.002	0.019	0.029	0.002	<0.002	<0.002	<0.002	<0.002
Calcium	46.	38.	130.	12.	35.	4.6	32.	30.	30.	41.	28.	31.	30.
Chloride-	3.9	4.2	2.7	3.0	2.1	<1.0	1.1	<1.0	<1.0	1.1	<1.0	<1.0	<1.0
Carbonate	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0	<1.0
Copper	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	0.049	0.043
Iron	36.	17.	0.22	0.20	6.9	<0.030	12.	<0.030	<0.030	<0.030	<0.030	<0.030	<0.030
Hardness	171.	139.	503.	54.	111.	16.	97.	112.	92.	125.	98.	116.	113.
Bicarb.	154.	22.	89.	61.	114.	21.	113.	168.	71.	81.	75.	74.	76.
Potassium	11.	24.	74.	5.3	1.1	<1.0	<1.0	2.4	<1.0	<1.0	<1.0	<1.0	<1.0
Magnesium	13.	11.	43.	5.6	5.7	1.1	3.9	9.0	4.1	5.6	6.7	9.5	9.1
Manganese	5.8	3.0	0.99	0.10	0.55	0.14	0.59	0.24	<0.005	<0.005	<0.005	0.040	0.037
Sodium	9.0	35.	53.	5.8	5.6	<1.0	5.4	23.	1.3	1.3	2.0	1.6	1.5
Ammonia	0.60	4.2	10.	<0.10	<0.10	0.20	0.30	0.42	<0.10	<0.10	<0.10	<0.10	<0.10
Nitrate+nitrite	0.19	0.31	<0.10	<0.10	0.32	0.13	<0.10	<0.10	1.1	1.6	0.79	1.6	1.4
Lead	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002	<0.002
pH	8.0	6.6	7.9	6.7	8.0	6.9	7.2	8.2	8.1	8.4	8.1	8.3	8.2
Antimony	<0.005	<0.005	<0.005	<0.005	<0.005	<0.005	<0.005	<0.005	<0.005	<0.005	<0.005	0.011	0.011
Silica	58.	43.	10.	54.	25.	10.	30.	13.	12.	9.8	9.8	8.1	8.3
Sulfate	9.2	292.	693.	14.	9.6	5.3	1.4	1.2	27.	41.	25.	39.	42.
TDS	284.	406.	1060.	150.	160.	31.	178.	197.	136.	164.	133.	171.	164.
Zinc	<0.010	<0.010	<0.010	<0.010	<0.010	0.013	<0.010	<0.010	<0.010	<0.010	<0.010	<0.010	<0.010

Note: All units (mg/L) except pH (std. units), alkalinity, carbonate, bicarbonate and hardness (all mg/L as CaCO₃), ammonia and nitrate + nitrite (mg/L as N). All metals are dissolved.

REVIEW DRAFT

TABLE 3. APRIL 2001 WATER QUALITY DATA (AFTER TRACER ADDITION PERIOD)

	DECANT POND-1	DECANT POND-2	MW-95-3	MW-95-7	IW-1
Alkalinity.	50.	76.	146.	220A	70.
Arsenic	<0.002	<0.002	0.002	<0.002	<0.002
Calcium	22.	30.	130.	32.	26.
Chloride-	96.	67.	2.6	1.3	<1.0
Carbonate	<1.0	<1.0	<1.0	<1.0	<1.0
Copper	0.040	0.036	<0.002	0.002	<0.002
Iron	<0.050	<0.050	<0.050	<0.050	<0.050
Hardness	82.	113.	502.	96.	90.
Bicarb.	50.	76.	146.	220A	70.
Potassium	1.6	1.9	79.	1.8	1.5
Magnesium	6.5	9.2	43.	4.2	6.2
Manganese	0.053	0.036	1.0	<0.005	<0.005
Sodium	51.	38.	56.	1.5	2.8
Ammonia	<0.10	<0.10	12.	0.12	<0.10
Nitrate+nitrite	0.80	1.5	<0.050	1.0	0.71
Lead	<0.002	<0.002	<0.002	<0.002	<0.002
pH	7.8	8.0	7.8	7.9	8.0
Antimony	0.007	0.010	<0.005	<0.005	<0.005
Silica	7.0	7.3	10.	11.	13.
Sulfate	33.	40.	730.	31.	24.
TDS	243.	246.	1051.	160.	135.
Zinc	<0.010	<0.010	<0.010	<0.010	<0.010

Note: All units (mg/L) except pH (std. units), alkalinity, carbonate, bicarbonate and hardness (all mg/L as CaCO₃), ammonia and nitrate + nitrite (mg/L as N). All metals are dissolved. Alkalinity and bicarbonate values for MW-95-7 are suspected to be in error (cation/anion balance is significantly off); 70 mg/L assumed for Piper diagram in Appendix E.

3.2 TRACER TEST

The addition of salt to the decant ponds caused the following changes in decant pond water chemistry:

- increased the specific conductivity (SC) of the decant ponds to approximately 400 to 700 micromhos/cm (2 to 3 times higher than ambient groundwater, i.e. IW-1, MW-95-7 and MW-95-8);
- increased the sodium concentration to 40 to 90 mg/L (20 to 50 times higher than ambient groundwater); and
- increased the chloride concentration to 60 to 160 mg/L (60 to 160 times higher than ambient groundwater).

The premise of the tracer test is that because SC, sodium, and chloride were elevated in the decant pond by the salt addition, the presence of increased concentrations of these parameters in ambient groundwater would demonstrate a connection (i.e., a groundwater flowpath) between the decant ponds and the monitoring wells. Because of residual process water within the tailings, the specific conductivity and sodium concentrations in water in well MW-95-3 (completed in the tailings west of the decant ponds) were similar to, or higher than decant pond water prior to, during, and after the tracer test. Therefore, the presence of increased SC and sodium in well MW-95-3 is unrelated to the addition of salt to decant pond water.

3.2.1 Chemical Effects of the Tracer on Groundwater

Results of the groundwater tracer test are shown in the trend plots (SC, sodium, chloride versus time) in Appendix F. Overall, there was no meaningful increase in concentrations of the tracer (salt) in any monitoring wells during the test. Therefore, the tracer test indicates poor hydrologic connection (low or slow rate of flow) between the decant ponds and the units in which the monitoring wells are completed. As noted above, groundwater quality data suggests that the groundwater in deeper gravel/sand units near the decant ponds (e.g., units monitored by wells IW-1, MW-95-7, and MW-95-8) is derived primarily from

REVIEW DRAFT

minewater disposed in the decant ponds. However, tracer test results suggests that these gravel units are not the primary groundwater flowpath for decant pond water.

3.2.2 Decant Pond Infiltration Rate

Water levels in the decant ponds were monitored over the course of the tracer addition period (24 days, from March 13 to April 6) while no water was discharged to the ponds. Based on the volume of water lost from the ponds over this period it is possible to calculate an infiltration rate for the ponds. Decant pond #1 and decant pond #2 have surface areas of approximately 3.2 and 1 acres, respectively and during the tracer addition period, water levels in the ponds dropped 10.7 and 8.8 feet, respectively. Total volume of water infiltrated through the ponds is therefore approximately 14 million gallons. Since the tracer addition period was 34,380 minutes, the average infiltration rate for the period was about 400 gpm or about 0.2 in/hr (0.4 ft/day).

The calculated infiltration rate is not particularly high and is in the range of what might be expected for a fine sand. However, this infiltration rate implies that the hydraulic conductivity of the groundwater aquifer that transports the minewater away from the decant pond is much larger, perhaps on the order of 500 ft/day or more, typical of coarse sands or gravel. The decant ponds have a large surface area for infiltration (about 180,000 ft²) and the hydraulic gradient in the pond is 1 (vertical gradient). In the groundwater system near the pond it is likely that the gradient is much less than 1 (i.e. flow is not strictly vertical but likely has a large horizontal component) and the cross-sectional area of the aquifer is likely a fraction of the surface area of the ponds. Given the likely much lower gradient and smaller area of the aquifer in the flowpath, the aquifer would have to have a much higher hydraulic conductivity to transmit the flow from the decant ponds.

4.0 PHASE 2 AND 3 WORKPLAN

4.1 CONCEPTUAL MODEL OF HYDROGEOLOGIC SYSTEM

Groundwater in the Lake Creek Valley occurs in glaciolacustrine deposits and alluvium (Levings et al, 1984). Within the glaciolacustrine deposits groundwater occurs primarily in the sand, gravel, and cobble zones. Silt and clay zones within the glaciolacustrine deposits are confining beds that restrict vertical ground-water movement. Although the general directions of groundwater movement in the area are likely towards the north (in the general direction of flow in Lake Creek) and towards the west (toward Lake Creek), local groundwater flow directions are likely controlled primarily by the location and orientation of permeable water bearing zones within the glaciolacustrine and alluvial deposits.

Previous monitoring well drilling (see Summit and MF&G, 1996) and geotechnical boring investigations (see Woodward-Clyde, 1977) in the area have identified the presence of sand and gravel zones within the glaciolacustrine deposits in the general vicinity of the decant ponds. As summarized in the Draft EIS (p. 103),

”The surficial geology of the pond [tailing impoundment] site reflects the complex geology of the area and varies from fine-grained clayey silts to coarse rock and gravel. The coarser deposits tend to be nearer the edge of the valley and along the lower reaches of Camp Creek. Test wells and test holes also reflect the complex geology of the area with substantial vertical variations in materials encountered.”

The most notable sand and gravel occurrences in wells and test holes are summarized in Table 4. Data from wells and test holes suggest permeable gravels are located at or near the surface near the decant ponds and extending to the east and south of the ponds. Gravels are also present at depth near the decant ponds and to the west of the impoundment near Lake Creek. Groundwater sampling and

REVIEW DRAFT

tracer testing in spring 2001 indicate that the groundwater in deeper gravel/sand units near the decant ponds (e.g., units monitored by wells IW-1, MW-95-7, and MW-95-8) is derived primarily from minewater disposed in the decant ponds. However, tracer test results suggests that these gravel units are not the primary groundwater flowpath for decant pond water. Based on Phase 1 results and review of the local stratigraphy (see Section 4), it is hypothesized that the shallow gravel unit may provide the primary flowpath for transport of water from the decant ponds. While this shallow unit is believed to transport water primarily laterally away from the decant pond, some slow transport of water downward to the deeper gravels (i.e., leakage through the silts/fine sands that underlie the shallow gravel) is believed to also occur. The slow rate of the vertical transport of water to the deeper gravels is believed to account for the lack of detection of the tracer in the deeper gravels during the relatively short tracer test.

REVIEW DRAFT

**TABLE 4. IDENTIFIED PERMEABLE ZONES IN THE VICINITY OF
THE DECANT PONDS**

Well/Boring	Location	Zone Description	Other
MW-95-7	Approx. 250 feet east southeast of decant pond #3	Gravel 0 to 18 ft bgs;	Well screened 79 to 89 ft bgs.
MW-95-8	Approx. 50 feet north of decant pond #1, north side of Section 2/3 divider dike.	Gravel, cobbles, sand 0 to 20 ft bgs;	Well screened 48 to 53 ft bgs.
IW-1	Approx. 20 feet north of decant pond #1.	No drilling logs can be found, Asarco internal memos refer to well as screened in “lower sand and gravel aquifer”.	Well screened 190 to 210 ft bgs.
MW-1	Approx. 4,000’ northwest of decant ponds between dike and Lake Creek	Clean sand and gravel 160 to 280 bgs	Well perforated 55’ to 100’ bgs in silt/clay and 160’ to 280’ t bgs
MW-3	Approx. 3,000’ west of decant ponds between dike and Lake Creek.	Sand and gravel 118 to 125’ bgs	Open-ended casing
TH-10	Section 2 approx. 1,000’ west of decant ponds	Very gravelly sand 8 to 40’ bgs	Field permeability 4,400 ft/yr
TH-11	In section 1 approx. 800’ south of decant ponds	Gravel 0 to 40’ bgs	Field permeability 8,100 ft/yr
TH-12	Southeast portion section 2; approx. 1200 feet southwest decant ponds	Gravelly sand to gravel 5 to 40’ bgs	Field permeability 1,100 ft/yr

*bgs = below ground surface.

4.2 PHASE 2 MONITORING WELL INSTALLATION AND TESTING

REVIEW DRAFT

The primary target for additional borings and well construction is shallow gravel and cobble units similar to the shallow materials encountered in well MW-95-8 (located 50 feet north of the decant pond), well MW-95-7 (located approximately 250 feet east of the decant pond), and exposed in the terraces immediately to the northeast of the decant pond. The quantity of water disposed in the decant pond (greater than 1,000 gpm or more at times) requires that the pond must be underlain by a relatively permeable unit in order to percolate and transmit the disposed water. The presence of approximately 20 feet of “gravel, cobbles and sand” material at MW-95-8 and MW-95-7 (Appendix II in MFG and Summit, 1996) suggests that surficial gravel materials are a likely conduit for mine water flows.

In summer 2001, excavations were made adjacent to the decant pond in an attempt to locate the surficial gravel materials and install monitoring wells as proposed in the first step in Phase II of the Fate and Transport Study (Hydrometrics, 2000). Although the surface of the tailings near the ponds is stable and workable by heavy equipment, it was not possible to maintain an open excavation sufficient to reach the base of the tailings. Therefore, future work will require drill rigs.

The proposed drilling program to install monitoring wells is anticipated to require one or two phases, Phases 2a and 2b. Phase 2a will include initial well installation to find and identify the groundwater flow path(s) from the decant ponds. Phase 2a will be comprised of wells relatively close to the ponds. Proposed Phase 2a monitoring well locations are shown on Figure 1. Wells are proposed to the west of decant pond #1, south of decant pond #1, south of decant pond #2, north of decant ponds #1 and #2 near the pond divider dike, and east of decant pond #3. Anticipated depth of completion of the wells is 20 to 40 feet, within the shallow gravel zone. Typical monitoring well construction details are shown on Figure 2. Actual well locations and construction details will be selected in the field and may be modified based on field conditions. Wells will be

REVIEW DRAFT

developed after well construction by pumping with a submersible pump, purging with rig air, bailing, or surging with a block.

Phase 2b of well installation will include wells completed further downgradient in the groundwater flowpath(s) identified in Phase 2a. The exact number and locations of Phase 2b wells will be determined based on Phase 2a results.

Monitoring wells completed in Phase 2 will be sampled after well development for the water quality parameters listed in Table 2.

4.3 PHASE 3 GROUNDWATER TRACER TEST

The need for another groundwater tracer test will be determined based on the results of monitoring well installation and sampling. If a chemical tracer from minewater (i.e., copper) can be found in the monitoring wells then tracer testing may not be needed. If it is believed that the minewater flowpath has been found but minewater can not be definitively identified in groundwater based on water quality analyses, then a tracer test will be conducted. Protocol for the tracer test would be developed at that time, similar to the protocol used in spring 2001 (e.g. see Appendix D).

REVIEW DRAFT

**TABLE 2 ANALYTICAL PARAMETERS FOR GROUNDWATER
MONITORING AT THE TROY MINE**

Analytical Parameter	Required Reporting Value * (mg/L)	Analytical Method **
<i>Common Ions and Physical Parameters</i>		
pH	0.1 S.U.	EPA 150.1
TDS	(10)	EPA 160.1
Calcium	(1)	EPA 6010
Magnesium	(1)	EPA 6010
Hardness	(1)	ASTM 2340B
Potassium	(2)	EPA 6010
Sodium	(1)	EPA 6010
Total Alkalinity as CaCO ₃	(1)	EPA 310.1
Bicarbonate	(1)	EPA 310.1
Carbonate	(1)	EPA 310.1
Chloride	(1)	EPA 300.0
Sulfate	(1)	ASTM 9036
Silica	(0.1)	EPA 6010
Temperature (field)	0.1 °C	(field)
Specific Conductivity (field)		(field)
Static Water Level (field)		(field)
<i>Nutrients</i>		
Ammonia as N	0.05	EPA 350.1
Nitrate + Nitrite as N	0.01	EPA 353.2
<i>Dissolved Metals (filtered sample)</i>		
Antimony	0.003	EPA 6020M
Arsenic	0.003	EPA 6020M
Copper	0.001	EPA 6020M
Iron	0.01	EPA 6010/6020
Lead	0.003	EPA 6020M
Manganese	0.005	EPA 6020M
Zinc	0.01	EPA 6020M

* Values in parentheses are detection limit goals, no RRVs exist for these parameters.

** Dissolved metals will be analyzed by ICP-MS (methods 6010, 6020, or 6020M).

REVIEW DRAFT

**FIGURE 2 TYPICAL SHALLOW MONITORING WELL
CONSTRUCTION**

REVIEW DRAFT

5.0 REFERENCES

- Hydrometrics, 2000. Minewater Fate and Transport Study Plan – Draft.
September, 2000
- Levings, G. W., R.F. Ferreira and J.H. Lambing, 1984. Water Resources of Lake
Creek Valley, Northwestern Montana
- Sterling Mining Co., 2000. Genesis Inc. Troy Mine Revised Reclamation Plan.
- Summit Envirosolutions and McCulley Frick & Gilman, Inc., 1996. Interim
Report of Findings, Hydrologic and Hydrogeologic Assessment, Asarco
Troy Mine, Troy Montana.
- Summit Envirosolutions, 1999 Site Instrumentation Status Report Troy Tailings
Impoundment, Lincoln County, Montana.
- Woodward-Clyde Consultants, 1977. Phase II Geotechnical Services, Troy
Pojectnear Troy Montana. Prepared for ASARCO, Inc. January 25, 1977.

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APPENDIX A

GROUNDWATER MONITORING PROPOSAL
(August 16, 2000 letter from Genesis Inc. to MDEQ)

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APPENDIX B

TRACER TEST PROPOSAL, REQUEST AND APPROVAL

(December 4, 2001 letter from Genesis Inc. to MDEQ)

(January 22, 2001 letter from MDEQ to Genesis Inc.)

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APPENDIX C

TRACER TEST PROTOCOL REVISION 1
(February 2, 2001 letter from Hydrometrics to MDEQ)

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February 2, 2001

Mr. Joe Gurrieri
MT Dept. of Environmental Quality
PO Box 200901
Helena, MT 59620-0901

RE: Groundwater Tracer Test at the Troy Mine

Dear Joe:

Sterling Mining Company received permission from MDEQ (Abe Horpestad) to proceed with the tracer test at the Genesis Inc. Troy Mine and was requested by Abe to inform you of our progress. Because minewater flows and conditions at the decant ponds have changed somewhat since the tracer test proposal was submitted in early December, we are planning to modify the test procedure. This letter describes the modifications to the tracer test and our final proposed monitoring schedule.

Schedule

We are tentatively planning to start the test the week of February 12, 2001 and would like you to observe the test, if possible. This schedule is contingent on our testing and identifying an appropriate method to add the salt to the ponds. As we have discussed, it will be important to make sure that all of the salt is dissolved quickly in the ponds. We will do some testing of salt additions prior to the start of the actual tracer test. We will inform you of any changes to this schedule.

Test Modification

Based on current conditions in the ponds and mine water flow rates, we now plan to add all of the salt to the ponds (approximately 10 tons) and cease all minewater flows to the ponds for the duration of the salt addition period (approximately 1 to 2 weeks).

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As described in Sterling's tracer test proposal, in December the decant ponds contained approximately 12 million gallons of water and mine water inflows to the ponds were approximately 470 gpm. Based on these water volumes, we intended to add approximately 5 tons of salt to the ponds to achieve the desired initial TDS concentration in the ponds. Then approximately 500 pounds of salt would be added daily to maintain this TDS.

Currently, the volume of water in the ponds is much higher (21 million gallons) and minewater flows are much lower (around 200 gpm). At these flow rates, Sterling can divert all minewater flows to the thickener and mill water pond for storage and thereby eliminate minewater flows to the decant ponds for a period of approximately 1 to 2 weeks (actual duration of minewater storage will depend on precipitation runoff to mill water pond during this period). Given these conditions, we believe that it will be easier to simply add the salt to the pond as a "slug" and eliminate the need for daily salt additions. Moreover, this will result in more constant TDS concentrations in the pond. Therefore, we intend to add all of the salt to the pond over an approximately 8 hour period. We plan to add the salt into the tailings reclaim water line (the buried 10-inch line that goes from the mill water pond to the decant pond) at the mill water pond. At the end of the salt addition the reclaim water line will be turned off and minewater flow will be diverted to the thickener and/or mill water pond

Proposed Monitoring Schedule

The overall schedule for monitoring is described in attached Table 1. In this table, "tracer addition period" means the period after the salt has been added to the pond and during the period that mine water is not discharged to the pond.

Prior to tracer addition, the decant pond and adjacent monitoring wells (Irrigation Well, MW-95-3, MW-95-7, and MW-95-8) will be sampled and analyzed for major ion, nutrient, metals and physical parameters according to the schedule in the attached Table 2. Wells around the perimeter of the impoundment will also be sampled for the complete analytical schedule.

During tracer addition, the specific conductivity (SC) of the various waters will be measured daily to track TDS concentrations and waters will be sampled weekly for sodium and chloride. Waters to be monitored include the decant pond and adjacent monitoring wells and wells that can be remotely monitored. Specific conductivity will be measured at four points (1 point along each side) in each pond. To document mixing in the ponds, specific conductivity in the ponds will be measured three times per day during the first 3 days of the test.

If the tracer is detected at adjacent monitoring wells during the tracer addition period, then the decant pond and adjacent monitoring wells will be sampled and analyzed for major ion, nutrient, and metal parameters according to the schedule in Table 2.

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At the end of the tracer addition period, the decant pond and adjacent monitoring wells will be sampled and analyzed for major ion, nutrient, and metal parameters (Table 2). Perimeter monitoring wells will be sampled and analyzed for specific conductivity, sodium, and chloride. If the specific conductivity of any perimeter well is more than 10 percent higher than pre-test baseline conductivity, then the well will be sampled for a full suite of analytical parameters (Table 2).

If the tracer has not been detected at adjacent monitoring wells by the end of the tracer addition period, monitoring of SC will be continued on a daily basis for an additional two weeks, then on a weekly basis thereafter. Sodium and chloride will also be monitored weekly in the adjacent wells until the test is terminated.

At the end of the tracer test, the decant pond and adjacent monitoring wells will be sampled and analyzed for major ion, nutrient, and metal parameters (Table 2). Perimeter monitoring wells will be sampled and analyzed for specific conductivity, sodium, and chloride. If the specific conductivity of any perimeter well is more than 10 percent higher than pre-test baseline conditions, then the well will be sampled for a full suite of analytical parameters (Table 2).

Water levels in the ponds will be monitored with staff gages throughout the test. Static water levels in the wells will be monitored at the time of sampling, or in instrumented wells, continuously throughout the test.

The decision of when to terminate the test will be based on monitoring results and will be coordinated with DEQ.

Please let me know if you would like any modifications to this test. I look forward to seeing you on Site for this tracer test.

Sincerely,

Scott Mason
Geochemist

c: Bruce Clark
 Frank Duval

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Table 1. Monitoring Schedule for Troy Tracer Test

	Pre-Test Baseline	Tracer Addition Period			End of Tracer Addition Period		Post Tracer Addition Period			End of Test	
Monitoring Location	Major Ions, Nutrients, & Metals (see Table 2)	Continuous Specific Conductivity	Daily Specific Conductivity & Water Level	Weekly Sodium and Chloride	Major Ions, Nutrients, & Metals (see Table 2)	Sodium, Chloride, & Specific Conductivity Only	Continuous Specific Conductivity	Daily Specific Conductivity & Water Level	Weekly Sodium and Chloride	Major Ions, Nutrients, & Metals (see Table 2)	Sodium, Chloride, & Specific Conductivity Only
Decant Pond 1 *	Yes		Yes **	Yes	Yes			Yes	Yes	Yes	
Decant Pond 2 *	Yes		Yes **	Yes	Yes			Yes	Yes	Yes	
Decant Pond 3 *	Yes		Yes **	Yes	Yes			Yes	Yes	Yes	
Irrigation Well	Yes		Yes	Yes	Yes			Yes	Yes	Yes	
MW-95-8	Yes	Yes	Yes	Yes	Yes		Yes	Yes	Yes	Yes	
MW-95-7	Yes	Yes	Yes	Yes	Yes		Yes	Yes	Yes	Yes	
MW-95-3	Yes	Yes	Yes	Yes	Yes		Yes	Yes	Yes	Yes	
MW-95-4	Yes	Yes	Yes		?	Yes	Yes	Yes		?	Yes
MW-97-12	Yes	Yes	Yes		?	Yes	Yes	Yes		?	Yes
MW-97-14	Yes	Yes	Yes		?	Yes	Yes	Yes		?	Yes
MW-95-5	Yes	Yes	Yes		?	Yes	Yes	Yes		?	Yes
MW-1	Yes				?	Yes				?	Yes
MW-2	Yes				?	Yes				?	Yes
MW-3	Yes				?	Yes				?	Yes
MW-4	Yes				?	Yes				?	Yes

Notes:

? Well will be sampled for major ions, nutrients, and metals per Table 2 only if specific conductivity is 10% or more greater than pre-test result.

* Specific conductivity will be measured at four points (1 point along each side) in each pond.

** Specific conductivity will be measured three times per day during the first 3 days of the test to document mixing in the ponds.

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**TABLE 2 ANALYTICAL PARAMETERS FOR GROUNDWATER
MONITORING AT THE TROY MINE**

Analytical Parameter	Required Reporting Value * (mg/L)	Analytical Method **
<i>Common Ions and Physical Parameters</i>		
pH	0.1 S.U.	EPA 150.1
TDS	(10)	EPA 160.1
Calcium	(1)	EPA 6010
Magnesium	(1)	EPA 6010
Hardness	(1)	ASTM 2340B
Potassium	(2)	EPA 6010
Sodium	(1)	EPA 6010
Total Alkalinity as CaCO ₃	(1)	EPA 310.1
Bicarbonate	(1)	EPA 310.1
Carbonate	(1)	EPA 310.1
Chloride	(1)	EPA 300.0
Sulfate	(1)	ASTM 9036
Silica	(0.1)	EPA 6010
Temperature (field)	0.1 °C	(field)
Specific Conductivity (field)		(field)
Static Water Level (field)		(field)
<i>Nutrients</i>		
Ammonia as N	0.05	EPA 350.1
Nitrate + Nitrite as N	0.01	EPA 353.2
<i>Dissolved Metals (filtered sample)</i>		
Antimony	0.003	EPA 6020M
Arsenic	0.003	EPA 6020M
Copper	0.001	EPA 6020M
Iron	0.01	EPA 6010/6020
Lead	0.003	EPA 6020M
Manganese	0.005	EPA 6020M
Zinc	0.01	EPA 6020M

* Values in parentheses are detection limit goals, no RRVs exist for these parameters.

** Dissolved metals will be analyzed by ICP-MS (methods 6010, 6020, or 6020M).

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APPENDIX D

TRACER TEST PROTOCOL REVISION 2
(March 6, 2001 letter from Hydrometrics to MDEQ)

REVIEW DRAFT



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March 6, 2001

Mr. Joe Gurrieri
MT Dept. of Environmental Quality
PO Box 200901
Helena, MT 59620-0901

RE: Groundwater Tracer Test at the Troy Mine

Dear Joe:

Sterling Mining Company received permission from MDEQ (Abe Horpestad) to proceed with the tracer test at the Genesis Inc. Troy Mine and was requested by Abe to inform you of our progress. Because the remote groundwater monitoring system was damaged and is no longer functioning, the planned monitoring schedule has changed somewhat since the tracer test proposal revision in early February. This letter describes the modifications to the proposed monitoring schedule. We plan to conduct pre-test sampling of the wells on March 8 and 9 and to start the tracer (salt) addition on March 12, 2001.

The revised schedule for monitoring is described in attached Table 1. In this table, "tracer addition period" means the period after the salt has been added to the pond and during the period that mine water is not discharged to the pond. Revisions to the Table have been done in "redline" mode (additions are italicized, deletions are strikeout) to clarify the changes that have been made. The primary change is that continuous monitoring of wells is not possible and will not be done. To offset the loss of continuous specific conductivity data, the frequency of sampling of wells near the decant pond for SC, sodium, and chloride has been increased from weekly to daily.

Please let me know if you would like any modifications to the planned monitoring schedule.

Sincerely,

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Scott Mason
Geochemist

c: Abe Horpestad
Bruce Clark
Frank Duval

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Table 1. Revised Monitoring Schedule for Troy Tracer Test

	Pre-Test Baseline	Tracer Addition Period			End of Tracer Addition Period		Post Tracer Addition Period			End of Test	
Monitoring Location	Major Ions, Nutrients, & Metals (see Table 2)		Daily Specific Conductivity, Sodium, Chloride & Water Level		Major Ions, Nutrients, & Metals (see Table 2)	Sodium, Chloride, Water Level & Specific Conductivity Only		Daily Specific Conductivity, Sodium, Chloride & Water Level		Major Ions, Nutrients, & Metals (see Table 2)	Sodium, Chloride, Water Level & Specific Conductivity Only
Decant Pond 1 *	Yes		Yes **		Yes			Yes		Yes	
Decant Pond 2 *	Yes		Yes **		Yes			Yes		Yes	
Decant Pond 3 *	Yes		Yes **		Yes			Yes		Yes	
Irrigation Well	Yes		Yes		Yes			Yes		Yes	
MW-95-8	Yes		Yes		Yes			Yes		Yes	
MW-95-7	Yes		Yes		Yes			Yes		Yes	
MW-95-3	Yes		Yes		Yes			Yes		Yes	
MW-95-4	Yes				?	Yes				?	Yes
MW-97-12	Yes				?	Yes				?	Yes
MW-97-14	Yes				?	Yes				?	Yes
MW-95-5	Yes				?	Yes				?	Yes
MW-1	Yes				?	Yes				?	Yes
MW-2	Yes				?	Yes				?	Yes
MW-3	Yes				?	Yes				?	Yes
MW-4	Yes				?	Yes				?	Yes

Notes:

? Well will be sampled for major ions, nutrients, and metals per Table 2 only if specific conductivity is 10% or more greater than pre-test result.

* Specific conductivity will be measured at four points (1 point along each side) in each pond.

** Specific conductivity will be measured three times per day during the first 3 days of the test to document mixing in the ponds.

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**TABLE 2 ANALYTICAL PARAMETERS FOR GROUNDWATER
MONITORING AT THE TROY MINE**

Analytical Parameter	Required Reporting Value * (mg/L)	Analytical Method **
<i>Common Ions and Physical Parameters</i>		
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TDS	(10)	EPA 160.1
Calcium	(1)	EPA 6010
Magnesium	(1)	EPA 6010
Hardness	(1)	ASTM 2340B
Potassium	(2)	EPA 6010
Sodium	(1)	EPA 6010
Total Alkalinity as CaCO ₃	(1)	EPA 310.1
Bicarbonate	(1)	EPA 310.1
Carbonate	(1)	EPA 310.1
Chloride	(1)	EPA 300.0
Sulfate	(1)	ASTM 9036
Silica	(0.1)	EPA 6010
Temperature (field)	0.1 °C	(field)
Specific Conductivity (field)		(field)
Static Water Level (field)		(field)
<i>Nutrients</i>		
Ammonia as N	0.05	EPA 350.1
Nitrate + Nitrite as N	0.01	EPA 353.2
<i>Dissolved Metals (filtered sample)</i>		
Antimony	0.003	EPA 6020M
Arsenic	0.003	EPA 6020M
Copper	0.001	EPA 6020M
Iron	0.01	EPA 6010/6020
Lead	0.003	EPA 6020M
Manganese	0.005	EPA 6020M
Zinc	0.01	EPA 6020M

* Values in parentheses are detection limit goals, no RRVs exist for these parameters.

** Dissolved metals will be analyzed by ICP-MS (methods 6010, 6020, or 6020M).

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APPENDIX E

PIPER DIAGRAMS

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APPENDIX F

TRACER TEST RESULTS – SC, Na, CL Trendplots

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Appendix I

**Assessment of Natural Attenuation of
Metals in a Decant Pond Disposal System**

**ASSESSMENT OF NATURAL ATTENUATION OF METALS IN A
DECANT POND DISPOSAL SYSTEM**

TROY MINE

Prepared for:

Genesis Inc. Troy Mine
P.O. Box 1660
Troy, MT 59935

Prepared by:

Hydrometrics, Inc.
3020 Bozeman Avenue
Helena, MT 59601

May 2010

TABLE OF CONTENTS

LIST OF TABLES	iv
LIST OF FIGURES.....	v
LIST OF APPENDICES	vii
EXECUTIVE SUMMARY	viii
1.0 INTRODUCTION.....	1-1
1.1 BACKGROUND.....	1-3
1.2 PURPOSE AND SCOPE OF EVALUATION	1-6
1.3 LIMITATIONS AND UNCERTAINTIES	1-7
2.0 OCCURRENCE OF METALS IN TROY MINE WATER	2-1
2.1 SOURCE OF METALS	2-1
2.2 CONCENTRATIONS OF METALS IN WATER	2-2
2.3 CONCENTRATION OF COMMON ION PARAMETERS IN WATER....	2-5
3.0 REVIEW OF METAL ATTENUATION LITERATURE	3-1
3.1 NATURAL ATTENUATION MECHANISMS.....	3-1
3.2 ANTIMONY	3-4
3.3 ARSENIC.....	3-4
3.4 CADMIUM.....	3-5
3.5 COPPER.....	3-6
3.6 LEAD	3-7
3.7 MERCURY	3-9
3.8 SILVER.....	3-10
3.9 URANIUM.....	3-10
4.0 FIELD EVIDENCE FOR NATURAL ATTENUATION	4-1
4.1 FIELD EVIDENCE OF ACTIVE METALS REMOVAL	4-1
4.1.1 Changes in Metal Concentrations in Groundwater	4-1
4.1.2 Changes in Metal Concentrations in Mixed Water Samples	4-3
4.1.3 Metals Enrichment of Decant Pond Sediment	4-5
4.2 FIELD EVIDENCE OF ATTENUATION MECHANISMS	4-6
4.2.1 Identification of Metal Phases in Sediment	4-7

4.2.2 Decant Pond/Mine Water and Groundwater Mixing Tests.....	4-8
5.0 GEOCHEMICAL EQUILIBRIUM MODELING	5-1
5.1 PROCESS MODEL FOR GEOCHEMICAL EQUILIBRIUM MODELING	5-2
5.2 GEOCHEMICAL MODELING SOFTWARE AND THERMODYNAMIC DATABASE	5-4
5.3 REPRESENTATIVE DISPOSAL WATER (MINE WATER AND DECANT POND WATER)	5-5
5.4 REPRESENTATIVE REDUCTION-OXIDATION CONDITIONS	5-6
5.5 AQUEOUS SPECIATION AND MINERAL SATURATION INDICES ...	5-11
5.6 STABILITY AND SOLUBILITY DIAGRAMS.....	5-16
5.6.1 Copper Oxides and Copper Iron Oxides	5-20
5.6.2 Copper Carbonates	5-22
5.6.3 Copper Hydroxysulfate	5-24
5.6.4 Copper Silicates	5-26
5.6.5 Iron	5-28
5.6.6 Lead.....	5-32
5.6.7 Manganese.....	5-34
6.0 SUMMARY SITE MODEL FOR NATURAL ATTENUATION	6-1
7.0 EVALUATION OF LONGEVITY OF NATURAL ATTENUATION MECHANISMS.....	7-1
7.1 ATTENUATION BY MINERAL PRECIPITATION.....	7-1
7.2 NEUTRAL TO ALKALINE PH.....	7-2
7.3 OXIDIZING REDOX CONDITIONS.....	7-2
7.3.1 Dissolved Silica.....	7-3
7.3.2 Dissolved Bicarbonate	7-4
7.4 ATTENUATION BY ADSORPTION TO MINERALS AND SEDIMENTS	7-5
7.4.1 Water pH	7-6
7.4.2 Quantity of Sorbent Material	7-7
7.4.3 Concentration of Competing Ions	7-7

8.0 MANAGEMENT PRACTICES AND MONITORING TO VERIFY	
CONTINUED EFFECTIVENESS OF METAL ATTENUATION	8-1
8.1 MONITORING	8-1
8.2 MANAGEMENT PRACTICES	8-1
8.3 POTENTIAL CORRECTIVE ACTIONS.....	8-2
9.0 SUMMARY AND CONCLUSIONS.....	9-1
9.1 EFFECTIVENESS AND LONGEVITY OF MONITORED NATURAL	
ATTENUATION	9-1
9.2 ANTIMONY	9-2
9.3 ARSENIC.....	9-3
9.4 CADMIUM.....	9-3
9.5 COPPER.....	9-4
9.6 LEAD	9-4
9.7 MERCURY	9-5
9.8 SILVER.....	9-5
9.9 URANIUM.....	9-5
10.0 REFERENCES.....	10-1

LIST OF TABLES

TABLE 1. METALS AND BICARBONATE CONCENTRATIONS	
IN TROY MINE WATER.....	2-3
TABLE 2. SUMMARY OF ATTENUATION PROCESSES	
FOR ANTIMONY, MERCURY AND SILVER	3-3
TABLE 3. SUMMARY OF ATTENUATION PROCESSES FOR ARSENIC,	
CADMIUM, COPPER, LEAD, URANIUM, AND ZINC.....	3-3
TABLE 4. RESULTS OF LABORATORY TESTING OF METAL	
ATTENUATION BY MIXING WITH AMBIENT	
GROUNDWATER.....	4-4

TABLE 5.	TOTAL COPPER AND IRON CONTENT OF DECANT POND SOIL SAMPLES	4-5
TABLE 6.	SUMMARY OF FIELD EVIDENCE OF ATTENUATION MECHANISMS	4-7
TABLE 7.	PRINCIPLE COMPONENTS OF THE PROCESS MODEL FOR GEOCHEMICAL EQUILIBRIUM MODELING.....	5-3
TABLE 8.	COMPARISON OF FIELD MEASURED EH AND EH CALCULATED FROM REDOX PAIRS	5-8
TABLE 9.	PRIMARY AQUEOUS SPECIES OF METALS	5-13
TABLE 10.	MINERALS SATURATED IN MINE WATER AND GROUNDWATER.....	5-14
TABLE 11.	SUMMARY OF FACTORS AFFECTING METAL ATTENUATION BY MINERAL PRECIPITATION	6-2
TABLE 12.	SUMMARY OF FACTORS AFFECTING METAL ATTENUATION BY CO-PRECIPITATION OR ADSORPTION WITH IRON AND MANGANESE MINERALS.....	6-3

LIST OF FIGURES

FIGURE 1.	LOCATION OF TROY MINE.....	1-2
FIGURE 2.	MAP OF TAILING IMPOUNDMENT SHOWING LOCATION OF DECANT POND AND MONITORING WELLS	1-4
FIGURE 3.	COMPARISON OF FIELD MEASURED EH AND EH CALCULATED FROM REDOX PAIRS	5-9
FIGURE 4.	EH-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6})	5-19
FIGURE 5.	EH-PH DIAGRAM FOR COPPER ACTIVITY = $10^{-7.5}$	5-20
FIGURE 6.	EH-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6}) WITH IRON	5-21
FIGURE 7.	EH-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$) WITH IRON	5-21

FIGURE 8.	BICARBONATE-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6})	5-23
FIGURE 9.	BICARBONATE-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$)	5-24
FIGURE 10.	EH-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6}) WITH SULFATE (ACTIVITY = $10^{-3.5}$)	5-25
FIGURE 11.	EH-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6}) WITH SULFATE (ACTIVITY = $10^{-4.25}$)	5-25
FIGURE 12.	EH-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$) WITH SULFATE (ACTIVITY = $10^{-3.5}$)	5-26
FIGURE 13.	EH-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$) WITH SULFATE (ACTIVITY = $10^{-4.25}$)	5-26
FIGURE 14.	SILICA-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6})	5-28
FIGURE 15.	SILICA-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$)	5-28
FIGURE 16.	EH-PH DIAGRAM FOR IRON (ACTIVITY = 10^{-6}) WITH SULFATE (ACTIVITY = $10^{-3.5}$) AND POTASSIUM (ACTIVITY = $10^{-3.2}$)	5-29
FIGURE 17.	EH-PH DIAGRAM FOR IRON (ACTIVITY = 10^{-6}) WITH SULFATE (ACTIVITY = $10^{-3.5}$) AND POTASSIUM (ACTIVITY = $10^{-3.2}$), GOETHITE SUPPRESSED	5-30
FIGURE 18.	EH-PH DIAGRAM FOR IRON (ACTIVITY = $10^{-6.2}$ WITH SULFATE (ACTIVITY = $10^{-4.25}$) AND POTASSIUM (ACTIVITY = $10^{-4.3}$)	5-30
FIGURE 19.	EH-PH DIAGRAM FOR IRON (ACTIVITY = $10^{-6.2}$ WITH SULFATE (ACTIVITY = $10^{-4.25}$) AND POTASSIUM (ACTIVITY = $10^{-4.3}$), GOETHITE SUPPRESSED	5-31
FIGURE 20.	LEAD ACTIVITY-PH DIAGRAM WITH BICARBONATE ACTIVITY = $10^{-2.6}$	5-33
FIGURE 21.	LEAD ACTIVITY-PH DIAGRAM WITH BICARBONATE ACTIVITY = $10^{-3.2}$	5-33

FIGURE 22.	MANGANESE-pH DIAGRAM WITH BICARBONATE (ACTIVITY = $10^{-2.6}$)	5-34
FIGURE 23.	MANGANESE-pH DIAGRAM WITH BICARBONATE (ACTIVITY = $10^{-3.2}$)	5-35

LIST OF APPENDICES

APPENDIX A	WATER QUALITY SAMPLE SUMMARIES AND LAB REPORTS
APPENDIX B	CONCENTRATION TIME SERIES GRAPHS
APPENDIX C	GEOCHEMICAL MODEL RESULTS

EXECUTIVE SUMMARY

Mine and tailings water have been directed to the Troy Mine tailings impoundment for over 30 years without appreciable transport of metals in groundwater away from the infiltration pond (decant pond). This is due in part to natural attenuation of metals in the sediments beneath the infiltration pond and in the groundwater aquifer. Field evidence for natural attenuation of metals is provided by the sampling and analysis of water and soil samples in the vicinity of the tailings impoundment infiltration pond and from laboratory testing.

Water quality data shows that the concentrations of parameters that are not attenuated (e.g. chemically conservative parameters such as nitrate) are similar between the mine water/decant pond water and the underlying groundwater. This similarity in concentrations indicates that mine water is not appreciably diluted in the groundwater system in close proximity to the infiltration pond. In contrast, the concentrations of metal parameters are much lower in the groundwater than in the mine water/decant pond, indicating that metal parameters are attenuated (removed) in the groundwater system.

Analyses of sediments in the subsurface of the decant pond reveal that metals are enriched in the sediments, indicating that the metals attenuated/removed from groundwater are retained on the sediments. Microscopic and electron microprobe examination of sediments document the occurrence of secondary copper minerals (i.e., non-ore forming minerals) in the sediments. These secondary minerals form by precipitation of copper from mine water. Geochemical equilibrium modeling confirms that copper, iron, lead, and manganese minerals are oversaturated in mine water and are favored to form from mine water.

The longevity of the natural attenuation mechanisms is dependent on the duration of the geochemical conditions conducive to attenuation and to the capacity of the tailing impoundment/aquifer sediments to continue to uptake metals. **The mineral precipitation and co-precipitation mechanisms are expected to last indefinitely or in perpetuity as long as geochemical conditions remain similar to current conditions.** The adsorption mechanisms are conservatively estimated to last a minimum of 600 years.

Potential monitoring and management practices that would help ensure the continued effectiveness of metal attenuation by maintaining the geochemical conditions that are conducive to attenuation are described. The goal of monitoring and management practices would be to maintain the following current conditions at the tailing impoundment:

1. Neutral or alkaline pH;
2. Oxidizing redox conditions (moderate to high Eh);
3. Presence of moderate amounts of dissolved silica;
4. Presence of moderate amounts of bicarbonate; and
5. Presence of low to moderate amounts of organic material.

Management practices should be aimed at maintaining the existing geochemical conditions by:

1. Maintaining exposure of disposal water to atmospheric oxygen and carbon dioxide;
2. Avoiding excessive use of phosphate fertilizer for reclamation (high levels of dissolved phosphate may suppress arsenic adsorption and limit arsenic attenuation);
3. Maintaining pH by monitoring or mitigation; and
4. Maintaining the decant pond/infiltration site as an aerobic environment by monitoring and if necessary removing excessive organic material that may build up.

ASSESSMENT OF NATURAL ATTENUATION OF METALS IN DECANT POND DISPOSAL SYSTEM

TROY MINE

1.0 INTRODUCTION

This report summarizes an assessment of the natural attenuation of selected metals (antimony, cadmium, copper, lead, mercury, silver, and uranium) in post-closure mine water that is proposed to be disposed by infiltration into the reclaimed tailing impoundment at the Troy Mine in Lincoln County, Montana (Figure 1). Although there are several different sources or types of water associated with the Troy Mine, “mine water” for purposes of this evaluation, refers to water that is pumped or that flows from the underground mine workings via the mine adits. After mine closure, mine water would be piped to the tailings pond area for disposal (Genesis Inc., 2006).

Water that is used in the milling circuit and/or to transport tailings to the tailings impoundment is termed “tailings water.” During operational periods, tailings water flows over the deposited tailings to the tailings decant pond where a portion of the water infiltrates to the subsurface and the remainder is pumped back to the mine for use in the mill. During times of excess mine water (especially during the mine dewatering phase of the mine restart in 2005), mine water may also flow to the decant pond directly via the tailings lines. Thus during operational periods, water in the decant pond (“decant pond water”) is composed of a varying mix of tailings water, mine water, and natural rainfall and snowmelt. Since decant pond water is partially composed of mine water and is similar in metal content and overall chemistry to mine water, the fate of metals in decant pond water that infiltrates from the decant pond during operations is also considered in this report to provide additional evaluation of attenuation mechanisms.

FIGURE 1. LOCATION OF TROY MINE



LOCATION MAP

LOCATION MAP
TROY MINE
SANDERS COUNTY, MT

FIGURE

1

1.1 BACKGROUND

The Troy Mine was developed by Asarco in 1979 and went into production in 1981. Asarco operated the mine until 1993 when the mine was placed on temporary shutdown. In 1999, the mine was purchased Sterling Mining Company (now Revett Minerals, Inc) and has been operated by Genesis Inc., a subsidiary of Revett. Mining restarted in 2005 with milling and tailings production beginning in November 2005 and continuing to this day.

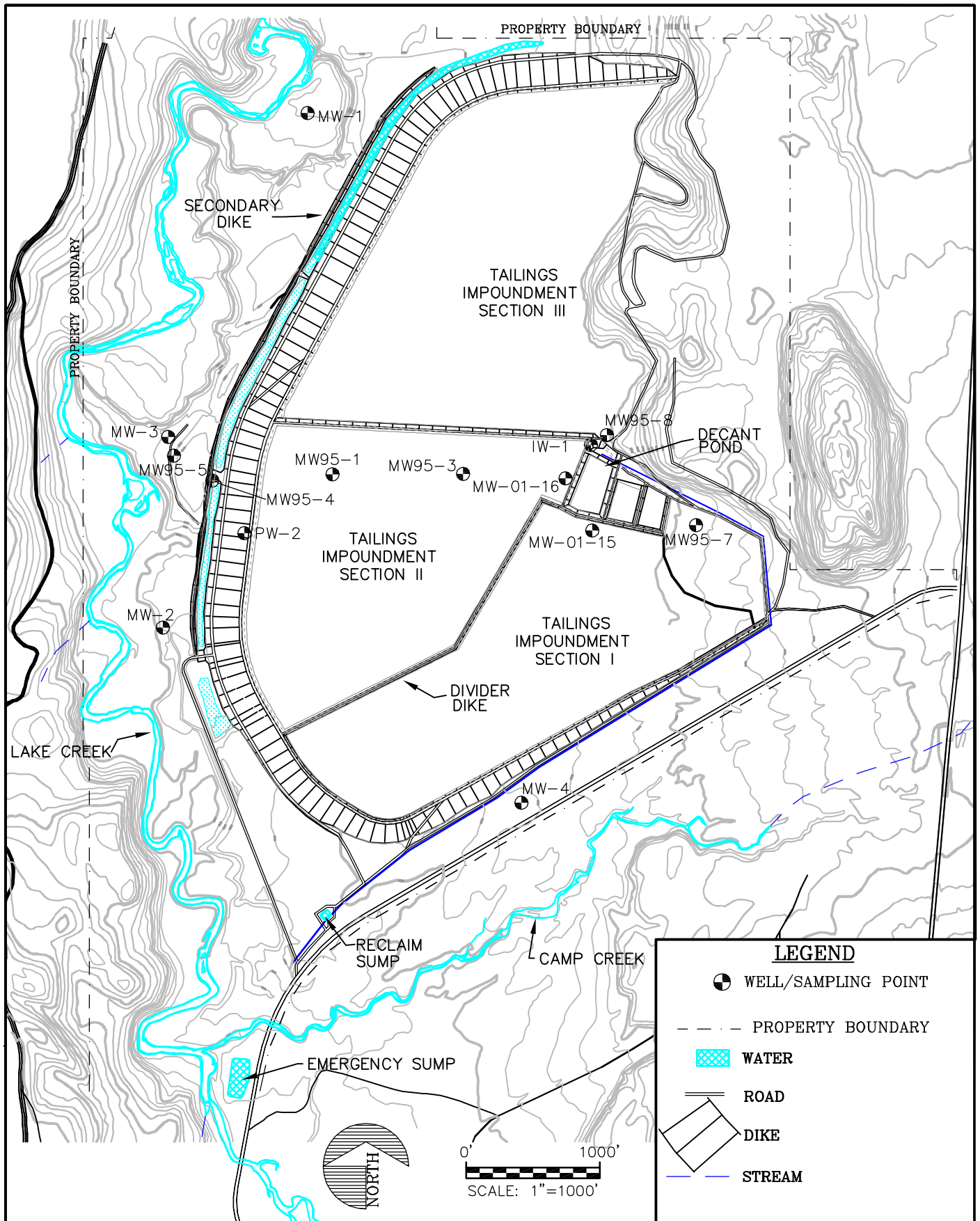
During the interim shutdown period of 1993 through 2004, Asarco and Genesis Inc. disposed of water from the underground workings by discharge to the tailings impoundment in one of three ways:

- Sprinkler irrigation over the impoundment (primarily during the growing season).
- Discharge to the decant pond (also known as the barge pond in some reports).
- Discharge to the Section III pond (an area in Section III of the impoundment where water ponds on the surface before flowing through a pipe to the decant pond).

The general layout of the tailings impoundment and associated ponds and monitoring wells is shown on Figure 2. Long-term disposal of mine water through the decant pond system after mine closure has been proposed as part of the Mine's final reclamation plan (Genesis, Inc., 2006) and is being evaluated by Montana Department of Environmental Quality (DEQ) as part of the ongoing environmental assessment of the final reclamation plan.

The primary constituent of concern is copper, which is typically present in mine water at concentrations of approximately 0.04 to 0.17 mg/L during shutdown periods. For comparison, copper concentrations in Lake Creek adjacent to the impoundment typically range from approximately <0.001 to 0.003 mg/L (higher values occur during high flow events), the chronic aquatic life standard for copper is approximately 0.004 mg/L (assuming hardness of 35 mg/L) and the Montana groundwater quality standard and Federal drinking water standard is 1.3 mg/L. In spite of the higher copper concentrations discharged to the impoundment, copper concentrations in Lake Creek and in groundwater monitoring wells installed in the impoundment vicinity have not increased measurably as a result of the

**FIGURE 2. MAP OF TAILING IMPOUNDMENT SHOWING LOCATION OF
DECANT POND AND MONITORING WELLS**



discharge of mine waters (see Summit, 1999; Parametrix, 2008). Similar to copper, concentrations of other metals (arsenic, antimony, lead, and uranium) are significantly reduced as decant water infiltrates the sediments beneath the pond and flows in the groundwater system. Other metals of interest (cadmium, lead, mercury, silver) are present at very low concentrations and are rarely detected in mine water and decant pond water.

The observed natural attenuation of copper and other metals (i.e., the reduction in metal concentrations through natural geochemical processes) in the decant pond and groundwater system led to the consideration of the infiltration of mine water at the tailing impoundment as a potential long-term disposal and treatment method for mine water following mine closure. From 2000 through 2004, a series of field and laboratory investigations were conducted to identify the geochemical processes responsible for the natural attenuation of copper in the decant pond and groundwater system (Hydrometrics, 2000 and 2001; Land & Water Consulting, 2004). The primary findings of these studies are:

1. Permeable sand and gravel layers beneath the decant pond provide the primary pathway for transport of mine water in the subsurface;
2. Copper concentrations are quickly attenuated (reduced) to levels below human health and aquatic life standards during groundwater transport within a short distance (ten to one hundred feet) of the decant ponds; and
3. Geochemical mechanisms responsible for copper attenuation include primarily the formation of copper minerals and secondarily adsorption by organic matter in soils and sediment beneath the decant pond.

The copper minerals forming in the pond sediments and groundwater are identical to the secondary copper minerals that form within the underground workings as the primary sulfide minerals are exposed to oxygen, become thermodynamically unstable, and oxidize to form new (i.e., secondary) minerals. Thus, the copper attenuation mechanism occurring in the decant pond sediments and groundwater is the completion of a chemical reaction started in the underground mine.

Subsequent to the 2000-2004 studies (that focused on copper), the natural attenuation of arsenic, antimony, cadmium, lead, mercury, silver, and uranium from mine water and decant pond water have been evaluated by additional investigations. CDM (2010) conducted geochemical modeling and lab experiments to demonstrate the natural attenuation provided by the precipitation of iron from mixing of mine water and decant pond water with natural ambient groundwater. These studies demonstrated that precipitation of naturally-occurring iron and manganese in groundwater can significantly reduce the concentrations of arsenic, antimony, copper, lead, and uranium in mine water and decant pond water.

1.2 PURPOSE AND SCOPE OF EVALUATION

The purpose and scope of this evaluation is to further evaluate the natural attenuation of copper as well as other metals in mine water and groundwater beneath the decant pond. The purpose and scope is based on addressing issues identified in the following communications/interactions with Montana DEQ and Kootenai National Forest:

- March 27, 2008 DEQ Task Order to CDM on copper attenuation;
- May 8, 2008 letter from DEQ to Doug Miller (Genesis);
- December 3, 2008 conference call with Wayne Jepson and Jim Castro (DEQ), Kent Whiting (CDM), Fess Foster (consultant to Genesis), Doug Parker and Scott Mason (Hydrometrics);
- December 10 meeting in Helena with DEQ, KNF (by phone), Genesis, Fess Foster, and Hydrometrics; and
- January 8, 2009 conference call with Wayne Jepson, Jim Castro, Lisa Boettcher, and Pat Plantenberg (DEQ), Kent Whiting (CDM), and Scott Mason (Hydrometrics).

The metal attenuation issues identified in these meetings and communications are:

- Evaluation of metals (antimony, cadmium, lead, mercury, silver, and uranium) in addition to copper that are present in mine water;
- Evaluation of the longevity of the natural metal attenuation mechanisms; and

- Identification of management practices and monitoring to ensure continued effectiveness of metal attenuation.

The Scope of Work for this evaluation was further clarified in the *Proposed Scope of Work for Copper Attenuation Study Addendum* (Hydrometrics, 2009). The proposed project is organized in a phased or tiered manner that follows the Tiered Analysis Approach recommended by EPA in *Monitored Natural Attenuation of Inorganic Contaminants in Groundwater, Volumes 1 and 2* (EPA, 2007a). EPA's tiered approach addresses four primary issues:

1. Demonstration of active contaminant removal from ground water;
2. Determination of the mechanism(s) and rate of attenuation;
3. Determination of the long-term capacity for attenuation and stability of immobilized contaminants; and
4. Design of performance monitoring program.

Some metal parameters that were initially identified for evaluation (mercury and silver) are not present at measurable concentrations in mine water and decant pond water. Moreover, potential concentrations (i.e., detection limit concentrations) in mine water and decant pond water are better than (below) applicable water quality standards. Therefore, this evaluation of the fate of these metals does not include a detailed evaluation of their attenuation behavior based on field evidence or geochemical modeling. For these parameters, the evaluation is limited to a summary of the available literature.

1.3 LIMITATIONS AND UNCERTAINTIES

As with all predictive studies, this study has certain limitations and uncertainties. They are listed below, and discussed in the following paragraphs. In spite of these uncertainties, it is felt that this study is a reliable prediction of post-mine closure water quality and the behavior

of metals in the infiltration pond and groundwater system. The reliability of these predictions is bolstered by the empirical evidence of attenuation that has occurred for over thirty years.

Limitations and uncertainties in this evaluation include:

1. The lack of actual post-closure mine water for evaluation;
2. The lack of the analysis of several metals in some previous investigations;
3. The very low or non-detectable concentrations of some metals;
4. The assumption of chemical equilibria or disequilibria in geochemical modeling; and
5. Uncertainties in the thermodynamic data upon which geochemical calculations and modeling is based.

After mine closure, disposal water would be composed exclusively of mine water plus rainfall and snowmelt that enters the disposal pond. Because the mine is not yet permanently closed, post-closure mine water does not yet exist and post-closure mine water quality cannot be directly measured. Therefore, it is necessary to base the evaluation of natural attenuation on the behavior of metals in analogous waters, namely interim shutdown mine water, operational mine water, and decant pond water. Although this adds some uncertainty to this analysis, the Troy Mine has been operating for over 30 years and there is an extensive record of mine water quality data that indicates that mine water quality is consistent over time. In addition, mine water quality data for the interim closure period (1995 through 2004) is believed to be a very good predictor of future post-closure water quality.

Concentrations of some metals (in particular arsenic, cadmium, lead, and uranium) in mine water, decant pond water, and groundwater are frequently below or just above the analytical reporting limit. The inability to determine actual concentrations in waters with reported “nondetect” results limits the extent of the evaluation of the attenuation behavior of these metals because known concentrations are required in order to quantify the amount of attenuation and to conduct geochemical models. In addition, detected results that are near the limits of analytical detection have a relatively high amount of uncertainty or imprecision.

However, because concentrations of these metals are very low and near or below detection limits, they also tend to be below applicable water quality standards and are of less environmental concern.

Extensive field and laboratory studies were conducted in 2000 through 2004 on the mechanisms of the natural attenuation of copper in mine water disposed at the decant pond (Land and Water Consulting, 2004). Metals in addition to copper were not identified as potential problems at the time and little data concerning the mechanisms of attenuation of other metals were collected. Although the evidence for natural attenuation of these other metals is strong (i.e., concentrations of arsenic, antimony, lead, and uranium in groundwater are demonstrated to be significantly reduced through natural geochemical processes), field evidence of the specific mechanisms of attenuation (e.g. adsorption, mineral precipitation, etc.) for metals other than copper is limited. In this report, the nature of attenuation mechanisms for copper and other metals is evaluated and demonstrated by additional site-specific laboratory testing and geochemical modeling of attenuation mechanisms. In addition, natural attenuation has been studied and demonstrated extensively at many other sites and a review of natural attenuation mechanisms based on scientific literature is also provided.

Virtually all geochemical modeling software including the geochemical model and software (Geochemist's Workbench; RockWare, 2010) used in this assessment are equilibrium models. Although complete chemical equilibrium is relatively rare in nature, equilibrium-based models remain powerful tools as they help to define geochemical conditions and indicate the types of geochemical changes that are likely to occur. In addition, if parts of the geochemical system are known to be in disequilibrium or partial equilibrium, the equilibrium models can be forced to maintain disequilibrium or partial equilibrium in portions of the modeled system. Instances where partial disequilibrium is assumed for modeling purposes are described in Section 5.

The thermodynamic data used for geochemical calculations and modeling in this report are from recognized sources (e.g. Lawrence Livermore National Laboratory, 2006) and are

generally considered to be valid. However, all measurements contain uncertainties and in the case of thermodynamic data such as equilibrium or solubility product constants, many of the numbers are extremely large or small (e.g. 10^{-50}) and small relative uncertainties in the data can yield relatively large uncertainties in the results of calculations (such as mineral saturation indices). To reduce the uncertainty in conclusions based on geochemical modeling, the model results are “field truthed” or corroborated with empirical evidence from field and laboratory studies to the extent possible.

2.0 OCCURRENCE OF METALS IN TROY MINE WATER

2.1 SOURCE OF METALS

The source of metals in the Troy mine water and decant pond water are the metal-bearing minerals present in and around the ore body and associated mineralized zones. As rainwater and snowmelt percolate through rock to the underground workings, the water contacts metal-bearing minerals and dissolves small amounts of copper and other trace metals. As described in the Rock Creek EIS (MDEQ and USFS, 2001):

“Sulfide mineral oxidation is a process that may occur in pH-neutral conditions, and is the probable source of the elevated copper in the water that is in the underground workings in the Troy mine. The Troy mine has developed and exposed the bornite-digenite-native silver ore zone. Bornite is a copper-iron-sulfur mineral and digenite is a copper-sulfur mineral. The underground pillars that support the overlying sedimentary rocks are composed of ore grade bornite-digenite-native silver. In a pH-neutral oxidizing environment, the bornite-digenite-native silver minerals that occur on fracture and joint surfaces in the rock are oxidized in place to a mixture of various copper oxide minerals. These copper oxide minerals are exposed in the outcrops of the Troy orebody on the south side of Mt. Vernon and underground in the mine workings. The oxide minerals include tenorite, chrysocolla, brochantite, malachite, and cupriferous goethite, Hayes and Balla, (1986). These are all secondary copper oxide minerals which occur in areas of low acidity. The bornite and digenite minerals that occur between the individual sand grains and that comprise the main ore zone are not leached, as they are encapsulated by the quartz overgrowths that occurred during diagenesis and burial metamorphism.

Subsequent melting snow water, percolating down through the various sedimentary rocks along these same fracture and joint surfaces, partially dissolve the oxidized copper oxide minerals, and the dissolved copper is seen and detected in the underground mine waters. It should be noted that in the Troy mine, the overlying galena halo zone and the pyrite halo zone were not mined and are not exposed in the Troy mine.

It should also be noted that the melting snow water, percolating down through the overlying galena and pyrite halo zones, is not creating acid rock conditions, as the sampling of the underground mine waters consistently show that the pH of the mine waters is 7.2 to 7.4, a near-

neutral to alkaline condition. These conditions are what is expected be similar at the Rock Creek project site.” (pg 3-29 Rock Creek FEIS Vol. 1. September 2001)

The source of other metals is similar to that of copper. Arsenic, antimony, cadmium, lead and mercury either form sulfide minerals or are present in sulfide minerals as substituting ions for sulfur or other metals. Oxidation of these sulfides leads to the release of trace amounts of these metals. Silver occurs as native silver and argentiferous sulfides in the ore body and is very slightly soluble (silver concentrations in mine water are below the analytical detection limit).

As described by Land and Water Consulting (2004) and summarized below, the copper attenuation mechanism operating in the decant pond and groundwater system today results in the formation of copper minerals that are identical to the secondary copper minerals that form within the underground workings as described in the Rock Creek EIS. Thus, the copper attenuation mechanism occurring in the decant pond sediments and groundwater is the continuation and completion of chemical reactions started in the underground mine.

2.2 CONCENTRATIONS OF METALS IN WATER

Mine water quality is routinely monitored by the collection of mine water from the mine adit at two locations: the Adit Ditch which includes wash down water from the Transfer Point and the Adit Pipe which is excess water pumped from the back of the mine. Over the course of the mine’s life the analytical detection limits attainable from laboratories has continually improved (minimum detection limits have decreased) and much of the mine water quality data collected in the early years of mining was reported at elevated detection limits, relative to current limits. Inclusion of this early data skews statistical calculations and misrepresents actual concentrations. Therefore, in this evaluation only more recent data (2000 through 2009) is considered.

The concentration of metals in recent mine water (2000 through 2009) for both the interim shutdown period (2000 though November 2004 data) and the recent operational period

**TABLE 1. METALS AND BICARBONATE CONCENTRATIONS
IN TROY MINE WATER**

Parameter (Dissolved)	Minimum (mg/L)	Maximum (mg/L)	Average* (mg/L)	Montana Groundwater Quality Standard** (mg/L)	Montana Surface Water Quality Standard** (mg/L)
Mine Water - Interim Shutdown Period 2000 through November 2004					
Arsenic	0.001	0.004	0.0021	0.010	0.010
Antimony	<0.003	0.016	0.0072	0.006	0.0056
Cadmium	No data	No data	No data	0.005	0.000097
Copper	0.044	0.17	0.076	1.3	0.00285
Iron	0.01	0.03	0.02	0.3	1
Lead	<0.002	0.003	0.0013	0.015	0.000545
Manganese	0.011	0.22	0.099	0.05	none
Mercury	No data	No data	No data	0.002	0.00005
Silver	No data	No data	No data	0.100	0.00037
Uranium	No data	No data	No data	0.03	0.03
Bicarbonate	20	113	60	none	None
Mine Water - Operational Period – December 2004 through 2009					
Arsenic	0.001	0.008	0.0026	0.010	0.010
Antimony	<0.003	0.015	0.0081	0.006	0.0056
Cadmium	0.00087	0.0022	0.0015	0.005	0.000097
Copper	0.041	0.986	0.339	1.3	0.00285
Iron	0.01	0.81	0.20	0.3	1
Lead	0.002	0.059	0.0204	0.015	0.000545
Manganese	0.025	0.312	0.177	0.05	None
Mercury	<0.000005	0.000005***	0.0000025	0.002	0.00005
Silver	<0.0005	<0.003	NC	0.100	0.00037
Uranium	<0.0003	0.0013	0.000091	0.03	0.03
Bicarbonate	49	112	84	none	none
Decant Pond Water – Operational Period – December 2004 through 2009					
Arsenic	<0.0005	0.005	0.0019	0.010	0.010
Antimony	0.008	0.062	0.028	0.006	0.0056
Cadmium	<0.0001	0.00126	0.00062	0.005	0.000097
Copper	0.006	0.20	0.0238	1.3	0.00285
Iron	0.008	0.38	0.082	0.3	1
Lead	<0.0005	0.0016	0.0013	0.015	0.000545
Manganese	0.101	0.791	0.4927	0.05	none
Mercury	<0.000005	<0.000005	<0.000005	0.002	0.00005
Silver	<0.0005	<0.003	NC	0.100	0.00037
Uranium	0.0029	0.0033	0.0031	0.03	0.03
Bicarbonate	52	121	93	none	none

Period of Record = 2000 through 2009 Data. Mine water includes Adit Ditch and Adit Pipe sampling locations.

NC = average not calculated as all values are less than analytical detection limit.

* One half of the detection limit value used for nondetect results in calculation of average.

**Circular DEQ-7 Montana Numeric Water Quality Standards, February 2008. Values for surface water standards are the lower of the human health standard or the chronic aquatic life criteria. Hardness dependent criteria based on hardness of 25 mg/L. Values for iron and manganese in groundwater is the Federal Secondary Drinking Water Standard. For silver there is no chronic aquatic life criterion, value shown is acute criterion.

***Laboratory QC results indicate that this result is biased high.

(December 2004 through 2009 data) are given in Table 1. Both the Adit Ditch and Adit Pipe are representative of mine water and therefore data from both sites is combined in Table 1. For comparison, metal concentrations in decant pond water and Montana water quality standards are also shown. Graphs in Appendix B show the concentrations of selected chemical constituents in mine water and decant pond water for the period of 2000 through 2009 and for nitrogen species for the course of mine life.

There are several key observations regarding mine water and decant pond water quality:

1. Mercury and silver are not detected, or if detected, are present at concentrations below relevant Montana water quality standards.
2. On average, metal concentrations are generally higher in mine water during the operational period (e.g. copper average concentration of 0.076 mg/L during shutdown versus 0.339 mg/L during operations; lead average concentration of 0.0013 mg/L during shutdown versus 0.0204 mg/L during operations). However, more recent (2009) operational data indicates that copper and lead concentrations have returned to near interim shutdown levels.
3. During the shutdown period, antimony is the only metal parameter that exceeds the groundwater standard (average concentration of 0.0072 mg/L in mine water compared to the standard of 0.006 mg/L). Antimony concentrations are similar in operational period mine water (average of 0.0081 mg/L). Antimony concentrations are steady in operational period mine water and in decant pond water through 2007 (about 0.02 mg/L) but increased in decant pond water in 2009 to about 0.06 mg/L. After the cessation of mining, mine disposal water is expected to be similar to current mine water (<0.003 to 0.016 mg/L).

2.3 CONCENTRATION OF COMMON ION PARAMETERS IN WATER

Graphs in Appendix B show the concentrations of selected common chemical constituents in mine water and decant pond water over the course of mine life. Notable observations regarding common ion concentrations include:

1. Much lower nitrogen concentrations during interim shutdown. The source of nitrogen in mine water is predominantly from the residues of explosives used in blasting and mining. After mining (blasting) ceases, nitrogen levels quickly fall to low levels.
2. Sulfate concentrations and pH levels in mine water are somewhat variable, probably from seasonal variations in rain and snowmelt, but relatively steady exhibiting no overall trends. The lack of trend suggests that sulfide oxidation rates within the mine are relatively steady as well, and therefore no significant change in sulfide oxidation rate and mine water pH is expected after final mine closure.
3. Common cation (calcium, potassium, sodium) concentrations are steady in mine water but exhibit some variation in decant pond water due to the addition of potassium and sodium-bearing milling reagents. Potassium and sodium concentrations in decant pond water are similar to mine water during the interim shutdown period, but are increased during active mining phases. Calcium concentrations are higher during the recent mining phase and correlate with increases in bicarbonate concentrations.
4. Bicarbonate (the principle contributor to alkalinity) concentrations during the interim mine closure typically varied between 20 and 80 mg/L and typically were about double calcium concentrations, suggesting some control of bicarbonate levels by calcite (CaCO_3) solubility. Since the mine re-start in 2005, bicarbonate levels have increased and in 2009 ranged from approximately 80 to 120 mg/L (Table 1). This increase is believed to reflect the geology of the current mining area where the ore body is higher in carbonate minerals.

3.0 REVIEW OF METAL ATTENUATION LITERATURE

This Section describes the environmental chemistry and behavior of selected metals in the natural environment. This information is summarized from the following EPA publications:

- Water-related Environmental Fate of 129 Priority Pollutants, Volume 1 and 2. (EPA, 1979);
- Monitored Natural Attenuation of Inorganic Contaminants in Ground Water, Volume 1 and 2, (EPA, 2007a);
- Metal Attenuation Processes at Mining Sites (EPA, 2007b);
- Issue Paper on the Environmental Chemistry of Metals (Langmuir et al., 2004); and
- Understanding Variation in Partition Coefficient, K_d , Values, Volumes II and III (EPA, 1999 and 2004).

3.1 NATURAL ATTENUATION MECHANISMS

Natural attenuation is typically defined as the removal or immobilization of dissolved chemical constituents (e.g. metals such as copper) from groundwater by interactions with natural aquifer materials such as soil, rock and organic matter. For inorganic constituents such as metals, the primary attenuation mechanisms are:

1. Sorption to aquifer solids;
2. Mineral precipitation; and
3. Co-precipitation.

These mechanisms may work independently or combined. For instance a metal may initially be primarily sorbed to a substrate at low surface concentrations and then form a surface precipitate as the surface concentration reaches a level where mineral nucleation and growth can occur. Langmuir et al (2004) describe this combined adsorption/precipitation as follows:

“Except for Al and sometimes Mn, concentrations of trace metals in the environment are generally too low for those metals to exceed the solubility products of their pure metal solids and thus to precipitate. Instead, toxic metal concentrations are generally limited by sorption onto the

surfaces of minerals, and onto organic matter including microbial cell wall surfaces. For aluminosilicate-rich soils, however, surface precipitation of Zn, Ni, and Co as layered double hydroxides was reported (e.g. Ford and Sparks, 2000). As metal concentrations further increase and fill available sorption sites, most metals tend to be incorporated in the structures of major mineral precipitates as “coprecipitates” in which they substitute for major metal cations, forming so-called solid solutions (Langmuir, 1997a). At higher metal concentrations, the metals may be precipitated in pure metal phases, limiting further increases in metal concentration.”

The extent of sorption of trace metals in groundwater is largely controlled by the following factors:

1. Groundwater pH;
2. Metal concentration in solution;
3. The quantity of sorbent material (principally hydrous ferric and manganese oxides, organic matter, and clays) within the aquifer; and
4. The concentration of competing ions in solution.

The extent of co-precipitation and precipitation of metals in groundwater is largely controlled by the following factors:

1. Groundwater pH;
2. Groundwater redox (e.g. Eh or pe) conditions;
3. Metal concentrations in solution;
4. The concentrations of common ions (e.g. sulfate, bicarbonate/carbonate, hydroxide) that can form minerals with the metal ions;
5. The solubility and stability of metal-bearing minerals; and
6. The concentrations of common ions that form soluble metal complexes in solution.

Summaries of the attenuation mechanisms that have been identified in the scientific literature as controlling factors for the metals of interest are provided in Tables 2 and 3.

**TABLE 2. SUMMARY OF ATTENUATION PROCESSES
FOR ANTIMONY, MERCURY AND SILVER**

Metal	Attenuation Mechanisms
Antimony	Sorption by iron, manganese, and aluminum oxides and clay
Mercury	Sorption by organic matter, iron and manganese oxides and clays
Silver	Sorption by iron and manganese oxides and organic matter

**TABLE 3. SUMMARY OF ATTENUATION PROCESSES FOR ARSENIC,
CADMIUM, COPPER, LEAD, URANIUM, AND ZINC**

Contaminant	Possible Attenuation Reactions	Relevant parameters
As	Sorption in aerobic environments Sorption/Precipitation in anaerobic environments	<ul style="list-style-type: none"> • Abundance/stability of hosts; typically Fe and Al (hydr)oxides • Solid-phase sulfide accumulation; redox buffer capacity, sulfate reducing activity
Cd	Sorption in aerobic environments Sorption/Coprecipitation carbonates Sorption/Precipitation in anaerobic environments	<ul style="list-style-type: none"> • Abundance/stability of hosts; typically Fe and Al (hydr)oxides • Abundance/stability of hosts; may require consideration of pH buffer capacity • Solid-phase sulfide accumulation; redox buffer capacity, sulfate reducing activity
Cu	Sorption in aerobic environments Sorption/Coprecipitation carbonates Sorption/Precipitation in anaerobic environments	<ul style="list-style-type: none"> • Abundance/stability of hosts; typically Fe and Al (hydr)oxides • Abundance/stability of hosts; may require consideration of pH buffer capacity • Solid-phase sulfide accumulation; redox buffer capacity, sulfate reducing activity
Pb	Sorption/Coprecipitation in aerobic environments Precipitation as hydroxycarbonate or sulfate Sorption/Precipitation in anaerobic environments	<ul style="list-style-type: none"> • Abundance/stability of hosts; typically Fe and Al (hydr)oxides • Aquifer pH buffer capacity • Solid-phase sulfide accumulation; redox buffer capacity, sulfate reducing activity
U	Reductive Precipitation Sorption	<ul style="list-style-type: none"> • Abundance/reactivity of electron donors • Abundance of hosts; typically metal (hydr)oxides
Ni	Sorption in aerobic environments Sorption Sorption/Precipitation in anaerobic environments	<ul style="list-style-type: none"> • Abundance/stability of hosts; typically Fe and Al (hydr)oxides • Abundance of hosts (clays) • Solid-phase sulfide accumulation; redox buffer capacity, sulfate reducing activity

Note: Measurement objectives and methodologies to support MNA investigations are documented in U.S. Environmental Protection Agency (2007). MNA assessments focus on both the aqueous phase and the solid phase in order to identify attenuation pathways.

Source: Table 2 in Metal Attenuation Processes at Mining Sites (EPA, 2007b).

3.2 ANTIMONY

The attenuation of antimony is primarily controlled by adsorption mechanisms as described by Langmuir et al (2004):

“[Antimony occurs] chiefly as oxyanions in oxidizing environments. As such they are relatively mobile, although they are adsorbed by ferrihydrite under acid to neutral conditions.” (p. 60)

“[A]ntimony may be too soluble for its concentration to be limited by mineral precipitation. The least soluble Sb phase is probably $\text{Sb}(\text{OH})_3$, which does not precipitate until Sb concentrations exceed about $10^{-6.67}$ mol/kg, or 26 ug/L (Barnes and Langmuir, 1978). Antimony is a weak complex former, except for its reaction with sulfur at low Eh to form sulfide complexes.” (p. 60)

As described by USEPA (1979):

“The extent to which sorption reduces the aqueous transport of antimony is unknown, but it is clear that sorption processes are normally the most important mechanisms resulting in the removal of antimony from solution. Antimony apparently has an affinity for clay and other mineral surfaces. Coprecipitation of antimony with hydrous iron, manganese, and aluminum oxides may exert a significant control on antimony mobility in areas where there is active precipitation of these metals.”(p. 5-3)

3.3 ARSENIC

The attenuation of arsenic in groundwater is primarily controlled by adsorption onto iron oxides and clays. As described by EPA (2004):

“Arsenic is a known carcinogen, and occurs in natural systems in the +5 (arsenate) and +3 (arsenite) valence states. Arsenic(III) ($\text{As}(\text{III})$) is more mobile (adsorbs less) and is many times more toxic than $\text{As}(\text{V})$. The pH and redox conditions are the two most important geochemical factors affecting the mobility of arsenic in the environment. Sorption studies indicate that the concentrations of dissolved $\text{As}(\text{V})$ and $\text{As}(\text{III})$ are controlled by adsorption on iron and aluminum oxides and clays. The adsorption of $\text{As}(\text{V})$ is high and independent of pH at acidic pH values, and decreases with increasing pH in the range of 7 to 9. Arsenic(V) adsorption will be decreased in soils with high phosphate concentrations because of anion competition.” (EPA, 2004, p. 5.14)

“Sorption studies have also been conducted to determine the effect on arsenic sorption by the presence of other anions, such as dissolve phosphate, sulfate, nitrate, and chloride (Reynolds et al., 1999; Livesey and Huang, 1981). The results of these studies indicate that dissolved phosphate competes with arsenic for adsorption sites, and thus suppressing arsenic adsorption and enhancing the mobility of arsenic. Livesey and Huang (1981) found that dissolved sulfate, nitrate, and chloride present at concentrations present in saline soils had little effect on arsenic adsorption.” (EPA, 2004, p. 5.14)

As described by USEPA (1979):

“Cycling of arsenic in the aquatic environment is dominated by adsorption and desorption to sediments. Arsenic may be sorbed onto clays, aluminum hydroxide, iron oxides, and organic material. ... Under most conditions, coprecipitation or sorption of arsenic with hydrous oxides of iron is probably the prevalent process in the removal of dissolved arsenic.” (p. 6-6)

3.4 CADMIUM

The attenuation of cadmium is primarily controlled by mineral precipitation and adsorption mechanisms as described by EPA (1999):

“The dominant cadmium aqueous species in groundwater at pH values less than 8.2 and containing moderate to low concentrations of sulfate ($<10^{-2.5}$ M SO_4^{2-}) is the uncomplexed Cd^{2+} species. The dominant cadmium solution species in groundwater at pH values greater than 8.2 are CdCO_3 (aq) and to a smaller extent CdCl^+ . Both precipitation/coprecipitation/dissolution and adsorption/desorption reactions control cadmium concentrations. Several researchers report that otavite (CdCO_3) limits cadmium solution concentrations in alkaline soils. The solid $\text{Cd}_3(\text{PO}_4)_2$ has also been reported to be a solubility-controlling solid for dissolved cadmium. Under low redox conditions, sulfide concentrations and the formation of Cd precipitates may play an important role in controlling the concentrations of dissolved cadmium. At high concentrations of dissolved cadmium ($>10^{-7}$ M Cd), either cation exchange or (co)precipitation are likely to control dissolved cadmium concentrations. Precipitation with carbonate is increasingly important in systems with a pH greater than 8, and cation exchange is more important in lower pH systems. At lower environmental concentrations of dissolved cadmium, surface complexation with calcite and aluminum- and iron-oxide minerals may be the primary process influencing retardation. Transition metals (e.g. copper, lead, zinc) and alkaline earth (e.g. calcium, magnesium) cations

reduce cadmium adsorption by competition for available specific adsorption and cation exchange sites. In conclusion, the key aqueous- and solid-phase parameters influencing cadmium adsorption include pH, cadmium concentration, competing cation concentrations, redox, cation exchange capacity (CEC), and mineral oxide concentrations.” (EPA, 1999)

Competition between cations for adsorption sites strongly influences the adsorption behavior of cadmium. The presence of calcium, magnesium, and trace metal cations reduce cadmium adsorption by soils (Cavallaro and McBride, 1978; Singh, 1979), iron oxides (Balistrieri and Murray, 1982), manganese oxides (Gadde and Laitinen, 1974), and aluminum oxides (Benjamin and Leckie, 1980). ... The addition of copper or lead, which are more strongly adsorbed, slightly reduces cadmium adsorption by iron and aluminum oxides ... In contrast, zinc almost completely displaces cadmium, indicating that cadmium and zinc compete for the same group of binding sites. (EPA, 1999, p.5.11)

Langmuir et al (2004) summarize cadmium mobility in the environment as follows:

“[Cadmium] sulfides are quite insoluble at low Eh, even at low metal and sulfide concentrations (Figures 18 through 23). ... Other important Cd and Zn minerals include their carbonates, although the carbonates are relatively soluble at pH values below 8. Pure metal-containing mineral phases generally do not control the dissolved concentrations of Cd, Ni, or Zn in aerobic soils. These metals are more often controlled through adsorption or coprecipitation by oxyhydroxides of iron, manganese, and aluminum.”

3.5 COPPER

Langmuir et al (2004) summarize copper mobility in the environment as follows:

“Silver, copper, and mercury are highly insoluble in reduced environments, where they precipitate as metals or as sulfides. In the absence of ligands other than OH- and depending on the copper concentration, the solubility of copper above pH 7–8 can be very low, due to the precipitation of tenorite (CuO). Moreover, in the presence of abundant carbonate, relatively insoluble Cu²⁺ carbonate minerals can precipitate. (p. 49)

Mercury and copper are strongly adsorbed by organic matter. All three metals are also strongly adsorbed by Fe(III) and Mn oxides, and secondarily by clays. Sorption of mercury is very fast

and practically irreversible (Bodek et al., 1988). By inhibiting mercury sorption, Hg-Cl complexing helps to mobilize the metal. Dimethyl mercury is very insoluble in water and tends to be volatilized from soils.” (p. 60)

As described by USEPA (1979):

“Sorption of copper by precipitating hydrous iron and manganese oxides is an effective control on dissolved copper concentrations where these metals are being actively weathered or otherwise introduced into unpolluted aquatic environments.” (p. 11-1)

“Copper has a strong affinity for hydrous iron and manganese oxides, clays, carbonate mineral, and organic matter. Sorption to these materials, both suspended in the water column and in the bed sediment, results in relative enrichment of the solid phase and reduction in dissolved levels. Hem and Skougstrad (1960) demonstrated that coprecipitation of copper with the hydrous oxides or iron effectively scavenges copper from solution.” (p. 11-6)

“Copper is adsorbed to clay and mineral surfaces and organic materials. ... The addition of various anions significantly increased adsorption. Humic acid was particularly effective in this regard.” (p 11-9)

3.6 LEAD

The mobility and attenuation of lead is primarily controlled by mineral precipitation and adsorption mechanisms as described by EPA (1999):

“Lead has 3 known oxidation states, 0, +2, and +4, and the most common redox state encountered in the environment is the divalent form. Total dissolved lead concentrations in natural waters are very low ($\sim 10^{-8}$ M). Dissolved lead in natural systems may exist in free ionic form and also as hydrolytic and complex species. Speciation calculations show that at pH values exceeding 7, aqueous lead exists mainly as carbonate complexes $[\text{PbCO}_3(\text{aq})]$, and $[\text{Pb}(\text{CO}_3)_2^{2-}]$. Important factors that control aqueous speciation of lead include pH, the types and concentrations of complexing ligands and major cationic constituents, and the magnitude of stability constants for lead-ligand aqueous complexes.

A number of studies and calculations show that under oxidizing conditions depending on pH and ligand concentrations, pure-phase lead solids, such as PbCO_3 , $\text{Pb}_3(\text{OH})_2(\text{CO}_3)_2$, PbSO_4 ,

$\text{Pb}_5(\text{PO}_4)_3(\text{Cl})$, and $\text{Pb}_4\text{SO}_4(\text{CO}_3)_2(\text{OH})_2$, may control aqueous lead concentrations. Under reducing conditions, galena (PbS) may regulate the concentrations of dissolved lead. It is also possible that lead concentrations in some natural systems are being controlled by solid solution phases such as barite ($\text{Ba}(1-x)\text{Pb}_x\text{SO}_4$), apatite [$\text{Ca}(1-x)\text{Pb}_x(\text{PO}_4)_3\text{OH}$], calcite ($\text{Ca}(1-x)\text{Pb}_x\text{CO}_3$), and iron sulfides ($\text{Fe}(1-x)\text{Pb}_x\text{S}$).

Lead is known to adsorb onto soil constituent surfaces such as clay, oxides, hydroxides, oxyhydroxides, and organic matter. In the absence of a distinct lead solid phase, natural lead concentrations would be controlled by adsorption/desorption reactions. Adsorption data show that lead has very strong adsorption affinity for soils as compared to a number of first transition metals. Lead adsorption studies on bulk soils indicate that the adsorption is strongly correlated with pH and the CEC values of soils. Properties that affect CEC of soils, such as organic matter content, clay content, and surface area, have greater affect on lead adsorption than soil pH.” (EPA, 1999)

Langmuir et al (2004) summarize lead mobility in the environment as follows:

“Lead is relatively immobile in soils, sediments, and ground waters. This reflects its strong tendency to be adsorbed by Fe and Mn oxides, but also the insolubility of a number of lead minerals including lead hydroxycarbonate, which limits lead concentrations in some public water systems, and pyromorphite, which controls lead concentrations in some soils adjacent to highways affected by road salt and leaded gasoline exhaust.”

3.7 MERCURY

Mercury is not present at measurable concentrations in mine water and decant pond water.

USEPA (1979) summarize the fate and mobility of mercury in the environment as follows:

“Mercury’s major removal mechanism from a natural water system is adsorption on the surfaces of particulate phases and subsequent settling to the bed sediment. The overwhelming majority of any dissolved mercury is removed in this manner with a relatively short time, generally in the immediate vicinity of the source.” (p. 14-1)

“Mercury is strongly sorbed to inorganic and organic particulates. Deposition of mercury-laden sediments in reducing zones can result in precipitation of the sulfide.”

“In summary, it is evident from environmental studies and theoretical considerations, that mercury adsorption onto the sediments is probably the most important process for determining the fate of mercury in the aquatic environment.” (p. 14-8.)

Langmuir et al (2004) summarize mercury mobility in the environment as follows:

“Bodek et al. (1988) offer a useful summary of the behavior of mercury. Hg(II) is usually complexed—in pure water as $\text{Hg}(\text{OH})_2^0$, and at chloride concentrations typical of fresh waters ($<10^{-2}$ mol/kg) as HgCl_2^0 . Both Hg(II) and Cu(II) form strong humate complexes, so that in soils $>99.9\%$ of the metals may be complexed.” (p. 49)

“Mercury and copper are strongly adsorbed by organic matter. All three metals are also strongly adsorbed by Fe(III) and Mn oxides, and secondarily by clays. Sorption of mercury is very fast and practically irreversible (Bodek et al., 1988). By inhibiting mercury sorption, Hg-Cl complexing helps to mobilize the metal. Dimethyl mercury is very insoluble in water and tends to be volatilized from soils.” (p. 60)

3.8 SILVER

Silver is not present at measurable concentrations in mine water and decant pond water. USEPA (1979) summarize the fate and mobility of silver in the environment as follows:

“Sorption by manganese dioxide and precipitation with halides are probably the dominant controls on the mobility of silver in the aquatic environment.” (p. 17-1)

“Sorption appears to be the dominant process leading to partitioning of silver into the sediments. It appears that manganese dioxide has a strong affinity for [silver], followed by ferric hydroxide and clay minerals. ... Organic materials also adsorb silver.” (p. 17-3.)

Langmuir et al (2004) summarize silver mobility in the environment as follows:

“As soft metals, silver and mercury form strong complexes with borderline soft Cl ion, which may dominate the solution chemistry of these metals. Silver and mercury form even stronger complexes with the soft halogens bromide and iodide. Silver, copper, and mercury are highly insoluble in reduced environments, where they precipitate as metals or as sulfides.” (p. 49)

“All three metals [copper, mercury, silver] are also strongly adsorbed by Fe(III) and Mn oxides, and secondarily by clays.” (p. 60)

3.9 URANIUM

Uranium attenuation is primarily controlled by adsorption. As described by EPA (1999):

“Uranium (VI) species dominate in oxidizing environments. Uranium(VI) retention by soils and rocks in alkaline conditions is poor because of the predominance of neutral or negatively charged species. An increase in CO₂ pressure in soil solutions reduces U(VI) adsorption by promoting the formation of poorly sorbing carbonate complexes.” (p. 5.66)

“Some of the sorption processes to which uranyl ion is subjected are not completely reversible. Sorption onto iron and manganese oxides can be a major process for extraction of uranium from solution (Hsi and Langmuir, 1985; Waite *et al.*, 1994). These oxide phases act as a somewhat irreversible sink for uranium in soils. ... Naturally occurring organic matter is another possible sink for U(VI) in soils and sediments. ... Uranium sorption to iron oxide minerals and smectite clay has been shown to be extensive in the absence of dissolved carbonate (Ames *et al.*, 1982; Hsi

and Langmuir, 1985; Kent *et al.*, 1988). However, in the presence of carbonate and organic complexants, sorption has been shown to be substantially reduced or severely inhibited (Hsi and Langmuir, 1985; Kent *et al.*, 1988).” (p. 5-73)

4.0 FIELD EVIDENCE FOR NATURAL ATTENUATION

As described in Section 3, natural attenuation of metals is well established in the scientific literature based on numerous field and laboratory investigations. This section presents the site-specific field evidence for natural attenuation of metals from decant water in the groundwater system at the Troy Mine. Field evidence for natural attenuation is provided by the sampling and analysis of water and soil samples in the vicinity of the decant pond and from laboratory testing. Water quality data shows that the concentrations of chemical parameters that are not attenuated (e.g. chemically conservative parameters such as nitrate) are similar between the mine water/decant pond water and the underlying groundwater. This similarity in concentrations indicates that mine water is not appreciably diluted in the groundwater system. In contrast, the concentrations of metal parameters are much lower in the groundwater than in the mine water/decant pond, indicating that metal parameters are attenuated (removed) in the groundwater system. Analyses of sediments in the subsurface of the decant pond reveal that the metals are enriched in the sediments, indicating that the metals attenuated/removed from groundwater are retained on the sediments. Geochemical modeling described in Section 5 corroborates the field evidence.

4.1 FIELD EVIDENCE OF ACTIVE METALS REMOVAL

Water and soil sampling demonstrates that metals are removed from water and retained on solid phases on the sediment matrix as water infiltrates the decant pond and moves through the groundwater system.

4.1.1 Changes in Metal Concentrations in Groundwater

Decreases in metal concentrations in water (absent dilution) as decant pond/mine water moves through the groundwater system is evidence of active metals removal/attenuation. Since the changes in metal concentrations are rapid and dramatic, the effects of the infiltration of mine water at the decant pond can only be observed in close proximity to the pond. As described by Parametrix (2008), no changes in metal concentrations are detectable in Lake Creek adjacent the tailings impoundment or in groundwater monitoring wells near the toe of the impoundment. Changes in metal concentration (but not conservative species)

are apparent by comparison of decant pond and mine water with the three wells adjacent the pond: IW-1 (screened at 65 to 150 and 165 to 280 feet bgs), MW-95-8 (screened at 48 to 53 feet bgs), and MW-01-15 (screened at 30 to 40 feet bgs). Locations of these wells are shown on Figure 2.

Graphs comparing nitrogen and metal concentrations over time for decant pond, mine water and groundwater in the wells are provided in Appendix B. During active mine operations, mine water and decant pond water are typically elevated in nitrogen compounds (ammonia and nitrate) due the use of nitrogen-based explosives for mining purposes. Because nitrate is commonly a conservative species that is not attenuated in groundwater systems, a comparison of nitrogen concentrations between mine water/decant pond water and groundwater in the wells can be used to determine the extent of dilution in the groundwater system. As shown by the graphs in Appendix B, ammonia and nitrate + nitrite concentrations in groundwater in the monitoring wells match the concentrations in mine water and decant pond very closely, demonstrating that there is no appreciable dilution of the decant pond/mine water by ambient groundwater. In other words, groundwater near the decant pond is composed of mine water/decant pond that infiltrates the decant pond.

Because there is no appreciable dilution of mine water/decant pond water in the groundwater system monitored by the wells, any difference in metal concentration between the mine water/decant pond water and groundwater in the monitoring wells must be caused by removal (attenuation) of metals from the groundwater. As shown by the graphs in Appendix B, significant attenuation of arsenic, antimony, copper, and uranium is demonstrated to occur:

- Arsenic concentrations in decant pond water are very low (0.001 to 0.008 mg/L) and are reduced further such that arsenic is not detectable in groundwater in the monitoring wells (<0.001 to <0.003 mg/L).
- Antimony concentrations are low in mine water but range up to 0.062 mg/L in decant pond water. Concentrations are reduced to about 50 percent of the decant pond water

in the shallowest groundwater (0.032 mg/L in well MW-01-15) and further reduced to 0.005 mg/L in deeper groundwater (well MW-95-8).

- Copper concentrations were attenuated strongly, by approximately an order of magnitude in the mine shutdown phase, and to a slightly lesser extent in the current mining phase.
- Lead concentrations in mine water were very low (<0.002 to 0.003 mg/L) during the interim shutdown period and slightly higher (0.002 to 0.008 mg/L in 2009) during the current active mining phase with the exception of the mine re-start period in 2005 (up to 0.059 mg/L in 2005). Lead concentrations range from <0.0005 to 0.0016 mg/L in decant pond water during the current active mining phase. Lead concentrations are attenuated and reduced in groundwater adjacent the decant pond as the highest groundwater concentration is 0.0011 mg/L.
- Uranium concentrations are reduced by approximately a third between the decant pond (0.003 mg/L) compared to groundwater in the closest, shallowest monitoring well (0.002 mg/L in well MW-01-15, located approximately 100 feet from the pond) and further reduced by approximately 85 percent in the deeper well (0.0005 mg/L in well MW-95-8, located approximately 150 feet from the decant pond).

Cadmium concentrations in decant pond water and mine water are low, typically <0.001 to 0.002 mg/L. Cadmium concentrations in nearby groundwater are similar to decant pond indicating that little attenuation occurs over short transport distances.

4.1.2 Changes in Metal Concentrations in Mixed Water Samples

CDM (2010) conducted geochemical modeling and a series of jar tests where mine water and decant pond water were mixed with natural groundwater from well MW-95-4 in varying ratios. The purpose of the test was to test the hypothesis that mixing of decant pond water with ambient groundwater with naturally high iron concentrations would result in the precipitation of iron oxyhydroxides and attenuation of metals by co-precipitation and adsorption mechanisms. Well MW-95-4 is located below the toe of the tailing impoundment, between the impoundment and Lake Creek as shown on Figure 2. Groundwater with naturally high dissolved iron concentrations is known to exist in the area, along, and on both

sides of Lake Creek (EPA, 1991). If water from the decant pond were to flow toward Lake Creek, it is likely that decant water would become mixed with ambient groundwater in the area, such as that monitored by well MW-95-4.

Geochemical modeling of the mixing of decant pond water with groundwater in MW-95-4 by CDM (2010) suggests that upon mixing, large amounts of iron and manganese oxyhydroxides would precipitate. The formation of iron oxyhydroxides is known to result in strong attenuation of other metals (see for instance, Langmuir, et al 2004). CDM conducted a series of mixing tests in jars to further test this effect. In these tests, samples of decant pond water and mine water were mixed with samples of groundwater from well MW-95-4, allowed to sit and react for a period of 22 days, and then were analyzed for metal concentrations. Results of the jar tests are summarized in Table 4, Section 4.1.2. In all of the jar tests, the mixing of the more oxidized decant pond and mine water caused large amounts of iron to precipitate in the groundwater resulting in strong decreases in antimony, copper, lead and uranium concentrations.

TABLE 4. RESULTS OF LABORATORY TESTING OF METAL ATTENUATION BY MIXING WITH AMBIENT GROUNDWATER

Sample	Arsenic (mg/L)	Antimony (mg/L)	Copper (mg/L)	Iron (mg/L)	Lead (mg/L)	Manganese (mg/L)	Uranium (mg/L)
Groundwater MW-95-4	<0.003	<0.003	0.006	15.5	<0.0005	5.34	0.0005
Decant Pond Water	<0.003	0.068	0.032	<0.05	0.0008	0.862	0.0033
Mine Water (7 East Dam)	<0.003	0.015	0.114	0.114	0.001	0.105	0.0011
25% groundwater 75% decant	<0.001	0.046	0.007	0.008	<0.0005	2.17	0.0017
50% groundwater 50% decant	<0.001	0.026	0.005	0.009	<0.0005	3.49	0.0009
75% groundwater 25% decant	<0.001	0.010	0.003	0.008	<0.0005	4.86	0.0004
25% groundwater 75% mine water	<0.001	0.010	0.012	<0.005	<0.0005	1.52	0.0007
50% groundwater 50% mine water	<0.001	0.006	0.005	0.007	<0.0005	2.74	0.0004
75% groundwater 25% mine water	<0.001	0.002	0.003	<0.005	<0.0005	4.47	<0.0005

Source: Table 4-1. CDM (2010). Mine water was collected within the mine behind the 7 East Dam.

4.1.3 Metals Enrichment of Decant Pond Sediment

Geochemical testing of aquifer materials was conducted in 2003 during the interim shutdown period (Land & Water Consulting, 2004). Testing was conducted on sediment materials collected from pits dug in the bottom of the decant pond while the pond was nearly dry. Laboratory testing consisted of total metal analyses, sequential extraction analyses, microscopic examination and electron microprobe analyses.

Total metal analyses were performed to determine the overall metal content of sediments in order to identify areas of metal enrichment. Comparison of copper concentration in decant pond sediment with copper concentrations in Troy tailings and native soils indicates that all of the sediment samples collected at depths of less than four inches are strongly enriched in copper relative to native soils and tailings (Table 5 and Figure 3). Some soil/sediment samples (DP-1 3 ¾ “and DP-1 4”) are enriched to the extent that they contain nearly twice the copper content as ore from the Troy mine. Total copper concentrations in decant pond sediments are evidence that copper is attenuated in surficial soil/sediments in the decant pond.

**TABLE 5. TOTAL COPPER AND IRON CONTENT OF
DECANT POND SOIL SAMPLES**

Sample Designation	Total Copper (mg/kg or ppm)	Total Iron (mg/kg or ppm)
DP-1 algae	3,881	23,380
DP-1 0-1”	5,672	21,520
DP-1 3 ¾”	14,010	18,170
DP-1 4”	14,850	17,580
DP-1 4” organic	5,987	19,450
DP-1 2-4” slime tails	1,441	10,720
DP-1 4-8” sand tails	698	3,724
DP-1 1-4’ sand tails	594	4,269
DP-1 native sand	127	23,660
Troy tailings (1)	630	7,600
Troy ore (2)	7,600	--
Native soil (3)	10 to 30	

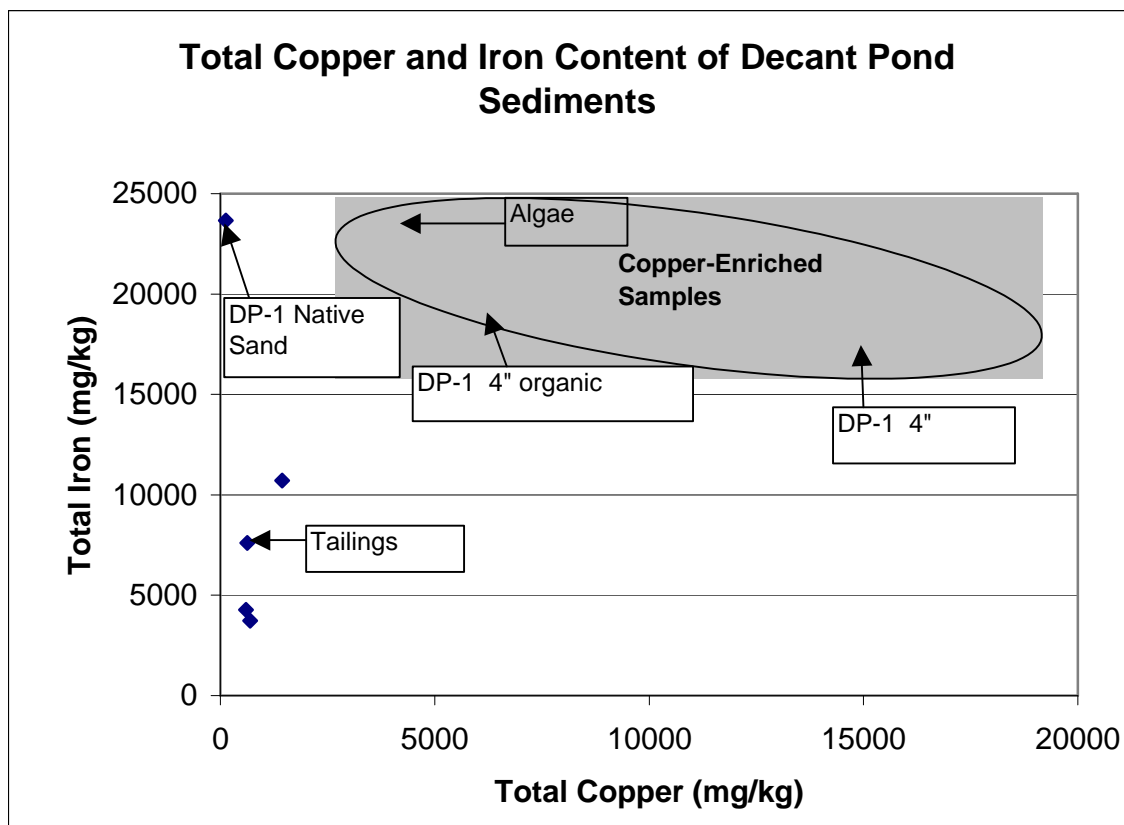
(1) Data source: Table 4-23, Rock Creek Final EIS, September 2001.

(2) Data source: Table 3-2, Rock Creek Final EIS, September 2001.

(3) Data source: Appendix D, Troy Unit Revised Reclamation Plan,
(Hydrometrics et al, December 2000).

Source of Table: (Land and Water Consulting, 2004)

FIGURE 3. COPPER AND IRON CONTENT OF DECANT POND SEDIMENTS



4.2 FIELD EVIDENCE OF ATTENUATION MECHANISMS

Evidence of attenuation mechanisms and rates of attenuation comes from the sampling and analysis of water and soil samples in the vicinity of the decant pond, laboratory mixing and testing of waters, and correlations between metal concentrations and water chemical indicators. Field evidence is corroborated by geochemical modeling as illustrated by Eh-pH and activity diagrams in Section 5. A summary of the identified attenuation mechanisms and the evidence supporting these mechanisms is provided in Table 6.

**TABLE 6. SUMMARY OF FIELD EVIDENCE
OF ATTENUATION MECHANISMS**

Mechanism	Metals Attenuated by Mechanism	Field Evidence
Carbonate Mineral Formation	Cu	Sequential extraction Visual identification Electron microprobe analysis
Ion Exchange/Adsorption	As, Cu, Sb, Pb and U	Sequential extraction Jar mixing tests
Organic Matter	Cu	Sequential extraction Visual identification
Silicate Mineral Formation	Cu	Sequential extraction Visual identification Electron microprobe analysis
Co-precipitation with Iron oxides	As, Cu, Sb, Pb and U	Sequential extraction Visual identification Electron microprobe Jar mixing tests

4.2.1 Identification of Metal Phases in Sediment

As decant pond water infiltrates sediments beneath the pond, metals are removed from the water and incorporated into the sediments as evidenced by the decrease in groundwater metal concentrations and the corresponding increase in sediment metal concentrations. The forms of occurrence or phases of metals in decant pond sediments are indicative of the mechanisms of attenuation that cause the forms or phases to occur. Thus, the attenuation mechanisms can be identified by identifying the forms of occurrence of metals in decant pond sediments. Corroboration of these mechanisms is also provided by geochemical modeling that indicates that mine water and decant pond are oversaturated with copper, iron, lead, and manganese minerals, thus metals are expected to precipitate and to be removed from these waters.

Geochemical testing of decant pond sediments was conducted in 2003 during the interim shutdown period (Land & Water Consulting, 2004). Testing was conducted on sediment materials collected from pits dug in the bottom of the decant pond while the pond was nearly dry. Laboratory testing consisted of total metal analyses, sequential extraction analyses, and

visual reflected light microscopy and electron microprobe analyses. As described in Section 4.1.3, total metal analyses demonstrated that sediment beneath the decant pond is highly enriched in copper due to the attenuation of copper as decant infiltrates the sediments. Sediments were also analyzed by sequential extraction and electron microprobe to identify the forms of copper in the sediments.

Sequential extraction analyses were performed to identify the sediment solid phases enriched in copper. These analyses determined the trace metal contents of six phases: water soluble, exchangeable and carbonate, manganese oxides, organic matter, amorphous iron oxides, silicates and crystalline oxides (see Gatehouse et al, 1977). Sequential extraction results indicate that the dominant copper-bearing phases in the decant pond soils are exchangeable ions and carbonates, organic matter and sulfides, and silicates. It should be noted that the sulfide minerals are believed to be ore minerals that were not captured during the milling and metal concentration process and are not believed to represent attenuation of metals by sulfide formation.

Visual (stereo microscope) and electron microprobe analyses of sediment samples were performed to identify mineral phases enriched in copper (Cannon Microprobe, 2003). Copper phases identified in the samples included:

1. Copper silicates (likely the blue-green mineral chrysocolla);
2. Copper carbonates (likely the green mineral malachite);
3. Copper manganese silicates;
4. Copper iron oxides; and
5. Copper iron sulfides (likely chalcopyrite and bornite).

4.2.2 Decant Pond/Mine Water and Groundwater Mixing Tests

Geochemical modeling of the mixing of decant pond water with groundwater in MW-95-4 by CDM (2010) indicates that upon mixing, large amounts of iron oxyhydroxides would precipitate. The formation of iron oxyhydroxides is known to result in strong attenuation of other metals (see for instance, Langmuir, 2004). To further test this adsorption/co-

precipitation effect, a series of water mixing tests were conducted in jars. In these tests, samples of decant pond water and mine water were mixed with samples of groundwater from well MW-95-4, allowed to sit and react for a period of 22 days and then were analyzed for metal concentrations. Results of the jar tests are summarized in Table 4, Section 4.1.2. In all of the jar tests, the mixing of the more oxidized decant pond and mine water caused large amounts of iron to precipitate in the groundwater resulting in strong decreases in antimony, copper, lead and uranium.

5.0 GEOCHEMICAL EQUILIBRIUM MODELING

The geochemical system considered in this report includes mine water, decant pond water, and groundwater and associated minerals (e.g. in sediments beneath the decant pond) and gases (e.g. oxygen and carbon dioxide in the atmosphere) that comprise the decant pond and adjacent groundwater system. In this assessment of natural attenuation mechanisms, the purpose of constructing thermodynamic equilibrium models of the geochemical system is to:

1. Determine the geochemical conditions of disposal water including redox conditions, the activities of metals and other ions in water, and the saturation indices of metal-bearing minerals that may be important in attenuation of metals;
2. Identify the geochemical controls on metal attenuation mechanisms and water quality;
3. Identify geochemical changes that are likely to occur to the water as water approaches geochemical equilibrium in the disposal pond and groundwater system; and
4. Identify the potential geochemical changes that could impede or preclude the attenuation of metals and potential mitigation methods.

The overall approach to geochemical modeling in this assessment is as follows:

1. Define the framework and goals of modeling;
2. Select the geochemical software and thermodynamic database;
3. Select the representative waters to be modeled and the system constraints;
4. Calculate the activities of the important aqueous metal species and the saturation indices of solid metal phases (i.e., minerals); and
5. Construct stability and solubility diagrams to graphically show the relationships between aqueous metal species, minerals, and the geochemical parameters (e.g. Eh and pH) controlling the relationships between the aqueous species and mineral phases.

5.1 PROCESS MODEL FOR GEOCHEMICAL EQUILIBRIUM MODELING

The first and most critical step in developing a geochemical model is conceptualizing the system or process of interest in a useful manner (Bethke, 2008). As described by EPA (2007a):

“The modeling effort begins with the careful identification of the processes that play significant roles in contaminant migration and attenuation at the site. In this way a conceptual model emerges that will eventually be coded into the input streams of the software packages that will produce the modeling results.” (p. 10)

The process model of the Troy Mine water disposal cycle is a description of the processes affecting water quality and forms a guide or general framework for formulating the computer-based geochemical model. In most cases, it is not necessary or practical to construct a geochemical model that includes all aspects of the process model. For this assessment of natural attenuation the primary area of interest is the water in the decant pond and groundwater system. Therefore, this portion of the water disposal cycle model is examined in detail using computer-based geochemical equilibrium models as described in this section.

The process model for geochemical equilibrium modeling is summarized in Table 7. In this model, water is initially formed in the atmosphere and falls as rain or snow. It then interacts with a variety of different phases (atmospheric gases, minerals, organic matter, etc.) present in soil, groundwater systems, and the tailings impoundment and disposal pond. Several trends or cycles in the major element composition of water are observed as water moves through the hydrologic system of the mine:

1. Oxygen – Meteoric water is fully oxygenated but as the water moves through the soil profile oxygen is consumed by microbes and the decomposition of organic matter. In the bedrock aquifer, oxygen is further consumed by sulfide minerals. Upon entering the mine voids water is once again exposed to atmospheric oxygen and becomes more oxygenated; this process of oxygenation continues as water is transported to the decant pond and infiltrates to the groundwater system. The formation of oxide

**TABLE 7. PRINCIPLE COMPONENTS OF THE PROCESS MODEL
FOR GEOCHEMICAL EQUILIBRIUM MODELING**

Water Type	General Water Quality	Principle Phases in Contact with Water	Principle Changes/Reactions Occurring
Meteoric Water	Dilute, slightly acidic, oxygenated rainwater	Atmospheric gases (O ₂ , CO ₂ , SO _x , NO _x) and particulates	Equilibrium with atmospheric gases
Soil Water	Less dilute, less oxygenated, high pCO ₂	Soil gases, minerals, organic matter	Equilibrium with soil gases Consumption of oxygen by microbes/organic matter Dissolution of soil minerals Ion exchange
Groundwater	Less dilute, less oxygenated, high pCO ₂ , some metals	Aquifer minerals (primarily bedrock)	Partial equilibrium with minerals Sulfide mineral oxidation
Mine Water	More oxygenated, higher metals	Atmospheric gases Waste rock and ore including sulfide minerals Explosive residues (temporary)	Partial equilibrium with atmospheric gases Partial equilibrium with waste rock and ore Sulfide mineral oxidation Dissolution of explosive residues (temporary)
Water Disposed to Impoundment surface	More oxygenated, higher silica	Atmospheric gases Tailings Soil (after reclamation) Organic material	Partial equilibrium with atmospheric gases Evaporation Dilution by rain/snowmelt Partial equilibrium with tailings
Decant/Disposal Pond Water	More oxygenated, decreased metals	Atmospheric gases Sediments Tailings	Partial equilibrium with atmospheric gases Evaporation Dilution by rain/snowmelt Partial equilibrium with tailings/sediment Precipitation of metallic minerals
Groundwater	More oxygenated, decreased metals	Aquifer minerals (primarily unconsolidated glacial and fluvial deposits)	Precipitation of metallic minerals Adsorption of metals to aquifer sediments

minerals is favored by the oxidizing nature of the water. Isolated from the atmosphere once again, the water becomes less oxygenated as it infiltrates to groundwater.

2. Carbon dioxide/carbonate – Meteoric water is in equilibrium with atmospheric CO₂ but contains relatively low amounts of carbonate. As water moves through the soil profile, microbes and the decomposition of organic matter generate CO₂ and reduce pH, water dissolves carbonate from rocks and soil particles and the partial pressure of CO₂ (pCO₂) and carbonate content rise. In the bedrock aquifer, carbonate content may increase further as carbonate minerals are contacted and dissolved. Upon entering the mine voids, water is once again exposed to the lower atmospheric pressure of CO₂, some degassing of CO₂ and reduction in carbonate alkalinity occurs, and pH may rise (depending on how much acid is contributed by sulfide mineral oxidation). As water is disposed at the decant pond and infiltrates sediments, carbonate content may rise again from contact with additional carbonate minerals in tailings and sediments and from decomposition of organic matter in sediments.
3. Silica – Silicate minerals are the major rock forming minerals. Silicate minerals generally have low solubilities and in addition the rate of silicate dissolution is slow. The silica content of meteoric water is initially low. As water moves through the soil, bedrock aquifer, and mine voids, the silica concentration slowly increases due to the dissolution of silicate minerals. This process continues as water contacts the tailings (predominately finely ground quartz silica) and moves through the decant pond sediments and groundwater system.

5.2 GEOCHEMICAL MODELING SOFTWARE AND THERMODYNAMIC DATABASE

Several computer software packages are available that are accepted and approved by the scientific community and that are commonly used for geochemical modeling (e.g. PHREEQC, WATEQ, and Geochemist's Workbench). Each of these software packages has its own nuances, strengths and weaknesses of use and the choice of the most appropriate software depends on the purpose of the model and to some extent personal preference of the modeler. For the purposes of this assessment, Geochemist's Workbench was selected due to

its ease of use and ability to generate a variety of visual aids (e.g. activity diagrams, solubility diagrams, Eh – pH diagrams) that are useful in interpreting and describing geochemical conditions.

Geochemist's Workbench includes several thermodynamic databases that can be employed, including Lawrence Livermore National Laboratory (LLNL) database, the PHREEQC database, the WATEQF database, and an updated version 8, release 6 LLNL database. All of these databases are generally accepted and approved by the scientific community and are commonly used for geochemical modeling. The choice of database for this assessment was based on the inclusion of data for the constituents of concern (arsenic, antimony, cadmium, copper, lead, uranium, and zinc) in the databases. The updated version 8, release 6 LLNL database (LLNL, 2006) is found to be the most complete and therefore is used in all modeling described in this report.

5.3 REPRESENTATIVE DISPOSAL WATER (MINE WATER AND DECANT POND WATER)

After mine closure, disposal water would be composed exclusively of mine water plus rainfall and snowmelt that enters the reclaimed tailing impoundment infiltration area. Because the mine is not yet permanently closed, post-closure mine water quality can not be directly measured but instead must be estimated based on past and present quality of mine water and decant pond water. Waters that are believed to be most representative of post-closure are mine water and decant pond during the interim shutdown period of 1993 through 2004 and recent (2009) mine and decant pond water, with some exceptions (i.e., nitrogen levels after closure will be much lower than during operations). However, the majority of water quality analyses conducted on mine water, decant pond water and groundwater did not contain a complete list of analytes sufficient to support geochemical modeling. Therefore, water samples used for geochemical modeling are limited to samples collected in 2003 during the investigation of attenuation mechanisms described in Land and Water Consulting (2004) and samples collected in 2009 specifically for geochemical modeling. Sampling summaries and lab reports for these samples are included in Appendix A.

5.4 REPRESENTATIVE REDUCTION-OXIDATION CONDITIONS

Reduction-oxidation conditions (commonly referred to as “redox”) control many of the primary equilibria or relationships that characterize water and affect the attenuation of metals. For instance, redox conditions control whether an aqueous species (such as Fe^{++} or Cu^+) is stable and soluble in water or whether the aqueous species is unstable and will react to form a different aqueous species (e.g. Fe^{+++} or Cu^{++}) or a solid mineral phase (such as hematite or cuprite). Some metals, such as arsenic, may exhibit different adsorption behavior depending on whether the aqueous species is more or less oxidized. Redox conditions are typically defined in terms of potential (i.e., Eh measured in millivolts). Redox relationships involve at least two aqueous species or phases (e.g. a gas or mineral) in which one species is more reduced, one species is more oxidized, and electrons are exchanged. The two related species are termed a “redox pair.” Common redox relationships or pairs in water include:

- Ammonia (NH_3^0) and nitrate (NO_3^{-2});
- Ferrous iron (Fe^{++}) and ferric hydroxide ($\text{Fe}(\text{OH})_3$) iron; and
- Sulfide (S^-) and sulfate (SO_4^{-2}).

Unfortunately redox conditions in water are difficult to assess or define and difficult to model as very few waters are actually in complete redox equilibrium where all redox pairs demonstrate equilibria with the same redox condition or Eh value. Lindberg and Runnells (1984) recommend evaluating all of the redox couples within a solution to determine if the system is in redox equilibrium. The definition of redox conditions for geochemical modeling in this assessment used a weight-of-evidence approach as follows:

- Step 1 – Field measurements of ORP (oxidation-reduction potential made with a platinum electrode) were made at the time of sample collection and converted to Eh values (adjustment of potential to be relative to the standard hydrogen electrode and corrected for temperature).
- Step 2 – Eh values were calculated based on measured concentrations of all available redox pairs:

- Assume measured concentration defines overall or bulk concentration of redox pair species.
- Calculate activities of redox pair species using Geochemist's Workbench software.
- Calculate Eh values for redox pairs using Geochemist's Workbench software.
- Step 3 – Consider chemical indicators of redox conditions and processes.
- Step 4 – Select Eh value for modeling:
 - Select system Eh value based on weight-of-evidence from Steps 1 through 3.
 - Allow selected redox pairs to remain in disequilibria.
- Step 5 – Construct Eh –pH diagrams to show metal solubility under a range of redox conditions.

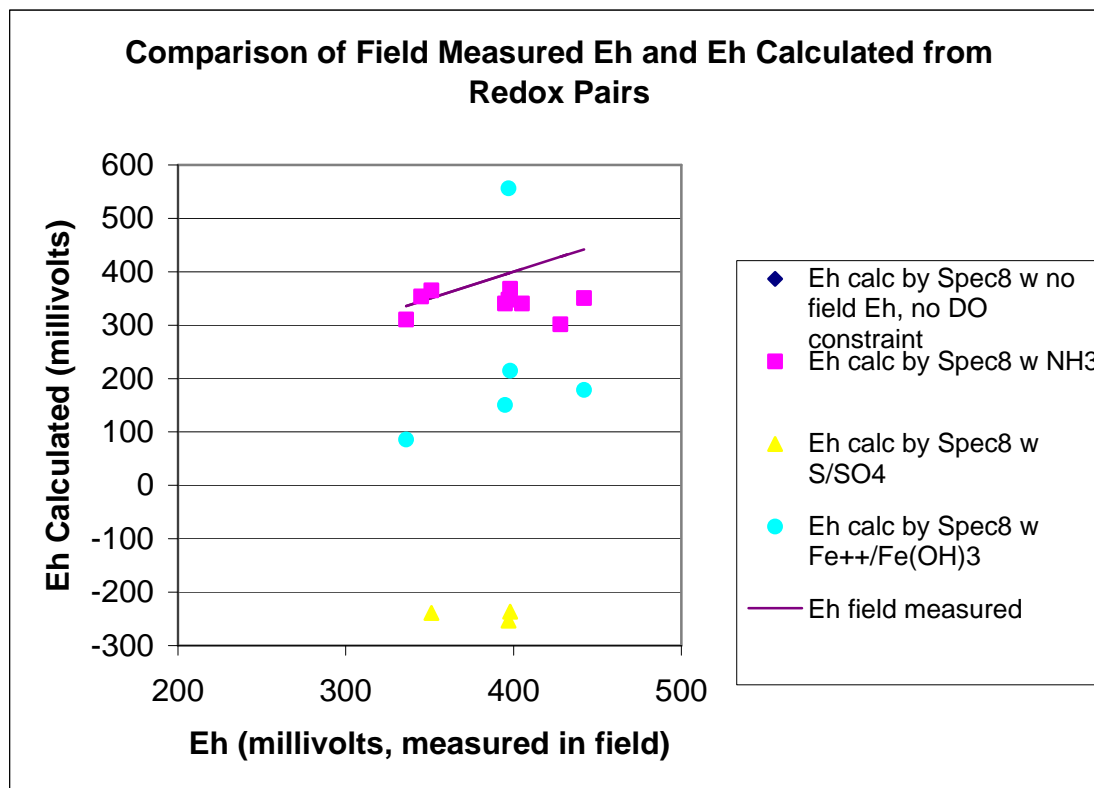
Table 8 summarizes field measurements and calculated Eh values for the ammonia/nitrate, sulfide/sulfate, and ferrous iron/ferric hydroxide redox pairs for the various water samples. As shown in Figure 3, field measured Eh and calculated Eh for sulfide/sulfate and ferrous/ferric hydroxide do not correlate well, indicating that the redox pairs are in disequilibrium. Field measured Eh values correlate fairly well with Eh calculated based on total ammonia and nitrate+nitrite concentrations. Although it is tempting to view this correlation as proof that field and ammonia/nitrate Eh values are the most correct, this correlation may simply reflect the fact that the platinum electrode used in the field measurements is more responsive to nitrogen species than to iron and sulfur species (perhaps due to the relatively high concentrations of the nitrogen species compared to sulfur and iron species). It should also be noted that the concentrations of iron and sulfide are generally low, just slightly higher than the analytical detection limits and may be subject to a relatively high degree of uncertainty; thus Eh calculated from iron and sulfur redox pairs may also be subject to a relatively high degree of uncertainty.

**TABLE 8. COMPARISON OF FIELD MEASURED EH
AND EH CALCULATED FROM REDOX PAIRS**

	Ferrous Iron (mg/L)	Diss. Oxygen (mg/L)	Sulfide (mg/L)	Nitrate (mg/L)	Eh field (milliV)	Eh calc. from redox pairs (millivolts)		
						NH_3/NO_3	S^2/SO_4	$\text{Fe}^{2+}/\text{Fe}(\text{OH})_3$
Mine Water Adit Ditch May 2009	ND	10.75	ND	14.25	405	340.4		
Mine Water Adit Pipe May 2009	0.05	11.11	ND	35.37	395	340.6		150.7
MW-01-15 May 2009	0.06	5.17	0.05	74.37	398	368.4	-236.2	214.6
Decant Pond May 2009	0.06	8.19	0.04	77.91	397	348	-253.1	556
MW-01-15 June 2009	ND	4.35	0.04	64.19	351	365	-238.7	
Decant Pond June 2009	0.05	6.86	ND	64.19	336	310.8		85.7
MW-95-8 June 2009		4.52	ND	108.5	345	354		
MW-95-8 July 2009		5.37	ND	97.83	432			
Decant Pond July 2009		6.26	ND	60.2	428	301.6		
MW-01-15 July 2009	0.06	3.34		62.42	442	350.6		178.8
Mine Water Pipe April 2003			Na	4.869				
Mine Water Ditch April 2003			Na	0.5755				
Transfer Point April 2003			Na					
Mine Water 9 East April 2003			Na	1.726		353.8		
Mine Water UQ-1 April 2003			Na	1.815		347		

ND = not detected.

**FIGURE 3. COMPARISON OF FIELD MEASURED EH
AND EH CALCULATED FROM REDOX PAIRS**



Chemical indicators of redox conditions (Table 8) include dissolved oxygen concentrations, presence/absence of sulfide, ammonia, and dissolved iron. McMahon and Chappelle (2008) propose that the dominant redox processes in groundwater can be classified based on dissolved oxygen (DO), ferrous iron, manganese, nitrate, sulfate, and sulfide concentrations. In this classification, all waters containing greater than 0.5 mg/L DO and less than 0.1 mg/L iron are considered to be oxidic, with oxygen as the principle electron acceptor. Classification as oxidic is also supported by comparison of the low concentrations of the reduced species sulfide, ammonia and ferrous iron. Westall (2002) notes that in the presence of dissolved oxygen, field Pt electrode measured Eh is typically high, in the range of 300 to 400 millivolts.

Although the data are not conclusive given the redox disequilibrium, the redox evidence points to a fairly high Eh indicative of an oxic environment. The field Eh measurements and ammonia/nitrate redox pairs both satisfy this requirement and are in good agreement. However, some water samples did not contain measurable amounts of ammonia. Therefore, field Eh was selected to define system Eh for purposes of modeling. For samples from 2003 in which field Eh is not available, Eh calculated from the ammonia/nitrate redox pair is assumed.

Given that none of the redox pairs exactly match the field Eh, the imposition of this Eh condition during modeling has the potential effect of causing chemical concentrations (and activities) to shift away from the measured concentrations toward the concentrations that are calculated to be in equilibrium with the imposed Eh. For instance, if a water sample has a field Eh of 400 millivolts and a calculated Eh based on the ammonia/nitrate pair of 300 millivolts then the ratio of ammonia and nitrate concentrations is not in equilibrium with the Eh of the system as defined in the model. If the model is configured to equilibrate aqueous redox species, then the activities (and concentrations) of ammonia and nitrate will be adjusted according to the redox reaction so that a portion of the ammonia is converted to nitrate and the calculated Eh of the ammonia/nitrate redox pair matches that of the system Eh, in this case field Eh. If this adjustment is not desired, the model may be configured to “decouple” the ammonia/nitrate redox pair from the system Eh and the redox disequilibria of the pair as compared to the system Eh will be retained.

The choice of whether to allow redox pairs to be “coupled” to the system Eh so that concentrations are in equilibria with system Eh or whether to decouple redox pairs to maintain the measured species concentrations depends on both the available data and the purpose of the modeling exercise. In this assessment, all metal redox pairs were coupled to system Eh (as defined by field Eh) for the following reasons:

- All metal concentrations are for bulk concentrations of the elements (e.g. iron as Fe, copper as Cu) and the concentrations of the aqueous ions (e.g. Fe^{2+} , Fe^{3+} , Cu^+ , Cu^{++}) are unknown. Coupling of the redox pairs is necessary to allow calculation of the

concentrations (and activities) of aqueous metal species and the saturation indices of metallic minerals (see discussion in Section 5.5) below.

- One of the purposes of modeling is to identify attenuation reactions that are occurring or are likely to occur in the groundwater. To do this it is necessary to identify the saturation state of the various metallic minerals and in many cases mineral stabilities are dependent on redox conditions.

Other redox pairs, such as ammonia/nitrate and sulfide/sulfate were not coupled with system Eh. The potential or Eh of ammonia/nitrate pair is in close agreement with the selected system Eh (field Eh) (see Figure 5) and so little would be gained by this coupling. The sulfide/sulfate redox pair was not coupled because given the high system Eh, essentially all of the sulfide would be converted to sulfate. Although this would have little effect on sulfate concentrations since sulfide concentrations are several orders of magnitude less than sulfate, it would greatly affect the potential stability of sulfide minerals. In order to allow the assessment of the potential for sulfide minerals to precipitate from waters, the sulfide/sulfate pair was not coupled, allowing sulfide to remain in solution in the model runs.

5.5 AQUEOUS SPECIATION AND MINERAL SATURATION INDICES

All laboratory water analyses report constituent concentrations in terms of the bulk concentrations (mass/unit volume) of the constituent element (e.g. 1 mg/L copper as copper). In actuality, elements are present in water in a variety of ionic forms or “aqueous species” (e.g. Cu^{++} , Cu^+) and ion complexes (e.g. $\text{Cu}_2(\text{OH})_2^{2+}$). Moreover, the equilibria of chemical reactions such as mineral precipitation, and therefore mineral saturation indices, are dependent on the activity of constituents rather than on concentration. Activity is defined as:

$$\text{Activity} = [\text{concentration} / \text{molecular weight}] \times \text{activity coefficient}$$

Where the activity coefficient is dependent on the amounts and types of other ionic species in water (i.e., the ionic strength of the solution).

The speciation of the various chemical constituents, the activities of all aqueous species, and mineral saturation indices for all mine water, decant pond, and groundwater samples were calculated using the Spec8 module of the Geochemist's Workbench program as follows:

- All model runs were constrained by field Eh or Eh calculated from the ammonia/nitrate redox pair if field Eh was not available.
- All redox pairs that include metals of interest (i.e., arsenic, copper, iron, manganese[DP1]) were coupled to system Eh.
- No minerals or gases were allowed to dissolve, precipitate, or exsolve (i.e., the system is closed).
- The original charge imbalance in the water analysis was retained (not balanced).

The primary aqueous species for the metals of concern are summarized in Table 9. Mineral saturation indices for saturated phases and select unsaturated phases containing metals of concern are summarized in Table 10. Complete outputs of the model runs are included in Appendix C.

Aqueous metal speciation is similar for all water types and is summarized as follows:

- Divalent metals cadmium, copper, lead, iron and manganese are predominately present as simple divalent cations (Cd^{+2} , Cu^{+2} , Pb^{+2} , Fe^{+2} , Mn^{+2}) and are not complexed to any significant extent by anionic species.
- Arsenic is present predominately as an oxide (monoprotic arsenic acid).
- Antimony is present predominately as a hydroxide complex.
- Uranium is present as a carbonate complex.

A "saturated" mineral phase is a mineral that is in equilibrium with water where the mineral ion activity product (IAP) is equal to the mineral solubility equilibrium constant (K). Mineral saturation index (SI) is defined as $\log \text{IAP/K}$. Oversaturated minerals are favored to precipitate or form from solution and have positive SI values (SI greater than zero). Undersaturated minerals are favored to be dissolved into solution, and have negative SI

TABLE 9. PRIMARY AQUEOUS SPECIES OF METALS

	Arsenic	Antimony	Cadmium	Copper	Lead	Iron	Manganese	Uranium
Mine Water Adit Ditch May 2009	HAsO_4^{-2}	Sb(OH)_3	Cd^{++}	Cu^{++}	Pb^{++}	Not detected	Mn^{++}	$\text{UO}_2(\text{CO}_3)_2^{-2}$
Mine Water Adit Pipe May 2009	HAsO_4^{-2}	Sb(OH)_3	Cd^{++}	Cu^{++}	Pb^{++}	Fe^{+2}	Mn^{++}	$\text{UO}_2(\text{CO}_3)_3^{-4}$
MW-01-15 May 2009	Not detected	Sb(OH)_3	Cd^{++}	Cu^{++}	Pb^{++}	Fe^{+2}	Mn^{++}	$\text{UO}_2(\text{CO}_3)_2^{-2}$
Decant Pond May 2009	HAsO_4^{-2}	Sb(OH)_3	Cd^{++}	Cu^{++}	Pb^{++}	Fe^{+2}	Mn^{++}	$\text{UO}_2(\text{CO}_3)_2^{-2}$
MW-01-15 June 2009	Not detected	Sb(OH)_3	Not detected		Not detected	Not detected	Mn^{++}	$\text{UO}_2(\text{CO}_3)_2^{-2}$
Decant Pond June 2009	HAsO_4^{-2}	Sb(OH)_3	Cd^{++}	Cu^{++}	Pb^{++}	Fe^{+2}	Mn^{++}	$\text{UO}_2(\text{CO}_3)_3^{-4}$
MW-95-8 June 2009	Not detected	Not detected	Cd^{++}	Cu^{++}	Not detected	Not detected	Not detected	$\text{UO}_2(\text{CO}_3)_3^{-4}$
MW-95-8 July 2009	Not detected	Not detected	Not detected	Cu^{++}	Not detected	Not detected	Not detected	$\text{UO}_2(\text{CO}_3)_3^{-4}$
Decant Pond July 2009	HAsO_4^{-2}	Sb(OH)_3	Cd^{++}	Cu^{++}	Not detected	Not detected	Mn^{++}	$\text{UO}_2(\text{CO}_3)_2^{-2}$
MW-01-15 July 2009	Not detected	Sb(OH)_3	Not detected	Cu^{++}	Not detected	Fe^{+2}	Mn^{++}	$\text{UO}_2(\text{CO}_3)_2^{-2}$
Mine Water 9 East April 2003	HAsO_4^{-2}	Not detected	Not analyzed	$\text{CuCO}_3(\text{aq})$	Not detected	Not detected	Mn^{++}	Not analyzed
Mine Water UQ-1 April 2003	Not detected	Not detected	Not analyzed	$\text{CuCO}_3(\text{aq})$	Not detected	Not detected	Mn^{++}	Not analyzed

TABLE 10. MINERALS SATURATED IN MINE WATER AND GROUNDWATER

	Antimony Cadmium	Copper	Lead	Iron	Manganese
Mine Water Adit Ditch May 2009	None	Brochantite Malachite Tenorite (Diopside)	Cerussite	None	(Rhodochrosite)
Mine Water Adit Pipe May 2009	None	Ferrite-Cu Malachite Tenorite Delafossite Brochantite (Diopside)	Cerussite	Hematite Ferrite-Cu Magnetite Delafossite Goethite Jarosite Fe(OH) ₃	(Rhodochrosite)
MW-01-15 May 2009	None	Ferrite-Cu Malachite Tenorite Delafossite	(Cerussite)	Hematite Ferrite-Cu Magnetite Delafossite Goethite Jarosite Fe(OH) ₃	(Rhodochrosite)
Decant Pond May 2009	None	Ferrite-Cu Malachite Tenorite Delafossite (Brochantite)	Cerussite	Hematite Ferrite-Cu Magnetite Delafossite Goethite Jarosite Fe(OH) ₃	Rhodochrosite
MW-01-15 June 2009	None	(Tenorite) (Malachite)	None	None	(Rhodochrosite)
Decant Pond June 2009		Brochantite Ferrite-Cu Malachite Tenorite Delafossite (Diopside)	Cerussite	Hematite Ferrite-Cu Magnetite Delafossite Goethite Jarosite Fe(OH) ₃	Rhodochrosite
MW-95-8 June 2009	None	Tenorite (Malachite)	None	None	None
MW-95-8 July 2009	None	Tenorite (Malachite)	None	None	None
Decant Pond July 2009	None	Brochantite Malachite Tenorite (Diopside)	None	None	Rhodochrosite (Hausmannite) (Pyrolusite)
MW-01-15 July 2009	None	(Tenorite) (Malachite)	None	None	(Rhodochrosite)
Mine Water 9 East April 2003	None	(Tenorite) (Malachite)	None	None	None
Mine Water UQ-1 April 2003	None	(Tenorite)	None	None	None

*Minerals in parentheses are near saturation (SI > approximately -1).

values. However, it is important to note that the calculated mineral saturation indices incorporate all of the uncertainties or errors associated with field collection of water samples and measurement of redox conditions, laboratory measurement of chemical concentrations of the metals and other elements, and thermodynamic data. Thus, the interpretation of saturation indices should consider the likelihood that some truly saturated minerals may have a calculated SI of <0 and conversely some truly unsaturated minerals may have a calculated SI of >0 .

Mineral phases that are calculated by Geochemists Workbench to be saturated ($SI>0$) are summarized in Table 10 and described as follows:

- Arsenic, antimony, cadmium, uranium - no phases saturated.
- Copper – oxide, hydroxysulfate, and hydroxycarbonate minerals saturated:
 - Brochantite: $\text{Cu}_4\text{SO}_4(\text{OH})_6$
 - Malachite: $\text{Cu}_2(\text{OH})_2(\text{CO}_3)_2$; presence of this mineral in decant pond sediments was confirmed by microscopic examination (Cannon Microprobe, 2003)
 - Tenorite: CuO
 - Delafossite: CuFeO_2 ; presence of this mineral in decant pond sediments was confirmed by microscopic examination (Cannon Microprobe, 2003)
 - Ferrite-Cu: CuFe_2O_4
- Iron – oxide minerals saturated:
 - Hematite: Fe_2O_3
 - Ferrite-Cu: CuFe_2O_4
 - Magnetite: Fe_3O_4
 - Delafossite: CuFeO_2
 - Jarosite: $\text{KFe}^{3+}_3(\text{OH})_6(\text{SO}_4)_2$
 - Iron hydroxide: $\text{Fe}(\text{OH})_3$
 - Goethite: $\text{FeO}(\text{OH})$

- Lead – carbonate mineral saturated:
 - Cerussite: PbCO_3
- Manganese – carbonate mineral saturated:
 - Rhodochrosite: MnCO_3

In addition to these phases that are calculated to be saturated, other important phases that are calculated to be undersaturated but near saturation (SI greater than approximately -1) include:

- Copper:
 - Dioptase: $\text{CuSiO}_2(\text{OH})_2$; the occurrence of a hydrated copper silicate mineral in decant pond sediments was confirmed by microscopic and microprobe examination by (Cannon Microprobe, 2003). The phase was tentatively identified as chrysocolla ($\text{CuSiO}_3(\text{H}_2\text{O})_n$), a similar hydrated copper silicate.
- Manganese:
 - Hausmannite: Mn_3O_4
 - Pyrolusite: MnO_2

5.6 STABILITY AND SOLUBILITY DIAGRAMS

One of the goals of this evaluation is to identify attenuation mechanisms and to assess the long-term reliability of the mechanisms. For metals that are attenuated by mineral precipitation mechanisms or are adsorbed to freshly precipitated mineral surfaces, the effectiveness of the attenuation mechanisms are controlled by the stability of the associated minerals. As described in Section 5.5, mine water, decant pond water and groundwater are calculated to be near saturation or oversaturation with respect to copper, iron, lead, and manganese mineral phases and therefore the formation or precipitation of these minerals from these waters is favored. In the case of copper and lead, the precipitation of copper and lead minerals is a direct form of attenuation (i.e., copper and lead are removed from solution by the mineral precipitation). For other metals of interest (arsenic, antimony, cadmium, and uranium), the precipitation of iron or manganese minerals provides fresh mineral surfaces upon which the metals of interest may be adsorbed. Iron and manganese oxides and

hydroxide minerals have a strong affinity for adsorption of many metals (see discussion in Section 3).

Stability and solubility diagrams provide a visual method of showing the conditions under which minerals are likely to form, the factors controlling the stability of the mineral phases, and the changes in water chemical conditions that could limit mineral precipitation or lead to dissolution of existing mineral phases. Stability diagrams and solubility diagrams in this assessment were developed for copper, iron, lead, and manganese using the Act2 module of the Geochemist's Workbench software (RockWare, 2010).

For some metals, there are multiple minerals that are oversaturated or nearly saturated in solution. Theoretically, only the most stable (most saturated) mineral would be formed if the solution were allowed to come to chemical equilibrium. However, chemical equilibrium is rare in nature and formation of the most stable phases may not be kinetically favored. Moreover, given the uncertainties in the water chemical analyses and thermodynamic data, it is possible that minerals that are not quite saturated may in actuality be stable and perhaps even favored. In order to evaluate the stability of minerals in addition to those thought to be most thermodynamically stable, some minerals were “suppressed” or purposely omitted from some stability and solubility diagrams. In these cases, the suppressed minerals are identified in the notes on the diagrams and the rationales for suppression of the minerals are discussed in the text below.

All diagrams except those containing the copper silicate mineral chrysocolla are for the average temperature (11.4 C°) in the mine water, decant pond water, and groundwater. For the copper silicate mineral chrysocolla, the only thermodynamic data available is for 25°C and the information needed to extrapolate to lower temperatures is not available; therefore the copper silicate diagrams are for 25°C.

Interpretation of the stability diagrams is illustrated by Figures 6 and 7 as follows:

- The diagrams show the Eh-pH conditions or aqueous species-pH conditions where mineral phases and aqueous species predominate for the given conditions.
- Solid mineral phases are yellow/tan in color and are labeled with the name of the mineral (e.g. tenorite).
- Aqueous phases are blue in color and are labeled with the name of the aqueous species (e.g. Cu^{++}).
- Lines on the diagrams denote phase boundaries where two or more distinct phases are in equilibrium. Lines may denote equilibrium between solid phases (e.g. the minerals tenorite and cuprite), between a solid phase and an aqueous phase (e.g. tenorite and the Cu^{++} ion at the specified aqueous activity), or between two or more aqueous phases (e.g. the redox equilibrium boundary between Cu^+ and Cu^{++}).

As an example of the interpretation of activity diagrams, consider Figures 4 and 5 that show the Eh and pH for copper at copper activities of 10^{-6} (a concentration of about 0.08 mg/L, slightly less than measured in the Adit Ditch in May 2009) and $10^{-7.5}$ (a concentration of about 0.006 mg/L, similar to the concentration routinely measured in groundwater near the decant pond). Plotted on the diagrams are the Eh-pH conditions for mine water and decant pond water as measured in May 2009. From these diagrams it is possible to make a number of important observations regarding the changes in copper activities (and concentrations) that are likely to occur to attenuate copper and conversely the geochemical conditions that might impair or preclude copper attenuation:

1. At a copper activity of 10^{-6} (Figure 4), the waters are not at equilibrium with tenorite (equilibrium is on the boundary line between tenorite and aqueous copper) but rather are oversaturated with respect to tenorite. Therefore, under these conditions tenorite would be expected to precipitate, reducing the copper activity in the water until equilibrium is achieved. If the copper activity were reduced to $10^{-7.5}$ (a concentration of about 0.006 mg/L) then the diagram would look like Figure 5. At this lower copper activity, the waters plot near the equilibrium line and no further tenorite

- precipitation is predicted. From these diagrams we would conclude that under current conditions of Eh and pH, copper is expected to be attenuated from mine and decant pond water until the copper concentration is approximately 0.006 mg/L, an observation confirmed by the concentration of copper in groundwater near the decant pond and the identification of copper oxide minerals in the decant pond sediments.
2. At the Eh typical of mine and decant pond waters (about 400 millivolts), there are only two dominant phases – the Cu^{++} ion and the mineral tenorite. The equilibrium relationship between these two phases is dependent on pH.
 - Higher pH conditions increase the stability of the solid phase (tenorite), therefore higher pH conditions will lead to more copper attenuation (lower copper concentration in water). Conversely, if conditions were to become more acidic (lower pH), copper attenuation would be less effective.
 - Therefore, the continuation or longevity of the copper attenuation mechanism is dependent on continued neutral pH conditions. If water were to become acidic, the attenuation of copper would be reduced. If this condition were to occur, an effective mitigation would be to take actions (e.g. addition of a base such as lime or limestone) to raise the pH of the water and restore the attenuation mechanism.

FIGURE 4. EH-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6})

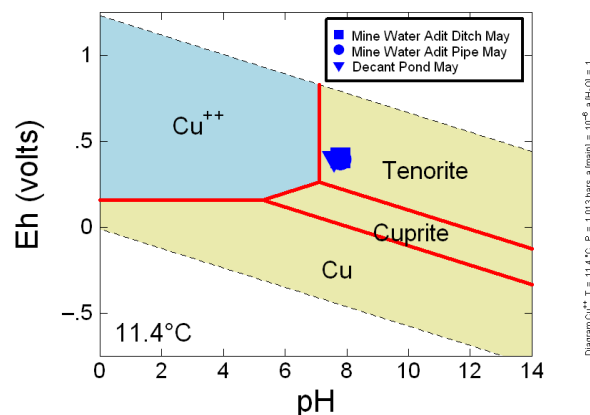
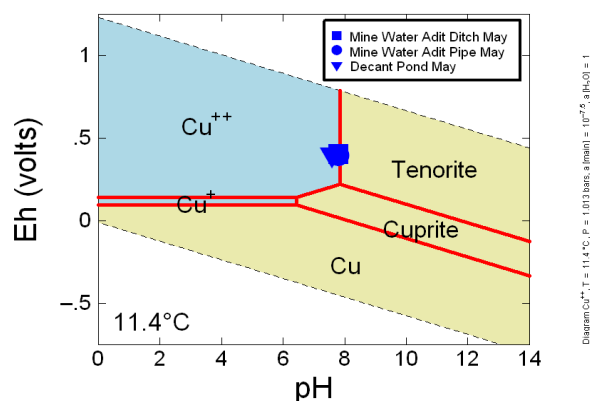


FIGURE 5. EH-PH DIAGRAM FOR COPPER ACTIVITY = $10^{-7.5}$



5.6.1 Copper Oxides and Copper Iron Oxides

Copper oxide minerals that are saturated or nearly saturated in mine water, decant water, and groundwater include tenorite (CuO), delafossite (CuFeO_2), and copper ferrite (CuFe_2O_4). These minerals are defined by a system containing copper, iron, oxygen (Eh) and hydrogen (pH). The stability and solubility of the minerals are constrained by copper and iron activity and the Eh and pH of the solution.

Eh-pH diagrams for copper for the range of copper concentrations (and activities) present in water at the Troy Mine (activity of 10^{-6} to $10^{-7.5}$) are shown in Figures 4 and 5 (above). The Eh-pH of mine water, decant pond water, and groundwater are plotted on the diagrams for reference. The majority of the mine water, decant pond water, and groundwater samples plot in the tenorite mineral stability field. If allowed to come to chemical equilibrium, tenorite would be expected to precipitate from these waters, reducing the concentration of copper in the waters. Further, these diagrams illustrate that formation of copper oxide minerals is favored by:

1. Higher pH;
2. Higher copper concentrations; and
3. Higher Eh (more oxidizing conditions).

At the neutral pH typical of mine water, decant pond water and groundwater near the decant pond, the solubility of minerals is low, regardless of Eh conditions.

Some water samples are found to contain measurable amounts of iron. The effects of the addition of iron to the copper oxide system are shown in Figures 6 and 7. In the presence of iron, the copper oxide stability fields are diminished and copper iron oxides (delafossite and ferrite copper) are the most stable mineral phases at neutral to alkaline pH conditions. As noted above, there is field evidence corroborating the precipitation of delafossite in decant pond sediments as this mineral was identified in pond sediments by microscopic and electron microprobe analyses. The majority of the mine water, decant pond water, and groundwater samples plot in the delafossite mineral stability field. If allowed to come to chemical equilibrium, delafossite would be expected to precipitate from these waters, reducing the concentration of copper in the waters. Further, these diagrams illustrate that attenuation of copper by the formation of copper iron oxide minerals is favored by:

1. Higher pH;
2. Higher iron and copper concentrations; and
3. Higher Eh (more oxidizing conditions).

FIGURE 6. EH-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6}) WITH IRON

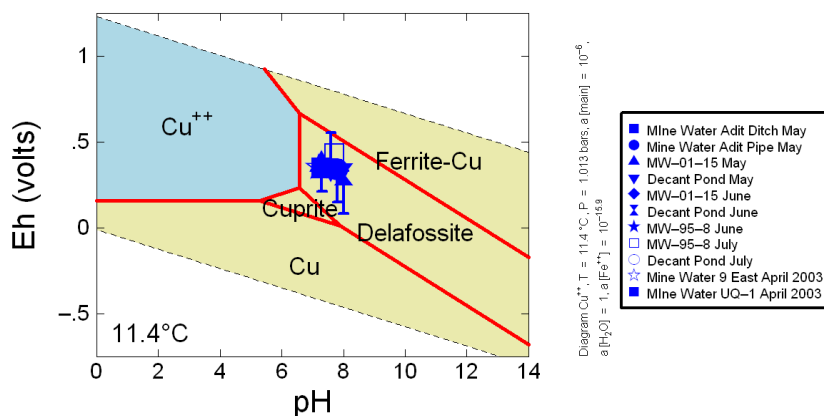
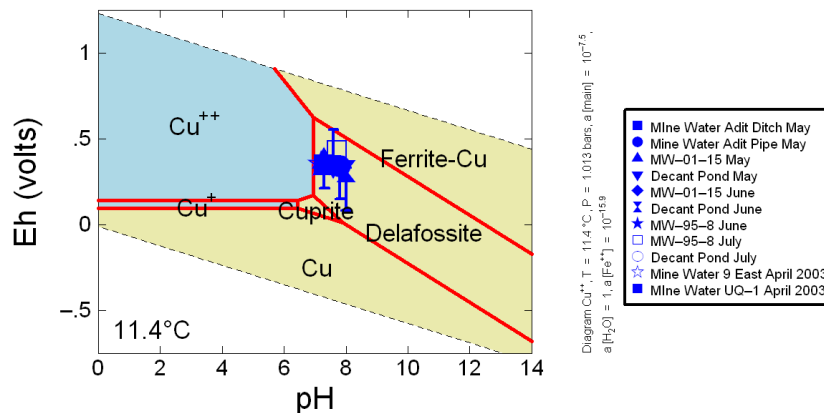


FIGURE 7. EH-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$) WITH IRON



At the neutral pH typical of mine water, decant pond water and groundwater near the decant pond, the solubility of copper iron oxide minerals is low, regardless of Eh conditions. If waters were to become acidic, the attenuation of copper by the formation of copper iron oxide minerals would be reduced. If this condition were to occur, an effective mitigation would be to take actions (e.g. addition of a base such as lime or limestone) to raise the pH of the water and restore the attenuation mechanism.

5.6.2 Copper Carbonates

Copper carbonate mineral (e.g. malachite $\text{Cu}_2(\text{OH})_2(\text{CO}_3)_2$) stability and solubility are constrained by copper and bicarbonate and pH of the solution. The presence of malachite in decant pond sediments was confirmed by microscopic examination (Cannon Microprobe, 2003). Stability diagrams for the copper bicarbonate system for the range of copper concentrations (and activities) present in water at the Troy Mine (activity of 10^{-6} to $10^{-7.5}$), average field Eh (393 millivolts) are shown in Figures 8 and 9. Because the stability and solubility of malachite is not directly dependent on Eh for the Eh range of interest, bicarbonate activity is shown on the y-axis instead of Eh. The bicarbonate activity and pH of mine water, decant pond water, and groundwater are plotted on the diagrams for reference.

These diagrams illustrate that formation of copper bicarbonate minerals is favored by:

1. Higher dissolved bicarbonate concentrations (at lower bicarbonate levels the copper oxide tenorite is stable);
2. Higher pH; and
3. Higher copper concentrations.

Attenuation of copper by the mechanism of copper carbonate mineral formation would be impaired by acidic (low pH) conditions or low dissolved bicarbonate concentrations. If waters were to become acidic, the attenuation of copper by the formation of copper carbonate minerals would be reduced. If this condition were to occur, an effective mitigation would be to take actions (e.g. addition of a base such as lime or limestone) to raise the pH and bicarbonate of the water and restore the attenuation mechanism.

FIGURE 8. BICARBONATE-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6})

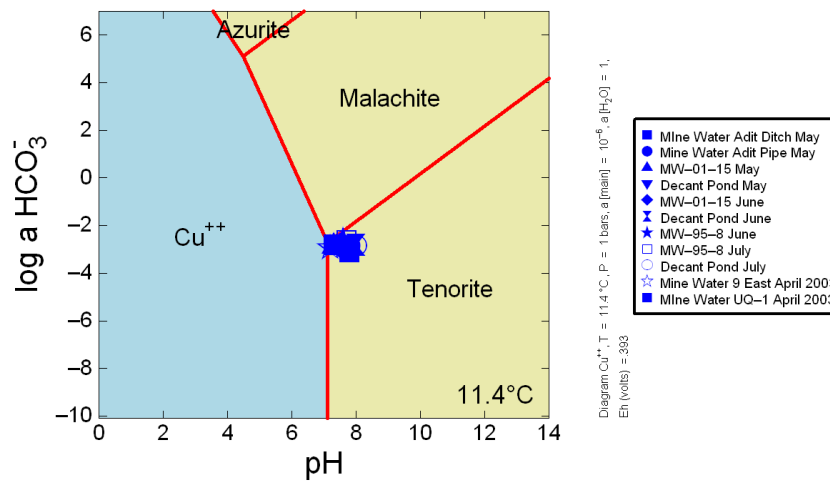
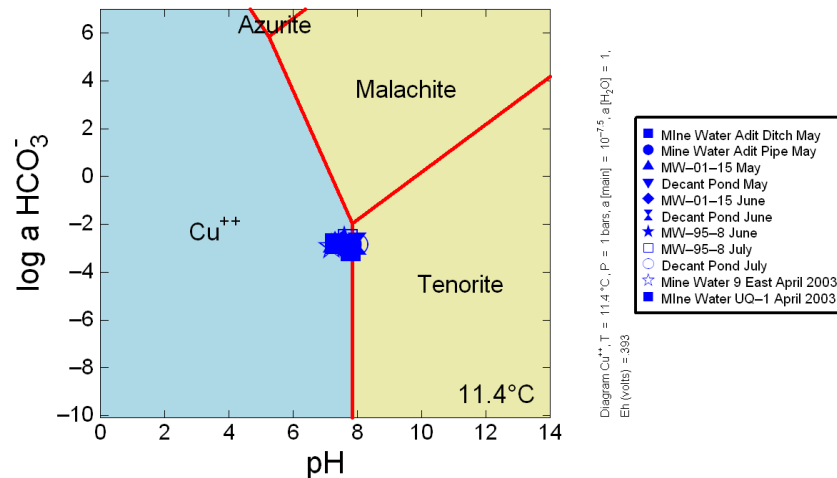


FIGURE 9. BICARBONATE-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$)

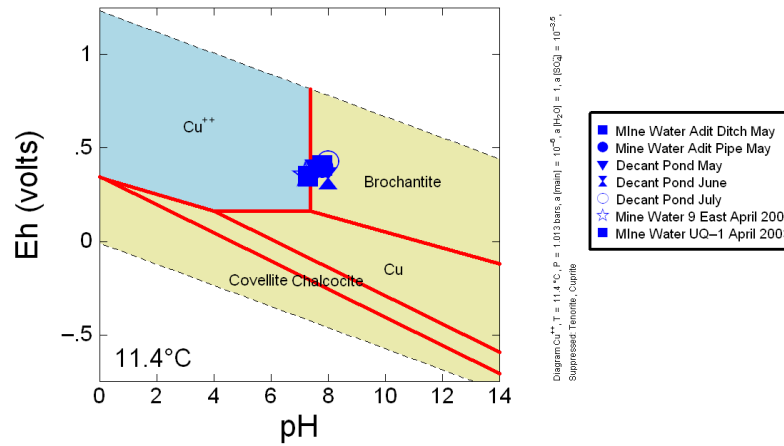


5.6.3 Copper Hydroxysulfate

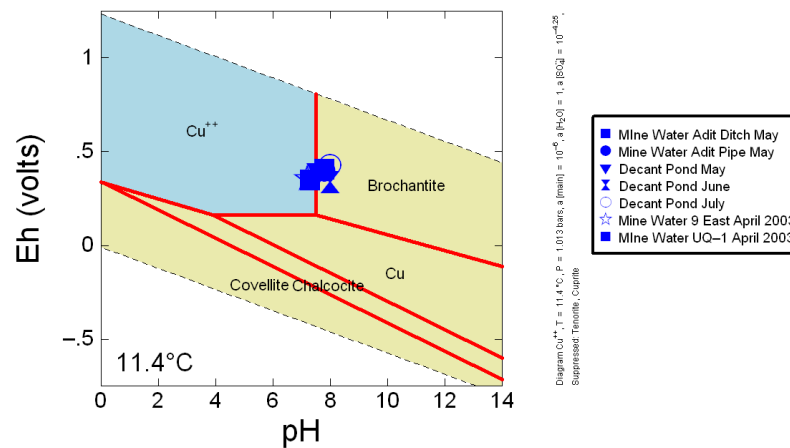
Copper hydroxysulfate (brochantite) mineral stability and solubility are constrained by copper activity, sulfate activity and the Eh-pH of the solution. Eh-pH diagrams for copper for the range of copper activities (10^{-6} to $10^{-7.5}$) and sulfate activities ($10^{-3.49}$ to $10^{-4.25}$) present in water at the Troy Mine are shown in Figures 10 through 13. The Eh-pH of mine water, decant pond water, and groundwater are plotted on the diagrams for reference. The copper oxide minerals tenorite and cuprite are suppressed in these diagrams in order to show brochantite stability. The majority of the mine water, decant pond water, and groundwater samples plot near or in the brochantite mineral stability field. If allowed to come to chemical equilibrium, brochantite would be expected to precipitate from these waters, reducing the concentration of copper in the waters. Further, these diagrams illustrate that formation of brochantite minerals is favored by:

1. Higher pH;
2. Higher sulfate and copper concentrations; and
3. Higher Eh (more oxidizing conditions).

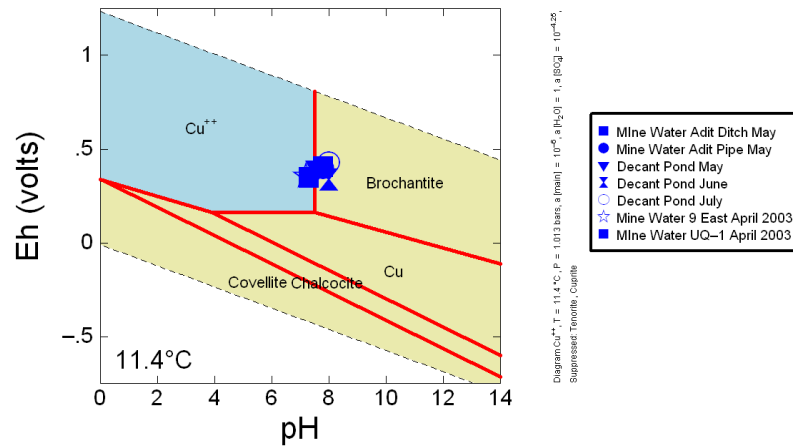
**FIGURE 10. EH-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6})
WITH SULFATE (ACTIVITY = $10^{-3.5}$)**



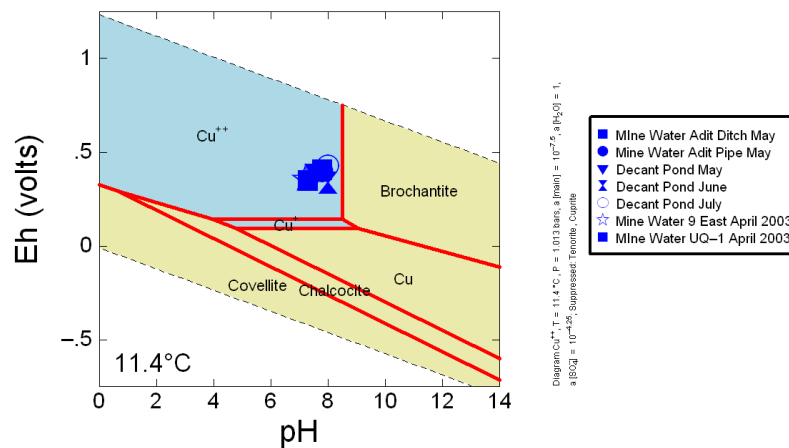
**FIGURE 11. EH-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6})
WITH SULFATE (ACTIVITY = $10^{-4.25}$)**



**FIGURE 12. EH-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$)
WITH SULFATE (ACTIVITY = $10^{-3.5}$)**



**FIGURE 13. EH-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$)
WITH SULFATE (ACTIVITY = $10^{-4.25}$)**



5.6.4 Copper Silicates

Copper silicate mineral stability and solubility are constrained by copper and silicate activity and pH of the solution. Although the copper silicate mineral diopside appears to be more saturated than chrysocolla in waters at the Troy Mine, diopside was “suppressed” and was omitted from the copper silicate stability diagrams so that the stability/solubility of

chrysocolla could be shown instead. The suppression of diopside was done for the following reasons. Field evidence (microscopic and electron microprobe evaluation of sediment pond samples, Land and Water (2004); Cannon Microprobe (2003)) indicates that the copper silicate mineral chrysocolla has precipitated from decant pond water and is a stable mineral phase whereas diopside was not found in the decant pond. Diopside is a relatively rare mineral. Some investigators have proposed that chrysocolla formation is kinetically favored over diopside and that diopside may be formed by aging of cryptocrystalline chrysocolla (Crane et al, 2001). Omission of diopside from the stability diagrams allows the stability and solubility of chrysocolla to be illustrated.

Stability diagrams for the copper-silica system for the range of copper concentrations (and activities) present in water at the Troy Mine (activity of 10^{-6} to $10^{-7.5}$), average field Eh (393 millivolts), and temperature of 25°C are shown in Figures 14 and 15. Because the stability and solubility of chrysocolla is not directly dependent on Eh, silica activity is shown on the y-axis instead of Eh. The silica activity and pH of mine water, decant pond water, and groundwater are plotted on the diagrams for reference. These diagrams illustrate that formation of copper silicate minerals is favored by:

1. Higher dissolved silica concentrations;
2. Higher pH; and
3. Higher copper concentrations.

Attenuation of copper by the mechanism of copper silicate mineral formation would be impaired by acidic (low pH) conditions or low dissolved silica conditions.

FIGURE 14. SILICA-PH DIAGRAM FOR COPPER (ACTIVITY = 10^{-6})

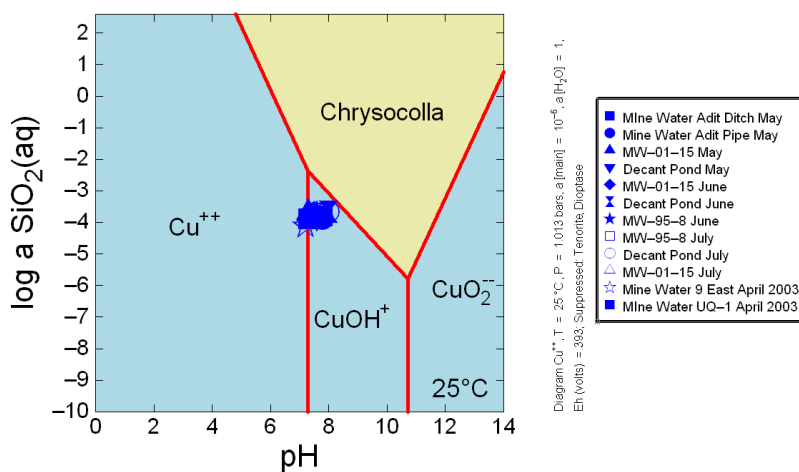
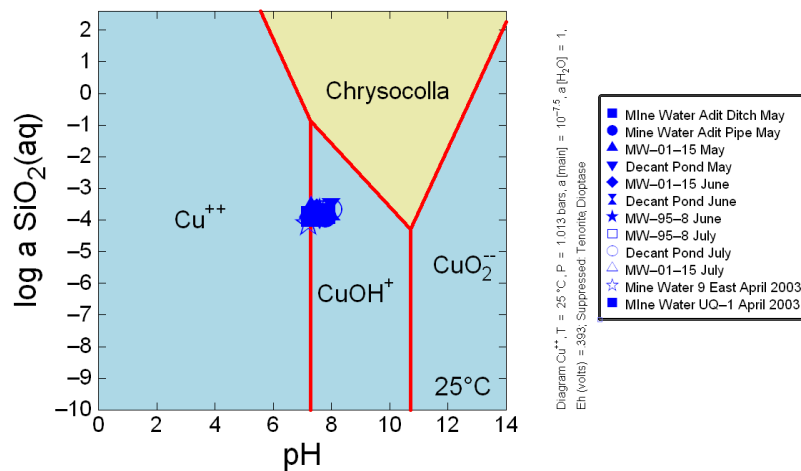


FIGURE 15. SILICA-PH DIAGRAM FOR COPPER (ACTIVITY = $10^{-7.5}$)

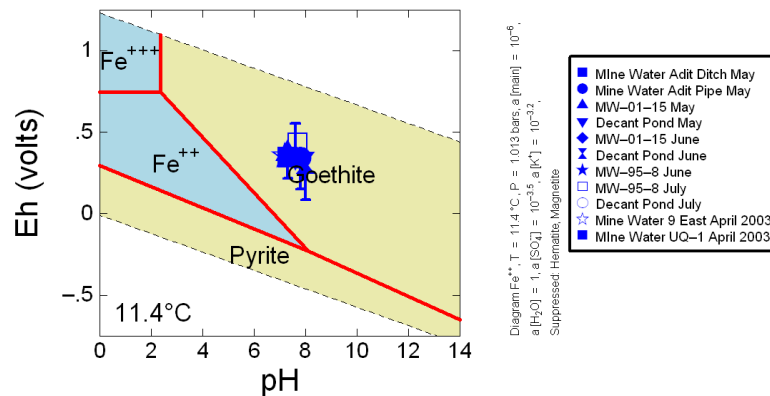


5.6.5 Iron

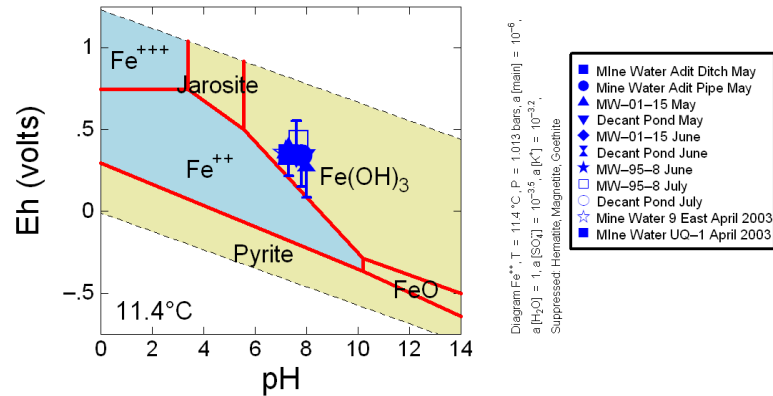
Iron minerals that are saturated or nearly saturated in mine water, decant water, and groundwater include iron oxides, copper iron oxides, iron hydroxides, and potassium iron hydroxysulfate minerals. These minerals are defined by systems containing iron, oxygen

(Eh), sulfate (sulfur), hydrogen (pH) and potassium. Copper iron oxides are discussed in Section 5.6.1 and are not further discussed in this section. Eh-pH diagrams for iron for the range of iron activities (10^{-6} to $10^{-6.2}$), sulfate ($10^{-3.5}$ to $10^{-4.25}$) and potassium ($10^{-3.2}$ to $10^{-4.3}$) activities present in water at the Troy Mine are shown in Figures 16 through 19. The Eh-pH of mine water, decant pond water, and groundwater are plotted on the diagrams for reference. The minerals hematite and magnetite were suppressed and omitted from the diagrams in order to show the more common secondary iron minerals ferric hydroxide ($\text{Fe}_3(\text{OH})_3$), goethite ($\text{FeO}(\text{OH})$), and jarosite ($\text{KFe}^{3+}_3(\text{OH})_6(\text{SO}_4)_2$). Goethite was also suppressed in some diagrams to show ferric hydroxide stability.

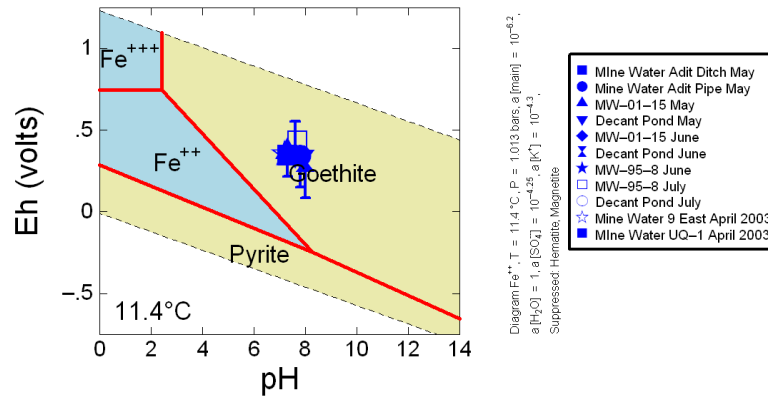
FIGURE 16. EH-PH DIAGRAM FOR IRON (ACTIVITY = 10^{-6}) WITH SULFATE (ACTIVITY = $10^{-3.5}$) AND POTASSIUM (ACTIVITY = $10^{-3.2}$)



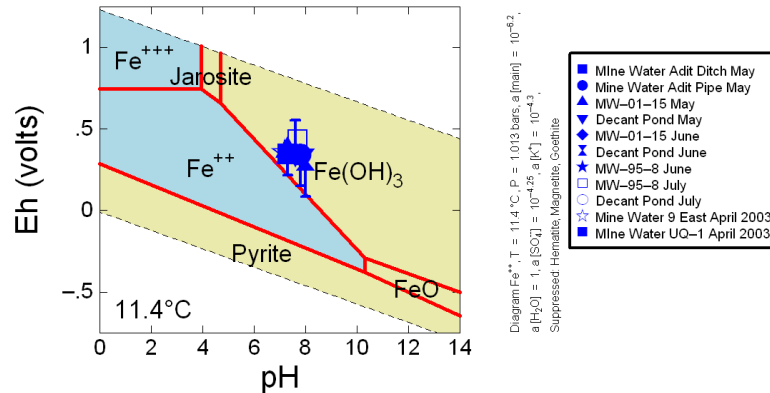
**FIGURE 17. EH-PH DIAGRAM FOR IRON (ACTIVITY = 10^{-6})
WITH SULFATE (ACTIVITY = $10^{-3.5}$) AND POTASSIUM
(ACTIVITY= $10^{-3.2}$), GOETHITE SUPPRESSED**



**FIGURE 18. EH-PH DIAGRAM FOR IRON (ACTIVITY = $10^{-6.2}$) WITH SULFATE
(ACTIVITY = $10^{-4.25}$) AND POTASSIUM (ACTIVITY= $10^{-4.3}$)**



**FIGURE 19. EH-PH DIAGRAM FOR IRON (ACTIVITY = $10^{-6.2}$)
WITH SULFATE (ACTIVITY = $10^{-4.25}$) AND POTASSIUM
(ACTIVITY = $10^{-4.3}$), GOETHITE SUPPRESSED**



These diagrams illustrate that all of the Troy Mine waters are in the goethite or ferric hydroxide stability field. Formation and stability of goethite and ferric hydroxide is favored by:

1. Higher Eh (more oxidizing);
2. Higher pH; and
3. Higher iron concentrations.

Jarosite is only stable at relatively low pH and high sulfate conditions such as might occur in acid rock drainage. Given the lack of acid rock drainage and neutral pH of water at the Troy Mine, jarosite formation is not likely to form in the decant pond or the groundwater. Attenuation of metals by co-precipitation or adsorption is favored by conditions that lead to precipitation and stability of iron oxide and hydroxide minerals. Conditions that might lead to a reduction in metal attenuation effectiveness due to the lack of stability of iron oxide or hydroxide sorbent phases include lower Eh (more reducing) and acidic (lower pH) conditions.

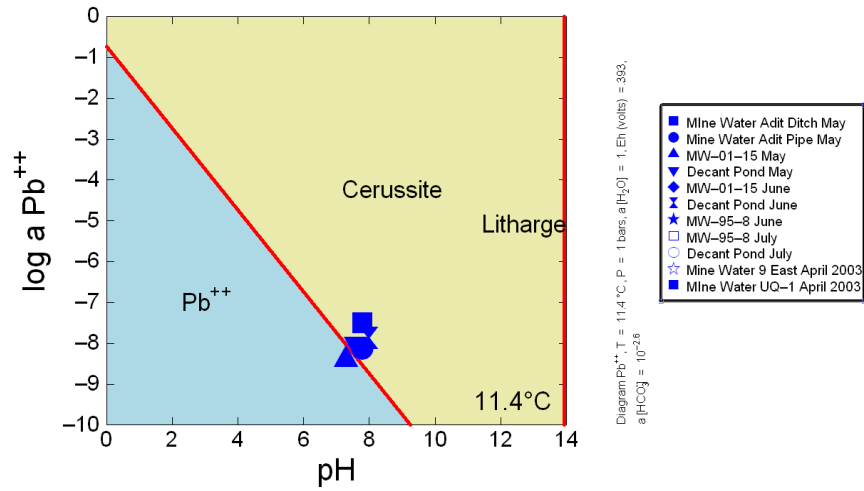
5.6.6 Lead

Cerussite (PbCO_3) is the only lead mineral that is saturated or nearly saturated in mine water, decant water, and groundwater. Stability and solubility of cerussite is defined by a system containing copper, hydrogen (pH) and carbonate. Because the stability and solubility of cerussite is not directly dependent on Eh for the Eh range of interest, bicarbonate activity is shown on the y-axis instead of Eh. Solubility diagrams showing lead solubility as a function of pH in the presence of bicarbonate for the range of lead activities ($10^{-7.5}$ to $10^{-8.4}$) and bicarbonate activities ($10^{-2.6}$ to $10^{-3.2}$) present in water at the Troy Mine are shown in Figures 20 and 21. The pH and bicarbonate concentrations of decant pond water, and groundwater are plotted on the diagrams for reference. These diagrams illustrate that formation of cerussite is favored by:

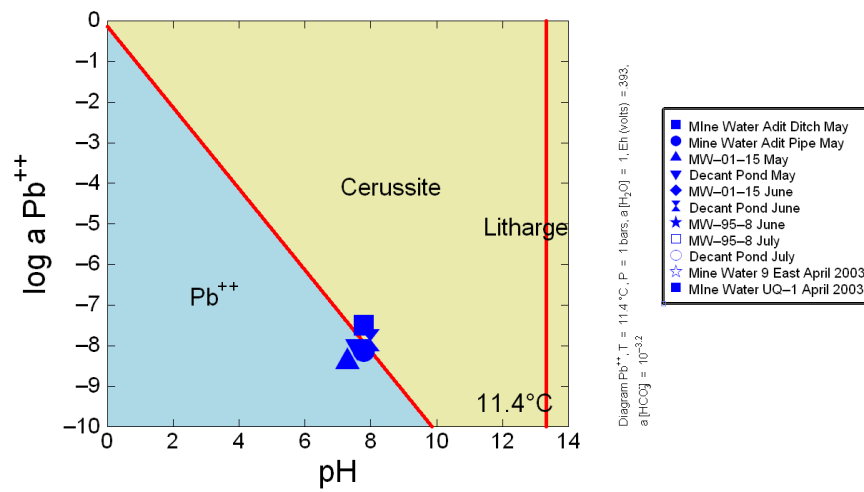
1. Higher bicarbonate silica concentrations;
2. Higher pH; and
3. Higher lead concentrations.

Attenuation of lead by the mechanism of lead carbonate mineral formation would be impaired by acidic (low pH) conditions or low dissolved bicarbonate concentrations. If waters were to become acidic, the attenuation of lead by the formation of lead carbonate minerals would be reduced. If this condition were to occur, an effective mitigation would be to take actions (e.g. addition of a base such as lime or limestone) to raise the pH and bicarbonate of the water and restore the attenuation mechanism.

**FIGURE 20. LEAD ACTIVITY-PH DIAGRAM
WITH BICARBONATE ACTIVITY = $10^{-2.6}$**



**FIGURE 21. LEAD ACTIVITY-PH DIAGRAM
WITH BICARBONATE ACTIVITY = $10^{-3.2}$**



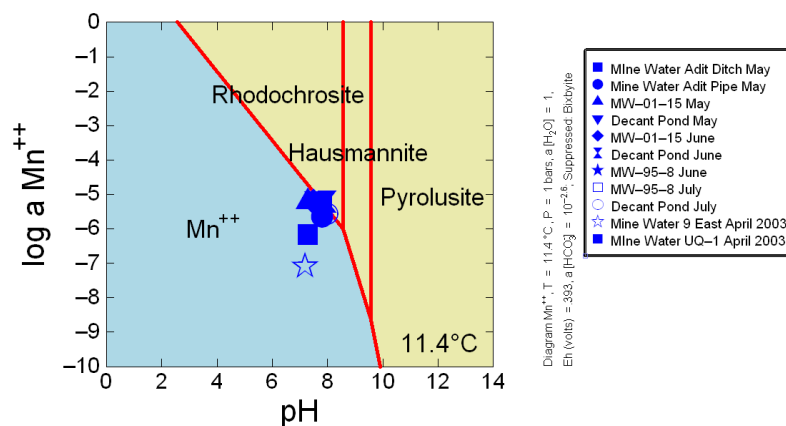
5.6.7 Manganese

Manganese minerals that are saturated or nearly saturated in mine water, decant water, and groundwater include manganese oxides (hausmannite (Mn_3O_4) and pyrolusite (MnO_2)) and the manganese carbonate mineral rhodochrosite (MnCO_3). These minerals are defined by a system containing manganese, oxygen (Eh), hydrogen (pH) and carbonate.

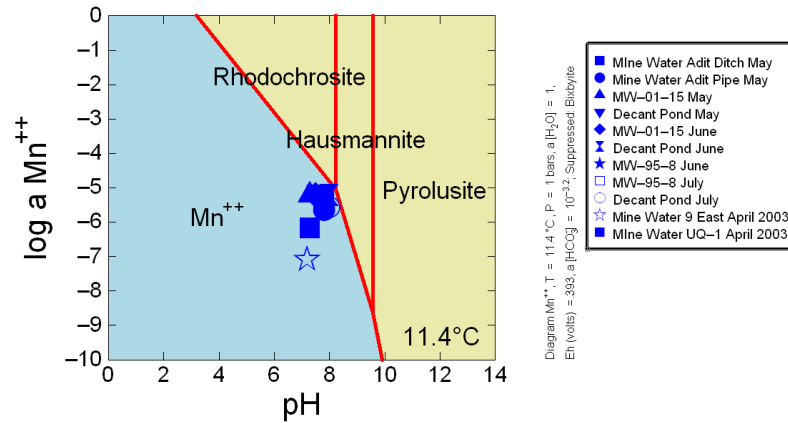
Solubility diagrams showing manganese solubility as a function of pH in the presence of bicarbonate for the range of bicarbonate activities ($10^{-2.6}$ to $10^{-3.2}$) present in water at the Troy Mine and the average field measured Eh (393 millivolts) are shown in Figures 22 and 23. Because the stability and solubility of the manganese minerals is not very sensitive to Eh for the Eh range of interest, bicarbonate activity is shown on the y-axis instead of Eh. The pH and bicarbonate concentrations of mine water, decant pond water, and groundwater are plotted on the diagrams for reference. These diagrams illustrate that formation and stability of manganese oxide minerals is favored by:

1. Higher pH; and
2. Higher manganese concentrations.

**FIGURE 22. MANGANESE-pH DIAGRAM
WITH BICARBONATE (ACTIVITY = $10^{-2.6}$)**



**FIGURE 23. MANGANESE-pH DIAGRAM
WITH BICARBONATE (ACTIVITY = $10^{-3.2}$)**



Formation and stability of manganese carbonate minerals is favored by:

1. Higher bicarbonate silica concentrations;
2. Higher pH; and
3. Higher manganese concentrations.

Attenuation of metals by co-precipitation or adsorption is favored by conditions that lead to precipitation and stability of manganese oxide and hydroxide minerals. Conditions that might lead to a reduction in metal attenuation effectiveness by this attenuation mechanism include conditions of lower Eh (more reducing), lower pH (more acidic), and lower bicarbonate concentration. If this condition were to occur, an effective mitigation would be to take actions (e.g. addition of a base such as lime or limestone) to raise the pH and bicarbonate of the water and restore the attenuation mechanism.

6.0 SUMMARY SITE MODEL FOR NATURAL ATTENUATION

The summary site model for natural attenuation at the Troy Mine is that metals are attenuated (removed) from water by a combination of mineral precipitation, co-precipitation with other metals, and adsorption by various solid materials including iron and manganese oxides and hydroxides, clays, and organic matter. Field and laboratory testing, supported by geochemical equilibrium modeling, evidence the removal of copper and lead from water (i.e., attenuation) by precipitation of copper and lead minerals in sediments beneath the decant pond. Table 11 summarizes the minerals formed and the geochemical controls on the stability and solubility of copper and lead minerals.

Field and laboratory testing and geochemical modeling also evidence the removal of arsenic, antimony, copper, lead and uranium from water (i.e., attenuation) by co-precipitation or adsorption of these metals with iron and/or manganese minerals. Table 12 summarizes the geochemical controls on the stability of iron and manganese minerals. The extent of sorption of trace metals in groundwater is largely controlled by the following factors:

1. Groundwater pH;
2. Metal concentration in solution;
3. The quantity of sorbent material (principally hydrous ferric and manganese oxides, organic matter, and clays) within the aquifer; and
4. The concentration of competing ions and metal-complexing ions in solution.

**TABLE 11. SUMMARY OF FACTORS AFFECTING METAL
ATTENUATION BY MINERAL PRECIPITATION**

Metal	Minerals	Geochemical Conditions Affecting Stability/Solubility	Factors Enhancing Attenuation	Factors Reducing Attenuation
Copper	Copper oxides (tenorite, ferrite-Cu, delafossite)	Redox: high Eh/oxygen preferred	Equilibration with atmospheric oxygen	Consumption of oxygen by organic material decomposition or sulfide minerals
		pH: high pH preferred	Buffering by carbonate rocks and minerals (e.g. calcite, limestone)	Sulfide mineral oxidation
	Copper carbonate (malachite)	pH: high pH preferred	Buffering by carbonate rocks and minerals (e.g. calcite, limestone)	Sulfide mineral oxidation
		Bicarbonate: high Bicarbonate preferred	Dissolution of carbonate rocks and minerals Equilibration with CO ₂ in atmosphere	
	Copper silicates (chrysocolla)	pH: high pH preferred	Buffering by carbonate rocks and minerals (e.g. calcite, limestone)	Sulfide mineral oxidation
		Silica: high silica preferred	Equilibration with silicate minerals	
	Copper hydroxysulfates (brochantite)	pH: high pH preferred	Buffering by carbonate rocks and minerals (e.g. calcite, limestone)	Sulfide mineral oxidation
		High sulfate	Sulfide mineral oxidation Dissolution of sulfate minerals (e.g. gypsum)	
Lead	Lead carbonate (cerussite)	pH: high pH preferred	Buffering by carbonate rocks and minerals (e.g. calcite, limestone)	Sulfide mineral oxidation
		Bicarbonate: high Bicarbonate preferred	Dissolution of carbonate rocks and minerals Equilibration with CO ₂ in atmosphere	

**TABLE 12. SUMMARY OF FACTORS AFFECTING METAL
ATTENUATION BY CO-PRECIPITATION OR ADSORPTION
WITH IRON AND MANGANESE MINERALS**

Metal	Minerals	Geochemical Conditions Affecting Stability/Solubility	Factors Enhancing Attenuation	Factors Reducing Attenuation
Iron	Iron oxides (Fe(OH) ₃ , goethite, ferrite- Cu, delafossite)	Redox: high Eh/oxygen preferred	Equilibration with atmospheric oxygen	Consumption of oxygen by organic material decomposition or sulfide minerals
		pH: high pH preferred	Buffering by carbonate rocks and minerals (e.g. calcite, limestone)	Sulfide mineral oxidation
Manganese	Manganese carbonate (rhodochrosite)	pH: high pH preferred	Buffering by carbonate rocks and minerals (e.g. calcite, limestone)	Sulfide mineral oxidation
		Bicarbonate: high Bicarbonate preferred	Dissolution of carbonate rocks and minerals Equilibration with CO ₂ in atmosphere	
	Manganese oxides (hausmannite, pyrolusite)	Redox: high Eh/oxygen preferred	Equilibration with atmospheric oxygen	Consumption of oxygen by organic material decomposition or sulfide minerals
		pH: high pH preferred	Buffering by carbonate rocks and minerals (e.g. calcite, limestone)	Sulfide mineral oxidation

7.0 EVALUATION OF LONGEVITY OF NATURAL ATTENUATION MECHANISMS

As described in the preceding sections, the active attenuation mechanisms for metals at the Troy tailing impoundment site include the precipitation of metallic minerals from water and the adsorption of metals onto precipitated minerals (iron and possibly manganese oxides and hydroxides) present in sediments beneath the decant pond. The expected duration of the geochemical conditions and capacity of aquifer sediment for attenuation is described in this section.

The longevity of the natural attenuation mechanisms is dependent on the duration of the geochemical conditions conducive to attenuation and to the capacity of the tailing impoundment/aquifer sediments to continue to uptake metals. The mineral precipitation and co-precipitation mechanisms are expected to last indefinitely or in perpetuity as long as geochemical conditions remain similar to current conditions. The adsorption mechanisms are conservatively estimated to last a minimum of 600 years.

7.1 ATTENUATION BY MINERAL PRECIPITATION

Although there are numerous metals and minerals of interest, the conditions necessary for continuation of attenuation are similar for most of these and are:

1. Neutral or alkaline pH;
2. Oxidizing redox conditions (moderate to high Eh);
3. Presence of moderate amounts of silica; and
4. Presence of moderate amounts of bicarbonate

An additional condition that only applies to sulfate minerals is:

1. Presence of moderate amounts of sulfate.

However, sulfate minerals are likely a minor attenuation mechanism given the generally low sulfate concentrations present and will not be discussed further. The durations or longevities of the primary attenuation conditions are presented below.

7.2 NEUTRAL TO ALKALINE PH

There is nearly 30 years of water quality data demonstrating that mine water is consistently neutral in pH (Parametrix, 2008). Additionally, acid-base accounting tests and long-term laboratory testing with humidity cells indicates that waste rock, ore, and tailings from the Troy Mine are not acid-generating but rather are acid-neutralizing (i.e., contain an excess of alkaline materials (MDEQ and USFS, 2001; Maxim, 2003)). Therefore, mine water pH is expected to remain near neutral to alkaline in perpetuity, and pH conditions are not expected to limit the longevity of attenuation mechanisms.

The pH of mine water, decant pond water, and groundwater are controlled by reactions between atmospheric gas (carbon dioxide) and the various rocks and minerals that the water contacts. As described in the Process Model for Geochemical Equilibrium Modeling (Section 5.1), meteoric water dissolves carbon dioxide from the atmosphere (forming carbonic acid) and becomes mildly acidic. Meteoric waters become neutral to alkaline from reaction with soils and rocks, particularly carbonate minerals and rocks that buffer the water. Although the oxidation of sulfide minerals in the underground mine releases small amounts of sulfuric acid to the water, the amount of acid released is small relative to the buffering capacity of the mine water, and mine water remains near neutral in pH.

7.3 OXIDIZING REDOX CONDITIONS

The dissolved oxygen content and redox condition (Eh) of mine water, decant pond water, and groundwater are controlled by reactions between atmospheric gas (oxygen), the various rocks and minerals that the water contacts, and decomposition of organic matter. As described in the Process Model for Geochemical Equilibrium Modeling (Section 5.1), meteoric water dissolves oxygen from the atmosphere and becomes well oxygenated. During infiltration/percolation of rainfall and groundwater to the mine workings the amount of oxygen in mine water is reduced, first by soil biological activity and later by reaction with

sulfide minerals present in the ore and waste rock. In the mine workings and during disposal of mine water in the decant pond, water is once again exposed to abundant atmospheric oxygen which raises the Eh or redox condition of the water and results in the formation of oxide minerals.

As the supply of oxygen in the atmosphere is essentially limitless, oxygen does not pose a limit to the longevity of the attenuation mechanisms as long as 1) the waters continue to receive ample exposure to the atmosphere; and 2) materials/chemicals other than metals do not consume excess amounts of oxygen. Other potential oxygen consumers are sulfide minerals and organic carbon such as plant material.

As described above for pH, oxidation of sulfide minerals in the mine is relatively slow and produces minor amounts of sulfate and acid. No change in the future rate of sulfide mineral oxidation is anticipated. Therefore, oxidation of sulfide minerals is not expected to affect the duration and longevity of attenuation mechanisms.

Currently the amount of organic carbon in mine water and decant pond sediments is low to moderate. The decant pond was dewatered in 2004 and pond bottom sediments were found to be primarily sand and silt with some interbedded layers of organic materials that appeared to be primarily algae and other very fine organic particles, with the occasional leaf or stick (see photos in Land and Water (2004)). Under these current conditions there is sufficient oxygen in the water and sediments to oxidize metals in the presence of this organic material and there is no predicted limit to the duration or longevity of the attenuation mechanism. However, if pond conditions were to change such that large amounts of organic matter were present, then excessive oxygen consumption by the decay of organic matter could result in redox conditions such that the precipitation of oxide and hydroxide minerals is limited (i.e., low Eh). Therefore, monitoring and management actions should include monitoring of pond sediments for redox conditions and carbon content, and if needed, removal of excessive organic material from the pond bottom.

7.3.1 Dissolved Silica

Silica is abundant in nature as silicates are the primary rock-forming minerals in the earth's crust. Silicate minerals are particularly abundant in and around the Troy Mine in the form of quartz and quartzite rocks, which are nearly pure silica. Moreover, tailings are predominately fine-grained quartz and sediments in the decant pond contain significant tailings material. Silica is dissolved in mine water by interaction of groundwater and mine water with silicate minerals. Given the overwhelming abundance of silica in general and in the Troy Mine and tailings impoundment in particular, the supply of dissolved silica is essentially limitless and the supply of silica does not pose a limit to the longevity of the attenuation mechanisms.

7.3.2 Dissolved Bicarbonate

The dissolved bicarbonate content of mine water, decant pond water, and groundwater are controlled by reactions between atmospheric gas (carbon dioxide), the various rocks and minerals that the water contacts, and decomposition of organic matter.

As described in the Process Model for Geochemical Equilibrium Modeling (Section 5.1), meteoric water dissolves carbon dioxide from the atmosphere, forming carbonic acid (H_2CO_3) and bicarbonate (HCO_3^-). During infiltration/percolation of rainfall and groundwater to the mine workings the amount of bicarbonate in mine water is increased, by soil biological activity and by dissolution of carbonate minerals in the ore and waste rock. Given the abundance of carbon dioxide gas in nature and carbonate rocks in the Troy Mine and tailings impoundment, the supply of dissolved bicarbonate is essentially limitless and bicarbonate does not pose a limit to the longevity of the attenuation mechanisms.

Bicarbonate concentrations of decant pond water have increased since the mine re-start in 2005. This increase is believed to be related to the higher carbonate mineral content of the ore zone that is currently being mined and milled, relative to the entire mine workings. Therefore, it is likely that bicarbonate concentrations of decant water will decrease to concentrations similar to the interim closure period after final mine closure. Decreases in bicarbonate concentrations of this magnitude are not expected to limit attenuation as

evidence shows that metal attenuation was effective at these lower concentrations during the interim shutdown period.

7.4 ATTENUATION BY ADSORPTION TO MINERALS AND SEDIMENTS

Attenuation of metals to minerals and organic matter in sediments beneath the decant pond and in the underlying aquifer is dependent on the presence of solid sorbent phases (e.g. minerals such as $\text{Fe}(\text{OH})_3$, clay minerals, and organic matter) and appropriate water chemistry (primarily pH). The extent, duration and longevity of sorption of trace metals in groundwater are largely controlled by the following factors:

1. Groundwater pH;
2. The quantity of sorbent material (principally hydrous ferric and manganese oxides, organic matter, and clays) within decant pond and aquifer sediments; and
3. The concentration of competing ions in solution.

Site-specific evidence of the longevity of the adsorption mechanisms is provided by the 30-year operating life of the decant pond disposal system. Groundwater quality data from monitoring wells within 100 to 150 of the decant pond demonstrate that metals are effectively attenuated over this short distance, and have been for over 30 years. In comparison the decant pond is located approximately 3,000 feet from Lake Creek, a distance 20 times greater than the current transport distance of the metals. At a minimum, it is estimated that greater than 600 years of adsorption attenuation capacity remain in the shortest potential flow path to Lake Creek. This estimate is based on the following assumptions and calculations:

1. Mine water concentrations are expected to remain similar to current and past levels.
2. The adsorption capacity of sediments in the flow path between the decant pond and the adjacent monitoring wells is similar to the capacity of sediments in the flow path between the monitoring wells and Lake Creek.
3. There are not significant processes providing a sustainable supply of fresh sorbent material. This assumption is known to be false but is made to simplify the calculation

and provide a conservative estimate of adsorption life. Organic matter will continue to be replenished by deposition of algae and detritus in the pond. Groundwater containing high concentrations of iron is known to exist between the tailing impoundment and Lake Creek. Geochemical modeling and laboratory testing demonstrate that this iron provides an effective source of iron oxide and hydroxides for co-precipitation and adsorption of metals (CDM, 2010).

4. The adsorption capacity of sediments in the flow path from the pond to the adjacent monitoring wells (150 feet flow path) is at least 30 years, and likely greater as it has not been depleted yet. To be conservative, the adsorption capacity of the sediments is calculated to be 30 years of attenuation per 150 feet of flow. Stated differently, the attenuation front would be calculated to move toward Lake Creek at the rate of 5 feet/year or less.
5. The time required to deplete the adsorption capacity in the shortest potential flow path between the monitoring wells and Lake Creek (a distance of 2,850 feet) is:

$$2,850 \text{ feet} \times (30 \text{ years attenuation} / 150 \text{ feet}) = 570 \text{ years};$$

or approximately 600 years.

It is important to note that this estimate is only for the adsorption attenuation mechanism. As described in Section 7.3, the mineral precipitation mechanism is expected to last indefinitely.

7.4.1 Water pH

As described in Section 7.1.1, there is nearly 30 years of water quality data demonstrating that mine water is consistently neutral in pH (Parametrix, 2008). Additionally, acid-base accounting tests and long-term laboratory testing with humidity cells indicates that waste rock, ore, and tailings from the Troy Mine are not acid-generating but rather are acid-neutralizing (i.e., contain an excess of alkaline materials) (MDEQ and USFS, 2001; Maxim, 2003). Therefore, mine water pH is expected to remain near neutral to alkaline in perpetuity, and pH conditions are not expected to limit the longevity of the adsorption attenuation mechanism.

7.4.2 Quantity of Sorbent Material

The quantities of inorganic sorbent materials within the aquifer are very large and are not expected to be exhausted. Iron and manganese are the fourth and twelfth most abundant elements in the earth's crust and the average rock contains about 5 percent iron and about 0.1 percent manganese. Although a portion of the iron and manganese occurs in forms such as silicate minerals that are not strong sorbent materials, the amount of iron and manganese oxide and hydroxides in the aquifer is still very large. Moreover, the supply of fresh iron and manganese oxides and hydroxides is continually replenished by the precipitation of iron and manganese minerals from mine water and groundwater near the decant pond. As demonstrated by geochemical modeling and field and laboratory testing by CDM (2010), there is an additional supplemental source of iron and manganese oxides downgradient of the tailing impoundment toe ponds. Groundwater near the impoundment toe naturally contains high levels of dissolved iron and manganese. Should metal-bearing water from the decant pond migrate to this area, mixing of the two groundwaters would result in the formation of fresh iron and manganese oxide/hydroxide minerals and that would adsorb and attenuate metals. This process has been demonstrated to occur in laboratory jar tests (CDM, 2010).

The quantities of organic matter in the aquifer and infiltration area sediments is likely to be low, probably less than 1 percent based on visual estimates of decant pond sediments. The organic matter that is present in the decant ponds was naturally deposited by wind and water transport of detrital material such as leaves and small pieces of tree branches and by the growth and subsequent death and deposition of algae on the pond surface. These processes are expected to continue after mine closure. Deposition of organic matter in the pond may be increased after full reclamation of the tailing impoundment surface as there will be much greater vegetative cover of the impoundment and a larger source of leaves, twigs, etc. Thus, organic matter is a renewable source of sorbent material that will not be depleted.

7.4.3 Concentration of Competing Ions

Adsorption may be viewed as an electrostatic reaction between a charge surface (the sorbent) and the charged ions in water (Cu^{++} , etc.). In some cases, there may be specific sites on the surface that are more suited for one species than another, so that ions may be seen as

“competing” for adsorption sites. A simple classic case of competing ions is the ion exchange reaction occurring in a home water softener where calcium in water reacts with zeolite minerals so that adsorbed sodium ions are forced off the zeolites and into water, while calcium ions take the adsorption sites formerly occupied by sodium. At low concentrations of ions in water (dilute solutions) competition between ions is low as there are a large number of adsorption sites relative to the number of ions desiring sites. Conversely, competition is high at high ionic concentrations.

If the quality of the disposal water were to change such that the total ionic content (e.g. total dissolved solids or TDS) or dissolved metal concentrations were to increase greatly, adsorption could be limited for some species. However, given the dilute nature of the disposal waters and groundwater (total dissolved solids generally less than 400 mg/L) the concentrations of ions in water are relatively low at present and are expected to decrease further after the cessation of mining and milling due the absence of potassium and sodium-containing milling reagents (see Section 2.3 for discussion of common ion trends). Moreover, mine water metal concentrations have been fairly consistent for over 30 years and are not expected to increase in the future.

Competition between ions is not expected to limit the duration or longevity of the adsorption attenuation mechanism, with a few possible exceptions as described in Section 3.3 and summarized as follows:

- Dissolved phosphate competes with arsenic for adsorption sites, and thus high levels of dissolved phosphate may suppress arsenic adsorption and limit arsenic attenuation. Phosphate concentrations in mine water are expected to remain low. One possible source of phosphate to the decant pond would be from fertilizer used during reclamation. Excessive phosphate fertilizer use should be avoided.
- Calcium, magnesium and trace metals (primarily zinc) compete with cadmium for adsorption sites, and thus high levels of these constituents may suppress cadmium adsorption.

- Uranium adsorption is shown to be greatly reduced in water with high carbonate concentrations due to the formation of uranium carbonate complexes that are poorly sorbed. Bicarbonate concentrations are expected to be reduced from current levels after final mine closure. Bicarbonate concentrations of decant pond water have increased since the mine re-start in 2005 (see Section 2.3). This increase is believed to be related to the higher carbonate mineral content of the ore zone that is currently being mined and milled, relative to the entire mine workings. Therefore, it is likely that bicarbonate concentrations of decant water will decrease to concentrations similar to the interim closure period after final mine closure.

Increased concentrations of ions and dissolved organic species can in some cases increase metal adsorption effectiveness. This occurs when the ions or organics form complexes with metals and the complexed form of the metal has a higher adsorption affinity than the uncomplexed metal ion. A notable example of this is copper, which adsorbs better in high ionic strength solutions, especially in water containing dissolved organic chemicals such as humic acid.

8.0 MANAGEMENT PRACTICES AND MONITORING TO VERIFY CONTINUED EFFECTIVENESS OF METAL ATTENUATION

Potential monitoring and management practices that would help ensure the continued effectiveness of metal attenuation by maintaining the geochemical conditions that are conducive to attenuation are described in this Section. The goal of monitoring and management practices would be to maintain the following current conditions at the tailing impoundment:

1. Neutral or alkaline pH;
2. Oxidizing redox conditions (moderate to high Eh);
3. Presence of moderate amounts of dissolved silica;
4. Presence of moderate amounts of bicarbonate; and
5. Presence of low to moderate amounts of organic material.

8.1 MONITORING

Monitoring activities to assess ongoing conditions related to metals attenuation should include continued monitoring of the water quality of mine water, decant pond water, and groundwater near the decant pond and conditions in the decant pond. Water monitoring parameters critical to attenuation include pH, redox (Eh and dissolved oxygen), dissolved silica, bicarbonate, and dissolved metals. The infiltration site should be monitored for excessive build up of organic matter that may reduce oxygen levels and change redox conditions in the groundwater system.

8.2 MANAGEMENT PRACTICES

Management practices should be aimed at maintaining the existing geochemical conditions by:

1. Maintaining exposure of disposal water to atmospheric oxygen and carbon dioxide;
2. Avoiding excessive use of phosphate fertilizer for reclamation (high levels of dissolved phosphate may suppress arsenic adsorption and limit arsenic attenuation);

3. Maintaining pH by monitoring or mitigation; and
4. Maintaining the decant pond/infiltration site as an aerobic environment by monitoring and if necessary removing excessive organic material that may build up.

8.3 POTENTIAL CORRECTIVE ACTIONS

Corrective actions could be done if needed to adjust or maintain geochemical conditions to maintain or enhance metal attenuation. Corrective actions could include:

1. Actions to raise pH
 - a. Add alkaline material to mine water and/or decant pond.
2. Actions to maintain oxidizing conditions
 - a. Remove organic matter from pond bottom;
 - b. Aeration of water; and
 - c. Addition of a chemical oxidant.
3. Actions to raise dissolved silica
 - a. Increase exposure to fine-grained silica material (e.g. place tailings in transport path or in pond bottom); and
 - b. Raise pH to increase silica solubility.
4. Actions to raise bicarbonate
 - a. Add carbonate material (crushed or powdered limestone) to mine water and/or decant pond.

9.0 SUMMARY AND CONCLUSIONS

Mine water and tailings impoundment water have been infiltrated to the groundwater system at the Troy Mine tailings impoundment for over 30 years without appreciable transport of metals in groundwater away from the infiltration pond (decant pond). The lack of transport of metals in groundwater is due to natural attenuation of metals in the sediments beneath the infiltration pond and in the groundwater aquifer. Field evidence for natural attenuation of metals is provided by the sampling and analysis of water and soil samples in the vicinity of the tailings impoundment infiltration pond and from laboratory testing.

Water quality data shows that the concentrations of chemical parameters that are not attenuated (e.g. chemically conservative parameters such as nitrate) are similar between the mine water/decant pond water and the underlying groundwater. This similarity in concentrations indicates that mine water is not appreciably diluted in the groundwater system in close proximity to the infiltration pond. In contrast, the concentrations of metal parameters are much lower in the groundwater than in the mine water/decant pond, indicating that metal parameters are attenuated (removed) in the groundwater system.

Analyses of sediments in the subsurface of the decant pond reveal that metals are enriched in the sediments, indicating that the metals attenuated/removed from groundwater are retained on the sediments. Microscopic and electron microprobe examination of sediments document the occurrence of secondary copper minerals (i.e., non-ore forming minerals) in the sediments. These secondary minerals form by precipitation of copper from mine water. Geochemical equilibrium modeling confirms that copper, iron, lead, and manganese minerals are oversaturated in mine water and are favored to form from mine water.

9.1 EFFECTIVENESS AND LONGEVITY OF MONITORED NATURAL ATTENUATION

The effectiveness of natural attenuation in reducing the concentrations of metals in mine water disposed at the Try tailing impoundment has been proven by extensive field, laboratory and modeling evidence. Disposal of mine water and/or decant water (a mixture of tailings

water, mine water and natural precipitation) at the tailing impoundment has been conducted for over 30 years with no detection of increased metal concentrations in Lake Creek or in groundwater near Lake Creek. Monitoring of groundwater in wells very close to the decant pond (within 150 feet) demonstrate the rapid removal of metals over relatively short distances.

The longevity of the natural attenuation mechanisms is dependent on the duration of the geochemical conditions conducive to attenuation and to the capacity of the decant pond/aquifer sediments to continue to uptake metals. The mineral precipitation and co-precipitation mechanisms are expected to last indefinitely or in perpetuity as long as geochemical conditions remain similar to current conditions. The adsorption mechanisms are conservatively estimated to last a minimum of 600 years.

9.2 ANTIMONY

Antimony concentrations are relatively low in mine water, ranging from <0.003 mg/L to 0.016 mg/L since 1999. Historically, decant water concentrations were similar to mine water but since the mine restart in 2005 antimony concentrations have ranged up to 0.062 mg/L in decant pond water. After the cessation of mining, mine disposal water is expected to be similar to current mine water (<0.003 to 0.016 mg/L).

The scientific literature documents that antimony attenuation occurs primarily by the adsorption of antimony by iron oxides and hydroxides. This effect was demonstrated for Troy Mine waters in a water mixing test by CDM (2010). Mixing of mine water and decant pond water with groundwater from the vicinity of the tailing impoundment resulted in the precipitation of iron and manganese oxides/hydroxides from the mixed waters and the concurrent removal of up to 59 percent of the antimony. Groundwater monitoring data from wells within 150 feet of the decant pond also demonstrates that antimony concentrations are reduced to about 50 percent of the decant pond water in the shallowest groundwater (0.032 mg/L in well MW-01-15). In deeper groundwater (well MW-95-8), antimony concentrations are further attenuated to 0.005 mg/L, a greater than 90 percent reduction.

9.3 ARSENIC

Arsenic concentrations in mine water and decant pond water are very low (0.001 to 0.008 mg/L) and are reduced further by attenuation mechanisms such that arsenic is not detectable in groundwater in the monitoring wells adjacent the decant pond (<0.001 to <0.003 mg/L). Literature reports and site-specific evidence suggests that the primary mechanism of arsenic attenuation is co-precipitation and adsorption to iron and manganese oxide/hydroxide minerals. Geochemical modeling of mine water and decant pond water suggests that pure arsenic mineral phases (i.e., phases containing only arsenic and no other metals) are not expected to precipitate from mine water, decant pond water, and groundwater. Laboratory tests by CDM (2010) in which mine water and decant pond water were mixed with ambient groundwater containing high concentrations of iron demonstrate that arsenic is attenuated simultaneously with the precipitation of iron and manganese oxide/hydroxide solid phases.

9.4 CADMIUM

Cadmium concentrations in decant pond water and mine water are low, typically <0.001 to 0.002 mg/L, and are probably similar to ambient cadmium concentrations in surface water and groundwater in the area. For comparison, Hem (1985, p. 142) estimates the median concentration of cadmium in surface water in the United States to be 0.001 mg/L. Because the cadmium concentrations in water are so low and are often unmeasurable, little is known about cadmium behavior in water at the Troy Mine. In the mine water, decant pond water, and groundwater laboratory mixing tests (CDM, 2010), cadmium concentrations in mine water and decant pond water were not detected at a detection limit of 0.00008 mg/L (i.e., <0.08 parts per billion). Cadmium attenuation could not be determined in these tests because of the extremely low and unmeasurable concentrations.

Cadmium concentrations in nearby groundwater are similar to decant pond water suggesting that little attenuation occurs over the short transport distance between the pond and the monitoring wells. If significantly higher concentrations of cadmium were to occur in mine disposal water after mine closure, it is likely that cadmium would be attenuated at these higher concentrations as cadmium attenuation by mineral precipitation/co-precipitation and adsorption as described in the literature (see Section 3.4).

9.5 COPPER

Copper concentrations ranged from 0.044 to 0.17 mg/L during the interim mine shutdown period of 2000 through 2004 and are expected to be similar after final mine closure. The fate of copper disposed in the decant pond and the attenuation of copper has been extensively studied (Cannon Microprobe (2003), Land Water Consulting (2004), CDM (2010)) by many methods including:

- Groundwater monitoring that demonstrates that copper concentrations are greatly reduced in close proximity to the decant pond;
- Microscopic and electron microprobe identification of secondary copper minerals that are formed within the sediments beneath the decant pond;
- Chemical analyses of sediments beneath the decant pond that demonstrate enrichment in copper by the attenuation mechanisms; and
- Sequential extraction analyses of sediments beneath the decant pond that demonstrate the fractionation of copper in solid phases formed by the attenuation mechanisms.

These studies demonstrate the effectiveness of the natural attenuation of copper in limiting the transport of copper in the groundwater system.

9.6 LEAD

Lead concentrations in mine water are very low (<0.002 to 0.003 during the interim shutdown period; 0.002 to 0.008 mg/L in 2009). Lead concentrations are significantly attenuated and reduced in groundwater adjacent the decant pond as the highest measured groundwater concentration is only 0.0011 mg/L. For comparison, the Montana Groundwater Quality Standard is an order of magnitude higher than groundwater concentration, 0.015 mg/L.

The attenuation of lead is most likely by adsorption and the precipitation of lead minerals from mine water. The scientific literature documents that adsorption and mineral precipitation are common methods of lead attenuation. Geochemical modeling of Troy mine

water, decant pond, and groundwater indicates that the waters are saturated or nearly saturated with cerussite (PbCO_3), meaning that precipitation of cerussite and removal of lead from solution is favored. The adsorption effect was demonstrated for Troy Mine waters in a water mixing test by CDM (2010). Mixing of mine water and decant pond water with groundwater from the vicinity of the tailing impoundment resulted in the precipitation of iron and manganese oxides/hydroxides from the mixed waters and the concurrent attenuation of lead concentrations to less than detectable amounts (<0.0005 mg/L).

9.7 MERCURY

Mercury has not been detected in mine water, decant pond water, or groundwater at the Troy Mine. The scientific literature documents that mercury is commonly attenuated by adsorption onto iron and manganese oxide minerals, clays, and organic matter. If significantly higher concentrations of mercury were to occur in mine disposal water after mine closure, it is likely that mercury would be attenuated by these mechanisms.

9.8 SILVER

Silver has not been detected in mine water, decant pond water, or groundwater at the Troy Mine. The scientific literature documents that silver is commonly attenuated by adsorption onto iron and manganese oxide minerals and clays. If significantly higher concentrations of silver were to occur in mine disposal water after mine closure, it is likely that silver would be attenuated by these mechanisms.

9.9 URANIUM

Uranium concentrations are low in mine water (ranging from <0.0003 to 0.0013 mg/L in 2009). For comparison the Montana Groundwater Quality Standard is significantly higher at 0.03 mg/L. Uranium is typically attenuated in groundwater by adsorption mechanisms, in particular by adsorption to iron and manganese oxides. This effect was demonstrated for Troy Mine waters in a test by CDM (2010) in which mixing of mine water and decant pond water with groundwater from the vicinity of the tailing impoundment resulted in the precipitation of iron and manganese oxides/hydroxides from the mixed waters and the concurrent removal of up to 85 percent of the uranium.

At the Troy Mine decant pond, attenuation of uranium is demonstrated by reductions in uranium concentrations as water flows from the decant pond to the adjacent groundwater monitoring wells. Uranium concentrations are reduced by approximately 30 percent during transport to the closest monitoring well (approximately 100 feet lateral distance) and further reduced by 85 percent during flow to the next closest well (approximately 150 feet lateral distance).

10.0 REFERENCES

- Bethke, CM, 2008. Geochemical and Biogeochemical Reaction Modeling, 2nd edition. Cambridge University Press, Cambridge. 543 pp.
- Camp, Dresser & McKee (CDM), 2010. Troy Mine Copper Attenuation Study – Secondary Processes. Final Draft Report. March 18, 2010. Prepared for Montana Department of Environmental Quality.
- Cannon Microprobe, 2003. Electron Microprobe and Scanning Electron Microscope Analysis of DP-2 Decant Pond Sediments. November 25, 2003.
- Crane, M.J.; JL Sharpe, and PA Williams, 2001. Formation of Chrysocolla and Secondary Copper Phosphates in the Highly Weathered Supergene Zones of Some Australian Deposits. Records of the Australian Museum (2001) Vol. 53: 49–56.
- Ford, R.G., and D.L. Sparks, 2000. The nature of Zn precipitates formed in the presence of pyrophyllite. Env. Sci. Technol. 34:2479-2483.
- Gatehouse, S., DW Russel and JC Van Moort, 1977. Sequential soil analysis in exploration geochemistry. Journal of Geochemical Exploration, 8:483-494.
- Genesis Inc., 2006. Genesis Incorporated Troy Mine Revised Reclamation Plan. Revised March 2006.
- Hem J. D., 1985. Study and Interpretation of the Chemical Characteristics of Natural Water. Water-Supply Paper 2254, U.S. Geological Survey, Reston, VA, 1985.
- Hydrometrics, Inc., 2000. Mine Water Fate and Transport Study Plan. Prepared for Genesis Inc./Sterling Mining Company. September, 2000.
- Hydrometrics, Inc., 2001. Mine Water Plume Location and Identification Phase 1 Results and Phase 2 & 3 Workplan. Prepared for Genesis Inc./Sterling Mining Company. October, 2001
- Hydrometrics, Inc., 2009. Proposed Scope of Work for Copper Attenuation Study Addendum. Prepared for Genesis Inc. January 23, 2009.
- Langmuir, D., 1997a. Aqueous environmental chemistry. Upper Saddle River, NJ: Prentice-Hall. 600 pp.
- Langmuir, D, P. Chrostowski, B. Vigneault and R. Chaney, 2004. Issue Paper on the Environmental Chemistry of Metals. Submitted to USEPA Risk Assessment Forum, August 2004.

- Land and Water Consulting, 2004. Assessment of Fate and Transport of Copper in Decant Pond Disposal System – Troy Mine. Prepared for Revett Mining Company. January 2004.
- Lindberg, R.D. and D.D. Runnells, 1984. Ground water redox reactions: An analysis of equilibrium state applied to Eh measurements and geochemical modeling. Science. v.225. p. 925-927.
- Lawrence Livermore National Laboratory, 2006. Thermodynamic database: thermo.com.V8.R6. Modified March 28, 2006. Database distributed with Geochemist's Workbench software (RockWare, 2010).
- Maxim Technologies, 2003. Rock Creek Project 2003 Geochemistry Review. Prepared for USFS Kootenai National Forest. May 2003.
- McMahon, P.B., Chapelle, F.H., 2008. Redox Processes and Water Quality of Selected Principal Aquifers, Ground Water, vol. 46(2), 259-271.
- Montana Dept. of Environmental Quality and US Forest Service, 2001. Final Environmental Impact Statement – Rock Creek Project. September 2001.
- Parametrix, 2008. Final Annual Report – Year 22; Comprehensive Water Quality Monitoring Program for Streams Adjacent to Genesis, Inc. – Troy Mine: Analysis of Spring 1986 through Fall 2006 Samples. Prepared for Genesis Inc. and MDEQ. January 2008.
- RockWare, 2010. The Geochemist's Workbench, Release 8.0.7. Copyright by Craig M. Bethke and The Board of Trustees of the University of Illinois.
- Summit Envirosolutions, 1999. Site Instrumentation Status Report Troy Tailings Impoundment, Lincoln County, Montana.
- US Environmental Protection Agency, 1979. Water-related Environmental Fate of 129 Priority Pollutants, Volume 1 and 2. EPA 440/4-79-029a.
- US Environmental Protection Agency, 1991. Mine Site Visit Report – Asarco/Troy Mine.
- US Environmental Protection Agency, 1999. Understanding Variation in Partition Coefficient, K_d , Values – Volume 2 - Review of Geochemistry and Available K_d Values for Cadmium, Cesium, Chromium, Lead, Plutonium, Radon, Strontium, Thorium, Tritium (3H), and Uranium. EPA 402-R-99-004B. August 1999.
- US Environmental Protection Agency, 2004. Understanding Variation in Partition Coefficient, K_d , Values – Volume III – Review of Geochemistry and Available K_d Values for Americium, Arsenic, Curium, Iodine, Neptunium, Radium, and Technetium. EPA 402-R-04-002C. July 2004.

- US Environmental Protection Agency, 2007a. Monitored Natural Attenuation of Inorganic Contaminants in Ground Water, Volume 1 and 2. EPA/600/R-07/139. October 2007.
- US Environmental Protection Agency, 2007b. Metal Attenuation Processes at Mining Sites EPA/600/R-07/092. September 2007
- Westall, J., 2002. Geochemical Equilibrium and the Interpretation of Eh in Workshop on Monitoring Oxidation-Reduction Processes for Ground-water Restoration. Workshop Summary. Dallas, Texas – April 25-27, 2000. EPA/600/R-02/002. January 2002

APPENDIX A

WATER QUALITY SAMPLE SUMMARIES AND LAB REPORTS



262 Burlington
Missoula, MT 59801
(406) 543-3660
Fax: (406) 721-4916
www.hydrometrics.com

May 15, 2009

MEMO TO FILE: Troy Mine (Genesis Mine) Copper Attenuation Samples.

Six of the eight proposed sites for collecting Copper attenuation samples and data were sampled on May 12 – 14, 2009. Seven samples (six sites plus a duplicate) were shipped from the mine site to Energy Laboratories in Billings, MT on May 14, 2009.

Two of the monitor wells proposed for sampling, MW01-16 and MW95-8, were not sampled. MW01-16 was dry. MW95-8 was not accessible because a new surface riser had been added to it when the tailings cell divider dike was raised and access to the well had not been reestablished.

Sites that were sampled include MW95-4, MW95-5, MW01-15, the Service Adit Ditch, Service Adit Pipe, and the decant pond. The Service Adit Ditch, Service Adit Pipe, and Decant samples were all grab samples. Monitor well samples were collected using a rediflo pump. Wells were purged at flow rates less than 0.1 gpm and routed through a flow through cell where groundwater parameters (pH, conductivity, dissolved oxygen, temperature, and oxidation-reduction potential) were measured. Samples were taken when the groundwater parameters stabilized per Hydrometrics standard operating procedure HF-SOP-11. In one case (MW95-4) the oxidation-reduction potential appeared to still be climbing when the sample was taken, however the other parameters had stabilized when the sample was taken.

Sampling field data sheets are attached.

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: TRAIL MINE
 Project Code: _____
 Sample Team Member(s): LTJPK
 Laboratory Used: ELI

Site Designation: Mu 45-4
 Sample Code Number: _____
 Sample Date: 5/12/09
 Sample Time: 10:30 (military)

If Duplicate Sample Collected,
Please Record Below

Duplicate Sample Code #: _____
 Duplicate Sample Time: _____

Site Conditions

New Site: Yes ☒ No ☐ Photo taken: Yes ☐ No ☐
 Site Type: DRY surface water process water

monitoring well domestic well adit seep
 spring-other: _____

Weather Conditions: calm ☒ breeze ☐ windy ☐
 no precip. ☒ rain ☐ snow ☐
 clear p. cloudy overcast

Air Temperature: _____ °C 55 °F

For Groundwater Samples

well volume formula: $V = (TD-SWL) \times (\text{Dia.}^2) \times 25$ Comments
 TD (ft): 17.7
 SWL (ft): 10.2 no access/pumping
 Casing Diameter (I.D.): 2"
 Water Volume (V) (gal): 1.2
 x 3=(gal.) 3.6
 Actual Vol. Removed (gal.): 2+/-
 Water Level Recovery: slow moderate ☒ rapid

For Surface Water Samples

Flow Method: Marsh McBirney Volumetric Flume Weir Estimate

Other Flow or Description: _____

Flow: _____ gpm _____ cfs Staff Gage: _____

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes
<u>1039</u>	<u>-21.3</u>	<u>0.12</u>	<u>6.36</u>	<u>254</u>	<u>—</u>	<u>5.87</u>	
<u>1048</u>	<u>-26.2</u>	<u>0.15</u>	<u>7.39</u>	<u>286</u>	<u>—</u>	<u>7.02</u>	
<u>1100</u>	<u>13.6</u>	<u>0.18</u>	<u>7.38</u>	<u>347</u>	<u>—</u>	<u>7.43</u>	
<u>1105</u>	<u>38.7</u>	<u>0.19</u>	<u>7.70</u>	<u>358</u>	<u>—</u>	<u>7.70</u>	
<u>1110</u>	<u>45.4</u>	<u>0.20</u>	<u>7.79</u>	<u>36.7</u>	<u>—</u>	<u>7.80</u>	
<u>1113</u>	<u>52.7</u>	<u>0.18</u>	<u>7.31</u>	<u>374</u>	<u>—</u>	<u>7.89</u>	<u>7.93</u>

Turbidity: clear ☐ moderate ☐
 (circle) slight ☒ very ☐

Sample Method: grab composite ☒ pump ☐ bailer ☐ other ☐
 (describe) 20 PLO 20.1 gpm

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>52.7</u>	
DO (mg/l)	<u>0.18</u>	
pH	<u>7.31</u>	
SC (µmhos/cm)	<u>374</u>	
Turbidity (ntu)		
H ₂ O Tmp. (°C)	<u>7.93</u>	
Color		
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1	500 ml	F or UF	—	Cond	<u>17.0 SAR AL</u>
1	500 ml	F or UF	—	TDS	
1	250 ml	UF	HNO ₃	Diss MET	
1	250 ml	UF	HCl	Hg	
1	250 ml	F or UF	HNO ₃	TOC	<u>20.0 BAC MET</u>
1	250 ml	F or UF	H ₂ SO ₄	As	<u>Asst-2000-SL</u>
1	250 ml	F or UF	NaOH-2000	C	<u>Sulfide</u>
1	500 ml	F or UF	HNO ₃	C	

Comments: 25' tubing, 1 filter

Sample Team Member Signature: Larry M. Jones

Page 1 of 1

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: Way Mine
 Project Code: _____
 Sample Team Member(s): LS/RK
 Laboratory Used: ECI

Site Designation: 11495 X3
 Sample Code Number: _____
 Sample Date: 3/12/09
 Sample Time: 1:15:45 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: _____
 Duplicate Sample Time: _____

Site Conditions

New Site: Yes ☐ No ☒ Photo taken: Yes ☐ No ☐
 Site Type: DRY surface water process water
monitoring well domestic well adit seep
 spring-other: _____

Weather Conditions: calm ☐ breeze ☒ windy ☐
 no precip. ☐ rain ☒ snow ☐
 clear ☐ p. cloudy ☒ overcast ☐
 Air Temperature: _____ °C 55 °F

For Groundwater Samples

well volume formula:	$V = (TD-SWL) \times (\text{Dia.}^2) / 25$	Comments
TD (ft):	<u>13.40 (sanded)</u>	no access/pumping
SWL (ft):	<u>1.25</u>	
Casing Diameter (I.D.):	<u>2</u>	
Water Volume (V) (gal):	<u>1</u>	
x 3=(gal.):	<u>3</u>	
Actual Vol. Removed (gal.):	<u>2.5</u>	
Water Level Recovery:	slow moderate rapid	

For Surface Water Samples

Flow Method:	Marsh McBirney Volumetric Flume Weir Estimate
Other Flow or Description:	_____
Flow:	gpm cfs Staff Gage:

Field Parameter Stabilization grab Pump @ 13:10

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes
<u>1319</u>	<u>167.7</u>	<u>1.86</u>	<u>6.10</u>	<u>78</u>	<u>—</u>	<u>6.98</u>	
<u>1330</u>	<u>175.2</u>	<u>1.88</u>	<u>6.99</u>	<u>75</u>	<u>—</u>	<u>8.10</u>	
<u>1335</u>	<u>177.9</u>	<u>1.76</u>	<u>5.44</u>	<u>74</u>	<u>—</u>	<u>8.51</u>	
<u>1343</u>	<u>174.5</u>	<u>1.77</u>	<u>6.04</u>	<u>712</u>	<u>—</u>	<u>8.8</u>	

Turbidity: clear ☐ slight ☒ moderate ☐ very ☐

Sample Method: grab composite ☒ pump ☐ bailer ☐ other ☐
 (describe) grab Pump @ 13:10

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>179.51</u>	
DO (mg/l)	<u>1.77</u>	
pH	<u>6.04</u>	
SC (µmhos/cm)	<u>72</u>	
Turbidity (ntu)	<u>7.2</u>	
H ₂ O Tmp. (°C)	<u>8.8</u>	
Color		
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1	250 ml	F or UF	—	Cond. H ⁺ , F, Fe, SAR, Alk	<u>P.H</u>
1	250 ml	F or UF	—	SPS	
1	250 ml	F or UF	HNO ₃	Diss Met	
1	250 ml	F or UF	HCl	Hg	
1	250 ml	F or UF	HNO ₃	For Bae Met	
1	250 ml	F or UF	H ₂ SO ₄	N	
1	250 ml	F or UF	H ₂ SO ₄	As, Pb, Cu, Zn, Cd, Ni, Cr, Mn, Co, Se, Sulfide	
1	500 ml	F or UF	H ₂ SO ₄	C	

Comments: 20' tubing, 1 filter

Sample Team Member Signature: Larry H. Fisher

Page 1 of 1

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: May Mine
 Project Code: _____
 Sample Team Member(s): LS/PK
 Laboratory Used: SLI

Site Designation: Adit Pipe
 Sample Code Number: _____
 Sample Date: 5/12/04
 Sample Time: 3:15 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: _____
 Duplicate Sample Time: _____

Site Conditions

New Site: Yes ☒ No ☐ Photo taken: Yes ☐ No ☐
 Site Type: DRY surface water process water
 monitoring well domestic well ☒ adit seep

spring-- other: _____

Weather Conditions: calm ☒ breeze ☐ windy ☐
 no precip. ☒ rain ☐ snow ☐
 clear p. cloudy ☒ overcast ☐

Air Temperature: _____ °C 57 °F

For Groundwater Samples

well volume formula: $V = (TD-SWL) \times (\text{Dia.}^2) / 25$ Comments
 TD (ft): _____
 SWL (ft): _____ no access/pumping
 Casing Diameter (I.D.): _____
 Water Volume (V) (gal): _____
 x 3=(gal.) _____
 Actual Vol. Removed (gal.): _____
 Water Level Recovery: slow moderate rapid

For Surface Water Samples

Flow Method: Marsh McBirney Volumetric Flume Weir Estimate

Other Flow or Description: _____

Flow: gpm cfs Staff Gage: _____

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes

Turbidity: clear ☐ moderate ☒ slight ☐ very ☐

Sample Method: ☒ grab composite ☐ pump ☐ bailer ☐ other

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>174.2</u>	
DO (mg/l)	<u>4.11</u>	
pH	<u>7.52</u>	
SC (µmhos/cm)	<u>245</u>	
Turbidity (ntu)		
H ₂ O Tmp. (°C)	<u>6.37</u>	
Color		
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1	300 ml	F or UF <input checked="" type="radio"/>		COND. #1, #1 #0	
1	500 ml	F or UF <input checked="" type="radio"/>		TPS	
1	250 ml	UF or UF <input checked="" type="radio"/>	HNO ₃	DISS MET	
1	250 ml	UF or UF <input checked="" type="radio"/>	HCl	HEI	Hg
1	250 ml	F or UF <input checked="" type="radio"/>	HNO ₃	DISS MET	
1	250 ml	F or UF <input checked="" type="radio"/>	H ₂ SO ₄	N	
1	250 ml	F or UF <input checked="" type="radio"/>	H ₂ O ₂	NO ₃	
1	250 ml	F or UF <input checked="" type="radio"/>	H ₂ PO ₄	H ₂ PO ₄	C

Comments: _____

Sample Team Member Signature: Lam H. Phan

Page 1 of 1

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: TRAY MINE
 Project Code: _____
 Sample Team Member(s): LS/PK
 Laboratory Used: ELI

Site Designation: ADIT DITCH
 Sample Code Number: _____
 Sample Date: 5/12/09
 Sample Time: 4:30 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: _____
 Duplicate Sample Time: _____

Site Conditions

New Site: Yes ☐ No ☒ Photo taken: Yes ☐ No ☐
 Site Type: DRY surface water process water
 monitoring well domestic well adit seep

Weather Conditions: calm ☐ breeze ☒ windy ☐
 no precip. ☐ rain ☒ snow ☐
 clear p. cloudy overcast

Air Temperature: _____ °C 55 °F

For Groundwater Samples

well volume formula: $V = (TD-SWL) \times (Dia.^2) / 25$	Comments
TD (ft): _____	no access/pumping
SWL (ft): _____	
Casing Diameter (I.D.): _____	
Water Volume (V) (gal): _____	
x 3=(gal.) _____	
Actual Vol. Removed (gal.): _____	
Water Level Recovery: slow moderate rapid	

For Surface Water Samples

Flow Method: Marsh McBirney Volumetric Flume Weir Estimate

Other Flow or Description: _____

Flow: gpm cfs Staff Gage: _____

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes
<u>3:30</u>	<u>182.7</u>	<u>10.75</u>	<u>7.53</u>	<u>146</u>		<u>5.25</u>	

Turbidity: clear ☐ slight ☐ moderate ☒ very ☐

Sample Method: grab composite pump bailer other
 (describe)

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>182.7</u>	
DO (mg/l)	<u>10.75</u>	
pH	<u>7.53</u>	
SC (µmhos/cm)	<u>146</u>	
Turbidity (ntu)		
H ₂ O Tmp. (°C)	<u>5.25</u>	
Color		
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1	500 ml	F or UF		CO ₂	u ⁺ , Alk, Fe
1	500 ml	F or UF		TDS	
1	250 ml	UF	HNO ₃	RSS	u ⁺
1	250 ml	UF	HCl	H ₂ O ₂	Hg
1	250 ml	F or UF	HNO ₃	H ₂ O ₂	10, 22 u ⁺
1	250 ml	F or UF	H ₂ SO ₄	H ₂ SO ₄	N
1	250 ml	F or UF	NaOH	NaOH	
1	500 ml	F or UF	H ₂ PO ₄	H ₂ PO ₄	u ⁺

Comments: _____

Sample Team Member Signature: Larry M. Jones

Page 1 of 1

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: Gray Muck
 Project Code: _____
 Sample Team Member(s): 1/5/14/14/14
 Laboratory Used: EL

Site Designation: MW-015
 Sample Code Number: _____
 Sample Date: 5/13/09
 Sample Time: 15:40 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: _____
 Duplicate Sample Time: _____

Site Conditions

New Site: Yes ☒ No ☐ Photo taken: Yes ☐ No ☐
 Site Type: ☒ DRY ☐ surface water ☐ process water
☐ monitoring well ☐ domestic well ☐ adit ☐ seep
☐ spring-- other: _____

Weather Conditions: calm ☒ breeze ☐ windy ☐
 no precip. ☐ rain ☒ snow ☐
 clear ☐ p. cloudy ☐ overcast ☐
 Air Temperature: _____ °C

well volume formula: $V = (TD-SWL) \times (\text{Dia.}^2) / 25$ Comments
 TD (ft): 43.9 53.25
 SWL (ft): 53 43.9 no access/pumping
 Casing Diameter (I.D.): 2"
 Water Volume (V) (gal): 1.5
 x 3=(gal.) 4.5
 Actual Vol. Removed (gal.): _____
 Water Level Recovery: slow moderate rapid

For Surface Water Samples

Flow Method: Marsh McBirney Volumetric Flume Weir Estimate
 Other Flow or Description: _____
 Flow: _____ gpm _____ cfs Staff Gage: _____

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes
<u>1426</u>	<u>183.7</u>	<u>3.80</u>	<u>7.18</u>	<u>390</u>	<u>—</u>	<u>8.68</u>	
<u>1449</u>	<u>181.8</u>	<u>3.49</u>	<u>7.20</u>	<u>390</u>	<u>—</u>	<u>9.97</u>	
<u>1501</u>	<u>180.9</u>	<u>4.98</u>	<u>7.30</u>	<u>390</u>	<u>—</u>	<u>10.9</u>	
<u>1516</u>	<u>183</u>	<u>5.17</u>	<u>7.24</u>	<u>390</u>	<u>—</u>	<u>10.85</u>	

Turbidity: clear ☐ moderate ☐ very ☒ slight ☐
 Sample Method: grab composite ☒ pump ☐ bailer other ☐
 (describe) 50-1 gpm

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>183</u>	
DO (mg/l)	<u>5.17</u>	
pH	<u>7.24</u>	
SC (µmhos/cm)	<u>390</u>	
Turbidity (ntu)		
H ₂ O Tmp. (°C)	<u>10.85</u>	
Color		
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
<u>1</u>	<u>500</u> ml	F or <input checked="" type="checkbox"/> UF	<u>—</u>	<u>COND</u>	<u>UF</u>
<u>1</u>	<u>500</u> ml	F or <input checked="" type="checkbox"/> UF	<u>—</u>	<u>TDS</u>	
<u>1</u>	<u>250</u> ml	<input checked="" type="checkbox"/> F or UF	<u>HNO3</u>	<u>DISS.</u>	<u>MT</u>
<u>1</u>	<u>250</u> ml	<input checked="" type="checkbox"/> F or UF	<u>HNO3</u>	<u>HS</u>	
<u>1</u>	<u>250</u> ml	F or <input checked="" type="checkbox"/> UF	<u>HNO3</u>	<u>Fe -</u>	<u>TOTAL</u>
<u>1</u>	<u>250</u> ml	F or <input checked="" type="checkbox"/> UF	<u>HNO3</u>	<u>N</u>	
<u>1</u>	<u>500</u> ml	F or <input checked="" type="checkbox"/> UF	<u>HNO3</u>	<u>S</u>	
<u>1</u>	<u>500</u> ml	F or <input checked="" type="checkbox"/> UF	<u>HNO3</u>	<u>CT</u>	

Comments: SET Pump @ 50' : 60' hose, 1 ft/sec
START Pump @ 14:15

Sample Team Member Signature: Gray Muck

Page 1 of 1

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: 1001 Nemo
 Project Code: _____
 Sample Team Member(s): LS/PA
 Laboratory Used: ELI

Site Designation: RECAUT 1
 Sample Code Number: _____
 Sample Date: 5/14/09
 Sample Time: 8:30 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: RECAUT 2
 Duplicate Sample Time: 8:45

Site Conditions

New Site: Yes ☐ No ☒ Photo taken: Yes ☐ No ☒
 Site Type: DRY surface water process water
 monitoring well domestic well adit seep
 spring- other: Recaut Pond
 Weather Conditions: calm breeze windy
 no precip. rain snow
 clear p. cloudy overcast
 Air Temperature: _____ °C 46 °F

For Groundwater Samples

well volume formula: $V = (TD-SWL) \times (\text{Dia.}^2) / 25$	Comments
TD (ft): _____	no access/pumping
SWL (ft): _____	
Casing Diameter (I.D.): _____	
Water Volume (V) (gal): _____	
Actual Vol. Removed (gal.): _____	
Water Level Recovery: slow moderate rapid	

For Surface Water Samples

Flow Method: Marsh McBirney Volumetric Flume Weir Estimate
Other Flow or Description: _____
Flow: _____ gpm _____ cfs Staff Gage: _____

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (μmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes

Turbidity: clear slight moderate very

Sample Method: grab composite pump bailer other
 (describe)

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>179.5</u>	<u>179.5</u>
DO (mg/l)	<u>8.19</u>	<u>8.19</u>
pH	<u>7.43</u>	<u>7.43</u>
SC (μmhos/cm)	<u>394</u>	<u>394</u>
Turbidity (ntu)	<u>9.50</u>	<u>9.5</u>
H ₂ O Tmp. (°C)		
Color		
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1+1	500 ml	F or UF		conduct	MF
1+1	500 ml	F or UF		TPS	
1+1	250 ml	F or UF	HNO ₃	Diss	MF
1+1	250 ml	F or UF	HCl	#6	
1+1	250 ml	F or UF	HNO ₃	#6	TOTAL C Fe
1+1	250 ml	F or UF	H ₂ SO ₄	N	
1+1	250 ml	F or UF	H ₂ SO ₄	5	
1+1	500 ml	F or UF	H ₂ PO ₄	24	

Comments: Taken from Recaut pond # from pump diff Reso dat.

Sample Team Member Signature: Long Refl

Page 1 of 1



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-001
Client Sample ID: MW 95-4

Report Date: 05/29/09
Collection Date: 05/12/09 10:30
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	6.5	s.u.		0.1		A4500-H B	05/18/09 14:52 / nlh
Conductivity	445	umhos/cm		1		A2510 B	05/18/09 15:51 / nlh
Solids, Total Dissolved TDS @ 180 C	292	mg/L		10		A2540 C	05/18/09 16:15 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	77	mg/L		4		A2320 B	05/18/09 22:56 / ehb
Bicarbonate as HCO ₃	94	mg/L		4		A2320 B	05/18/09 22:56 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	05/18/09 22:56 / ehb
Chloride	1	mg/L		1		E300.0	05/20/09 20:08 / kh
Sulfate	111	mg/L		1		E300.0	05/20/09 20:08 / kh
Fluoride	ND	mg/L		0.1		A4500-F C	05/18/09 22:56 / ehb
Sulfide	0.22	mg/L		0.04		A4500 S-D	05/19/09 09:30 / afb
Hardness as CaCO ₃	108	mg/L		1		A2340 B	05/19/09 16:20 / rls
Sodium Adsorption Ratio (SAR)	0.53			0.01		Calculation	05/19/09 16:20 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	4.8	mg/L	D	0.8		A5310 C	05/20/09 11:12 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	3.72	mg/L		0.05		E350.1	05/19/09 10:48 / sam
Nitrogen, Nitrate+Nitrite as N	0.02	mg/L		0.01		E353.2	05/18/09 14:27 / bls
Phosphorus, Orthophosphate as P	0.014	mg/L	LH	0.005		E365.1	05/19/09 15:49 / mlm
- The sample was received after the holding time had expired for Orthophosphat.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5		E245.7	05/27/09 17:30 / eli-c
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.8	05/19/09 17:32 / aje
Antimony	ND	mg/L		0.003		E200.8	05/19/09 17:32 / aje
Arsenic	0.002	mg/L		0.001		E200.8	05/19/09 17:32 / aje
Barium	0.114	mg/L		0.005		E200.8	05/19/09 17:32 / aje
Beryllium	ND	mg/L		0.001		E200.8	05/19/09 17:32 / aje
Cadmium	0.00093	mg/L		0.00008		E200.8	05/19/09 17:32 / aje
Calcium	29	mg/L		1		E200.7	05/19/09 16:20 / tao
Chromium	0.002	mg/L		0.001		E200.8	05/19/09 17:32 / aje
Copper	ND	mg/L		0.001		E200.8	05/19/09 17:32 / aje
Iron	7.84	mg/L		0.05		E200.7	05/19/09 16:20 / tao
Iron, Ferrous	7.84	mg/L		0.05		E200.7	05/19/09 16:20 / tao
Lead	ND	mg/L		0.0005		E200.8	05/19/09 17:32 / aje
Magnesium	8	mg/L		1		E200.7	05/19/09 16:20 / tao
Manganese	2.43	mg/L		0.005		E200.7	05/19/09 16:20 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-001
Client Sample ID: MW 95-4

Report Date: 05/29/09
Collection Date: 05/12/09 10:30
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	05/19/09 16:20 / tao
Potassium	17	mg/L		1		E200.7	05/19/09 16:20 / tao
Selenium	0.003	mg/L		0.001		E200.8	05/19/09 17:32 / aje
Silicon	21.7	mg/L		0.1		E200.7	05/20/09 13:06 / tao
Silica	46.4	mg/L		0.2		E200.7	05/20/09 13:06 / tao
Silver	ND	mg/L		0.0005		E200.8	05/19/09 17:32 / aje
Sodium	13	mg/L		1		E200.7	05/19/09 16:20 / tao
Thallium	ND	mg/L		0.0002		E200.8	05/19/09 17:32 / aje
Uranium	ND	mg/L		0.0003		E200.8	05/19/09 17:32 / aje
Zinc	ND	mg/L		0.01		E200.8	05/19/09 17:32 / aje
METALS, TOTAL							
Iron, Ferric	3.46	mg/L		0.03		Calculation	05/27/09 09:16 / rls
METALS, TOTAL RECOVERABLE							
Iron	11.3	mg/L		0.05		E200.7	05/21/09 11:56 / tao
FIELD PARAMETERS							
Field pH, su	7.3	s.u.				FIELD	05/12/09 10:30 / ---
Field Temperature, C	8.0	C				FIELD	05/12/09 10:30 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-002
Client Sample ID: MW 95-5

Report Date: 05/29/09
Collection Date: 05/12/09 13:45
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	5.9	s.u.		0.1		A4500-H B	05/18/09 14:52 / nlh
Conductivity	82	umhos/cm		1		A2510 B	05/19/09 11:13 / nlh
Solids, Total Dissolved TDS @ 180 C	97	mg/L		10		A2540 C	05/19/09 15:29 / afb
INORGANICS							
Alkalinity, Total as CaCO3	31	mg/L		4		A2320 B	05/18/09 23:04 / ehb
Bicarbonate as HCO3	38	mg/L		4		A2320 B	05/18/09 23:04 / ehb
Carbonate as CO3	ND	mg/L		4		A2320 B	05/18/09 23:04 / ehb
Chloride	ND	mg/L		1		E300.0	05/20/09 20:43 / kh
Sulfate	3	mg/L		1		E300.0	05/20/09 20:43 / kh
Fluoride	ND	mg/L		0.1		A4500-F C	05/18/09 23:04 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	05/19/09 09:30 / afb
Hardness as CaCO3	20	mg/L		1		A2340 B	05/19/09 16:24 / rls
Sodium Adsorption Ratio (SAR)	0.51			0.01		Calculation	05/19/09 16:24 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	2.4	mg/L		0.5		A5310 C	05/20/09 11:23 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	ND	mg/L		0.05		E350.1	05/19/09 10:51 / sam
Nitrogen, Nitrate+Nitrite as N	0.09	mg/L		0.01		E353.2	05/18/09 14:28 / bls
Phosphorus, Orthophosphate as P	0.016	mg/L	LH	0.005		E365.1	05/19/09 15:52 / mlm
- The sample was received after the holding time had expired for Orthophosphat.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5		E245.7	05/27/09 17:32 / eli-c
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.7	05/19/09 16:24 / tao
Antimony	ND	mg/L		0.003		E200.8	05/19/09 18:30 / aje
Arsenic	ND	mg/L		0.001		E200.8	05/19/09 18:30 / aje
Barium	0.016	mg/L		0.005		E200.8	05/19/09 18:30 / aje
Beryllium	ND	mg/L		0.001		E200.8	05/19/09 18:30 / aje
Cadmium	0.00089	mg/L		0.00008		E200.8	05/19/09 18:30 / aje
Calcium	6	mg/L		1		E200.7	05/19/09 16:24 / tao
Chromium	0.003	mg/L		0.001		E200.8	05/19/09 18:30 / aje
Copper	0.002	mg/L		0.001		E200.8	05/19/09 18:30 / aje
Iron	0.18	mg/L		0.05		E200.7	05/19/09 16:24 / tao
Iron, Ferrous	0.18	mg/L		0.05		E200.7	05/19/09 16:24 / tao
Lead	ND	mg/L		0.0005		E200.8	05/19/09 18:30 / aje
Magnesium	1	mg/L		1		E200.7	05/19/09 16:24 / tao
Manganese	0.457	mg/L		0.005		E200.7	05/19/09 16:24 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
H - Analysis performed past recommended holding time.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
L - Lowest available reporting limit for the analytical method used.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-002
Client Sample ID: MW 95-5

Report Date: 05/29/09
Collection Date: 05/12/09 13:45
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	05/19/09 16:24 / tao
Potassium	5	mg/L		1		E200.7	05/19/09 16:24 / tao
Selenium	0.001	mg/L		0.001		E200.8	05/20/09 14:06 / aje
Silicon	28.4	mg/L		0.1		E200.7	05/20/09 13:18 / tao
Silica	60.8	mg/L		0.2		E200.7	05/20/09 13:18 / tao
Silver	ND	mg/L		0.0005		E200.8	05/19/09 18:30 / aje
Sodium	5	mg/L		1		E200.7	05/19/09 16:24 / tao
Thallium	ND	mg/L		0.0002		E200.8	05/19/09 18:30 / aje
Uranium	ND	mg/L		0.0003		E200.8	05/19/09 18:30 / aje
Zinc	ND	mg/L		0.01		E200.8	05/19/09 18:30 / aje
METALS, TOTAL							
Iron, Ferric	0.20	mg/L		0.03		Calculation	05/27/09 09:16 / rls
METALS, TOTAL RECOVERABLE							
Iron	0.38	mg/L		0.05		E200.7	05/21/09 12:00 / tao
FIELD PARAMETERS							
Field pH, su	6.04	s.u.				FIELD	05/12/09 13:45 / ---
Field Temperature, C	8.8	C				FIELD	05/12/09 13:45 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-003
Client Sample ID: Adit Pipe

Report Date: 05/29/09
Collection Date: 05/12/09 15:15
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.8	s.u.		0.1		A4500-H B	05/18/09 14:52 / nlh
Conductivity	269	umhos/cm		1		A2510 B	05/18/09 15:56 / nlh
Solids, Total Dissolved TDS @ 180 C	156	mg/L		10		A2540 C	05/18/09 16:21 / afb
INORGANICS							
Alkalinity, Total as CaCO3	79	mg/L		4		A2320 B	05/18/09 23:12 / ehb
Bicarbonate as HCO3	96	mg/L		4		A2320 B	05/18/09 23:12 / ehb
Carbonate as CO3	ND	mg/L		4		A2320 B	05/18/09 23:12 / ehb
Chloride	ND	mg/L		1		E300.0	05/20/09 20:55 / kh
Sulfate	17	mg/L		1		E300.0	05/20/09 20:55 / kh
Fluoride	ND	mg/L		0.1		A4500-F C	05/18/09 23:12 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	05/19/09 09:30 / afb
Hardness as CaCO3	107	mg/L		1		A2340 B	05/19/09 16:28 / rls
Sodium Adsorption Ratio (SAR)	0.18			0.01		Calculation	05/19/09 16:28 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	ND	mg/L		0.5		A5310 C	05/20/09 11:33 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	2.33	mg/L		0.05		E350.1	05/19/09 10:52 / sam
Nitrogen, Nitrate+Nitrite as N	7.99	mg/L	D	0.02		E353.2	05/18/09 14:29 / bls
Phosphorus, Orthophosphate as P	0.096	mg/L	LH	0.005		E365.1	05/19/09 15:53 / mlm
- The sample was received after the holding time had expired for Orthophosphat.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5		E245.7	05/27/09 17:34 / eli-c
METALS, DISSOLVED							
Aluminum	0.09	mg/L		0.03		E200.7	05/19/09 16:28 / tao
Antimony	0.015	mg/L		0.003		E200.8	05/19/09 18:37 / aje
Arsenic	0.002	mg/L		0.001		E200.8	05/19/09 18:37 / aje
Barium	0.070	mg/L		0.005		E200.8	05/19/09 18:37 / aje
Beryllium	ND	mg/L		0.001		E200.7	05/19/09 16:28 / tao
Cadmium	0.00220	mg/L		0.00008		E200.8	05/19/09 18:37 / aje
Calcium	32	mg/L		1		E200.7	05/19/09 16:28 / tao
Chromium	ND	mg/L		0.001		E200.8	05/19/09 18:37 / aje
Copper	0.041	mg/L		0.001		E200.8	05/19/09 18:37 / aje
Iron	0.05	mg/L		0.05		E200.7	05/19/09 16:28 / tao
Iron, Ferrous	0.05	mg/L		0.05		E200.7	05/19/09 16:28 / tao
Lead	0.0021	mg/L		0.0005		E200.8	05/19/09 18:37 / aje
Magnesium	7	mg/L		1		E200.7	05/19/09 16:28 / tao
Manganese	0.163	mg/L		0.005		E200.7	05/19/09 16:28 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-003
Client Sample ID: Adit Pipe

Report Date: 05/29/09
Collection Date: 05/12/09 15:15
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	05/19/09 16:28 / tao
Potassium	2	mg/L		1		E200.7	05/19/09 16:28 / tao
Selenium	0.003	mg/L		0.001		E200.8	05/20/09 14:33 / aje
Silicon	3.6	mg/L		0.1		E200.8	05/20/09 14:33 / aje
Silica	7.7	mg/L		0.2		E200.8	05/20/09 14:33 / aje
Silver	ND	mg/L		0.0005		E200.8	05/19/09 18:37 / aje
Sodium	4	mg/L		1		E200.7	05/19/09 16:28 / tao
Thallium	ND	mg/L		0.0002		E200.8	05/19/09 18:37 / aje
Uranium	0.0013	mg/L		0.0003		E200.8	05/19/09 18:37 / aje
Zinc	ND	mg/L		0.01		E200.8	05/19/09 18:37 / aje
METALS, TOTAL							
Iron, Ferric	4.24	mg/L		0.03		Calculation	05/27/09 09:16 / rls
METALS, TOTAL RECOVERABLE							
Iron	4.29	mg/L		0.05		E200.7	05/21/09 12:04 / tao
FIELD PARAMETERS							
Field pH, su	7.52	s.u.				FIELD	05/12/09 15:15 / ---
Field Temperature, C	6.04	C				FIELD	05/12/09 15:15 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-004
Client Sample ID: Adit Ditch

Report Date: 05/29/09
Collection Date: 05/12/09 15:30
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.8	s.u.		0.1		A4500-H B	05/18/09 14:52 / nlh
Conductivity	169	umhos/cm		1		A2510 B	05/18/09 15:55 / nlh
Solids, Total Dissolved TDS @ 180 C	96	mg/L		10		A2540 C	05/18/09 16:20 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	42	mg/L		4		A2320 B	05/18/09 23:20 / ehb
Bicarbonate as HCO ₃	52	mg/L		4		A2320 B	05/18/09 23:20 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	05/18/09 23:20 / ehb
Chloride	ND	mg/L		1		E300.0	05/20/09 21:06 / kh
Sulfate	17	mg/L		1		E300.0	05/20/09 21:06 / kh
Fluoride	ND	mg/L		0.1		A4500-F C	05/18/09 23:20 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	05/19/09 09:30 / afb
Hardness as CaCO ₃	57	mg/L		1		A2340 B	05/19/09 16:39 / rls
Sodium Adsorption Ratio (SAR)	0.20			0.01		Calculation	05/19/09 16:39 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	ND	mg/L		0.5		A5310 C	05/20/09 11:42 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	1.59	mg/L		0.05		E350.1	05/19/09 10:54 / sam
Nitrogen, Nitrate+Nitrite as N	3.22	mg/L		0.01		E353.2	05/18/09 14:31 / bls
Phosphorus, Orthophosphate as P	0.051	mg/L	LH	0.005		E365.1	05/19/09 15:54 / mlm
- The sample was received after the holding time had expired for Orthophosphat.							
METALS - DISSOLVED							
Mercury	5	ng/L		5		E245.7	05/27/09 17:36 / eli-c
METALS, DISSOLVED							
Aluminum	0.06	mg/L		0.03		E200.7	05/19/09 16:39 / tao
Antimony	0.008	mg/L		0.003		E200.8	05/19/09 18:44 / aje
Arsenic	0.001	mg/L		0.001		E200.8	05/19/09 18:44 / aje
Barium	0.065	mg/L		0.005		E200.8	05/19/09 18:44 / aje
Beryllium	ND	mg/L		0.001		E200.8	05/19/09 18:44 / aje
Cadmium	0.00087	mg/L		0.00008		E200.8	05/19/09 18:44 / aje
Calcium	17	mg/L		1		E200.7	05/19/09 16:39 / tao
Chromium	ND	mg/L		0.001		E200.8	05/19/09 18:44 / aje
Copper	0.084	mg/L		0.001		E200.8	05/19/09 18:44 / aje
Iron	ND	mg/L		0.05		E200.7	05/19/09 16:39 / tao
Iron, Ferrous	ND	mg/L		0.05		E200.7	05/19/09 16:39 / tao
Lead	0.0080	mg/L		0.0005		E200.8	05/19/09 18:44 / aje
Magnesium	3	mg/L		1		E200.7	05/19/09 16:39 / tao
Manganese	0.312	mg/L		0.005		E200.7	05/19/09 16:39 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
H - Analysis performed past recommended holding time.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
L - Lowest available reporting limit for the analytical method used.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-004
Client Sample ID: Adit Ditch

Report Date: 05/29/09
Collection Date: 05/12/09 15:30
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	05/19/09 16:39 / tao
Potassium	2	mg/L		1		E200.7	05/19/09 16:39 / tao
Selenium	0.001	mg/L		0.001		E200.8	05/20/09 14:40 / aje
Silicon	4.4	mg/L		0.1		E200.8	05/20/09 14:40 / aje
Silica	9.5	mg/L		0.2		E200.8	05/20/09 14:40 / aje
Silver	ND	mg/L		0.0005		E200.8	05/19/09 18:44 / aje
Sodium	3	mg/L		1		E200.7	05/19/09 16:39 / tao
Thallium	ND	mg/L		0.0002		E200.8	05/19/09 18:44 / aje
Uranium	ND	mg/L		0.0003		E200.8	05/19/09 18:44 / aje
Zinc	ND	mg/L		0.01		E200.8	05/19/09 18:44 / aje
METALS, TOTAL							
Iron, Ferric	8.51	mg/L		0.03		Calculation	05/27/09 09:16 / rls
METALS, TOTAL RECOVERABLE							
Iron	8.51	mg/L		0.05		E200.7	05/21/09 12:08 / tao
FIELD PARAMETERS							
Field pH, su	7.53	s.u.				FIELD	05/12/09 15:30 / ---
Field Temperature, C	5.25	C				FIELD	05/12/09 15:30 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-005
Client Sample ID: MW-01-15

Report Date: 05/29/09
Collection Date: 05/13/09 15:20
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.3	s.u.		0.1		A4500-H B	05/18/09 14:52 / nlh
Conductivity	424	umhos/cm		1		A2510 B	05/18/09 15:57 / nlh
Solids, Total Dissolved TDS @ 180 C	256	mg/L		10		A2540 C	05/18/09 16:23 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	83	mg/L		4		A2320 B	05/18/09 23:28 / ehb
Bicarbonate as HCO ₃	102	mg/L		4		A2320 B	05/18/09 23:28 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	05/18/09 23:28 / ehb
Chloride	4	mg/L		1		E300.0	05/20/09 21:18 / kh
Sulfate	31	mg/L		1		E300.0	05/20/09 21:18 / kh
Fluoride	0.2	mg/L		0.1		A4500-F C	05/18/09 23:28 / ehb
Sulfide	0.05	mg/L		0.04		A4500 S-D	05/19/09 09:30 / afb
Hardness as CaCO ₃	71	mg/L		1		A2340 B	05/19/09 16:43 / rls
Sodium Adsorption Ratio (SAR)	1.66			0.01		Calculation	05/19/09 16:43 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	1.6	mg/L		0.5		A5310 C	05/20/09 11:53 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	3.83	mg/L		0.05		E350.1	05/19/09 10:55 / sam
Nitrogen, Nitrate+Nitrite as N	16.8	mg/L	D	0.02		E353.2	05/18/09 14:32 / bls
Phosphorus, Orthophosphate as P	0.073	mg/L	LH	0.005		E365.1	05/19/09 15:55 / mlm
- The sample was received after the holding time had expired for Orthophosphat.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5		E245.7	05/27/09 17:38 / eli-c
METALS, DISSOLVED							
Aluminum	0.04	mg/L		0.03		E200.8	05/19/09 18:51 / aje
Antimony	0.032	mg/L		0.003		E200.8	05/19/09 18:51 / aje
Arsenic	ND	mg/L		0.001		E200.8	05/19/09 18:51 / aje
Barium	0.101	mg/L		0.005		E200.8	05/19/09 18:51 / aje
Beryllium	ND	mg/L		0.001		E200.7	05/19/09 16:43 / tao
Cadmium	0.00090	mg/L		0.00008		E200.8	05/19/09 18:51 / aje
Calcium	22	mg/L		1		E200.7	05/19/09 16:43 / tao
Chromium	0.002	mg/L		0.001		E200.8	05/19/09 18:51 / aje
Copper	0.068	mg/L		0.001		E200.8	05/19/09 18:51 / aje
Iron	0.06	mg/L		0.05		E200.7	05/19/09 16:43 / tao
Iron, Ferrous	0.06	mg/L		0.05		E200.7	05/19/09 16:43 / tao
Lead	0.0011	mg/L		0.0005		E200.8	05/19/09 18:51 / aje
Magnesium	4	mg/L		1		E200.7	05/19/09 16:43 / tao
Manganese	0.455	mg/L		0.005		E200.7	05/19/09 16:43 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-005
Client Sample ID: MW-01-15

Report Date: 05/29/09
Collection Date: 05/13/09 15:20
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	05/19/09 16:43 / tao
Potassium	22	mg/L		1		E200.7	05/19/09 16:43 / tao
Selenium	ND	mg/L		0.001		E200.8	05/19/09 18:51 / aje
Silicon	5.1	mg/L		0.1		E200.7	05/20/09 13:34 / tao
Silica	11.0	mg/L		0.2		E200.7	05/20/09 13:34 / tao
Silver	ND	mg/L		0.0005		E200.8	05/19/09 18:51 / aje
Sodium	32	mg/L		1		E200.7	05/19/09 16:43 / tao
Thallium	ND	mg/L		0.0002		E200.8	05/19/09 18:51 / aje
Uranium	0.0020	mg/L		0.0003		E200.8	05/19/09 18:51 / aje
Zinc	0.01	mg/L		0.01		E200.7	05/19/09 16:43 / tao
METALS, TOTAL							
Iron, Ferric	3.31	mg/L		0.03		Calculation	05/27/09 09:16 / rls
METALS, TOTAL RECOVERABLE							
Iron	3.37	mg/L		0.05		E200.7	05/21/09 12:12 / tao
FIELD PARAMETERS							
Field pH, su	7.24	s.u.				FIELD	05/13/09 15:20 / ---
Field Temperature, C	10.85	C				FIELD	05/13/09 15:20 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-006
Client Sample ID: Decant 1

Report Date: 05/29/09
Collection Date: 05/14/09 08:30
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.6	s.u.		0.1		A4500-H B	05/18/09 14:52 / nlh
Conductivity	427	umhos/cm		1		A2510 B	05/18/09 16:00 / nlh
Solids, Total Dissolved TDS @ 180 C	305	mg/L		10		A2540 C	05/18/09 16:26 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	82	mg/L		4		A2320 B	05/18/09 23:36 / ehb
Bicarbonate as HCO ₃	99	mg/L		4		A2320 B	05/18/09 23:36 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	05/18/09 23:36 / ehb
Chloride	4	mg/L		1		E300.0	05/20/09 21:30 / kh
Sulfate	33	mg/L		1		E300.0	05/20/09 21:30 / kh
Fluoride	0.2	mg/L		0.1		A4500-F C	05/18/09 23:36 / ehb
Sulfide	0.04	mg/L		0.04		A4500 S-D	05/19/09 09:30 / afb
Hardness as CaCO ₃	59	mg/L		1		A2340 B	05/19/09 16:47 / rls
Sodium Adsorption Ratio (SAR)	1.73			0.01		Calculation	05/19/09 16:47 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	1.6	mg/L		0.5		A5310 C	05/20/09 12:04 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	6.0	mg/L	D	0.1		E350.1	05/21/09 10:47 / sam
Nitrogen, Nitrate+Nitrite as N	17.6	mg/L	D	0.02		E353.2	05/18/09 14:33 / bls
Phosphorus, Orthophosphate as P	0.134	mg/L	LH	0.005		E365.1	05/19/09 15:56 / mlm
- The sample was received after the holding time had expired for Orthophosphat.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5		E245.7	05/27/09 17:40 / eli-c
METALS, DISSOLVED							
Aluminum	0.11	mg/L		0.03		E200.8	05/19/09 18:58 / aje
Antimony	0.046	mg/L		0.003		E200.8	05/19/09 18:58 / aje
Arsenic	0.002	mg/L		0.001		E200.8	05/19/09 18:58 / aje
Barium	0.097	mg/L		0.005		E200.8	05/19/09 18:58 / aje
Beryllium	ND	mg/L		0.001		E200.8	05/19/09 18:58 / aje
Cadmium	0.00089	mg/L		0.00008		E200.8	05/19/09 18:58 / aje
Calcium	17	mg/L		1		E200.7	05/19/09 16:47 / tao
Chromium	ND	mg/L		0.001		E200.8	05/19/09 18:58 / aje
Copper	0.028	mg/L		0.001		E200.8	05/19/09 18:58 / aje
Iron	0.06	mg/L		0.05		E200.7	05/19/09 16:47 / tao
Iron, Ferrous	0.06	mg/L		0.05		E200.7	05/19/09 16:47 / tao
Lead	0.0026	mg/L		0.0005		E200.8	05/19/09 18:58 / aje
Magnesium	4	mg/L		1		E200.7	05/19/09 16:47 / tao
Manganese	0.554	mg/L		0.005		E200.7	05/19/09 16:47 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-006
Client Sample ID: Decant 1

Report Date: 05/29/09
Collection Date: 05/14/09 08:30
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	05/19/09 16:47 / tao
Potassium	27	mg/L		1		E200.7	05/19/09 16:47 / tao
Selenium	0.001	mg/L		0.001		E200.8	05/20/09 15:15 / aje
Silicon	4.0	mg/L		0.1		E200.8	05/20/09 15:15 / aje
Silica	8.5	mg/L		0.2		E200.8	05/20/09 15:15 / aje
Silver	ND	mg/L		0.0005		E200.8	05/19/09 18:58 / aje
Sodium	31	mg/L		1		E200.7	05/19/09 16:47 / tao
Thallium	ND	mg/L		0.0002		E200.8	05/19/09 18:58 / aje
Uranium	0.0029	mg/L		0.0003		E200.8	05/19/09 18:58 / aje
Zinc	ND	mg/L		0.01		E200.8	05/19/09 18:58 / aje
METALS, TOTAL							
Iron, Ferric	2.96	mg/L		0.03		Calculation	05/27/09 09:16 / rls
METALS, TOTAL RECOVERABLE							
Iron	3.02	mg/L		0.05		E200.7	05/21/09 12:24 / tao
FIELD PARAMETERS							
Field pH, su	7.45	s.u.				FIELD	05/14/09 08:30 / ---
Field Temperature, C	9.5	C				FIELD	05/14/09 08:30 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-007
Client Sample ID: Decant

Report Date: 05/29/09
Collection Date: 05/14/09 08:45
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.7	s.u.		0.1		A4500-H B	05/18/09 14:52 / nlh
Conductivity	429	umhos/cm		1		A2510 B	05/18/09 16:01 / nlh
Solids, Total Dissolved TDS @ 180 C	314	mg/L		10		A2540 C	05/18/09 16:27 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	82	mg/L		4		A2320 B	05/18/09 23:59 / ehb
Bicarbonate as HCO ₃	100	mg/L		4		A2320 B	05/18/09 23:59 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	05/18/09 23:59 / ehb
Chloride	4	mg/L		1		E300.0	05/21/09 00:01 / kh
Sulfate	33	mg/L		1		E300.0	05/21/09 00:01 / kh
Fluoride	0.2	mg/L		0.1		A4500-F C	05/18/09 23:59 / ehb
Sulfide	0.10	mg/L		0.04		A4500 S-D	05/19/09 09:30 / afb
Hardness as CaCO ₃	59	mg/L		1		A2340 B	05/19/09 16:51 / rls
Sodium Adsorption Ratio (SAR)	1.76			0.01		Calculation	05/19/09 16:51 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	1.6	mg/L		0.5		A5310 C	05/20/09 12:16 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	6.0	mg/L	D	0.1		E350.1	05/21/09 10:48 / sam
Nitrogen, Nitrate+Nitrite as N	19.4	mg/L	D	0.02		E353.2	05/18/09 14:34 / bls
Phosphorus, Orthophosphate as P	0.169	mg/L	LH	0.005		E365.1	05/19/09 15:57 / mlm
- The sample was received after the holding time had expired for Orthophosphat.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5		E245.7	05/27/09 17:44 / eli-c
METALS, DISSOLVED							
Aluminum	0.12	mg/L		0.03		E200.7	05/19/09 16:51 / tao
Antimony	0.047	mg/L		0.003		E200.8	05/19/09 19:05 / aje
Arsenic	0.002	mg/L		0.001		E200.8	05/19/09 19:05 / aje
Barium	0.099	mg/L		0.005		E200.8	05/19/09 19:05 / aje
Beryllium	ND	mg/L		0.001		E200.7	05/19/09 16:51 / tao
Cadmium	0.00091	mg/L		0.00008		E200.8	05/19/09 19:05 / aje
Calcium	17	mg/L		1		E200.7	05/19/09 16:51 / tao
Chromium	ND	mg/L		0.001		E200.8	05/19/09 19:05 / aje
Copper	0.024	mg/L		0.001		E200.8	05/19/09 19:05 / aje
Iron	ND	mg/L		0.05		E200.7	05/19/09 16:51 / tao
Iron, Ferrous	ND	mg/L		0.05		E200.7	05/19/09 16:51 / tao
Lead	0.0016	mg/L		0.0005		E200.8	05/19/09 19:05 / aje
Magnesium	4	mg/L		1		E200.7	05/19/09 16:51 / tao
Manganese	0.543	mg/L		0.005		E200.7	05/19/09 16:51 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Copper Attenuation - Scott Mason, Hydrometrics
Lab ID: B09051541-007
Client Sample ID: Decant

Report Date: 05/29/09
Collection Date: 05/14/09 08:45
Date Received: 05/18/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	05/19/09 16:51 / tao
Potassium	27	mg/L		1		E200.7	05/19/09 16:51 / tao
Selenium	ND	mg/L		0.001		E200.8	05/20/09 15:22 / aje
Silicon	3.9	mg/L		0.1		E200.8	05/20/09 15:22 / aje
Silica	8.3	mg/L		0.2		E200.8	05/20/09 15:22 / aje
Silver	ND	mg/L		0.0005		E200.8	05/19/09 19:05 / aje
Sodium	31	mg/L		1		E200.7	05/19/09 16:51 / tao
Thallium	ND	mg/L		0.0002		E200.8	05/19/09 19:05 / aje
Uranium	0.0030	mg/L		0.0003		E200.8	05/19/09 19:05 / aje
Zinc	ND	mg/L		0.01		E200.7	05/19/09 16:51 / tao
METALS, TOTAL							
Iron, Ferric	2.95	mg/L		0.03		Calculation	05/27/09 09:16 / rls
METALS, TOTAL RECOVERABLE							
Iron	2.95	mg/L		0.05		E200.7	05/21/09 12:28 / tao
FIELD PARAMETERS							
Field pH, su	7.45	s.u.				FIELD	05/14/09 08:45 / ---
Field Temperature, C	9.5	C				FIELD	05/14/09 08:45 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



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www.hydrometrics.com

June 10, 2009

MEMO TO FILE

RE: COPPER ATTENUATION SAMPLING – TROY MINE

Three samples were taken to provide additional data for the copper attenuation modeling of the tailing facility at the Genesis Mine (Troy Mine). One set of samples was collected from each of the monitor wells MW01-15 and MW95-8. Monitor well MW01-16 was also checked and found to be dry. The third sample was taken from the Decant Pond. Sample field sheets are attached.

Recovery rates in both of sampled wells (MW01-15 and MW95-8) were moderate to slow within the screened interval, but full recovery is extremely slow. Well MW01-15 didn't produce enough water to maintain SWL while pumping at a rate of <0.1 gpm. Details of the purging and sampling of the wells are included below.

Samples were shipped to Energy Labs in Billings with a request for a rush turnaround.

Monitor Well MW01-15

The original completion notes for this well indicate a total well depth of 45 feet with screen from 30 to 40 feet below the ground surface. A lift has been put on the tailing divider dike on which the well is located since the well was installed. The sounded depth for the well is 53.25 feet with a 1.5 foot stick up, implying the screen is now at about 37 to 47 feet. Static water level (SWL) in the well was 44.54. The Rediflo pump was set at 50 feet. Initial pumping of the well at approximately 0.25 gpm drew the SWL to 50.5 feet. The SWL recovered to 44.7 feet in 42 minutes. The well was pumped down and allowed to recover five times yielding approximately 3.5 gallons. Initial recovery rates after purging the well were up to one foot of SWL recovery in four minutes. Recovery rates significantly declined after 2 to 2.5 feet of SWL recovery.

The sample was taken under low flow (<0.1 gpm) conditions. The SWL in the well was drawn down approximately 2 feet while sampling. Field parameters were taken after the sample was taken because it was not known if the well could be pumped long enough to allow stabilization of field parameters before the sample was taken without drawing the SWL down to the pump.

Monitor Well MW95-8

Completion notes for this well indicate a total depth of 54.5 feet with the screen at 48 to 53 feet (presumably below ground surface). The sounded depth of the well was 64.18

feet, reflecting additional casing added to the well because of a lift being put on the divider dike where the well is located. The initial SWL was 50.06 feet. Approximately 3.5 gallons of water were purged with a bailer, drawing the SWL down to 59 feet. The well was left to recover overnight and recovered to 51.0 feet. Another 2 gallons were bailed, drawing the SWL down to a SWL of 59.0 feet. The pump was put in the well at a depth of 60 feet, in preparation for sampling. SWL was measured at 58.25 after installing the pump. SWL showed no recovery after 2.5 hours. Pump was pulled from the well due to a lack of water in the well and poor recovery and a sample was taken using a bailer. SWL at initiation of bailing for the sample was 58.27 feet, and taking the sample drew the well down to 59.24 feet. Well recovered to 59.01 feet in ten minutes after taking the sample.

Monitoring Well MW01-16

Completion notes for this well indicate a total depth of 42 feet with screen from 32 to 42 feet. The well was sounded to a depth of 42.88 feet and was dry.

Decant Pond

A sample from the decant pond was taken from a point adjacent to the access ramp to the pump dock. The pumps were not running at the time of sampling.



Sampling MW01-15: Decant pond in background. Irrigation well visible on far left side of decant pond.



Monitor Well MW01-16



MW01-16 next to dike road, looking toward M95-8 on dike road on far side of decant pond



Monitor Well MW01-16. Looking back toward MW00-15

Water Sampling Form ~ HF-430

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: TRC
Project Code: _____
Sample Team Member(s): LS/DK
Laboratory Used: ECI

Site Designation: WW 58
Sample Code Number: _____
Sample Date: 6/8/09
Sample Time: 12:15 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: _____
Duplicate Sample Time: _____

Site Conditions

New Site: Yes No Photo taken: Yes No
Site Type: DRY surface water process water
monitoring well domestic well adit seep
spring- other: _____

Weather Conditions: calm breeze windy
no precip. rain snow
clear p. cloudy overcast
Air Temperature: _____ °C _____ °F

For Groundwater Samples

well volume formula:	$V = (TD-SWL) \times (\text{Dia.})^2 / 25$	Comments
TD (ft):	<u>60.18</u>	
SWL (ft):	<u>50.10</u>	no access/pumping
Casing Diameter (I.D.):	<u>7"</u>	
Water Volume (V) (gal):	<u>2.3</u>	
x 3=(gal.)	<u>6.8</u>	
Actual Vol. Removed (gal.):	<u>5.5</u>	
Water Level Recovery:	slow moderate rapid	

For Surface Water Samples

Flow Method:	Marsh McBirney Volumetric Flume Weir Estimate
Other Flow or Description:	_____
Flow:	gpm cfs Staff Gage:

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes

Turbidity: clear moderate slight very

Sample Method: grab composite pump bailer other

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>128.8</u>	
DO (mg/l)	<u>4.52</u>	
pH	<u>7.57</u>	
SC (µmhos/cm)	<u>532</u>	
Turbidity (ntu)		
H ₂ O Tmp. (°C)	<u>10.2</u>	
Color	<u>6.8</u>	
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1	500 ml	F or UF		COLD	11/11/09
1	500 ml	F or UF		TDS	
1	250 ml		HNO ₃	DISS. MET	
1	250 ml		HCl		
1	250 ml	F or UF	HNO ₃	TOT ALK	
1	250 ml	F or UF	H ₂ SO ₄	SULF	
1	250 ml	F or UF	H ₂ SO ₄	SULF	
1	500 ml	F or UF	HNO ₃	CT	

Comments: Bailed ~ 3.5 gal : SW → 59.00 : let well sit overnight → 59.00
a pter 1 bailer : Bailed to 51" - let sit for 2.5 hours → 10
Pressure: Run pump 75 sec @ 58.27" - took sample → 59.27
Water 10 min SWL → 59.01

Sample Team Member Signature: _____

Page 1 of 1

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: Day MCB
 Project Code: _____
 Sample Team Member(s): LS/PE
 Laboratory Used: 21

Site Designation: MMU-01-15
 Sample Code Number: _____
 Sample Date: 6/2/09
 Sample Time: 17:00 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: _____
 Duplicate Sample Time: _____

Site Conditions

New Site: Yes ☐ No ☒ Photo taken: Yes ☐ No ☒
 Site Type: DRY surface water process water
 Monitoring well, domestic well, adit, seep
 spring, other: _____
 Weather Conditions: calm breeze windy
 no precip. rain snow
 clear p. cloudy overcast
 Air Temperature: _____ °C 50 °F

For Groundwater Samples

well volume formula: $V = (TD-SWL) \times (\text{Dia.}^2) / 25$	Comments
TD (ft): <u>53.25</u>	no access/pumping
SWL (ft): <u>44.54</u>	
Casing Diameter (I.D.): <u>2</u>	
Water Volume (V) (gal): <u>1.4</u>	
x 3=(gal.): <u>4.2</u>	
Actual Vol. Removed (gal.): <u>3.5</u>	
Water Level Recovery: slow moderate rapid	

For Surface Water Samples

Flow Method: Marsh McBirney Volumetric Flume Weir Estimate
 Other Flow or Description: _____
 Flow: gpm cfs Staff Gage: _____

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes
16:31	154.4	5.86	7.85	431	—	8.70	
16:34	153.0	5.79	7.86	423	—	8.79	
17:00	127.9	5.10	7.66	424	—	14.10	X AFTER
17:11	142.6	4.35	7.31	421	—	15.69	SAMPLES

Turbidity: (circle) clear slight moderate very

Sample Method: grab composite pump bailer other
 (describe) ROPER

Field Parameters

	Sample	Duplicate
ORP (mV)	127.9	
DO (mg/l)	4.35	
pH	7.31	
SC (µmhos/cm)	421	
Turbidity (ntu)		
H ₂ O Tmp. (°C)	14.10	
Color		
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1	500 ml	F or UF	—	Calc. M, Al, Fe, P	
1	500 ml	F or UF	—	DS	
1	250 ml	F or UF	HNO ₃	DISS M	
1	250 ml	F or UF	HNO ₃	Hg	
1	250 ml	F or UF	HNO ₃	TOTAL M	
1	250 ml	F or UF	HNO ₃	SE	ALERT'S
1	250 ml	F or UF	HNO ₃	SE	
1	500 ml	F or UF	HNO ₃	SE	

Comments: *Same additional csg has been added since construction i
pumped a 2.25 gal flow due to 50' pump set to 50' flow 12:28
13:10 well rec'd to 44.7 from down to 44.13:15: pumped to 50' csg
3 times 3.5 gal. Pumping @ 40.00 gpm & sue csgs up or is 5' csg
 Sample Team Member Signature: Long Page 1 of 1

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: TRC
 Project Code: _____
 Sample Team Member(s): LS JK
 Laboratory Used: ELI

Site Designation: MW 38
 Sample Code Number: _____
 Sample Date: 6/8/09
 Sample Time: 12:15 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: _____
 Duplicate Sample Time: _____

Site Conditions

New Site: Yes No Photo taken: Yes No
 Site Type: DRY surface water process water
 monitoring well domestic well adit seep
 spring- other: _____

Weather Conditions: calm breeze windy
 no precip. rain snow
 clear p. cloudy overcast

Air Temperature: _____ °C _____ °F

For Groundwater Samples

well volume formula: $V = (TD-SWL) \times (\text{Dia.})^2$	25	Comments
TD (ft):	<u>60.18</u> <u>64.18</u>	no access/pumping
SWL (ft):	<u>50.18</u> <u>50.06</u>	
Casing Diameter (I.D.):	<u>7"</u>	
Water Volume (V) (gal):	<u>2.3</u>	
x 3=(gal.)	<u>6.8</u>	
Actual Vol. Removed (gal.):		
Water Level Recovery:	slow moderate rapid	

For Surface Water Samples

Flow Method:	Marsh McBirney Volumetric Flume Weir Estimate
Other Flow or Description:	
Flow:	gpm cfs Staff Gage:

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes

Turbidity: (circle) clear slight moderate very

Sample Method: (describe) grab composite pump bailer other

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>128.8</u>	
DO (mg/l)	<u>4.52</u>	
pH	<u>7.57</u>	
SC (µmhos/cm)	<u>532</u>	
Turbidity (ntu)	<u>2</u>	
H ₂ O Tmp. (°C)	<u>10.2</u>	
Color	<u>448</u>	
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1	500 ml	F or UF	—	COND	MT, A, Fe
1	500 ml	F or UF	—	TDS	
1	250 ml	UF	HNO ₃	DISS	MT
1	250 ml	UF	HCl	Fe	
1	250 ml	F or UF	HNO ₃	TOT	
1	250 ml	F or UF	H ₂ SO ₄	SO ₄	
1	250 ml	F or UF	H ₂ SO ₄	SO ₄	
1	500 ml	F or UF	HNO ₃	CT	

Comments: Failed ~ 3.5 gal : SW → 59.00 : LET USE SW OVERWITE → 59.00
a flar 1 bailer: BAILER TO ~ 51' - let set for 2.5 hours → NO
Pressure: RUN PUMP → 59.00 @ 58.27 : TOOK SAMPLE → 59.00
WATT 10 MIN SWL → 59.01

Sample Team Member Signature: LS JK

Page 1 of 1

STANDARD OPERATING PROCEDURE WATER SAMPLING FORM HF-FORM-430

Project Name: TRAP

Project Code: _____

Sample Team Member(s): CS/PLK

Laboratory Used: _____

Site Designation: MWDI-16

Sample Code Number: _____

Sample Date: 6/8/09Sample Time: 12:50 (military)

If Duplicate Sample Collected,
Please Record Below

Duplicate Sample Code #: _____
Duplicate Sample Time: _____

Site Conditions

New Site: Yes ☒ No ☐ Photo taken: Yes ☒ No ☐
Site Type: DRY ☒ surface water ☐ process water ☐

monitoring well ☒ domestic well ☐ adit ☐ seep ☐
spring other: _____

Weather Conditions: calm ☒ breeze ☐ windy ☐
no precip. ☒ rain ☐ snow ☐
clear ☐ p. cloudy ☐ overcast ☐

Air Temperature: _____ °C _____ °F

For Groundwater Samples

well volume formula: $V = (TD-SWL) \times (\text{Dia})^2$
25
TD (ft): 42.89
SWL (ft): 742.88
Casing Diameter (I.D.): 2"
Water Volume (V) (gal): _____
x 3=(gal.) _____
Actual Vol. Removed (gal.): _____
Water Level Recovery: slow moderate rapid

Comments

no access/pumping

For Surface Water SamplesFlow Method: Marsh-McBirney Volumetric ☐ Flume ☐ Weir ☐ Estimate ☐

Other Flow or Description: _____

Flow: _____ gpm _____ cfs Staff Gage: _____

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes

Turbidity: clear ☐ slight ☐ moderate ☐ very ☐

Sample Method: grab ☐ composite ☐ pump ☐ bailer ☐ other ☐

Field Parameters

Sample Duplicate

ORP (mV) _____
DO (mg/l) _____
pH _____
SC (µmhos/cm) _____
Turbidity (ntu) _____
H₂O Temp. (°C) _____
Color _____
Other: _____

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
	ml	F or UF			
	ml	F or UF			
	ml	F or UF			
	ml	F or UF			
	ml	F or UF			
	ml	F or UF			
	ml	F or UF			
	ml	F or UF			

Comments: well dry at 42.89 TDSample Team Member Signature: [Signature]

Page 1 of 1

Project Name: RAV
Project Code: _____
Sample Team Member(s): LJH/K
Laboratory Used: ELI

Site Designation: DELAUT
Sample Code Number: _____
Sample Date: 6/9/09
Sample Time: 0935 (military)

**If Duplicate Sample Collected,
Please Record Below**

Duplicate Sample Code #: _____
Duplicate Sample Time: _____

Site Conditions

New Site: Yes No Photo taken: Yes No
Site Type: DRY surface water process water
monitoring well domestic well adit seep
spring-other: DELAUT POND
Weather Conditions: calm breeze windy
no precip. rain snow
clear p. cloudy overcast
Air Temperature: _____ °C 60 °F

For Groundwater Samples

well volume formula: $V = (TD-SWL) \times (\text{Dia.}^2) / 25$	Comments
TD (ft): _____	no access/pumping
SWL (ft): _____	
Casing Diameter (I.D.): _____	
Water Volume (V) (gal): _____	
$\times 3 =$ (gal.) _____	
Actual Vol. Removed (gal.): _____	
Water Level Recovery: slow moderate rapid	

For Surface Water Samples

Flow Method: Marsh McBirney Volumetric Flume Weir Estimate
Other Flow or Description: _____
Flow: gpm cfs Staff Gage: _____

Field Parameter Stabilization

Time (military)	Oxidation Reduction Potential (mV)	Dissolved Oxygen (mg/l)	pH	S.C. (µmhos/cm)	Turbidity (n.t.u.)	Temperature (°C)	Additional Parameters or Notes

Turbidity: (circle) clear slight moderate very Sample Method: (describe) grab composite pump bailer other

Field Parameters

	Sample	Duplicate
ORP (mV)	<u>125.4</u>	
DO (mg/l)	<u>6.86</u>	
pH	<u>6.92</u>	
SC (µmhos/cm)	<u>412</u>	
Turbidity (ntu)		
H ₂ O Tmp. (°C)	<u>15.1</u>	
Color	<u>white</u>	
Other:		

Bottles Collected

Quantity	Size	Filter or Unfilt.	Preservative	Parameter	Additional Notes
1	500 ml	F or UF			COMMON TRS
1	500 ml	F or UF			
1	250 ml	For UF	H ₂ SO ₄	DISSALC	
1	250 ml	For UF	HCl	Hg	
1	250 ml	F or UF	H ₂ SO ₄	TOTAL MC	
1	250 ml	F or UF	H ₂ SO ₄	ALUM	
1	250 ml	F or UF	H ₂ SO ₄	5	
1	500 ml	F or UF	H ₂ SO ₄	CH	

Comments: SAMPLE FROM POND DOCK PUMPS NOT RUNNING

Sample Team Member Signature: LJH/K

STANDARD OPERATING PROCEDURE
INSTRUMENT CALIBRATION FORM
HF-FORM-500

Hydrometrics, Inc.
Consulting Scientists and Engineers

Project Name: TRC/ MUE

Project Number: _____

Date: 6/3/09

Personnel: LS/ PK

Weather Conditions:

calm ☒ breeze ☒ windy
no precip ☒ rain ☐ snow
clear ☐ p. cloudy ☒ overcast

Air Temperature:

°F 50 °C

Dissolved Oxygen

Meter Type: _____

SN: _____

Time: 11:47

Calibration Results:

9.12 mg/L

Redox Potential (Eh)

Meter Type: _____

SN: _____

Electrode: Ag/AgCl

Other: _____

Time: 12:02

ZoBell Solution Potential (mV):

226

Temperature:

16.7

pH

Meter Type: _____

SN: _____

Time	Buffer	Temperature	Reading (Temp Corrected)
<u>11:52</u>	<u>4</u>	<u>18.5</u>	<u>4.00</u>
<u>11:55</u>	<u>7</u>	<u>18.1</u>	<u>7.00</u>
	<u>10</u>		
Calibration check:			

Specific Conductance

Meter Type: _____

SN: _____

Time	Standard	Temperature	Reading (Temp Corrected)	Cell Factor
<u>11:54</u>	<u>14.13</u>	<u>18</u>	<u>14.13</u>	
Calibration check:				

Turbidity

Meter Type: _____

SN: _____

Standard: _____

Reading: _____

PID/FID

Meter Type: _____

SN: _____

Span Gas Type: _____

Time	Span Gas		Calibration Memory	Response Factor	Reading
	Gauge Pressure	Concentration			
Calibration Check:					

Notes/Additional Information:

Hydrometrics, Inc.
Consulting Scientists and Engineers

Personnel: 43

°F °C

Calibration Results: 10:47

Temperature: 6.9

Calibration check:

Calibration check:

Reading:

Generation Check:

11/9/2008 12:35 PM
Revised 8/00



ANALYTICAL SUMMARY REPORT

July 31, 2009

Paul Kukay

Genesis Inc Troy Mine

PO Box 1660

Troy, MT 59935-1660

Workorder No.: B09072237

Quote ID: B1850 - Troy Mine Geochemical Modeling

Project Name: Cu Attenuation Study-Scott Mason Hydrometrics

Energy Laboratories Inc received the following 4 samples for Genesis Inc Troy Mine on 7/24/2009 for analysis.

Sample ID	Client Sample ID	Collect Date	Receive Date	Matrix	Test
B09072237-001	MW-95-8	07/21/09 9:10	07/24/09	Aqueous	Metals by ICP/ICPMS, Dissolved Metals by ICP/ICPMS, Tot. Rec. Alkalinity Anions by ion chromatography Conductivity Carbon, Total Organic Mercury, Dissolved Field Parameters Fluoride Hardness as CaCO3 Iron, ferric Nitrogen, Ammonia Nitrogen, Nitrate + Nitrite pH Metals Preparation by EPA 200.2 Phosphorus, Orthophosphate as P Sodium Adsorption Ratio Solids, Total Dissolved Sulfide, Methylene Blue Colorimetric
B09072237-002	Decant Pond	07/21/09 10:15	07/24/09	Aqueous	Same As Above
B09072237-003	MW-01-15	07/21/09 12:50	07/24/09	Aqueous	Same As Above
B09072237-004	MW-00	07/21/09 13:10	07/24/09	Aqueous	Same As Above

Any exceptions or problems with the analyses are noted in the Laboratory Analytical Report, the QA/QC Summary Report, or the Case Narrative.

If you have any questions regarding these tests results, please call.

Report Approved By: _____



CLIENT: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Sample Delivery Group: B09072237

Date: 31-Jul-09

CASE NARRATIVE

Tests Associated with Analyst identified as ELI-CA were subcontracted to Energy Laboratories Casper Branch, EPA Number WY00002.

LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Lab ID: B09072237-001
Client Sample ID: MW-95-8

Report Date: 07/31/09
Collection Date: 07/21/09 09:10
Date Received: 07/24/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.7	s.u.		0.1		A4500-H B	07/24/09 13:17 / nlh
Conductivity	534	umhos/cm		1		A2510 B	07/24/09 13:17 / nlh
Solids, Total Dissolved TDS @ 180 C	346	mg/L		10		A2540 C	07/24/09 15:41 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	139	mg/L		4		A2320 B	07/24/09 16:11 / ehb
Bicarbonate as HCO ₃	169	mg/L		4		A2320 B	07/24/09 16:11 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	07/24/09 16:11 / ehb
Chloride	4	mg/L		1		E300.0	07/27/09 18:07 / kh
Sulfate	41	mg/L		1		E300.0	07/27/09 18:07 / kh
Fluoride	0.1	mg/L		0.1		A4500-F C	07/24/09 16:11 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	07/24/09 12:15 / afb
Hardness as CaCO ₃	204	mg/L		1		A2340 B	07/27/09 11:58 / rls
Sodium Adsorption Ratio (SAR)	0.78			0.01		Calculation	07/27/09 11:58 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY000002)							
Organic Carbon, Total (TOC)	ND	mg/L		0.5		A5310 C	07/28/09 15:58 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	ND	mg/L		0.05		E350.1	07/24/09 15:18 / sam
Nitrogen, Nitrate+Nitrite as N	22.1	mg/L	D	0.03		E353.2	07/24/09 14:28 / bls
Phosphorus, Orthophosphate as P	ND	mg/L	LH	0.005		E365.1	07/24/09 12:22 / sam
- The sample was received after the holding time had expired for Ortho Phosphate.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5.0		E245.7	07/30/09 15:08 / eli-ca
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.8	07/27/09 21:03 / jjw
Antimony	ND	mg/L		0.003		E200.8	07/27/09 21:03 / jjw
Arsenic	ND	mg/L		0.001		E200.8	07/27/09 21:03 / jjw
Barium	0.135	mg/L		0.005		E200.7	07/27/09 11:58 / tao
Beryllium	ND	mg/L		0.001		E200.7	07/27/09 11:58 / tao
Cadmium	0.00148	mg/L		0.00008		E200.8	07/27/09 21:03 / jjw
Calcium	68	mg/L		1		E200.7	07/27/09 11:58 / tao
Chromium	ND	mg/L		0.001		E200.8	07/27/09 21:03 / jjw
Copper	0.007	mg/L		0.001		E200.8	07/27/09 21:03 / jjw
Iron	ND	mg/L		0.05		E200.7	07/27/09 11:58 / tao
Iron, Ferrous	ND	mg/L		0.05		E200.7	07/27/09 11:58 / tao
Lead	ND	mg/L		0.0005		E200.8	07/27/09 21:03 / jjw
Magnesium	8	mg/L		1		E200.7	07/27/09 11:58 / tao
Manganese	ND	mg/L		0.005		E200.7	07/27/09 11:58 / tao

Report Definitions:
 RL - Analyte reporting limit.
 QCL - Quality control limit.
 D - RL increased due to sample matrix interference.
 L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
 ND - Not detected at the reporting limit.
 H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Lab ID: B09072237-001
Client Sample ID: MW-95-8

Report Date: 07/31/09
Collection Date: 07/21/09 09:10
Date Received: 07/24/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	07/27/09 11:58 / tao
Potassium	8	mg/L		1		E200.7	07/27/09 11:58 / tao
Selenium	0.002	mg/L		0.001		E200.8	07/27/09 21:03 / jjw
Silicon	5.4	mg/L		0.1		E200.7	07/27/09 11:58 / tao
Silica	11.7	mg/L		0.2		E200.7	07/27/09 11:58 / tao
Silver	ND	mg/L		0.0005		E200.8	07/27/09 21:03 / jjw
Sodium	26	mg/L		1		E200.7	07/27/09 11:58 / tao
Thallium	ND	mg/L		0.0002		E200.8	07/27/09 21:03 / jjw
Uranium	0.0005	mg/L		0.0003		E200.8	07/27/09 21:03 / jjw
Zinc	ND	mg/L		0.01		E200.7	07/27/09 11:58 / tao
METALS, TOTAL							
Iron, Ferric	6.94	mg/L		0.05		Calculation	07/29/09 16:23 / mep
METALS, TOTAL RECOVERABLE							
Iron	6.94	mg/L		0.05		E200.8	07/27/09 21:10 / jjw
FIELD PARAMETERS							
Field pH, su	6.85	s.u.				FIELD	07/21/09 09:10 / ---
Field Temperature, C	10.32	C				FIELD	07/21/09 09:10 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.

LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Lab ID: B09072237-002
Client Sample ID: Decant Pond

Report Date: 07/31/09
Collection Date: 07/21/09 10:15
Date Received: 07/24/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	8.0	s.u.		0.1		A4500-H B	07/24/09 13:22 / nlh
Conductivity	398	umhos/cm		1		A2510 B	07/24/09 13:22 / nlh
Solids, Total Dissolved TDS @ 180 C	255	mg/L		10		A2540 C	07/24/09 15:40 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	81	mg/L		4		A2320 B	07/24/09 16:36 / ehb
Bicarbonate as HCO ₃	99	mg/L		4		A2320 B	07/24/09 16:36 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	07/24/09 16:36 / ehb
Chloride	4	mg/L		1		E300.0	07/27/09 18:19 / kh
Sulfate	42	mg/L		1		E300.0	07/27/09 18:19 / kh
Fluoride	0.2	mg/L		0.1		A4500-F C	07/24/09 16:36 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	07/24/09 12:15 / afb
Hardness as CaCO ₃	76	mg/L		1		A2340 B	07/27/09 12:05 / rls
Sodium Adsorption Ratio (SAR)	1.34			0.01		Calculation	07/27/09 12:05 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY000002)							
Organic Carbon, Total (TOC)	0.9	mg/L		0.5		A5310 C	07/28/09 16:09 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	3.14	mg/L		0.05		E350.1	07/24/09 15:19 / sam
Nitrogen, Nitrate+Nitrite as N	13.6	mg/L	D	0.03		E353.2	07/24/09 14:29 / bls
Phosphorus, Orthophosphate as P	0.006	mg/L	LH	0.005		E365.1	07/24/09 12:23 / sam
- The sample was received after the holding time had expired for Ortho Phosphate.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5.0		E245.7	07/30/09 15:10 / eli-ca
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.7	07/27/09 12:05 / tao
Antimony	0.062	mg/L		0.003		E200.8	07/27/09 21:17 / jjw
Arsenic	0.005	mg/L		0.001		E200.8	07/27/09 21:17 / jjw
Barium	0.111	mg/L		0.005		E200.7	07/27/09 12:05 / tao
Beryllium	ND	mg/L		0.001		E200.7	07/27/09 12:05 / tao
Cadmium	0.00097	mg/L		0.00008		E200.8	07/27/09 21:17 / jjw
Calcium	22	mg/L		1		E200.7	07/27/09 12:05 / tao
Chromium	ND	mg/L		0.001		E200.8	07/27/09 21:17 / jjw
Copper	0.013	mg/L		0.001		E200.8	07/27/09 21:17 / jjw
Iron	ND	mg/L		0.05		E200.7	07/27/09 12:05 / tao
Iron, Ferrous	ND	mg/L		0.05		E200.7	07/27/09 12:05 / tao
Lead	ND	mg/L		0.0005		E200.8	07/27/09 21:17 / jjw
Magnesium	5	mg/L		1		E200.7	07/27/09 12:05 / tao
Manganese	0.209	mg/L		0.005		E200.7	07/27/09 12:05 / tao

Report Definitions:
 RL - Analyte reporting limit.
 QCL - Quality control limit.
 D - RL increased due to sample matrix interference.
 L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
 ND - Not detected at the reporting limit.
 H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Lab ID: B09072237-002
Client Sample ID: Decant Pond

Report Date: 07/31/09
Collection Date: 07/21/09 10:15
Date Received: 07/24/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	ND	mg/L		0.01		E200.7	07/27/09 12:05 / tao
Potassium	26	mg/L		1		E200.7	07/27/09 12:05 / tao
Selenium	0.002	mg/L		0.001		E200.8	07/27/09 21:17 / jjw
Silicon	6.3	mg/L		0.1		E200.7	07/27/09 12:05 / tao
Silica	13.5	mg/L		0.2		E200.7	07/27/09 12:05 / tao
Silver	ND	mg/L		0.0005		E200.8	07/27/09 21:17 / jjw
Sodium	27	mg/L		1		E200.7	07/27/09 12:05 / tao
Thallium	ND	mg/L		0.0002		E200.8	07/27/09 21:17 / jjw
Uranium	0.0029	mg/L		0.0003		E200.8	07/27/09 21:17 / jjw
Zinc	ND	mg/L		0.01		E200.7	07/27/09 12:05 / tao
METALS, TOTAL							
Iron, Ferric	0.55	mg/L		0.05		Calculation	07/29/09 16:23 / mep
METALS, TOTAL RECOVERABLE							
Iron	0.55	mg/L		0.05		E200.7	07/28/09 17:30 / tao
FIELD PARAMETERS							
Field pH, su	7.66	s.u.				FIELD	07/21/09 10:15 / ---
Field Temperature, C	20.9	C				FIELD	07/21/09 10:15 / ---

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.

LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Lab ID: B09072237-003
Client Sample ID: MW-01-15

Report Date: 07/31/09
Collection Date: 07/21/09 12:50
Date Received: 07/24/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.3	s.u.		0.1		A4500-H B	07/24/09 13:24 / nlh
Conductivity	418	umhos/cm		1		A2510 B	07/24/09 13:24 / nlh
Solids, Total Dissolved TDS @ 180 C	239	mg/L		10		A2540 C	07/24/09 15:40 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	95	mg/L		4		A2320 B	07/24/09 16:44 / ehb
Bicarbonate as HCO ₃	115	mg/L		4		A2320 B	07/24/09 16:44 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	07/24/09 16:44 / ehb
Chloride	4	mg/L		1		E300.0	07/27/09 18:54 / kh
Sulfate	41	mg/L		1		E300.0	07/27/09 18:54 / kh
Fluoride	0.2	mg/L		0.1		A4500-F C	07/24/09 16:44 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	07/24/09 12:15 / afb
Hardness as CaCO ₃	82	mg/L		1		A2340 B	07/27/09 12:09 / rls
Sodium Adsorption Ratio (SAR)	1.49			0.01		Calculation	07/27/09 12:09 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY000002)							
Organic Carbon, Total (TOC)	0.9	mg/L		0.5		A5310 C	07/28/09 16:19 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	3.19	mg/L		0.05		E350.1	07/24/09 15:20 / sam
Nitrogen, Nitrate+Nitrite as N	14.1	mg/L	D	0.03		E353.2	07/24/09 14:30 / bls
Phosphorus, Orthophosphate as P	0.010	mg/L	LH	0.005		E365.1	07/24/09 12:24 / sam
- The sample was received after the holding time had expired for Ortho Phosphate.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5.0		E245.7	07/30/09 15:12 / eli-ca
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.7	07/27/09 12:09 / tao
Antimony	0.026	mg/L		0.003		E200.8	07/27/09 21:31 / jjw
Arsenic	ND	mg/L		0.001		E200.8	07/27/09 21:31 / jjw
Barium	0.106	mg/L		0.005		E200.7	07/27/09 12:09 / tao
Beryllium	ND	mg/L		0.001		E200.7	07/27/09 12:09 / tao
Cadmium	0.00193	mg/L		0.00008		E200.8	07/27/09 21:31 / jjw
Calcium	25	mg/L		1		E200.7	07/27/09 12:09 / tao
Chromium	0.002	mg/L		0.001		E200.8	07/27/09 21:31 / jjw
Copper	0.007	mg/L		0.001		E200.8	07/27/09 21:31 / jjw
Iron	0.06	mg/L		0.05		E200.7	07/27/09 12:09 / tao
Iron, Ferrous	0.06	mg/L		0.05		E200.7	07/27/09 12:09 / tao
Lead	ND	mg/L		0.0005		E200.8	07/27/09 21:31 / jjw
Magnesium	5	mg/L		1		E200.7	07/27/09 12:09 / tao
Manganese	0.491	mg/L		0.005		E200.7	07/27/09 12:09 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Lab ID: B09072237-003
Client Sample ID: MW-01-15

Report Date: 07/31/09
Collection Date: 07/21/09 12:50
Date Received: 07/24/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	0.01	mg/L		0.01		E200.7	07/27/09 12:09 / tao
Potassium	22	mg/L		1		E200.7	07/27/09 12:09 / tao
Selenium	0.002	mg/L		0.001		E200.8	07/27/09 21:31 / jjw
Silicon	5.8	mg/L		0.1		E200.7	07/27/09 12:09 / tao
Silica	12.3	mg/L		0.2		E200.7	07/27/09 12:09 / tao
Silver	ND	mg/L		0.0005		E200.8	07/27/09 21:31 / jjw
Sodium	31	mg/L		1		E200.7	07/27/09 12:09 / tao
Thallium	ND	mg/L		0.0002		E200.8	07/27/09 21:31 / jjw
Uranium	0.0020	mg/L		0.0003		E200.8	07/27/09 21:31 / jjw
Zinc	ND	mg/L		0.01		E200.7	07/27/09 12:09 / tao
METALS, TOTAL							
Iron, Ferric	6.40	mg/L		0.05		Calculation	07/29/09 16:23 / mep
METALS, TOTAL RECOVERABLE							
Iron	6.46	mg/L		0.05		E200.8	07/27/09 22:05 / jjw
FIELD PARAMETERS							
Field pH, su	7.21	s.u.				FIELD	07/21/09 12:50 / ---
Field Temperature, C	22.1	C				FIELD	07/21/09 12:50 / ---

Report
Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.

LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Lab ID: B09072237-004
Client Sample ID: MW-00 duplicate of MW-01-15

Report Date: 07/31/09
Collection Date: 07/21/09 13:10
Date Received: 07/24/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.4	s.u.		0.1		A4500-H B	07/24/09 13:26 / nlh
Conductivity	414	umhos/cm		1		A2510 B	07/24/09 13:26 / nlh
Solids, Total Dissolved TDS @ 180 C	248	mg/L		10		A2540 C	07/24/09 15:44 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	94	mg/L		4		A2320 B	07/24/09 16:52 / ehb
Bicarbonate as HCO ₃	115	mg/L		4		A2320 B	07/24/09 16:52 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	07/24/09 16:52 / ehb
Chloride	4	mg/L		1		E300.0	07/27/09 19:05 / kh
Sulfate	42	mg/L		1		E300.0	07/27/09 19:05 / kh
Fluoride	0.2	mg/L		0.1		A4500-F C	07/24/09 16:52 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	07/24/09 12:15 / afb
Hardness as CaCO ₃	86	mg/L		1		A2340 B	07/27/09 12:37 / rls
Sodium Adsorption Ratio (SAR)	1.43			0.01		Calculation	07/27/09 12:37 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	0.7	mg/L		0.5		A5310 C	07/28/09 16:29 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	3.12	mg/L		0.05		E350.1	07/24/09 15:22 / sam
Nitrogen, Nitrate+Nitrite as N	12.8	mg/L	D	0.03		E353.2	07/24/09 14:31 / bls
Phosphorus, Orthophosphate as P	0.012	mg/L	LH	0.005		E365.1	07/24/09 12:25 / sam
- The sample was received after the holding time had expired for Ortho Phosphate.							
METALS - DISSOLVED							
Mercury	ND	ng/L		5.0		E245.7	07/30/09 15:14 / eli-ca
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.7	07/27/09 12:37 / tao
Antimony	0.026	mg/L		0.003		E200.8	07/28/09 12:50 / jjw
Arsenic	ND	mg/L		0.001		E200.8	07/28/09 12:50 / jjw
Barium	0.115	mg/L		0.005		E200.7	07/27/09 12:37 / tao
Beryllium	ND	mg/L		0.001		E200.7	07/27/09 12:37 / tao
Cadmium	0.00137	mg/L		0.00008		E200.8	07/28/09 12:50 / jjw
Calcium	26	mg/L		1		E200.7	07/27/09 12:37 / tao
Chromium	0.003	mg/L		0.001		E200.8	07/28/09 12:50 / jjw
Copper	0.012	mg/L		0.001		E200.8	07/28/09 12:50 / jjw
Iron	0.06	mg/L		0.05		E200.7	07/27/09 12:37 / tao
Iron, Ferrous	0.06	mg/L		0.05		E200.7	07/27/09 12:37 / tao
Lead	0.0014	mg/L		0.0005		E200.8	07/27/09 22:12 / jjw
Magnesium	5	mg/L		1		E200.7	07/27/09 12:37 / tao
Manganese	0.516	mg/L		0.005		E200.7	07/27/09 12:37 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics
Lab ID: B09072237-004
Client Sample ID: MW-00

Report Date: 07/31/09
Collection Date: 07/21/09 13:10
Date Received: 07/24/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Nickel	0.01	mg/L		0.01		E200.7	07/27/09 12:37 / tao
Potassium	23	mg/L		1		E200.7	07/27/09 12:37 / tao
Selenium	0.002	mg/L		0.001		E200.8	07/27/09 22:12 / jjw
Silicon	6.1	mg/L		0.1		E200.7	07/27/09 12:37 / tao
Silica	13.1	mg/L		0.2		E200.7	07/27/09 12:37 / tao
Silver	ND	mg/L		0.0005		E200.8	07/28/09 12:50 / jjw
Sodium	31	mg/L		1		E200.7	07/27/09 12:37 / tao
Thallium	ND	mg/L		0.0002		E200.8	07/27/09 22:12 / jjw
Uranium	0.0021	mg/L		0.0003		E200.8	07/27/09 22:12 / jjw
Zinc	ND	mg/L		0.01		E200.7	07/27/09 12:37 / tao
METALS, TOTAL							
Iron, Ferric	1.87	mg/L		0.05		Calculation	07/29/09 16:23 / mep
METALS, TOTAL RECOVERABLE							
Iron	1.93	mg/L		0.05		E200.8	07/27/09 22:19 / jjw
FIELD PARAMETERS							
Field pH, su	7.28	s.u.				FIELD	07/21/09 13:10 / ---
Field Temperature, C	18.56	C				FIELD	07/21/09 13:10 / ---

Report RL - Analyte reporting limit.
Definitions: QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.

QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: A2320 B									Batch: R133312
Sample ID: MB	Method Blank					Run: MAN-TECH_090724B			07/24/09 15:11
Alkalinity, Total as CaCO ₃	ND	mg/L	4						
Bicarbonate as HCO ₃	ND	mg/L	4						
Carbonate as CO ₃	ND	mg/L	4						
Sample ID: LCS	Laboratory Control Sample					Run: MAN-TECH_090724B			07/24/09 15:17
Alkalinity, Total as CaCO ₃	97.4	mg/L	4.0	97	90	110			
Sample ID: B09072237-001AMS	Sample Matrix Spike					Run: MAN-TECH_090724B			07/24/09 16:19
Alkalinity, Total as CaCO ₃	262	mg/L	4.0	93	80	120			
Sample ID: B09072237-001AMSD	Sample Matrix Spike Duplicate					Run: MAN-TECH_090724B			07/24/09 16:28
Alkalinity, Total as CaCO ₃	256	mg/L	4.0	88	80	120	2.5	20	
Method: A2510 B									Batch: R133226
Sample ID: SC 150	Laboratory Control Sample					Run: PHSC _101-B_090724A			07/24/09 09:00
Conductivity	156	umhos/cm	1.0	104	90	110			
Sample ID: SC 5000	Laboratory Control Sample					Run: PHSC _101-B_090724A			07/24/09 09:02
Conductivity	4930	umhos/cm	1.0	99	90	110			
Sample ID: B09072235-002ADUP	Sample Duplicate					Run: PHSC _101-B_090724A			07/24/09 13:20
Conductivity	472	umhos/cm	1.0				1.1	10	
Method: A2540 C									Batch: TDS090724B
Sample ID: MBLK3	Method Blank					Run: CPA124S_090724B			07/24/09 15:38
Solids, Total Dissolved TDS @ 180 C	ND	mg/L	10						
Sample ID: LFB3	Laboratory Fortified Blank					Run: CPA124S_090724B			07/24/09 15:39
Solids, Total Dissolved TDS @ 180 C	1100	mg/L	10	99	90	110			
Sample ID: B09072237-001A MS	Sample Matrix Spike					Run: CPA124S_090724B			07/24/09 15:42
Solids, Total Dissolved TDS @ 180 C	2350	mg/L	10	101	80	120			
Sample ID: B09072237-001A MSD	Sample Matrix Spike Duplicate					Run: CPA124S_090724B			07/24/09 15:42
Solids, Total Dissolved TDS @ 180 C	2360	mg/L	10	101	80	120	0.5	20	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: A4500 S-D							Analytical Run: SPEC_090724A		
Sample ID: ICV1_090724A	Initial Calibration Verification Standard								07/24/09 09:00
Sulfide	0.190	mg/L	0.040	95	90	110			
Method: A4500 S-D							Batch: 090724A-SULFIDE-MB-W		
Sample ID: B09072120-001DMS	Sample Matrix Spike								07/24/09 09:00
Sulfide	0.192	mg/L	0.040	82	70	130			
Sample ID: B09072120-001DMSD	Sample Matrix Spike Duplicate								07/24/09 09:00
Sulfide	0.195	mg/L	0.040	84	70	130	1.8	30	
Sample ID: MBLK1_090724A	Method Blank								07/24/09 09:00
Sulfide	ND	mg/L	0.02						
Sample ID: LFB1_090724A	Laboratory Fortified Blank								07/24/09 09:00
Sulfide	0.190	mg/L	0.040	93	70	130			
Method: A4500-F C							Analytical Run: MAN-TECH_090724B		
Sample ID: ICV	Initial Calibration Verification Standard								07/24/09 15:30
Fluoride	1.01	mg/L	0.10	101	90	110			
Method: A4500-F C							Batch: R133312		
Sample ID: MB	Method Blank								07/24/09 15:24
Fluoride	ND	mg/L	0.05						
Sample ID: LFB	Laboratory Fortified Blank								07/24/09 15:27
Fluoride	0.960	mg/L	0.10	96	90	110			
Sample ID: B09072237-001AMS	Sample Matrix Spike								07/24/09 16:19
Fluoride	1.08	mg/L	0.10	97	80	120			
Sample ID: B09072237-001AMSD	Sample Matrix Spike Duplicate								07/24/09 16:28
Fluoride	1.07	mg/L	0.10	96	80	120	0.9	10	
Method: A4500-H B							Batch: R133226		
Sample ID: pH 4	Laboratory Control Sample								07/24/09 08:56
pH	4.06	s.u.	0.10	101	97	103			
Sample ID: pH 7	Laboratory Control Sample								07/24/09 08:57
pH	7.08	s.u.	0.10	101	98	102			
Sample ID: B09072235-002ADUP	Sample Duplicate								07/24/09 13:20
pH	7.91	s.u.	0.10				0.3	10	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: A5310 C							Analytical Run: SUB-C121564		
Sample ID: ICV-23206	Initial Calibration Verification Standard								07/28/09 09:44
Organic Carbon, Dissolved (DOC)	10.4	mg/L	0.10	104	90	110			
Method: A5310 C							Batch: C_R121564		
Sample ID: MBLK	Method Blank								07/28/09 15:00
Organic Carbon, Total (TOC)	0.05	mg/L	0.04						
Sample ID: C09070914-002AMS	Sample Matrix Spike								07/28/09 16:40
Organic Carbon, Total (TOC)	9.10	mg/L	0.50	98	85	115			
Sample ID: C09070914-002AMSD	Sample Matrix Spike Duplicate								07/28/09 16:51
Organic Carbon, Total (TOC)	9.25	mg/L	0.50	101	85	115	1.7	10	
Sample ID: LCS-5085	Laboratory Control Sample								07/28/09 17:02
Organic Carbon, Total (TOC)	10.1	mg/L	0.50	100	90	110			
Method: E245.7							Batch: C_23241		
Sample ID: QCS-23241	Quality Control Sample								07/30/09 14:16
Mercury	22.8	ng/L	5.0	91	90	110			
Sample ID: MB-23241	Method Blank								07/30/09 14:31
Mercury	1	ng/L	1						
Sample ID: LCS-23241	Laboratory Control Sample								07/30/09 14:34
Mercury	22.9	ng/L	5.0	87	85	115			
Sample ID: C09070682-001CMS	Sample Matrix Spike								07/30/09 14:44
Mercury	13.6	ng/L	5.0	112	63	111			S
Sample ID: C09070682-001CMSD	Sample Matrix Spike Duplicate								07/30/09 14:46
Mercury	12.8	ng/L	5.0	104	63	111	6.1	18	

Qualifiers:

RL - Analyte reporting limit.

S - Spike recovery outside of advisory limits.

ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study (Scott Mason, Hydrometrics)
Lab ID: B09061251-001
Client Sample ID: MW-01-15

Report Date: 06/22/09
Collection Date: 06/08/09 17:00
Date Received: 06/12/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.5	s.u.		0.1		A4500-H B	06/12/09 12:17 / kh
Conductivity	405	umhos/cm		1		A2510 B	06/12/09 12:17 / kh
Solids, Total Dissolved TDS @ 180 C	247	mg/L		10		A2540 C	06/12/09 16:23 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	85	mg/L		4		A2320 B	06/12/09 17:22 / ehb
Bicarbonate as HCO ₃	104	mg/L		4		A2320 B	06/12/09 17:22 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	06/12/09 17:22 / ehb
Chloride	4	mg/L		1		E300.0	06/15/09 16:31 / kh
Sulfate	41	mg/L		1		E300.0	06/15/09 16:31 / kh
Fluoride	0.3	mg/L		0.1		A4500-F C	06/15/09 15:32 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	06/15/09 09:30 / afb
Hardness as CaCO ₃	78	mg/L		1		A2340 B	06/15/09 13:02 / rls
Sodium Adsorption Ratio (SAR)	1.57			0.01		Calculation	06/15/09 13:02 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	0.8	mg/L		0.5		A5310 C	06/17/09 11:37 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	3.64	mg/L		0.05		E350.1	06/13/09 12:43 / sam
Nitrogen, Nitrate+Nitrite as N	14.5	mg/L	D	0.02		E353.2	06/15/09 14:28 / bls
Phosphorus, Orthophosphate as P	0.025	mg/L	LH	0.005		E365.1	06/13/09 13:56 / mlm
METALS - DISSOLVED							
Mercury	ND	ng/L		10		E245.7	06/19/09 11:32 / eli-ca
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.7	06/15/09 13:02 / tao
Antimony	0.029	mg/L		0.003		E200.8	06/15/09 20:43 / jjw
Arsenic	ND	mg/L		0.001		E200.8	06/15/09 20:43 / jjw
Barium	0.107	mg/L		0.005		E200.7	06/15/09 13:02 / tao
Beryllium	ND	mg/L		0.001		E200.7	06/15/09 13:02 / tao
Cadmium	0.00136	mg/L		0.00008		E200.8	06/15/09 20:43 / jjw
Calcium	23	mg/L		1		E200.7	06/15/09 13:02 / tao
Chromium	ND	mg/L		0.001		E200.8	06/15/09 20:43 / jjw
Copper	0.009	mg/L		0.001		E200.8	06/15/09 20:43 / jjw
Iron	ND	mg/L		0.05		E200.7	06/15/09 13:02 / tao
Iron, Ferrous	ND	mg/L		0.05		E200.7	06/15/09 13:02 / tao
Lead	ND	mg/L		0.0005		E200.8	06/15/09 20:43 / jjw
Magnesium	5	mg/L		1		E200.7	06/15/09 13:02 / tao
Manganese	0.467	mg/L		0.005		E200.7	06/15/09 13:02 / tao
Nickel	ND	mg/L		0.01		E200.7	06/15/09 13:02 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study (Scott Mason, Hydrometrics)
Lab ID: B09061251-001
Client Sample ID: MW-01-15

Report Date: 06/22/09
Collection Date: 06/08/09 17:00
Date Received: 06/12/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Potassium	23	mg/L		1		E200.7	06/15/09 13:02 / tao
Selenium	0.002	mg/L		0.001		E200.8	06/15/09 20:43 / jjw
Silicon	5.3	mg/L		0.1		E200.7	06/15/09 13:02 / tao
Silica	11.4	mg/L		0.2		E200.7	06/15/09 13:02 / tao
Silver	ND	mg/L		0.0005		E200.8	06/15/09 20:43 / jjw
Sodium	32	mg/L		1		E200.7	06/15/09 13:02 / tao
Thallium	ND	mg/L		0.0002		E200.8	06/15/09 20:43 / jjw
Uranium	0.0018	mg/L		0.0003		E200.8	06/15/09 20:43 / jjw
Zinc	ND	mg/L		0.01		E200.7	06/15/09 13:02 / tao
METALS, TOTAL							
Iron, Ferric	0.17	mg/L		0.03		Calculation	06/18/09 10:52 / mep
METALS, TOTAL RECOVERABLE							
Iron	0.17	mg/L		0.05		E200.7	06/15/09 22:47 / tao
FIELD PARAMETERS							
Field pH, su	7.3	s.u.				FIELD	06/08/09 17:00 / ---
Field Temperature, C	14.1	C				FIELD	06/08/09 17:00 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study (Scott Mason, Hydrometrics)
Lab ID: B09061251-002
Client Sample ID: Decant Pond

Report Date: 06/22/09
Collection Date: 06/09/09 09:35
Date Received: 06/12/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	8.0	s.u.		0.1		A4500-H B	06/12/09 12:19 / kh
Conductivity	411	umhos/cm		1		A2510 B	06/12/09 12:19 / kh
Solids, Total Dissolved TDS @ 180 C	300	mg/L		10		A2540 C	06/12/09 16:25 / afb
INORGANICS							
Alkalinity, Total as CaCO ₃	83	mg/L		4		A2320 B	06/12/09 17:31 / ehb
Bicarbonate as HCO ₃	101	mg/L		4		A2320 B	06/12/09 17:31 / ehb
Carbonate as CO ₃	ND	mg/L		4		A2320 B	06/12/09 17:31 / ehb
Chloride	4	mg/L		1		E300.0	06/15/09 16:42 / kh
Sulfate	45	mg/L		1		E300.0	06/15/09 16:42 / kh
Fluoride	0.2	mg/L		0.1		A4500-F C	06/15/09 15:35 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	06/15/09 09:30 / afb
Hardness as CaCO ₃	73	mg/L		1		A2340 B	06/15/09 13:10 / rls
Sodium Adsorption Ratio (SAR)	1.51			0.01		Calculation	06/15/09 13:10 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	1.2	mg/L		0.5		A5310 C	06/17/09 11:48 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	4.4	mg/L	D	0.1		E350.1	06/13/09 12:44 / sam
Nitrogen, Nitrate+Nitrite as N	14.5	mg/L	D	0.03		E353.2	06/15/09 12:40 / bls
Phosphorus, Orthophosphate as P	0.083	mg/L	LH	0.005		E365.1	06/13/09 13:59 / mlm
METALS - DISSOLVED							
Mercury	ND	ng/L		10		E245.7	06/19/09 11:38 / eli-ca
METALS, DISSOLVED							
Aluminum	0.18	mg/L		0.03		E200.7	06/15/09 13:10 / tao
Antimony	0.050	mg/L		0.003		E200.8	06/15/09 20:51 / jjw
Arsenic	0.002	mg/L		0.001		E200.8	06/15/09 20:51 / jjw
Barium	0.117	mg/L		0.005		E200.7	06/15/09 13:10 / tao
Beryllium	ND	mg/L		0.001		E200.7	06/15/09 13:10 / tao
Cadmium	0.00126	mg/L		0.00008		E200.8	06/15/09 20:51 / jjw
Calcium	21	mg/L		1		E200.7	06/15/09 13:10 / tao
Chromium	ND	mg/L		0.001		E200.8	06/15/09 20:51 / jjw
Copper	0.024	mg/L		0.001		E200.8	06/15/09 20:51 / jjw
Iron	0.05	mg/L		0.05		E200.7	06/15/09 13:10 / tao
Iron, Ferrous	0.05	mg/L		0.05		E200.7	06/15/09 13:10 / tao
Lead	0.0038	mg/L		0.0005		E200.8	06/15/09 20:51 / jjw
Magnesium	5	mg/L		1		E200.7	06/15/09 13:10 / tao
Manganese	0.458	mg/L		0.005		E200.7	06/15/09 13:10 / tao
Nickel	ND	mg/L		0.01		E200.7	06/15/09 13:10 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study (Scott Mason, Hydrometrics)
Lab ID: B09061251-002
Client Sample ID: Decant Pond

Report Date: 06/22/09
Collection Date: 06/09/09 09:35
Date Received: 06/12/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Potassium	28	mg/L		1		E200.7	06/15/09 13:10 / tao
Selenium	0.001	mg/L		0.001		E200.8	06/15/09 20:51 / jjw
Silicon	6.2	mg/L		0.1		E200.7	06/15/09 13:10 / tao
Silica	13.2	mg/L		0.2		E200.7	06/15/09 13:10 / tao
Silver	ND	mg/L		0.0005		E200.8	06/15/09 20:51 / jjw
Sodium	30	mg/L		1		E200.7	06/15/09 13:10 / tao
Thallium	ND	mg/L		0.0002		E200.8	06/15/09 20:51 / jjw
Uranium	0.0033	mg/L		0.0003		E200.8	06/15/09 20:51 / jjw
Zinc	ND	mg/L		0.01		E200.7	06/15/09 13:10 / tao
METALS, TOTAL							
Iron, Ferric	0.77	mg/L		0.03		Calculation	06/18/09 10:52 / mep
METALS, TOTAL RECOVERABLE							
Iron	0.82	mg/L		0.05		E200.7	06/16/09 20:42 / tao
FIELD PARAMETERS							
Field pH, su	6.92	s.u.				FIELD	06/09/09 09:35 / ---
Field Temperature, C	15.1	C				FIELD	06/09/09 09:35 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study (Scott Mason, Hydrometrics)
Lab ID: B09061251-003
Client Sample ID: MW 95-8

Report Date: 06/22/09
Collection Date: 06/09/09 12:15
Date Received: 06/12/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
PHYSICAL PROPERTIES							
pH	7.6	s.u.		0.1		A4500-H B	06/12/09 12:22 / kh
Conductivity	540	umhos/cm		1		A2510 B	06/12/09 12:22 / kh
Solids, Total Dissolved TDS @ 180 C	353	mg/L		10		A2540 C	06/12/09 16:26 / afb
INORGANICS							
Alkalinity, Total as CaCO3	134	mg/L		4		A2320 B	06/12/09 17:39 / ehb
Bicarbonate as HCO3	163	mg/L		4		A2320 B	06/12/09 17:39 / ehb
Carbonate as CO3	ND	mg/L		4		A2320 B	06/12/09 17:39 / ehb
Chloride	4	mg/L		1		E300.0	06/15/09 17:17 / kh
Sulfate	36	mg/L		1		E300.0	06/15/09 17:17 / kh
Fluoride	0.1	mg/L		0.1		A4500-F C	06/15/09 15:38 / ehb
Sulfide	ND	mg/L		0.04		A4500 S-D	06/15/09 09:30 / afb
Hardness as CaCO3	203	mg/L		1		A2340 B	06/15/09 13:14 / rls
Sodium Adsorption Ratio (SAR)	0.77			0.01		Calculation	06/15/09 13:14 / rls
AGGREGATE ORGANICS (CONTRACT LAB WY00002)							
Organic Carbon, Total (TOC)	1.3	mg/L		0.5		A5310 C	06/17/09 12:00 / eli-ca
NUTRIENTS							
Nitrogen, Ammonia as N	0.77	mg/L		0.05		E350.1	06/13/09 12:46 / sam
Nitrogen, Nitrate+Nitrite as N	24.5	mg/L	D	0.03		E353.2	06/15/09 12:50 / bls
Phosphorus, Orthophosphate as P	ND	mg/L	LH	0.005		E365.1	06/13/09 14:00 / mlm
METALS - DISSOLVED							
Mercury	ND	ng/L		10		E245.7	06/19/09 11:40 / eli-ca
METALS, DISSOLVED							
Aluminum	ND	mg/L		0.03		E200.7	06/15/09 13:14 / tao
Antimony	0.005	mg/L		0.003		E200.8	06/15/09 20:58 / jjw
Arsenic	ND	mg/L		0.001		E200.8	06/15/09 20:58 / jjw
Barium	0.149	mg/L		0.005		E200.7	06/15/09 13:14 / tao
Beryllium	ND	mg/L		0.001		E200.7	06/15/09 13:14 / tao
Cadmium	0.00296	mg/L		0.00008		E200.8	06/15/09 20:58 / jjw
Calcium	68	mg/L		1		E200.7	06/15/09 13:14 / tao
Chromium	ND	mg/L		0.001		E200.8	06/15/09 20:58 / jjw
Copper	0.011	mg/L		0.001		E200.8	06/15/09 20:58 / jjw
Iron	ND	mg/L		0.05		E200.7	06/15/09 13:14 / tao
Iron, Ferrous	ND	mg/L		0.05		E200.7	06/15/09 13:14 / tao
Lead	ND	mg/L		0.0005		E200.8	06/15/09 20:58 / jjw
Magnesium	8	mg/L		1		E200.7	06/15/09 13:14 / tao
Manganese	ND	mg/L		0.005		E200.7	06/15/09 13:14 / tao
Nickel	ND	mg/L		0.01		E200.7	06/15/09 13:14 / tao

Report Definitions:
RL - Analyte reporting limit.
QCL - Quality control limit.
D - RL increased due to sample matrix interference.
L - Lowest available reporting limit for the analytical method used.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.
H - Analysis performed past recommended holding time.



LABORATORY ANALYTICAL REPORT

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study (Scott Mason, Hydrometrics)
Lab ID: B09061251-003
Client Sample ID: MW 95-8

Report Date: 06/22/09
Collection Date: 06/09/09 12:15
Date Received: 06/12/09
Matrix: Aqueous

Analyses	Result	Units	Qualifiers	RL	MCL/ QCL	Method	Analysis Date / By
METALS, DISSOLVED							
Potassium	9	mg/L		1		E200.7	06/15/09 13:14 / tao
Selenium	0.004	mg/L		0.001		E200.8	06/15/09 20:58 / jjw
Silicon	5.2	mg/L		0.1		E200.7	06/15/09 13:14 / tao
Silica	11.1	mg/L		0.2		E200.7	06/15/09 13:14 / tao
Silver	ND	mg/L		0.0005		E200.8	06/15/09 20:58 / jjw
Sodium	25	mg/L		1		E200.7	06/15/09 13:14 / tao
Thallium	ND	mg/L		0.0002		E200.8	06/15/09 20:58 / jjw
Uranium	0.0005	mg/L		0.0003		E200.8	06/15/09 20:58 / jjw
Zinc	ND	mg/L		0.01		E200.8	06/15/09 20:58 / jjw
METALS, TOTAL							
Iron, Ferric	1.62	mg/L		0.03		Calculation	06/18/09 10:52 / mep
METALS, TOTAL RECOVERABLE							
Iron	1.62	mg/L		0.05		E200.7	06/16/09 20:50 / tao
FIELD PARAMETERS							
Field pH, su	7.57	s.u.				FIELD	06/09/09 12:15 / ---
Field Temperature, C	10.2	C				FIELD	06/09/09 12:15 / ---

Report Definitions: RL - Analyte reporting limit.
QCL - Quality control limit.

MCL - Maximum contaminant level.
ND - Not detected at the reporting limit.

QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E300.0							Analytical Run: IC202-B_090727A		
Sample ID: ICV072709-11	Initial Calibration Verification Standard								07/27/09 12:19
Chloride	24.1	mg/L	1.0	96	90	110			
Sulfate	103	mg/L	1.0	103	90	110			
Method: E300.0							Batch: R133371		
Sample ID: ICB072709-12	Method Blank								07/27/09 12:30
Chloride	ND	mg/L	0.03						
Sulfate	ND	mg/L	0.06						
Sample ID: LFB072709-13	Laboratory Fortified Blank								07/27/09 12:42
Chloride	22.8	mg/L	1.0	91	90	110			
Sulfate	101	mg/L	1.0	101	90	110			
Sample ID: B09072237-002AMS	Sample Matrix Spike								07/27/09 18:30
Chloride	28.0	mg/L	1.0	98	90	110			
Sulfate	152	mg/L	1.0	110	90	110			
Sample ID: B09072237-002AMSD	Sample Matrix Spike Duplicate								07/27/09 18:42
Chloride	28.4	mg/L	1.0	100	90	110	1.4	20	
Sulfate	154	mg/L	1.0	112	90	110	1.1	20	S
Method: E350.1							Analytical Run: FIA202-B_090724C		
Sample ID: ICV	Initial Calibration Verification Standard								07/24/09 13:23
Nitrogen, Ammonia as N	5.52	mg/L	0.11	101	90	110			
Method: E350.1							Batch: R133268		
Sample ID: MBLK	Method Blank								07/24/09 13:24
Nitrogen, Ammonia as N	ND	mg/L	0.02						
Sample ID: LFB	Laboratory Fortified Blank								07/24/09 13:26
Nitrogen, Ammonia as N	1.03	mg/L	0.050	104	90	110			
Sample ID: B09072233-003BMS	Sample Matrix Spike								07/24/09 15:16
Nitrogen, Ammonia as N	1.43	mg/L	0.050	104	90	110			
Sample ID: B09072233-003BMDS	Sample Matrix Spike Duplicate								07/24/09 15:17
Nitrogen, Ammonia as N	1.43	mg/L	0.050	105	90	110	0.2	10	

Qualifiers:

RL - Analyte reporting limit.

S - Spike recovery outside of advisory limits.

ND - Not detected at the reporting limit.

QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E353.2							Analytical Run: FIA203-B_090724A		
Sample ID: ICV	Initial Calibration Verification Standard								07/24/09 08:53
Nitrogen, Nitrate+Nitrite as N	37.4	mg/L	0.032	106	90	110			
Method: E353.2							Batch: R133234		
Sample ID: MBLK	Method Blank								07/24/09 08:54
Nitrogen, Nitrate+Nitrite as N	ND	mg/L	0.002						
Sample ID: LFB	Laboratory Fortified Blank								07/24/09 08:55
Nitrogen, Nitrate+Nitrite as N	1.00	mg/L	0.010	102	90	110			
Sample ID: B09072233-003BMS	Sample Matrix Spike								07/24/09 14:22
Nitrogen, Nitrate+Nitrite as N	0.979	mg/L	0.010	98	90	110			
Sample ID: B09072233-003BMSD	Sample Matrix Spike Duplicate								07/24/09 14:23
Nitrogen, Nitrate+Nitrite as N	0.981	mg/L	0.010	98	90	110	0.2	10	
Sample ID: B09071811-003AMS	Sample Matrix Spike								07/24/09 14:38
Nitrogen, Nitrate+Nitrite as N	0.971	mg/L	0.010	99	90	110			
Sample ID: B09071811-003AMSD	Sample Matrix Spike Duplicate								07/24/09 14:40
Nitrogen, Nitrate+Nitrite as N	0.980	mg/L	0.010	100	90	110	0.9	10	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E365.1			Analytical Run: FIA202-B_090724B						
Sample ID: ICV	Initial Calibration Verification Standard								07/24/09 11:11
Phosphorus, Orthophosphate as P	0.245	mg/L	0.0050	98	90	110			
Method: E365.1			Batch: R133252						
Sample ID: MBLK	Method Blank								07/24/09 11:14
Phosphorus, Orthophosphate as P	ND	mg/L	0.005						
Sample ID: LFB	Laboratory Fortified Blank								07/24/09 11:15
Phosphorus, Orthophosphate as P	0.0940	mg/L	0.0050	94	90	110			
Sample ID: B09072221-002AMS	Sample Matrix Spike								07/24/09 11:31
Phosphorus, Orthophosphate as P	0.134	mg/L	0.0050	95	90	110			
Sample ID: B09072221-002AMSD	Sample Matrix Spike Duplicate								07/24/09 11:32
Phosphorus, Orthophosphate as P	0.135	mg/L	0.0050	96	90	110	0.7	10	
Sample ID: B09072233-002AMS	Sample Matrix Spike								07/24/09 12:19
Phosphorus, Orthophosphate as P	0.165	mg/L	0.0050	99	90	110			
Sample ID: B09072233-002AMSD	Sample Matrix Spike Duplicate								07/24/09 12:20
Phosphorus, Orthophosphate as P	0.164	mg/L	0.0050	98	90	110	0.6	10	
Sample ID: B09072231-003AMS	Sample Matrix Spike								07/24/09 12:31
Phosphorus, Orthophosphate as P	0.101	mg/L	0.0050	101	90	110			
Sample ID: B09072231-003AMSD	Sample Matrix Spike Duplicate								07/24/09 12:32
Phosphorus, Orthophosphate as P	0.101	mg/L	0.0050	101	90	110	0	10	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7									Batch: 40344
Sample ID: MB-40344	Method Blank					Run: ICP202-B_090728A			07/28/09 16:30
Iron	0.04	mg/L	0.008						
Sample ID: LCS5-40344	Laboratory Control Sample					Run: ICP202-B_090728A			07/28/09 16:34
Iron	2.61	mg/L	0.030	103	85	115			
Sample ID: B09072246-001BMS5	Sample Matrix Spike					Run: ICP202-B_090728A			07/28/09 17:45
Iron	15.1	mg/L	0.030		70	130			A
Sample ID: B09072246-001BMSD5	Sample Matrix Spike Duplicate					Run: ICP202-B_090728A			07/28/09 17:49
Iron	15.1	mg/L	0.030		70	130	0.3	20	A

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

A - The analyte level was greater than four times the spike level. In accordance with the method % recovery is not calculated.



QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7					Analytical Run: ICP202-B_090727A				
Sample ID: ICV	Continuing Calibration Verification Standard								07/27/09 10:48
Aluminum	2.47	mg/L	0.10	99	95	105			
Barium	2.59	mg/L	0.10	104	95	105			
Beryllium	1.21	mg/L	0.010	96	95	105			
Calcium	24.2	mg/L	1.0	97	95	105			
Iron	2.60	mg/L	0.030	104	95	105			
Magnesium	25.2	mg/L	1.0	101	95	105			
Manganese	2.41	mg/L	0.010	96	95	105			
Nickel	2.60	mg/L	0.050	104	95	105			
Potassium	23.7	mg/L	1.0	95	95	105			
Silicon	5.08	mg/L	0.10	102	95	105			
Sodium	24.6	mg/L	1.0	98	95	105			
Zinc	2.55	mg/L	0.010	102	95	105			
Method: E200.7					Batch: R133358				
Sample ID: MB-SPDIS090727A	Method Blank								07/27/09 10:13
Aluminum	ND	mg/L	0.008						
Barium	0.001	mg/L	0.0001						
Beryllium	0.0005	mg/L	9E-05						
Calcium	0.03	mg/L	0.009						
Iron	ND	mg/L	0.002						
Magnesium	0.01	mg/L	0.01						
Manganese	0.0005	mg/L	0.0003						
Nickel	ND	mg/L	0.002						
Potassium	ND	mg/L	0.01						
Silicon	ND	mg/L	0.01						
Sodium	ND	mg/L	0.03						
Zinc	0.002	mg/L	0.002						
Sample ID: LFB-SPDIS090727A	Laboratory Fortified Blank								07/27/09 10:16
Aluminum	5.05	mg/L	0.10	101	85	115			
Barium	0.972	mg/L	0.10	97	85	115			
Beryllium	0.484	mg/L	0.010	97	85	115			
Calcium	48.3	mg/L	1.0	97	85	115			
Iron	5.19	mg/L	0.030	104	85	115			
Magnesium	51.1	mg/L	1.0	102	85	115			
Manganese	4.83	mg/L	0.010	97	85	115			
Nickel	1.05	mg/L	0.050	105	85	115			
Potassium	50.4	mg/L	1.0	101	85	115			
Silicon	11.0	mg/L	0.10	110	85	115			
Sodium	49.4	mg/L	1.0	99	85	115			
Zinc	1.02	mg/L	0.010	102	85	115			

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7							Batch: R133358		
Sample ID: B09072234-001BMS2	Sample Matrix Spike		Run: ICP202-B_090727A				07/27/09 11:51		
Aluminum	4.79	mg/L	0.10	96	70	130			
Barium	1.04	mg/L	0.10	94	70	130			
Beryllium	0.461	mg/L	0.0010	92	70	130			
Calcium	82.9	mg/L	1.0	84	70	130			
Iron	4.95	mg/L	0.030	99	70	130			
Magnesium	107	mg/L	1.0	89	70	130			
Manganese	4.68	mg/L	0.010	94	70	130			
Nickel	1.01	mg/L	0.010	100	70	130			
Potassium	50.8	mg/L	1.0	99	70	130			
Silicon	14.7	mg/L	0.10	94	70	130			
Sodium	52.4	mg/L	1.0	98	70	130			
Zinc	0.981	mg/L	0.010	98	70	130			
Iron, Ferrous	4.95	mg/L	0.030	99	70	130			
Silica	31.4	mg/L	0.21	94	70	130			
Sample ID: B09072234-001BMSD2	Sample Matrix Spike Duplicate		Run: ICP202-B_090727A				07/27/09 11:55		
Aluminum	4.84	mg/L	0.10	97	70	130	1	20	
Barium	1.03	mg/L	0.10	93	70	130	1.3	20	
Beryllium	0.465	mg/L	0.0010	93	70	130	0.9	20	
Calcium	83.6	mg/L	1.0	86	70	130	0.9	20	
Iron	5.03	mg/L	0.030	100	70	130	1.7	20	
Magnesium	109	mg/L	1.0	92	70	130	1.4	20	
Manganese	4.72	mg/L	0.010	94	70	130	0.8	20	
Nickel	1.01	mg/L	0.010	101	70	130	0.4	20	
Potassium	50.3	mg/L	1.0	98	70	130	1	20	
Silicon	14.9	mg/L	0.10	97	70	130	1.6	20	
Sodium	52.1	mg/L	1.0	98	70	130	0.6	20	
Zinc	0.988	mg/L	0.010	98	70	130	0.7	20	
Iron, Ferrous	5.03	mg/L	0.030	100	70	130	1.7	20	
Silica	31.9	mg/L	0.21	97	70	130	1.6	20	
Sample ID: B09072251-005BMS2	Sample Matrix Spike		Run: ICP202-B_090727A				07/27/09 13:02		
Aluminum	4.86	mg/L	0.10	97	70	130			
Barium	1.45	mg/L	0.10	92	70	130			
Beryllium	0.471	mg/L	0.0010	94	70	130			
Calcium	123	mg/L	1.0	78	70	130			
Iron	109	mg/L	0.030		70	130			A
Magnesium	67.8	mg/L	1.0	94	70	130			
Manganese	6.98	mg/L	0.010	90	70	130			
Nickel	0.994	mg/L	0.010	98	70	130			
Potassium	51.9	mg/L	1.0	98	70	130			
Silicon	27.5	mg/L	0.10	92	70	130			
Sodium	70.7	mg/L	1.0	97	70	130			

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

A - The analyte level was greater than four times the spike level. In accordance with the method % recovery is not calculated.

QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.7							Batch: R133358		
Sample ID: B09072251-005BMS2	Sample Matrix Spike		Run: ICP202-B_090727A				07/27/09 13:02		
Zinc	0.962	mg/L	0.010	96	70	130			
Iron, Ferrous	109	mg/L	0.030		70	130			A
Silica	58.8	mg/L	0.21	92	70	130			
Sample ID: B09072251-005BMSD2	Sample Matrix Spike Duplicate		Run: ICP202-B_090727A				07/27/09 13:05		
Aluminum	5.14	mg/L	0.10	103	70	130	5.6	20	
Barium	1.47	mg/L	0.10	95	70	130	1.6	20	
Beryllium	0.482	mg/L	0.0010	96	70	130	2.4	20	
Calcium	125	mg/L	1.0	83	70	130	1.7	20	
Iron	110	mg/L	0.030		70	130	1.3	20	A
Magnesium	69.3	mg/L	1.0	97	70	130	2.3	20	
Manganese	7.13	mg/L	0.010	93	70	130	2	20	
Nickel	1.05	mg/L	0.010	104	70	130	5.5	20	
Potassium	54.9	mg/L	1.0	104	70	130	5.6	20	
Silicon	28.0	mg/L	0.10	97	70	130	1.8	20	
Sodium	72.9	mg/L	1.0	102	70	130	3.2	20	
Zinc	1.01	mg/L	0.010	101	70	130	5.3	20	
Iron, Ferrous	110	mg/L	0.030		70	130	1.3	20	A
Silica	59.9	mg/L	0.21	97	70	130	1.8	20	
Method: E200.7							Analytical Run: ICP202-B_090728A		
Sample ID: ICV	Continuing Calibration Verification Standard						07/28/09 10:54		
Iron	2.57	mg/L	0.030	103	95	105			
Method: E200.8							Batch: 40344		
Sample ID: MB-40344	Method Blank		Run: ICPMS203-B_090727A				07/27/09 20:43		
Iron	0.08	mg/L	0.0008						
Sample ID: LCS5-40344	Laboratory Control Sample		Run: ICPMS203-B_090727A				07/27/09 20:50		
Iron	2.55	mg/L	0.030	99	85	115			
Sample ID: B09072246-001BMS5	Sample Matrix Spike		Run: ICPMS203-B_090727A				07/27/09 22:32		
Iron	15.7	mg/L	0.030		70	130			A
Sample ID: B09072246-001BMSD5	Sample Matrix Spike Duplicate		Run: ICPMS203-B_090727A				07/27/09 22:39		
Iron	15.8	mg/L	0.030		70	130	0.7	20	A

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

A - The analyte level was greater than four times the spike level. In accordance with the method % recovery is not calculated.



QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8			Analytical Run: ICPMS203-B_090727A						
Sample ID: QCS-090602A,090609B,0 Initial Calibration Verification Standard			07/27/09 19:40						
Aluminum	0.245	mg/L	0.10	98	90	110			
Antimony	0.0480	mg/L	0.050	96	90	110			
Arsenic	0.0490	mg/L	0.0050	98	90	110			
Cadmium	0.0257	mg/L	0.0010	103	90	110			
Chromium	0.0500	mg/L	0.010	100	90	110			
Copper	0.0512	mg/L	0.010	102	90	110			
Iron	0.273	mg/L	0.030	107	90	110			
Lead	0.0493	mg/L	0.010	99	90	110			
Selenium	0.0497	mg/L	0.0050	99	90	110			
Silver	0.0261	mg/L	0.0050	104	90	110			
Thallium	0.0501	mg/L	0.10	100	90	110			
Uranium	0.0191	mg/L	0.0010	95	90	110			
Method: E200.8			Batch: R133351						
Sample ID: LRB Method Blank			Run: ICPMS203-B_090727A 07/27/09 11:55						
Aluminum	0.005	mg/L	0.00010						
Antimony	ND	mg/L	1E-05						
Arsenic	ND	mg/L	2E-05						
Cadmium	ND	mg/L	4E-06						
Chromium	ND	mg/L	2E-05						
Copper	ND	mg/L	3E-05						
Lead	ND	mg/L	3E-06						
Selenium	ND	mg/L	4E-05						
Silver	3E-05	mg/L	9E-06						
Thallium	ND	mg/L	8E-06						
Uranium	ND	mg/L	2E-06						
Sample ID: LFB Laboratory Fortified Blank			Run: ICPMS203-B_090727A 07/27/09 12:15						
Aluminum	0.0530	mg/L	0.10	96	85	115			
Antimony	0.0496	mg/L	0.050	99	85	115			
Arsenic	0.0495	mg/L	0.0050	99	85	115			
Cadmium	0.0495	mg/L	0.0010	99	85	115			
Chromium	0.0499	mg/L	0.010	100	85	115			
Copper	0.0485	mg/L	0.010	97	85	115			
Lead	0.0497	mg/L	0.010	99	85	115			
Selenium	0.0487	mg/L	0.0050	97	85	115			
Silver	0.0191	mg/L	0.0050	95	85	115			
Thallium	0.0506	mg/L	0.10	101	85	115			
Uranium	0.0491	mg/L	0.0010	98	85	115			
Sample ID: B09072057-005CMS Sample Matrix Spike			Run: ICPMS203-B_090727A 07/28/09 00:29						
Aluminum	0.0942	mg/L	0.10	94	70	130			
Antimony	0.100	mg/L	0.0050	100	70	130			

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc Troy Mine
Project: Cu Attenuation Study-Scott Mason Hydrometrics

Report Date: 07/31/09
Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8							Batch: R133351		
Sample ID: B09072057-005CMS	Sample Matrix Spike		Run: ICPMS203-B_090727A				07/28/09 00:29		
Arsenic	0.101	mg/L	0.0050	101	70	130			
Cadmium	0.0988	mg/L	0.0010	99	70	130			
Chromium	0.102	mg/L	0.010	102	70	130			
Copper	0.156	mg/L	0.010	94	70	130			
Lead	0.101	mg/L	0.010	100	70	130			
Selenium	0.104	mg/L	0.0050	103	70	130			
Silver	0.0389	mg/L	0.0050	97	70	130			
Thallium	0.101	mg/L	0.0050	101	70	130			
Uranium	0.109	mg/L	0.0010	100	70	130			
Sample ID: B09072057-005CMSD	Sample Matrix Spike Duplicate		Run: ICPMS203-B_090727A				07/28/09 01:03		
Aluminum	0.0933	mg/L	0.10	93	70	130		20	
Antimony	0.100	mg/L	0.0050	100	70	130	0.3	20	
Arsenic	0.102	mg/L	0.0050	102	70	130	1	20	
Cadmium	0.0982	mg/L	0.0010	98	70	130	0.6	20	
Chromium	0.0996	mg/L	0.010	100	70	130	2.2	20	
Copper	0.156	mg/L	0.010	94	70	130	0.3	20	
Lead	0.102	mg/L	0.010	100	70	130	0.4	20	
Selenium	0.102	mg/L	0.0050	101	70	130	2.2	20	
Silver	0.0385	mg/L	0.0050	96	70	130	1.1	20	
Thallium	0.101	mg/L	0.0050	101	70	130	0.6	20	
Uranium	0.109	mg/L	0.0010	101	70	130	0.3	20	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.



QA/QC Summary Report

Client: Genesis Inc Troy Mine

Report Date: 07/31/09

Project: Cu Attenuation Study-Scott Mason Hydrometrics

Work Order: B09072237

Analyte	Result	Units	RL	%REC	Low Limit	High Limit	RPD	RPDLimit	Qual
Method: E200.8			Analytical Run: ICPMS202-B_090728A						
Sample ID: QCS - 090602A,090609B, Initial Calibration Verification Standard			07/28/09 09:41						
Antimony	0.0465	mg/L	0.050	93	90	110			
Arsenic	0.0499	mg/L	0.0050	100	90	110			
Cadmium	0.0247	mg/L	0.0010	99	90	110			
Chromium	0.0493	mg/L	0.010	99	90	110			
Copper	0.0501	mg/L	0.010	100	90	110			
Silver	0.0258	mg/L	0.0050	103	90	110			
Method: E200.8			Batch: R133463						
Sample ID: LRB			Method Blank		Run: ICPMS202-B_090728A		07/28/09 12:22		
Antimony	ND	mg/L	5E-05						
Arsenic	ND	mg/L	3E-05						
Cadmium	ND	mg/L	1E-05						
Chromium	ND	mg/L	0.0002						
Copper	ND	mg/L	8E-05						
Silver	ND	mg/L	2E-05						
Sample ID: LFB			Laboratory Fortified Blank		Run: ICPMS202-B_090728A		07/28/09 12:26		
Antimony	0.0507	mg/L	0.050	101	85	115			
Arsenic	0.0495	mg/L	0.0050	99	85	115			
Cadmium	0.0487	mg/L	0.0010	97	85	115			
Chromium	0.0477	mg/L	0.010	95	85	115			
Copper	0.0470	mg/L	0.010	94	85	115			
Silver	0.0175	mg/L	0.0050	88	85	115			
Sample ID: B09072234-001BMS			Sample Matrix Spike		Run: ICPMS202-B_090728A		07/28/09 12:38		
Antimony	0.0512	mg/L	0.0050	100	70	130			
Arsenic	0.0519	mg/L	0.0050	100	70	130			
Cadmium	0.0488	mg/L	0.0010	97	70	130			
Chromium	0.0491	mg/L	0.010	97	70	130			
Copper	0.0493	mg/L	0.010	94	70	130			
Silver	0.0177	mg/L	0.0050	88	70	130			
Sample ID: B09072234-001BMDS			Sample Matrix Spike Duplicate		Run: ICPMS202-B_090728A		07/28/09 12:42		
Antimony	0.0510	mg/L	0.0050	100	70	130	0.4	20	
Arsenic	0.0519	mg/L	0.0050	100	70	130	0.1	20	
Cadmium	0.0484	mg/L	0.0010	97	70	130	0.9	20	
Chromium	0.0491	mg/L	0.010	97	70	130	0.1	20	
Copper	0.0491	mg/L	0.010	94	70	130	0.3	20	
Silver	0.0180	mg/L	0.0050	90	70	130	1.7	20	

Qualifiers:

RL - Analyte reporting limit.

ND - Not detected at the reporting limit.

Energy Laboratories Inc

Workorder Receipt Checklist



B09072237

Genesis Inc Troy Mine

Login completed by: Gina McCartney

Date and Time Received: 7/24/2009 9:00 AM

Reviewed by: Denise Ruby

Received by: Ig

Reviewed Date: 7/24/2009 5:35:00 PM

Carrier name: Return-UPS
Ground

Shipping container/cooler in good condition?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	Not Present <input type="checkbox"/>
Custody seals intact on shipping container/cooler?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	Not Present <input type="checkbox"/>
Custody seals intact on sample bottles?	Yes <input type="checkbox"/>	No <input type="checkbox"/>	Not Present <input checked="" type="checkbox"/>
Chain of custody present?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Chain of custody signed when relinquished and received?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Chain of custody agrees with sample labels?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Samples in proper container/bottle?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Sample containers intact?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
Sufficient sample volume for indicated test?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	
All samples received within holding time?	Yes <input type="checkbox"/>	No <input checked="" type="checkbox"/>	
Container/Temp Blank temperature:	5°C On Ice		
Water - VOA vials have zero headspace?	Yes <input type="checkbox"/>	No <input type="checkbox"/>	No VOA vials submitted <input checked="" type="checkbox"/>
Water - pH acceptable upon receipt?	Yes <input checked="" type="checkbox"/>	No <input type="checkbox"/>	Not Applicable <input type="checkbox"/>

Contact and Corrective Action Comments:

All samples for Ortho Phosphorus received past the 48 hour hold time. Analyze, per ELI's project manager.



Chain of Custody and Analytical Request Record

PLEASE PRINT- Provide as much information as possible.

Company Name: Genetics Inc. - Troy Mini		Project Name, PWS, Permit, Etc. Cu Attenuation Study - Hydromed		Sample Origin State: MD		EPA/State Compliance: Yes <input type="checkbox"/> No <input checked="" type="checkbox"/>	
Report Mail Address: PO Box 1660 Troy, MT 59935		Contact Name: Paul Kikay		Email: pkikay@mtel.com		Sampler: (Please Print) Paul Kikay / Johnson	
Invoice Address: Same		Invoice Contact & Phone: Same		Purchase Order:		Quote/Bottle Order: 1850	
Special Report/Formats - ELI must be notified prior to sample submittal for the following: <input type="checkbox"/> DW <input type="checkbox"/> A2LA <input type="checkbox"/> EDD/EDT (Electronic Data) <input type="checkbox"/> GSA <input type="checkbox"/> POTW/WWTP <input type="checkbox"/> State: <input type="checkbox"/> LEVEL IV <input type="checkbox"/> Other: <input type="checkbox"/> NELAC				ANALYSIS REQUESTED SEE ATTACHED Normal Turnaround (TAT)			
Number of Containers Sample Type: A W S V B O Vegetation Bioassay Other		MATRIX Collection Date Collection Time		Contact ELI prior to RUSH sample submittal for charges and scheduling - See Instruction Page Comments:		Shipped by: Cooler ID(s):	
MWT-75-B Decade Pond MW-01-15 MW-00		21 July 09 10:00 AM 12:30 PM 1:10 PM		RUSH X X X X		Receipt Temp On Ice: <input checked="" type="checkbox"/> Yes <input type="checkbox"/> No Custody Seal: <input checked="" type="checkbox"/> N Bottles/Coolers: <input checked="" type="checkbox"/> B/C Intact: <input checked="" type="checkbox"/> Y/N Signature Match: <input checked="" type="checkbox"/> Y/N	
1 2 3 4 5		9:00 AM 10:00 AM 12:30 PM 1:10 PM		X X X X		18-6.85 Temp - 10.32°C pH 7.66 Temp 20.9°C pH 7.21 Temp 22.1°C pH 7.28 Temp 18.58°C	
6 7 8 9 10						LABORATORY USE ONLY 80907237-00 -002 -003 -004	
Relinquished by (print) Paul Kikay		Date/Time: 22 July 09 8:00 AM		Received by (print): Paul Kikay		Date/Time: 7-24-09 9:00 AM	
Relinquished by (print) Paul Kikay		Date/Time: 22 July 09 8:00 AM		Received by (print): Paul Kikay		Date/Time: 7-24-09 9:00 AM	
Sample Disposal:		Return to Client:		Lab Disposal:		Received by Laboratory:	
Custody Record MUST be Signed		Signature:					

In certain circumstances, samples submitted to Energy Laboratories, Inc. may be subcontracted to other certified laboratories in order to complete the analysis requested. This service is not available for all tests. All such subcontracted tests will be performed in accordance with the applicable regulatory requirements. This service is not available for all tests. All such subcontracted tests will be performed in accordance with the applicable regulatory requirements.

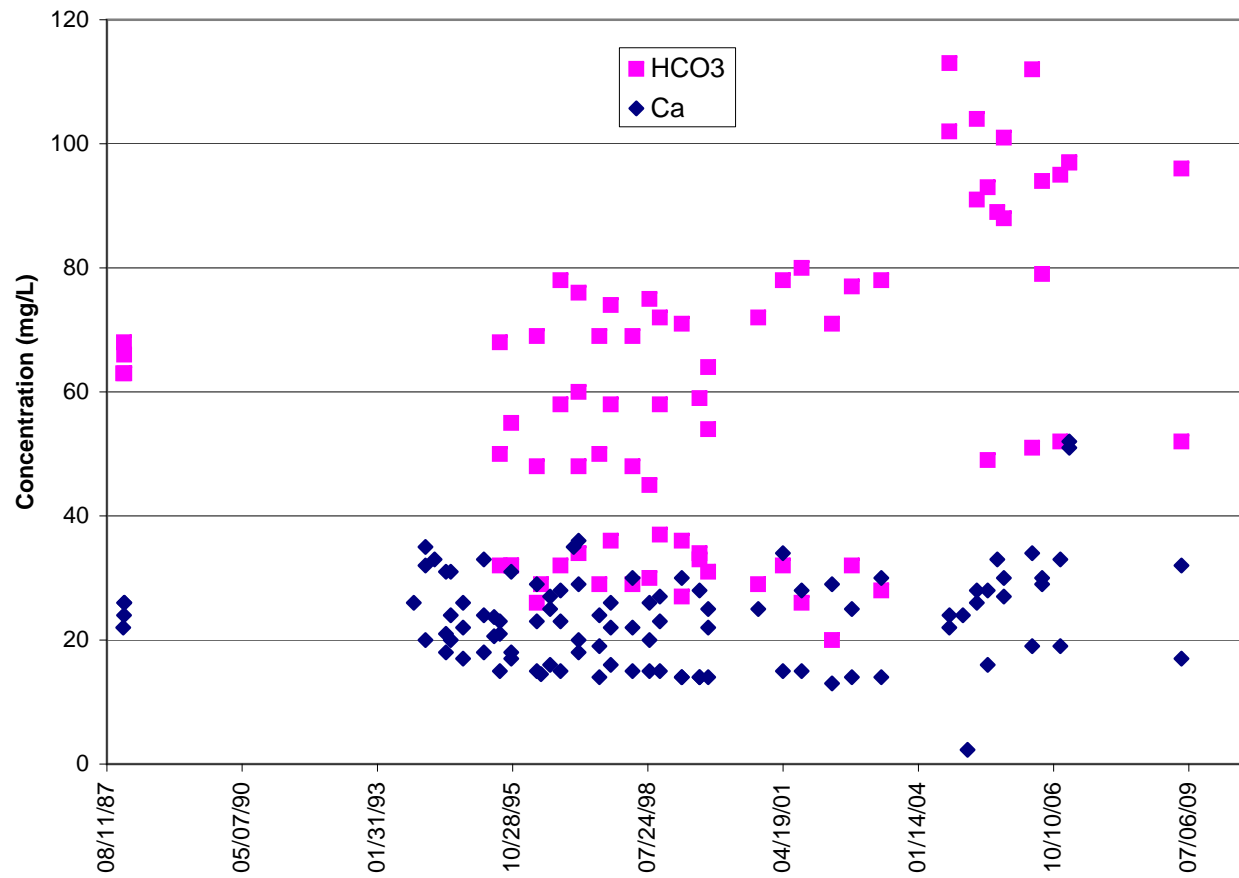
This serves as notice of this possibility. All sub-contract data will be clearly notated on your analytical report.

Visit our web site at www.energylab.com for additional information, downloadable fee schedule, forms, and links.

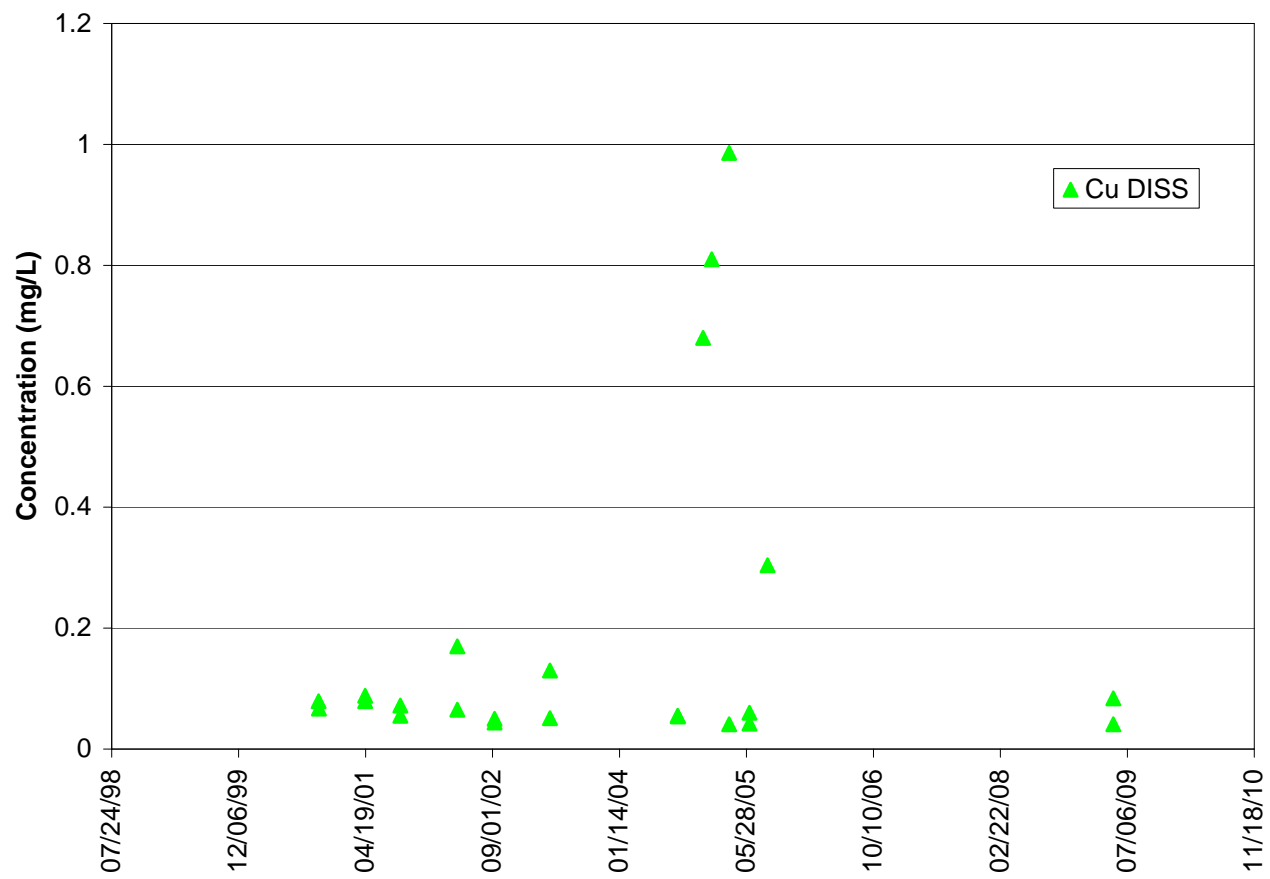
APPENDIX B

CONCENTRATION TIME SERIES GRAPHS

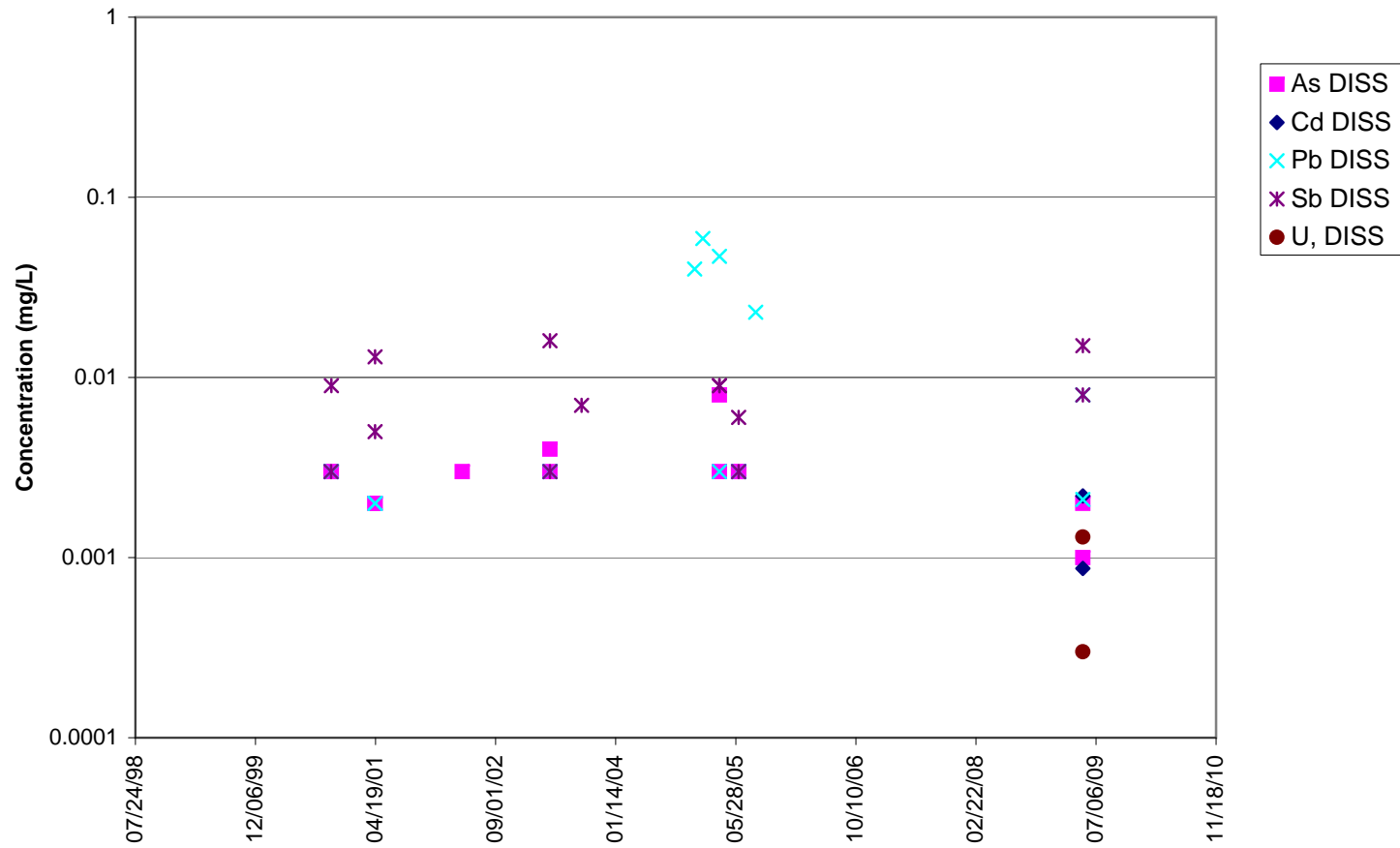
Troy Mine Adit Water Calcium and Bicarbonate Concentrations



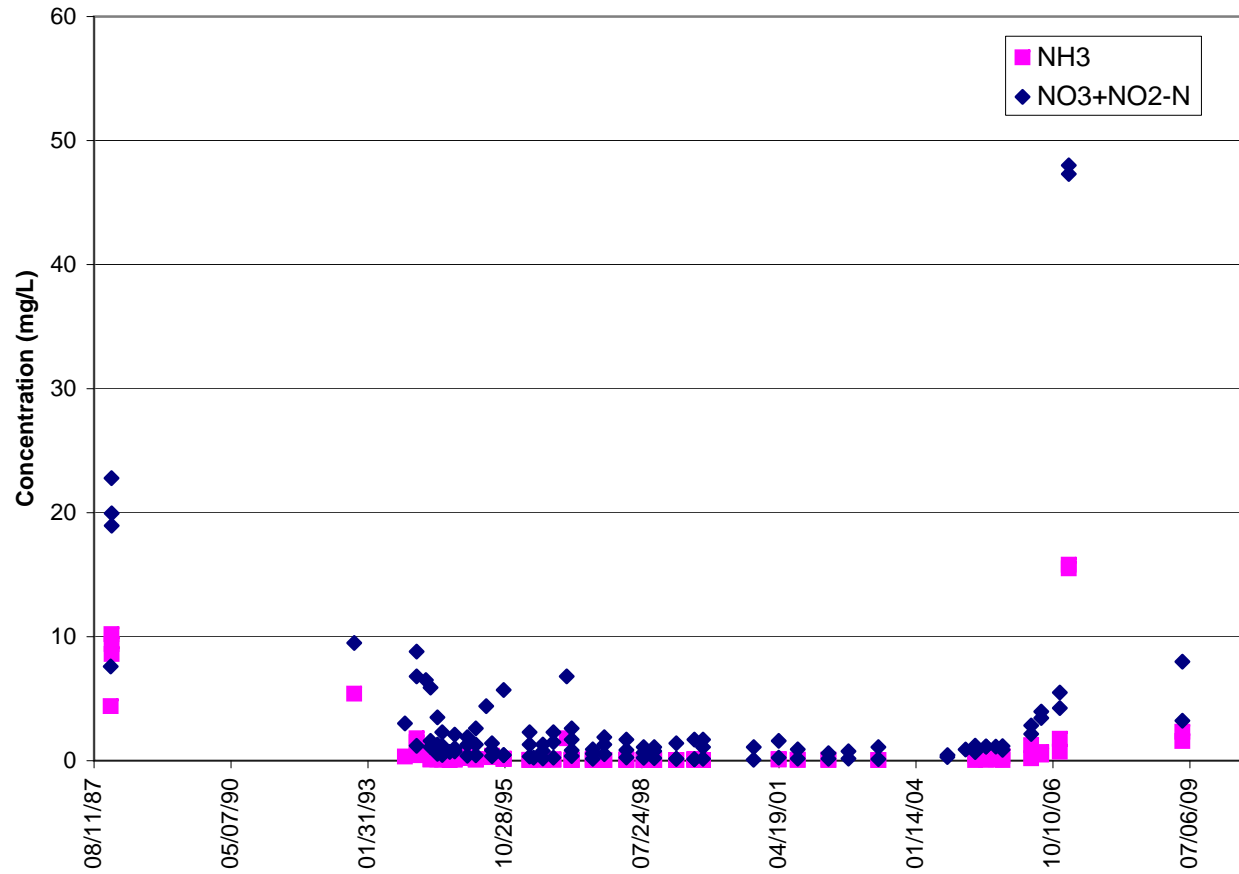
Troy Mine Adit Water Copper Concentrations



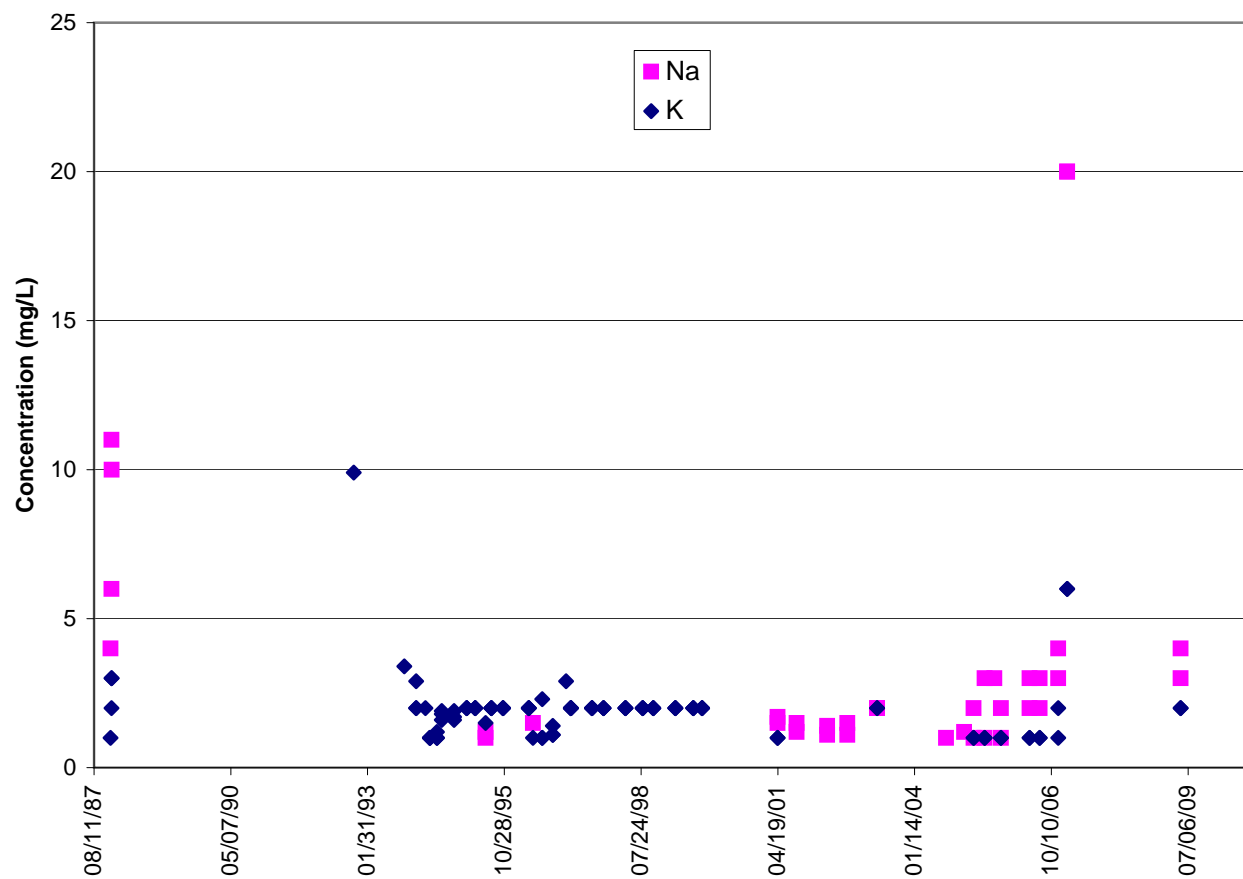
Troy Mine Adit Water Metal Concentrations



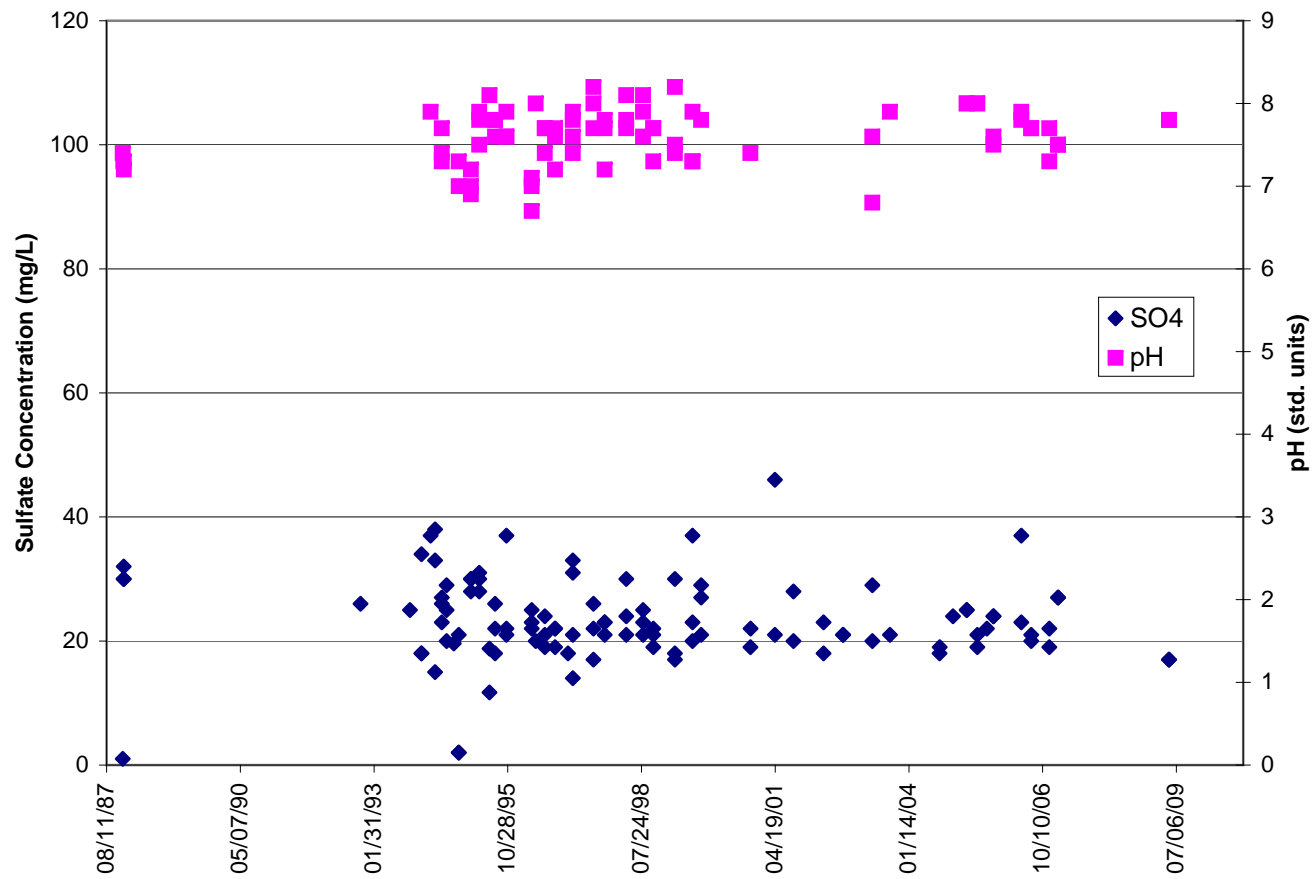
Troy Mine Adit Water Nitrogen Concentrations



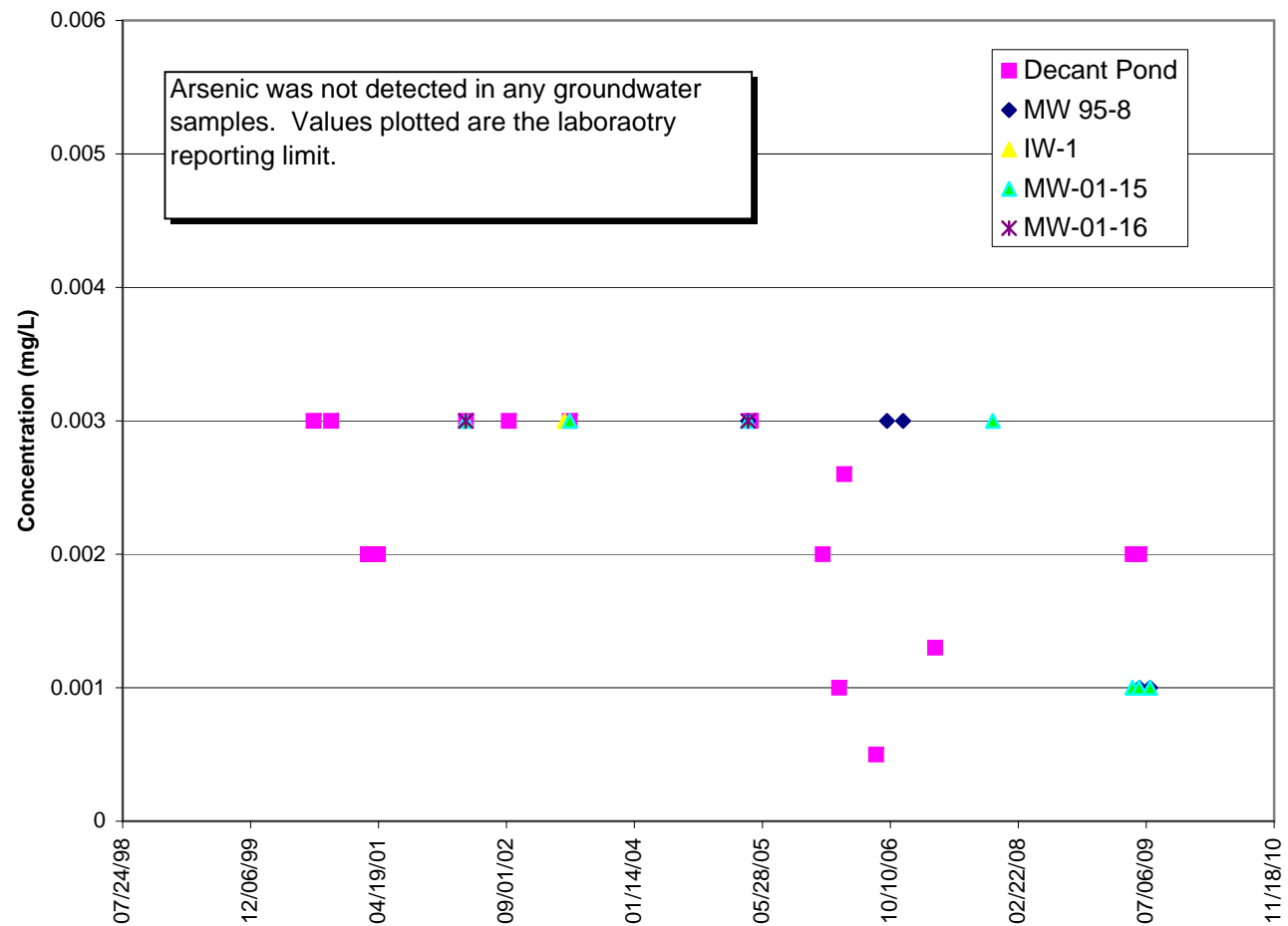
Troy Mine Adit Water Sodium and Potassium Concentrations



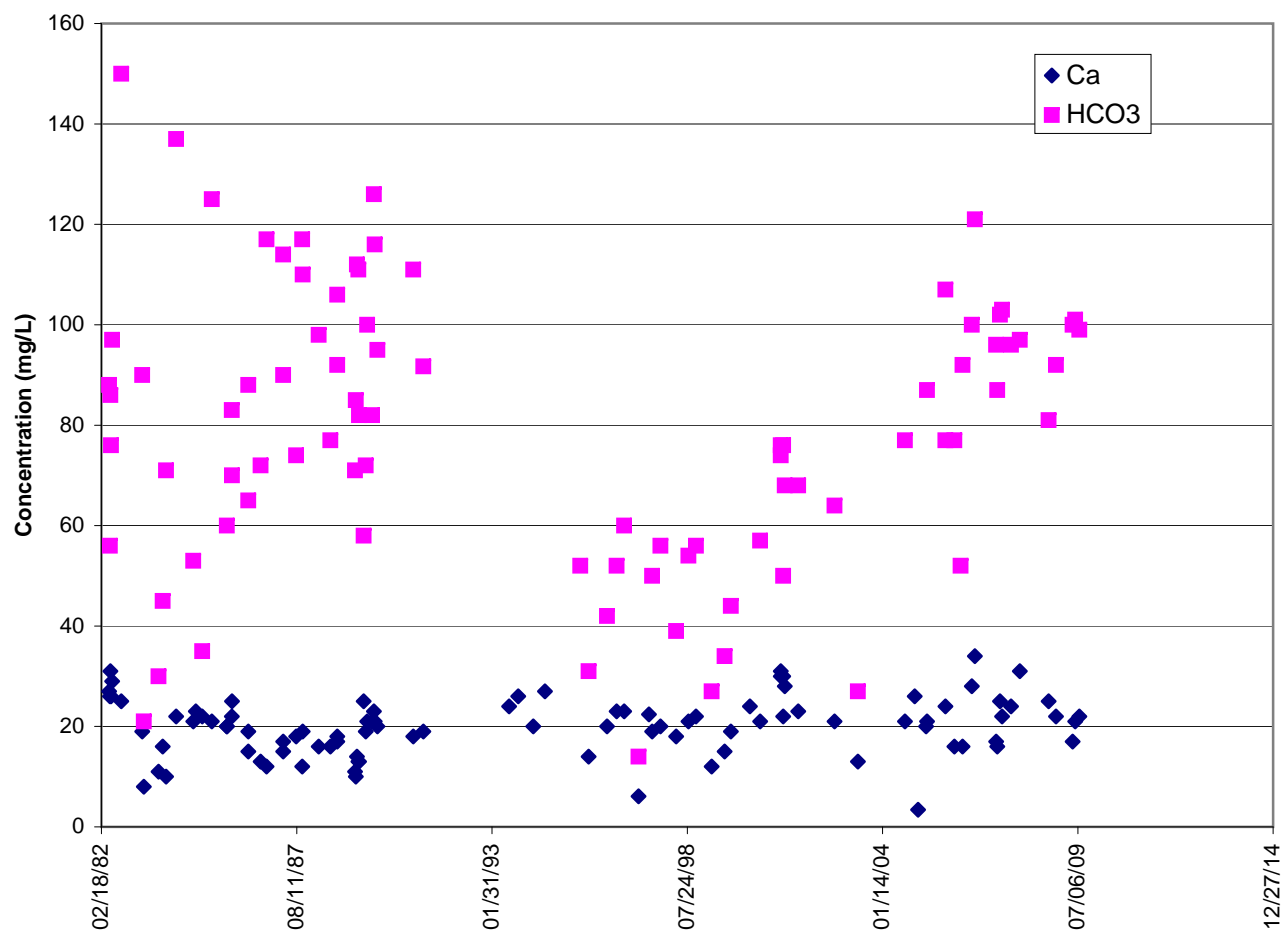
Troy Mine Adit Water pH and Sulfate Concentrations



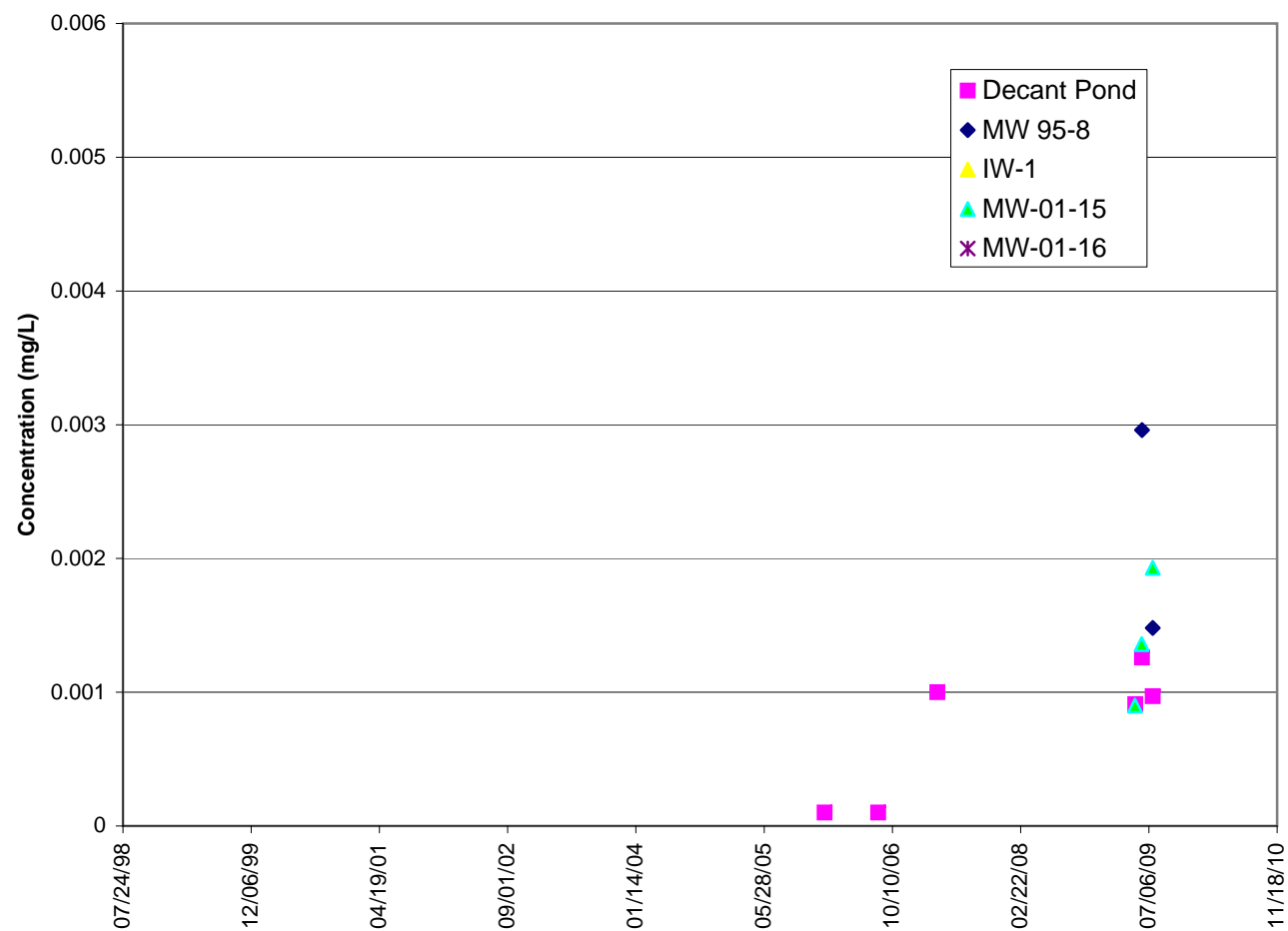
Troy Mine Decant Pond and Groundwater Dissolved Arsenic Concentrations



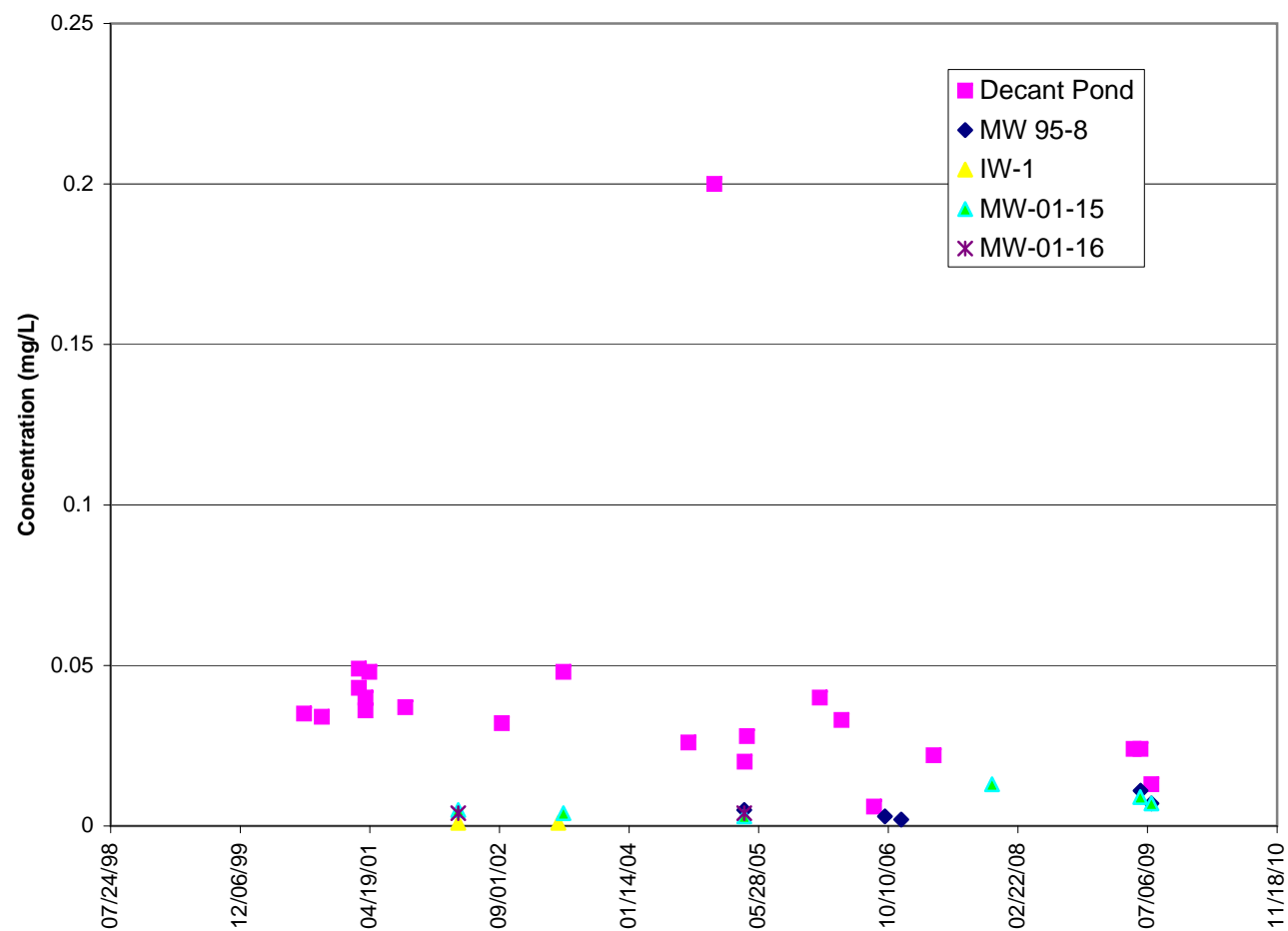
Troy Mine Decant Pond Calcium and Bicarbonate Concentrations



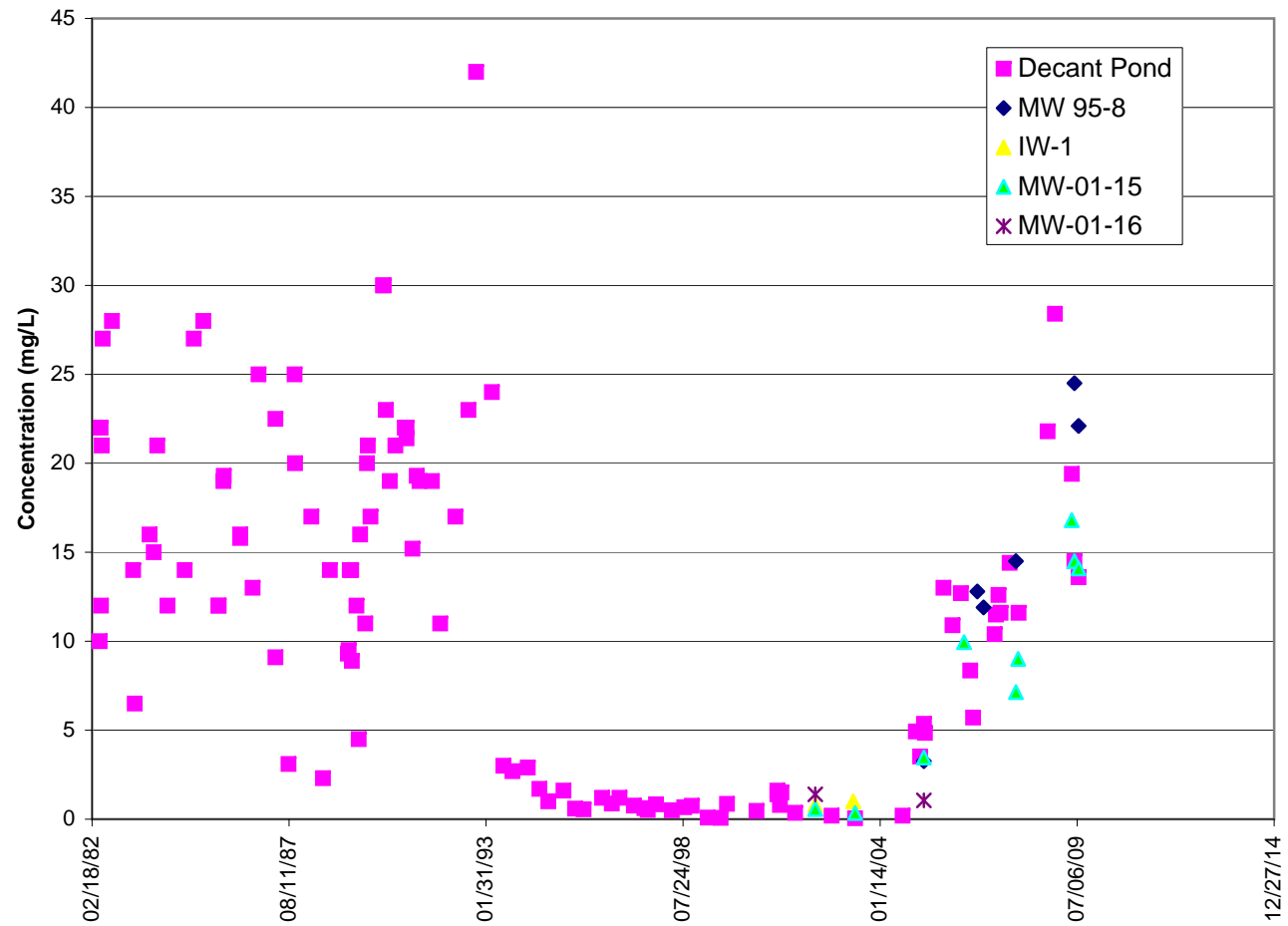
Troy Mine Decant Pond and Groundwater Dissolved Cadmium Concentrations



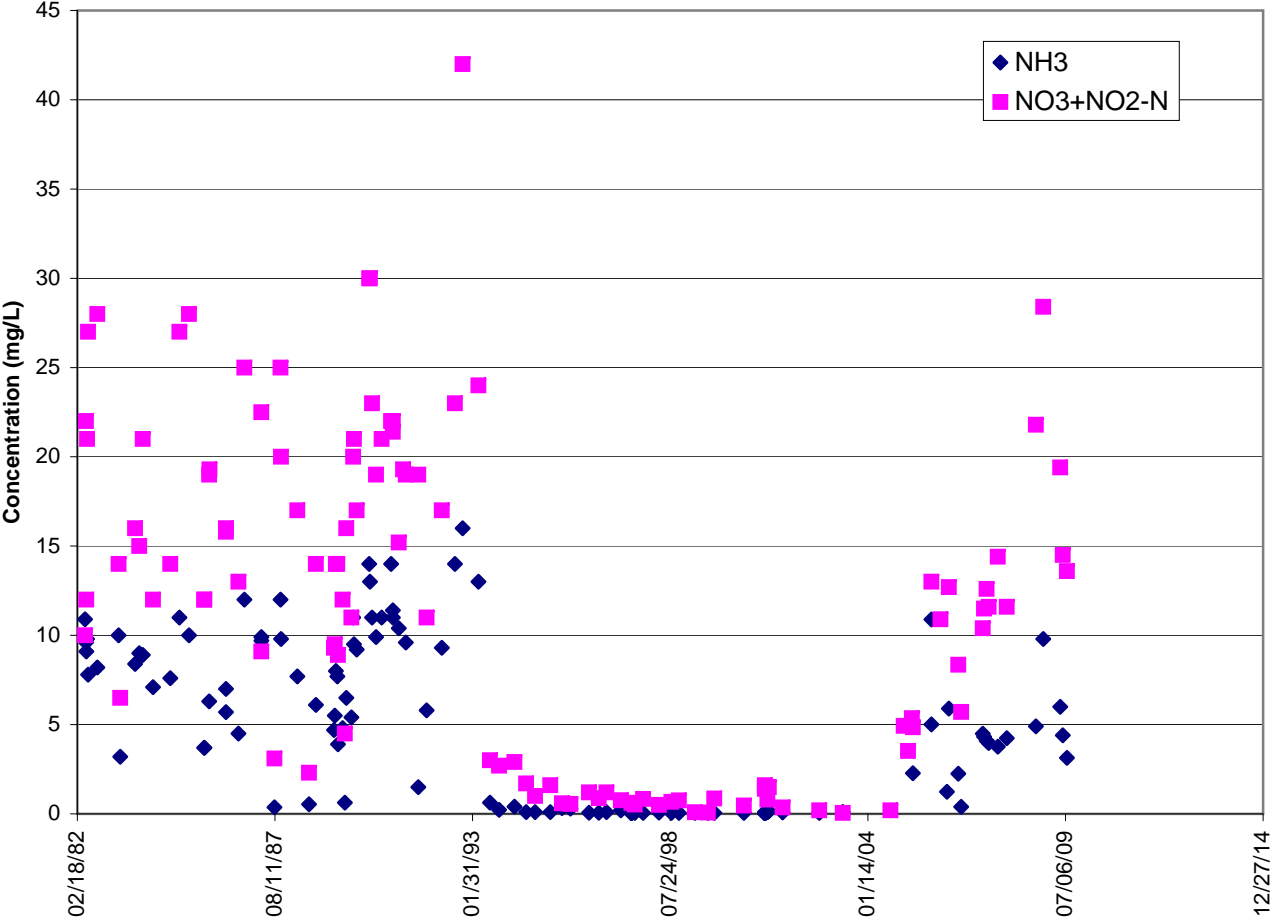
Troy Mine Decant Pond and Groundwater Dissolved Copper Concentrations



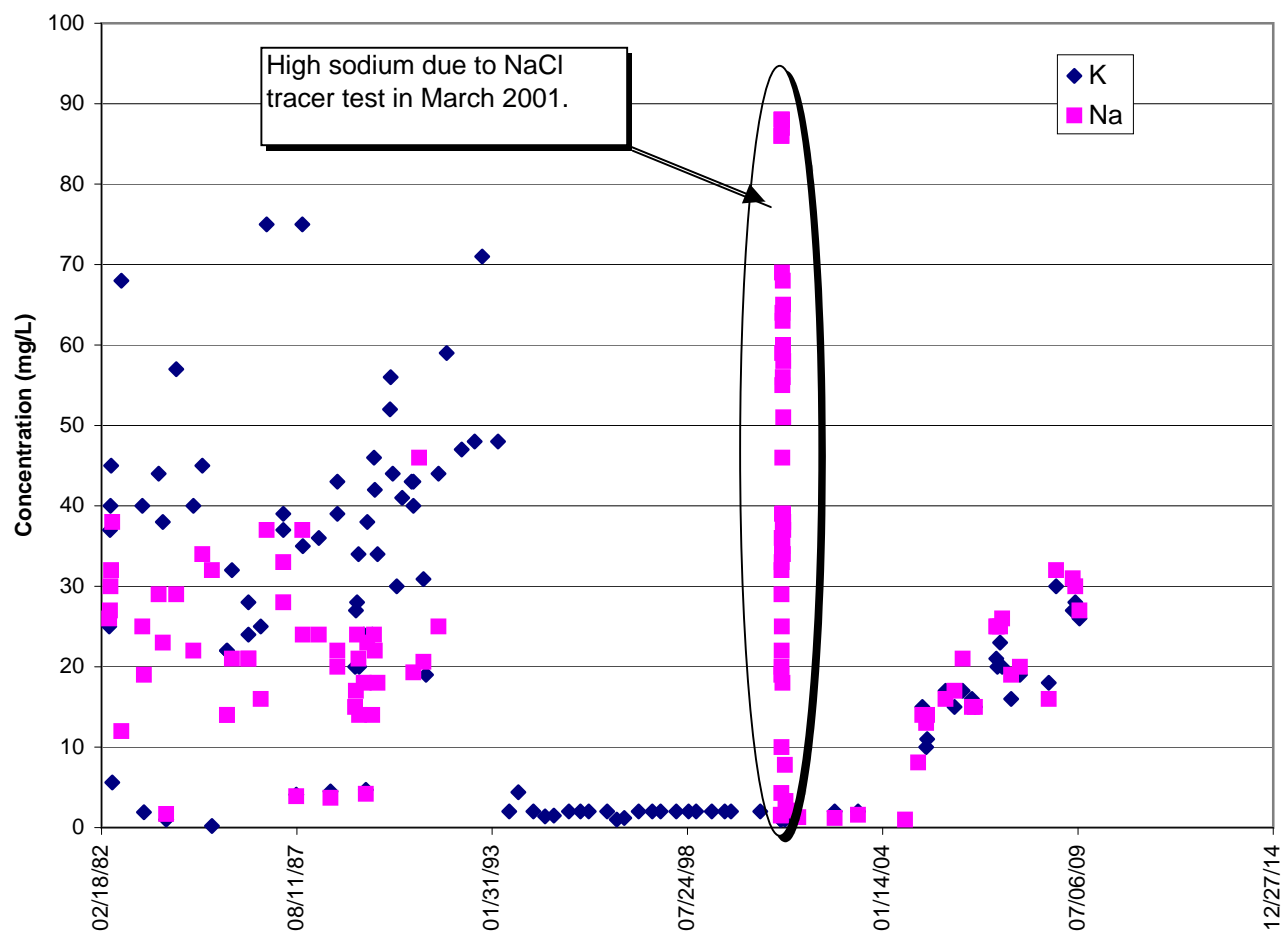
Troy Mine Decant Pond and Groundwater Nitrate+Nitrite Concentrations



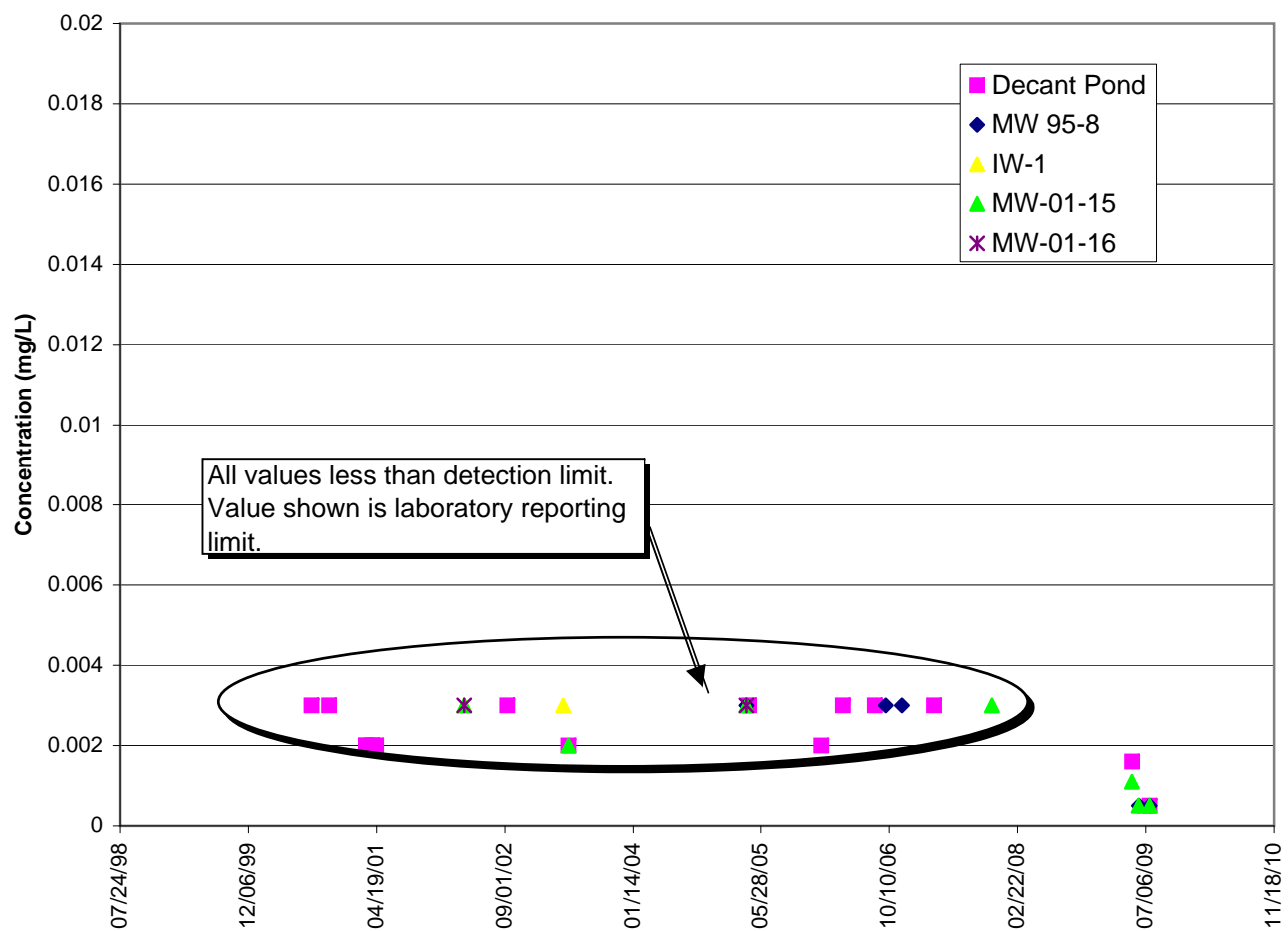
Troy Mine Decant Pond Nitrogen Concentrations



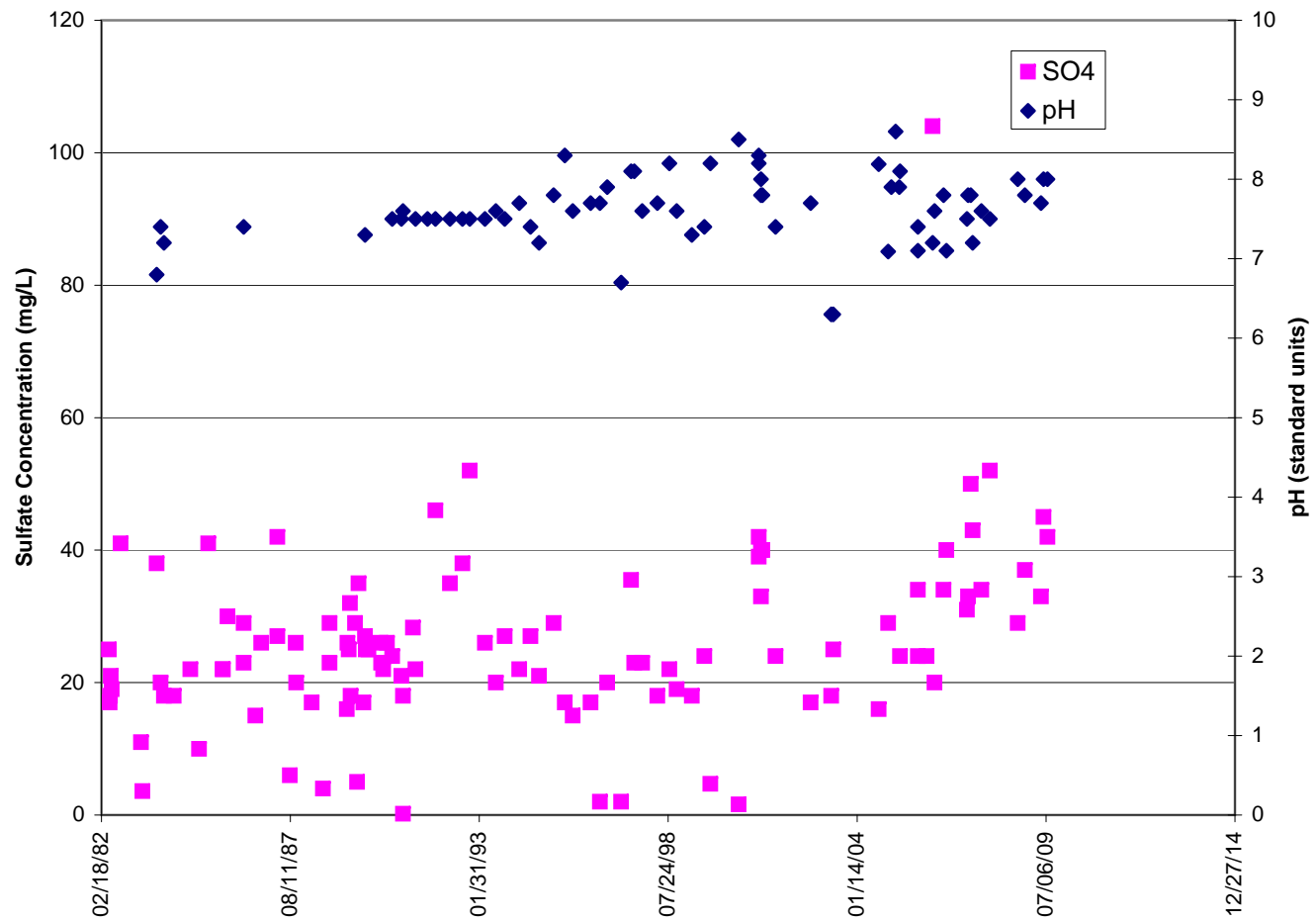
Troy Mine Decant Pond Sodium and Potassium Concentrations



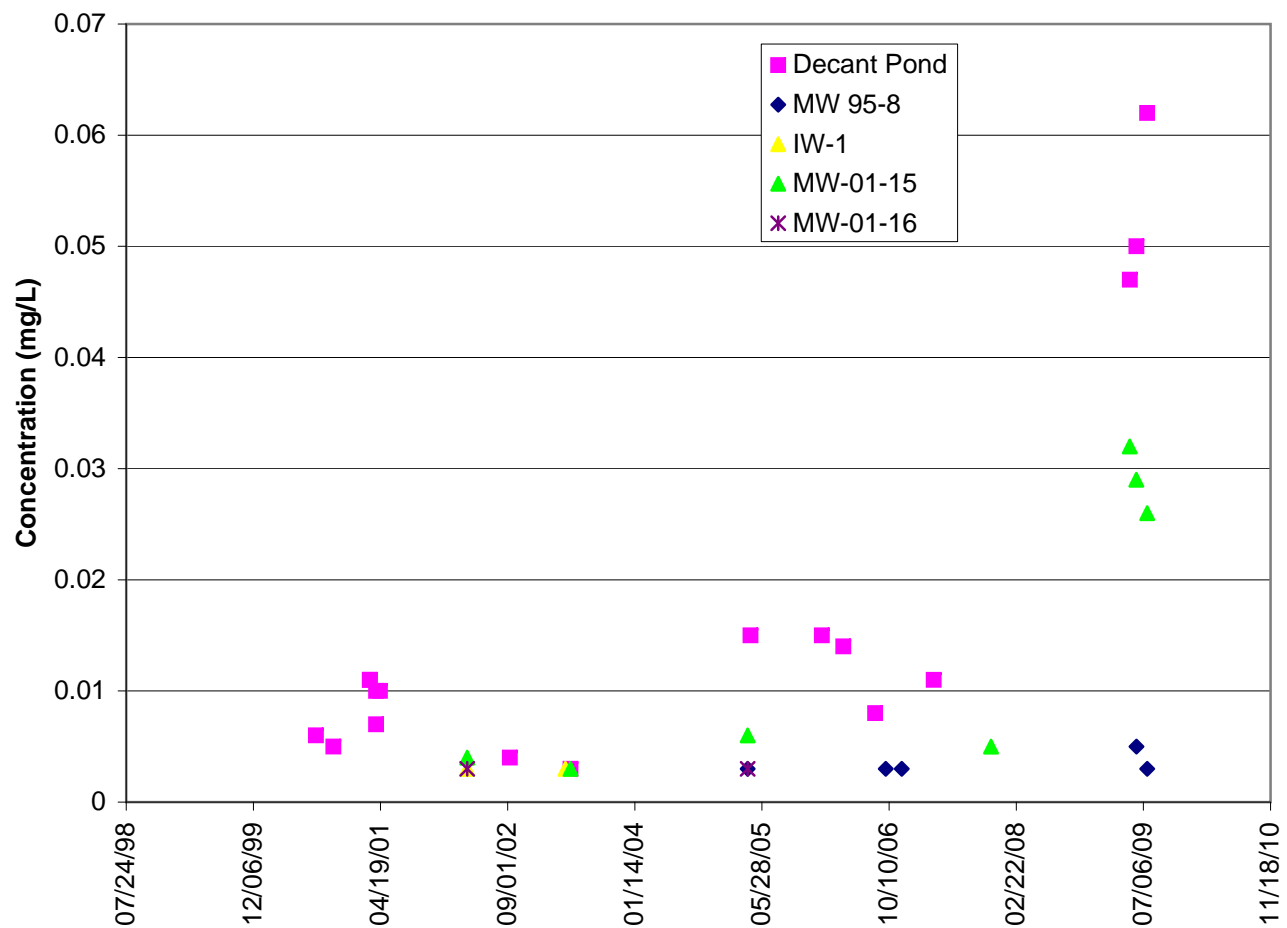
Troy Mine Decant Pond and Groundwater Dissolved Lead Concentrations



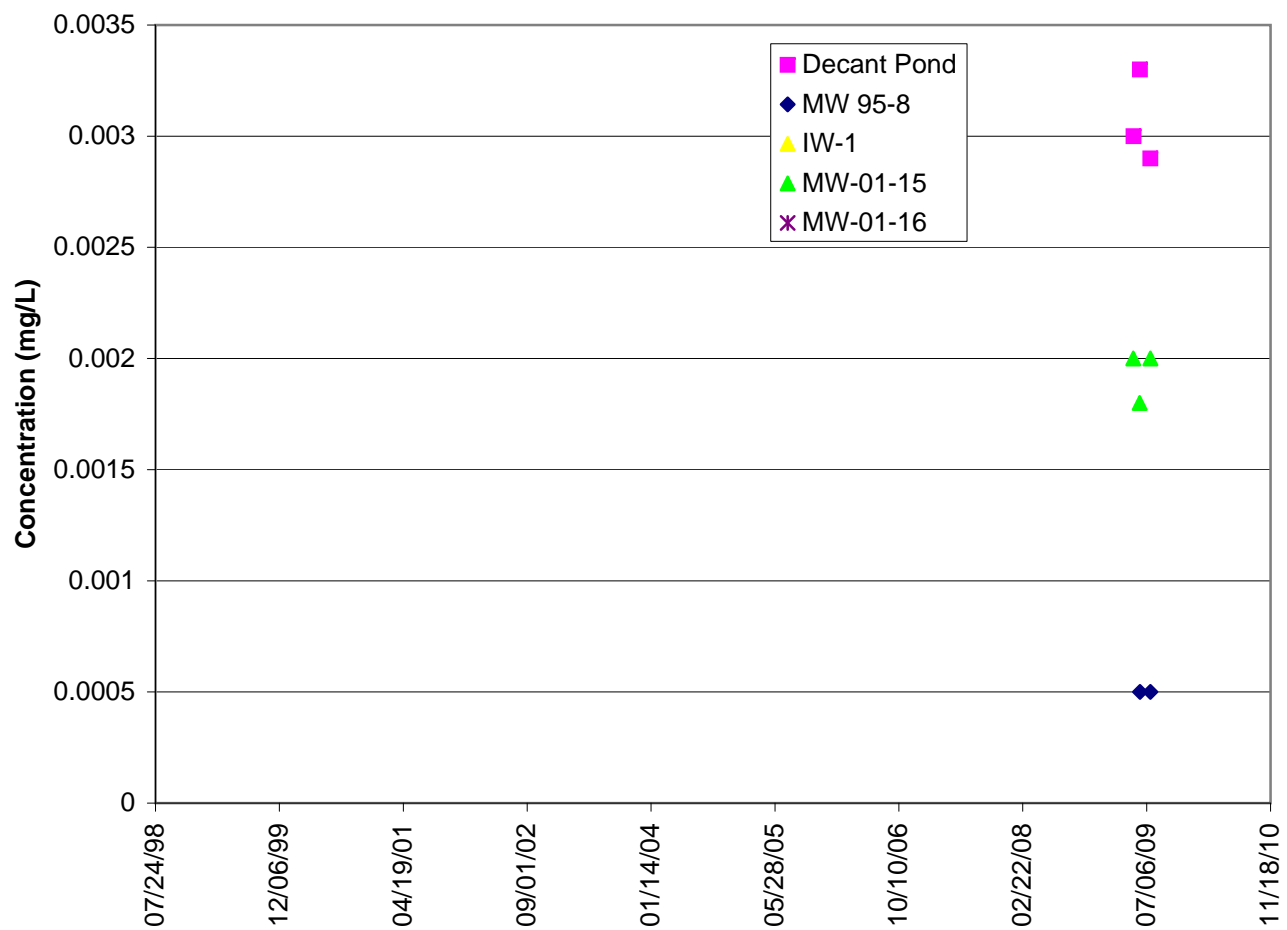
Troy Mine Decant Pond pH and Sulfate Concentrations



Troy Mine Decant Pond and Groundwater Dissolved Antimony Concentrations



Troy Mine Decant Pond and Groundwater Dissolved Uranium Concentrations



APPENDIX C

GEOCHEMICAL MODEL RESULTS

SpecE8_output_GSS_Decant Pond July.txt

Temperature = 20.9 C Pressure = 1.013 bars
 pH = 8.000 log fO2 = -22.996
 Eh = 0.4280 volts pe = 7.3361
 Ionic strength = 0.004457
 Charge imbalance = -0.000212 eq/kg (-6.354% error)
 Activity of water = 0.999996
 Solvent mass = 1.000000 kg
 Solution mass = 1.000303 kg
 Solution density = 1.016 g/cm3
 Chlorinity = 0.000111 molal
 Dissolved solids = 303 mg/kg sol'n
 Elect. conductivity = 345.82 uS/cm (or umho/cm)
 Hardness = 74.36 mg/kg sol'n as CaCO3
 carbonate = 74.36 mg/kg sol'n as CaCO3
 non-carbonate = 0.00 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 78.60 mg/kg sol'n as CaCO3
 Water type = Na-HCO3

Nernst redox couples	Eh (volts)	pe
$e^- + .25*O_2(aq) + H^+ = .5*H_2O$	0.4280	7.3361
$8*e^- + 9*H^+ + NO_3^- = 3*H_2O + NH_3(aq)$	0.5686	9.7454

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HCO3-	0.001538	93.82	0.9312	-2.8439
Na+	0.001154	26.52	0.9312	-2.9687
NO3-	0.0009562	59.27	0.9299	-3.0510
K+	0.0006533	25.54	0.9299	-3.2164
Ca++	0.0005140	20.60	0.7595	-3.4085
SO4--	0.0004036	38.76	0.7511	-3.5184
SiO2(aq)	0.0002188	13.14	1.0000	-3.6600
Mg++	0.0001890	4.592	0.7674	-3.8385
N2(aq)	0.0001169	3.274	1.0000	-3.9321
Cl-	0.0001111	3.936	0.9299	-3.9860
CO2(aq)	3.444e-005	1.515	1.0000	-4.4629
CaSO4(aq)	1.504e-005	2.047	1.0000	-4.8228
F-	1.031e-005	0.1959	0.9306	-5.0179
MgSO4(aq)	1.025e-005	1.233	1.0000	-4.9894
CO3--	8.015e-006	0.4808	0.7533	-5.2191
CaHCO3+	6.817e-006	0.6890	0.9312	-5.1973
CaCO3(aq)	4.634e-006	0.4637	1.0000	-5.3340
Mn++	3.567e-006	0.1959	0.7595	-5.5672
MgHCO3+	2.475e-006	0.2111	0.9312	-5.6374
NaHCO3(aq)	2.386e-006	0.2004	1.0000	-5.6223
HSiO3-	2.376e-006	0.1831	0.9312	-5.6552
KS04-	1.506e-006	0.2035	0.9312	-5.8531
MgCO3(aq)	8.065e-007	0.06798	1.0000	-6.0934
Ba++	7.945e-007	0.1091	0.7554	-6.2218
OH-	7.732e-007	0.01315	0.9306	-6.1430
Sb(OH)3(aq)	5.014e-007	0.08661	1.0000	-6.2998
Cu++	2.015e-007	0.01280	0.7595	-6.8153
MnSO4(aq)	1.787e-007	0.02697	1.0000	-6.7480
HP04--	1.681e-007	0.01612	0.7511	-6.8988
NaHSiO3(aq)	1.165e-007	0.01166	1.0000	-6.9336
HAsO4--	6.269e-008	0.008770	0.7511	-7.3271
MgF+	3.420e-008	0.001480	0.9312	-7.4970
NaCO3-	2.564e-008	0.002128	0.9312	-7.6220
H2PO4-	2.269e-008	0.002200	0.9312	-7.6751

	SpecE8_output_GSS_Decant	Pond	Jul y. txt	
CaF+	1. 915e-008	0. 001131	0. 9312	-7. 7488
NaCl (aq)	1. 819e-008	0. 001063	1. 0000	-7. 7401
MgCl +	1. 219e-008	0. 0007281	0. 9312	-7. 9450
H+	1. 067e-008	1. 075e-005	0. 9374	-8. 0000
CaCl +	8. 781e-009	0. 0006630	0. 9312	-8. 0874
Cd++	8. 495e-009	0. 0009547	0. 7554	-8. 1926
UO2(CO3) 2--	6. 367e-009	0. 002483	0. 7511	-8. 3204
UO2(CO3) 3----	5. 581e-009	0. 002511	0. 3179	-8. 7510
H2AsO4-	3. 004e-009	0. 0004233	0. 9312	-8. 5532
KCl (aq)	1. 852e-009	0. 0001380	1. 0000	-8. 7324
BaCO3(aq)	1. 474e-009	0. 0002907	1. 0000	-8. 8316
NaF(aq)	1. 007e-009	4. 228e-005	1. 0000	-8. 9968
MnCl +	5. 536e-010	5. 003e-005	0. 9312	-9. 2877
HSO4-	2. 831e-010	2. 748e-005	0. 9312	-9. 5789
HF(aq)	1. 321e-010	2. 641e-006	1. 0000	-9. 8792
NaOH(aq)	1. 270e-010	5. 079e-006	1. 0000	-9. 8961
UO2CO3(aq)	4. 893e-011	1. 615e-005	1. 0000	-10. 3104
Sb(OH) 4-	4. 654e-011	8. 829e-006	0. 9312	-10. 3631
BaCl +	2. 011e-011	3. 474e-006	0. 9312	-10. 7275
AsO4---	1. 953e-011	2. 713e-006	0. 5250	-10. 9891
PO4---	1. 034e-011	9. 819e-007	0. 5250	-11. 2653
BaF+	3. 915e-012	6. 119e-007	0. 9312	-11. 4382
Cu+	3. 531e-012	2. 243e-007	0. 9312	-11. 4830
CdSO4(aq)	1. 965e-012	4. 094e-007	1. 0000	-11. 7067
CaCl 2(aq)	1. 010e-012	1. 121e-007	1. 0000	-11. 9955
UO2OH+	8. 973e-013	2. 575e-007	0. 9312	-12. 0780
Sb(OH) 2F(aq)	5. 990e-013	1. 047e-007	1. 0000	-12. 2226
HN03(aq)	4. 065e-013	2. 561e-008	1. 0000	-12. 3909
HCl (aq)	2. 128e-013	7. 756e-009	1. 0000	-12. 6721
Sb(OH) 2+	1. 950e-013	3. 036e-008	0. 9312	-12. 7410
CdCl 2(aq)	1. 452e-013	2. 662e-008	1. 0000	-12. 8379
H3PO4(aq)	3. 007e-014	2. 946e-009	1. 0000	-13. 5219
KHSO4(aq)	1. 049e-014	1. 428e-009	1. 0000	-13. 9792
MgP2O7--	7. 430e-015	1. 473e-009	0. 7511	-14. 2533
H3AsO4(aq)	4. 727e-015	6. 708e-010	1. 0000	-14. 3254
UO2++	2. 313e-015	6. 245e-010	0. 7533	-14. 7588
UO2F+	2. 005e-015	5. 793e-010	0. 9312	-14. 7288
UO2SO4(aq)	5. 624e-016	2. 058e-010	1. 0000	-15. 2499
HF2-	3. 137e-016	1. 223e-011	0. 9312	-15. 5345
HP2O7---	7. 509e-017	1. 313e-011	0. 5250	-16. 4043
UO2F2(aq)	5. 572e-017	1. 716e-011	1. 0000	-16. 2540
P2O7----	7. 591e-018	1. 320e-012	0. 3179	-17. 6174
CdCl 3-	3. 558e-018	7. 781e-013	0. 9312	-17. 4797
H2P2O7--	2. 309e-018	4. 061e-013	0. 7511	-17. 7609
UO2(SO4) 2--	1. 699e-018	7. 851e-013	0. 7511	-17. 8940
KP2O7---	5. 703e-019	1. 215e-013	0. 5250	-18. 5238
UO2Cl +	2. 658e-019	8. 116e-014	0. 9312	-18. 6064
UO2F3-	9. 968e-020	3. 259e-014	0. 9312	-19. 0323
(UO2) 2(OH) 2++	7. 451e-020	4. 276e-014	0. 7533	-19. 2508
(UO2) 3(OH) 5+	8. 158e-021	7. 300e-015	0. 9312	-20. 1193
(UO2) 3(CO3) 6(6-)	4. 328e-021	5. 063e-015	0. 0759	-21. 4837
UO2+	2. 410e-021	6. 505e-016	0. 9312	-20. 6490
UO2F4--	6. 889e-024	2. 383e-018	0. 7511	-23. 2862
H3P2O7-	4. 158e-024	7. 356e-019	0. 9312	-23. 4120
Mn+++	2. 126e-024	1. 167e-019	0. 5317	-23. 9469
UO2Cl 2(aq)	1. 289e-024	4. 394e-019	1. 0000	-23. 8897
NO2-	1. 449e-025	6. 664e-021	0. 9299	-24. 8705
U(OH) 4(aq)	1. 336e-025	4. 087e-020	1. 0000	-24. 8742
HAsO2(aq)	5. 108e-026	5. 511e-021	1. 0000	-25. 2918
As(OH) 3(aq)	4. 414e-026	5. 558e-021	1. 0000	-25. 3551
O2(aq)	1. 391e-026	4. 449e-022	1. 0000	-25. 8567
AsO2-	2. 552e-027	2. 728e-022	0. 9312	-26. 6240
H2AsO3-	2. 522e-027	3. 150e-022	0. 9312	-26. 6292

	SpecE8_output_GSS_Decant	Pond	Jul y. txt	
HN02(aq)	2. 615e-030	1. 229e-025	1. 0000	-29. 5824
As020H--	2. 529e-030	3. 133e-025	0. 7511	-29. 7214
H4P207(aq)	1. 154e-030	2. 053e-025	1. 0000	-29. 9378
Mn04--	7. 769e-033	9. 238e-028	0. 7511	-32. 2339
Cl O-	6. 945e-033	3. 572e-028	0. 9312	-32. 1893
Formate	4. 587e-034	2. 064e-029	0. 9306	-33. 3697
H02-	2. 881e-034	9. 507e-030	0. 9312	-33. 5714
H2(aq)	1. 373e-034	2. 766e-031	1. 0000	-33. 8625
Mn04-	3. 368e-035	4. 005e-030	0. 9306	-34. 5039
Ca(For)+	5. 014e-036	4. 266e-031	0. 9312	-35. 3307
Mg(For)+	1. 971e-036	1. 366e-031	0. 9312	-35. 7363
Na(For)(aq)	5. 277e-037	3. 588e-032	1. 0000	-36. 2776
K(For)(aq)	2. 767e-037	2. 327e-032	1. 0000	-36. 5580
Mn(For)+	6. 631e-038	6. 626e-033	0. 9312	-37. 2094
Formi c_aci d(aq)	2. 451e-038	1. 128e-033	1. 0000	-37. 6107
S03--	1. 733e-038	1. 387e-033	0. 7533	-37. 8843
Cu(For)+	7. 300e-039	7. 923e-034	0. 9312	-38. 1676
Ba(For)+	6. 564e-039	1. 197e-033	0. 9312	-38. 2138
HS03-	2. 191e-039	1. 776e-034	0. 9312	-38. 6902
Si F6--	7. 276e-040	1. 033e-034	0. 7511	-39. 2624
Cd(For)+	2. 178e-040	3. 427e-035	0. 9312	-39. 6930
NH4+	5. 474e-041	9. 872e-037	0. 9292	-40. 2936
CO(aq)	5. 096e-041	1. 427e-036	1. 0000	-40. 2927
Oxal ate	1. 725e-041	1. 518e-036	0. 7511	-40. 8875
NH3(aq)	2. 163e-042	3. 682e-038	1. 0000	-41. 6650
HS05-	9. 080e-043	1. 026e-037	0. 9312	-42. 0729
U(C03)5(6-)	8. 077e-044	4. 345e-038	0. 0759	-44. 2128
U0H+++	3. 124e-045	7. 965e-040	0. 5317	-44. 7796
H-Oxal ate	2. 464e-045	2. 193e-040	0. 9312	-44. 6393
S02(aq)	1. 368e-045	8. 763e-041	1. 0000	-44. 8638
UF3+	2. 245e-046	6. 623e-041	0. 9312	-45. 6796
UF2++	1. 377e-046	3. 801e-041	0. 7533	-45. 9840
UF4(aq)	1. 867e-047	5. 860e-042	1. 0000	-46. 7289
UF+++	2. 516e-048	6. 466e-043	0. 5317	-47. 8736
U(S04)2(aq)	1. 329e-049	5. 715e-044	1. 0000	-48. 8765
US04++	9. 484e-050	3. 168e-044	0. 7533	-49. 1460
Oxal i c_aci d(aq)	4. 240e-052	3. 816e-047	1. 0000	-51. 3726
U+++	2. 345e-052	5. 580e-047	0. 3288	-52. 1130
Cl 02-	6. 132e-054	4. 135e-049	0. 9312	-53. 2433
UCl +++	8. 519e-055	2. 329e-049	0. 5317	-54. 3440
S208--	1. 538e-059	2. 954e-054	0. 7511	-58. 9373
Cl 03-	6. 091e-061	5. 082e-056	0. 9306	-60. 2465
S206--	5. 262e-063	8. 424e-058	0. 7511	-62. 4031
Ca(For)2(aq)	1. 569e-068	2. 041e-063	1. 0000	-67. 8042
Mg(For)2(aq)	6. 558e-069	7. 496e-064	1. 0000	-68. 1832
U+++	1. 465e-069	3. 487e-064	0. 5317	-69. 1084
Mn(For)2(aq)	3. 523e-070	5. 105e-065	1. 0000	-69. 4531
Na(For)2-	1. 403e-070	1. 585e-065	0. 9312	-69. 8839
K(For)2-	6. 758e-071	8. 725e-066	0. 9312	-70. 2011
Cu(For)2(aq)	6. 743e-071	1. 035e-065	1. 0000	-70. 1712
N3-	5. 954e-071	2. 501e-066	0. 9312	-70. 2561
Ba(For)2(aq)	1. 775e-071	4. 035e-066	1. 0000	-70. 7508
Cd(For)2(aq)	9. 697e-072	1. 963e-066	1. 0000	-71. 0133
Cl 04-	2. 827e-072	2. 811e-067	0. 9306	-71. 5799
Formal dehyde(aq)	1. 294e-073	3. 884e-069	1. 0000	-72. 8881
HN3(aq)	3. 171e-074	1. 364e-069	1. 0000	-73. 4988
U02Cl 03+	3. 376e-075	1. 193e-069	0. 9312	-74. 5025
S205--	8. 522e-083	1. 228e-077	0. 7511	-82. 1937
Cu(NH3)2++	3. 527e-083	3. 441e-078	0. 7533	-82. 5757
Cd(NH3)2++	3. 559e-087	5. 212e-082	0. 7533	-86. 5717
HCN(aq)	7. 301e-088	1. 973e-083	1. 0000	-87. 1366
Urea(aq)	3. 579e-088	2. 149e-083	1. 0000	-87. 4462
CN-	3. 452e-089	8. 978e-085	0. 9299	-88. 4936

	SpecE8_output_GSS_Decant	Pond	Jul y. txt	
Methanol (aq)	1. 840e-092	5. 894e-088	1. 0000	-91. 7352
Gl ycol ate	1. 466e-098	1. 100e-093	0. 9312	-97. 8647
Ca(Gl yc)+	2. 701e-100	3. 109e-095	0. 9312	-99. 5993
HS-	5. 647e-101	1. 867e-096	0. 9306	-100. 2794
Mg(Gl yc)+	4. 776e-101	4. 743e-096	0. 9312	-100. 3519
Na(Gl yc) (aq)	1. 701e-101	1. 667e-096	1. 0000	-100. 7693
K(Gl yc) (aq)	8. 918e-102	1. 018e-096	1. 0000	-101. 0497
H2S(aq)	5. 951e-102	2. 028e-097	1. 0000	-101. 2254
Cu(Gl yc)+	1. 973e-102	2. 733e-097	0. 9312	-101. 7359
Mn(Gl yc)+	1. 623e-102	2. 108e-097	0. 9312	-101. 8207
Gl ycol i c_aci d(aq)	9. 419e-103	7. 161e-098	1. 0000	-102. 0260
Ba(Gl yc)+	9. 023e-104	1. 916e-098	0. 9312	-103. 0756
Cd(Gl yc)+	7. 530e-105	1. 411e-099	0. 9312	-104. 1541
S2O4--	3. 383e-105	4. 333e-100	0. 7554	-104. 5925
S--	6. 050e-106	1. 939e-101	0. 7554	-105. 3401
Methane(aq)	3. 863e-106	6. 196e-102	1. 0000	-105. 4131
S2O3--	1. 387e-107	1. 555e-102	0. 7511	-106. 9821
Acetate	1. 393e-111	8. 221e-107	0. 9319	-110. 8868
CaCH3COO+	4. 766e-114	4. 722e-109	0. 9312	-113. 3528
MgCH3COO+	4. 194e-114	3. 495e-109	0. 9312	-113. 4083
NaCH3COO(aq)	1. 130e-114	9. 268e-110	1. 0000	-113. 9469
Aceti c_aci d(aq)	7. 443e-115	4. 468e-110	1. 0000	-114. 1283
KCH3COO(aq)	4. 264e-115	4. 183e-110	1. 0000	-114. 3702
MnCH3COO+	6. 573e-116	7. 490e-111	0. 9312	-115. 2132
CuCH3COO+	4. 028e-116	4. 937e-111	0. 9312	-115. 4258
BaCH3COO+	8. 068e-117	1. 584e-111	0. 9312	-116. 1242
CdCH3COO+	8. 189e-118	1. 404e-112	0. 9312	-117. 1177
Mal onate	2. 581e-120	2. 633e-115	0. 7511	-119. 7125
Cu(NH3)3++	5. 739e-122	6. 577e-117	0. 7533	-121. 3642
H-Mal onate	1. 013e-122	1. 044e-117	0. 9312	-122. 0253
CuCH3COO(aq)	9. 601e-123	1. 177e-117	1. 0000	-122. 0177
ASH3(aq)	4. 362e-123	3. 399e-118	1. 0000	-122. 3603
Mal oni c_aci d(aq)	6. 757e-128	7. 029e-123	1. 0000	-127. 1703
Methanami ne(aq)	1. 057e-131	3. 283e-127	1. 0000	-130. 9757
Gl yci ne(aq)	4. 580e-135	3. 437e-130	1. 0000	-134. 3391
Cu(Gl y)+	4. 286e-135	5. 896e-130	0. 9312	-134. 3988
Mg(Gl y)+	2. 785e-137	2. 739e-132	0. 9312	-136. 5861
Mn(Gl y)+	8. 271e-139	1. 067e-133	0. 9312	-138. 1134
Ca(Gl y)+	5. 249e-139	5. 989e-134	0. 9312	-138. 3109
S3O6--	6. 336e-140	1. 217e-134	0. 7511	-139. 3225
Cd(Gl y)+	2. 144e-140	3. 996e-135	0. 9312	-139. 6998
Ba(Gl y)+	1. 070e-141	2. 261e-136	0. 9312	-141. 0016
Acetal dehyde(aq)	5. 669e-149	2. 497e-144	1. 0000	-148. 2465
SCN-	1. 708e-150	9. 920e-146	0. 9306	-149. 7987
NH4CH3COO(aq)	1. 090e-151	8. 398e-147	1. 0000	-150. 9626
Acetami de(aq)	1. 017e-151	6. 007e-147	1. 0000	-150. 9926
UO2SCN+	7. 487e-164	2. 456e-158	0. 9312	-163. 1566
Cd(NH3)4++	4. 911e-168	8. 863e-163	0. 7533	-167. 4319
Ethyne(aq)	2. 511e-170	6. 536e-166	1. 0000	-169. 6002
Ethanol (aq)	2. 532e-172	1. 166e-167	1. 0000	-171. 5965
Ethyl ene(aq)	7. 077e-177	1. 985e-172	1. 0000	-176. 1502
Lactate	2. 276e-178	2. 027e-173	0. 9312	-177. 6737
Ca(Lac)+	2. 481e-180	3. 203e-175	0. 9312	-179. 6363
Mg(Lac)+	7. 986e-181	9. 052e-176	0. 9312	-180. 1286
Na(Lac) (aq)	2. 670e-181	2. 991e-176	1. 0000	-180. 5735
K(Lac) (aq)	1. 400e-181	1. 794e-176	1. 0000	-180. 8538
Mn(Lac)+	1. 787e-182	2. 573e-177	0. 9312	-181. 7787
Lacti c_aci d(aq)	1. 558e-182	1. 403e-177	1. 0000	-181. 8074
Cu(Lac)+	1. 431e-182	2. 183e-177	0. 9312	-181. 8753
Ba(Lac)+	6. 055e-184	1. 370e-178	0. 9312	-183. 2488
Cd(Lac)+	7. 946e-185	1. 601e-179	0. 9312	-184. 1308
Ethane(aq)	3. 432e-190	1. 032e-185	1. 0000	-189. 4644
Propanoate	1. 643e-193	1. 200e-188	0. 9312	-192. 8152

	SpecE8_output_GSS_Decant	Pond	Jul y. txt	
Ca(Prop)+	3. 133e-196	3. 544e-191	0. 9312	-195. 5350
Na(Prop) (aq)	1. 900e-196	1. 825e-191	1. 0000	-195. 7213
Mg(Prop)+	1. 351e-196	1. 315e-191	0. 9312	-195. 9003
Propanoi c_aci d(a	1. 187e-196	8. 793e-192	1. 0000	-195. 9254
K(Prop) (aq)	9. 962e-197	1. 117e-191	1. 0000	-196. 0016
Ca(Gl yc) 2(aq)	6. 460e-197	1. 228e-191	1. 0000	-196. 1898
Mn(Prop)+	8. 953e-198	1. 146e-192	0. 9312	-197. 0790
Mg(Gl yc) 2(aq)	6. 364e-198	1. 109e-192	1. 0000	-197. 1963
Cu(Prop)+	4. 609e-198	6. 295e-193	0. 9312	-197. 3673
Cu(Gl yc) 2(aq)	1. 675e-198	3. 577e-193	1. 0000	-197. 7760
USCN+++	2. 559e-199	7. 575e-194	0. 5317	-198. 8663
Na(Gl yc) 2-	2. 254e-199	3. 900e-194	0. 9312	-198. 6779
Mn(Gl yc) 2(aq)	1. 577e-199	3. 232e-194	1. 0000	-198. 8022
Ba(Prop)+	1. 336e-199	2. 811e-194	0. 9312	-198. 9050
K(Gl yc) 2-	1. 094e-199	2. 069e-194	0. 9312	-198. 9920
Cd(Prop)+	8. 809e-200	1. 633e-194	0. 9312	-199. 0860
Succi nate	1. 119e-200	1. 299e-195	0. 7511	-200. 0753
S406--	8. 713e-201	1. 953e-195	0. 7511	-200. 1841
Ba(Gl yc) 2(aq)	5. 866e-201	1. 686e-195	1. 0000	-200. 2316
Cd(Gl yc) 2(aq)	2. 781e-201	7. 298e-196	1. 0000	-200. 5558
Seri ne(aq)	2. 014e-202	2. 115e-197	1. 0000	-201. 6960
H-Succi nate	3. 882e-203	4. 543e-198	0. 9312	-202. 4419
Succi ni c_aci d(aq	5. 978e-207	7. 057e-202	1. 0000	-206. 2235
Ethanami ne(aq)	1. 887e-213	8. 506e-209	1. 0000	-212. 7242
Al ani ne(aq)	9. 728e-216	8. 665e-211	1. 0000	-215. 0120
Cu(Al a)+	2. 068e-216	3. 135e-211	0. 9312	-215. 7154
Mg(Al a)+	4. 701e-220	5. 282e-215	0. 9312	-219. 3588
Ca(Al a)+	2. 225e-220	2. 851e-215	0. 9312	-219. 6835
Mn(Al a)+	1. 673e-220	2. 392e-215	0. 9312	-219. 8074
Cd(Al a)+	2. 591e-221	5. 194e-216	0. 9312	-220. 6174
Ba(Al a)+	1. 739e-223	3. 920e-218	0. 9312	-222. 7905
Ca(CH3C00) 2(aq)	1. 002e-223	1. 584e-218	1. 0000	-222. 9993
Mg(CH3C00) 2(aq)	3. 348e-224	4. 766e-219	1. 0000	-223. 4752
Cu(CH3C00) 2(aq)	1. 389e-225	2. 522e-220	1. 0000	-224. 8574
Na(CH3C00) 2-	7. 289e-226	1. 028e-220	0. 9312	-225. 1683
Mn(CH3C00) 2(aq)	6. 072e-226	1. 050e-220	1. 0000	-225. 2167
Acetone(aq)	4. 530e-226	2. 630e-221	1. 0000	-225. 3439
K(CH3C00) 2-	1. 959e-226	3. 079e-221	0. 9312	-225. 7388
Asparti c_aci d(aq	1. 336e-226	1. 777e-221	1. 0000	-225. 8743
Ba(CH3C00) 2(aq)	3. 227e-227	8. 239e-222	1. 0000	-226. 4913
Cd(CH3C00) 2(aq)	1. 842e-227	4. 245e-222	1. 0000	-226. 7347
Propanal (aq)	2. 405e-230	1. 396e-225	1. 0000	-229. 6189
Cu(CH3C00) 2-	1. 430e-233	2. 596e-228	0. 9312	-232. 8757
1-Propyne(aq)	3. 629e-248	1. 453e-243	1. 0000	-247. 4402
1-Propanol (aq)	3. 379e-254	2. 030e-249	1. 0000	-253. 4712
S3--	1. 084e-254	1. 042e-249	0. 7511	-254. 0894
1-Propene(aq)	1. 476e-256	6. 211e-252	1. 0000	-255. 8308
2-Hydroxybutanoa	1. 153e-260	1. 188e-255	0. 9312	-259. 9692
NH4(CH3C00) 2-	1. 226e-262	1. 669e-257	0. 9312	-261. 9423
Asparagi ne(aq)	2. 301e-263	3. 039e-258	1. 0000	-262. 6381
Cu(Gl y) 2(aq)	5. 266e-264	1. 114e-258	1. 0000	-263. 2785
2-Hydroxybutanoi	7. 017e-265	7. 303e-260	1. 0000	-264. 1538
Mg(Gl y) 2(aq)	1. 725e-270	2. 973e-265	1. 0000	-269. 7633
Di gl yci ne(aq)	5. 237e-272	6. 916e-267	1. 0000	-271. 2810
Mn(Gl y) 2(aq)	4. 483e-272	9. 101e-267	1. 0000	-271. 3484
Propane(aq)	1. 927e-272	8. 497e-268	1. 0000	-271. 7150
Cd(Gl y) 2(aq)	6. 626e-273	1. 726e-267	1. 0000	-272. 1787
Di ketopi perazi ne	3. 291e-274	3. 754e-269	1. 0000	-273. 4826
Ca(Gl y) 2(aq)	1. 598e-274	3. 007e-269	1. 0000	-273. 7963
Butanoate	6. 496e-276	5. 656e-271	0. 9312	-275. 2183
Ba(Gl y) 2(aq)	4. 091e-277	1. 167e-271	1. 0000	-276. 3881
Ca(But)+	8. 547e-279	1. 087e-273	0. 9312	-278. 0991
Na(But) (aq)	7. 187e-279	7. 910e-274	1. 0000	-278. 1434

	SpecE8_output_GSS_Decant	Pond	Jul y. txt	
Butanoi c_aci d(aq)	3. 867e-279	3. 406e-274	1. 0000	-278. 4126
K(But) (aq)	3. 769e-279	4. 755e-274	1. 0000	-278. 4238
Mg(But)+	3. 519e-279	3. 919e-274	0. 9312	-278. 4845
Mn(But)+	3. 015e-280	4. 281e-275	0. 9312	-279. 5517
Cu(But)+	1. 444e-280	2. 175e-275	0. 9312	-279. 8713
Ba(But)+	3. 558e-282	7. 983e-277	0. 9312	-281. 4797
Gl utarate	3. 533e-282	4. 596e-277	0. 7511	-281. 5761
Cd(But)+	1. 508e-282	3. 009e-277	0. 9312	-281. 8524
Threoni ne(aq)	8. 315e-285	9. 902e-280	1. 0000	-284. 0801
H-Gl utarate	7. 403e-285	9. 703e-280	0. 9312	-284. 1615
Ethyl acetate(aq)	5. 529e-288	4. 870e-283	1. 0000	-287. 2574
Gl utari c_aci d(aq)	1. 518e-288	2. 006e-283	1. 0000	-287. 8186
S506--	8. 593e-291	2. 202e-285	0. 7511	-290. 1902
1-Propanami ne(aq)	8. 385e-295	4. 955e-290	1. 0000	-294. 0765
a-Ami nobutyri c_a	8. 241e-298	8. 495e-293	1. 0000	-297. 0840
Gl utami c_aci d(aq)	5. 243e-307	7. 712e-302	1. 0000	-300. 0000
U02(SCN)2(aq)	7. 148e-314	2. 760e-308	1. 0000	-300. 0000
Butanal (aq)	3. 779e-314	2. 724e-309	1. 0000	-300. 0000
Cd(CH3COO)3-	0. 0000	0. 0000	0. 9312	-300. 0000
Azel ai c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Cd(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
K(But)2-	0. 0000	0. 0000	0. 9312	-300. 0000
I sol euci ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Hexanoi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Cd(Al a)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Hexanoate	0. 0000	0. 0000	0. 9312	-300. 0000
Hexanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Heptanoi c_aci d(a	0. 0000	0. 0000	1. 0000	-300. 0000
Heptanoate	0. 0000	0. 0000	0. 9312	-300. 0000
Heptanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Hexanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexyne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
HSb2S4-	0. 0000	0. 0000	0. 9312	-300. 0000
p-Tol uate	0. 0000	0. 0000	0. 9312	-300. 0000
o-Tol ui c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
o-Tol uate	0. 0000	0. 0000	0. 9312	-300. 0000
2-Heptanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Propyl benzene(0. 0000	0. 0000	1. 0000	-300. 0000
n-Pentyl benzene(0. 0000	0. 0000	1. 0000	-300. 0000
n-Pentane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Octane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Hexyl benzene(a	0. 0000	0. 0000	1. 0000	-300. 0000
n-Hexane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Heptane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Butane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
m-Tol ui c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
m-Tol uate	0. 0000	0. 0000	0. 9312	-300. 0000
Val i ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Butanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Butyne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Undecanoi c_aci d(0. 0000	0. 0000	1. 0000	-300. 0000
Undecanoate	0. 0000	0. 0000	0. 9312	-300. 0000
Ca(Pent)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Pent)+	0. 0000	0. 0000	0. 9312	-300. 0000
Al anyl gl yci ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Lac)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
U02(SCN)3-	0. 0000	0. 0000	0. 9312	-300. 0000
Adi pi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
H-Suberate	0. 0000	0. 0000	0. 9312	-300. 0000
H-Sebacate	0. 0000	0. 0000	0. 9312	-300. 0000
H-Pi mel ate	0. 0000	0. 0000	0. 9312	-300. 0000

	SpecE8_output_GSS_Decant	Pond	July.txt	
Adipate	0.0000	0.0000	0.7511	-300.0000
1-Hexanol (aq)	0.0000	0.0000	1.0000	-300.0000
U(SCN)2++	0.0000	0.0000	0.7533	-300.0000
Tyrosine(aq)	0.0000	0.0000	1.0000	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
Toluene(aq)	0.0000	0.0000	1.0000	-300.0000
H-Azelaate	0.0000	0.0000	0.9312	-300.0000
H-Adipate	0.0000	0.0000	0.9312	-300.0000
Suberate	0.0000	0.0000	0.7511	-300.0000
Sebacic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacate	0.0000	0.0000	0.7511	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Sb2S4--	0.0000	0.0000	0.7511	-300.0000
Glutamine(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
S5--	0.0000	0.0000	0.7511	-300.0000
S4--	0.0000	0.0000	0.7511	-300.0000
1-Hexanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butene(aq)	0.0000	0.0000	1.0000	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
Pimelic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Pimelate	0.0000	0.0000	0.7511	-300.0000
Phenylalanine(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol (aq)	0.0000	0.0000	1.0000	-300.0000
Pentanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Pentanoate	0.0000	0.0000	0.9312	-300.0000
Pentanal (aq)	0.0000	0.0000	1.0000	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoate	0.0000	0.0000	0.9312	-300.0000
Octanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9312	-300.0000
Octanal (aq)	0.0000	0.0000	1.0000	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoate	0.0000	0.0000	0.9312	-300.0000
Nonanal (aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9312	-300.0000
Benzoinic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Decanal (aq)	0.0000	0.0000	1.0000	-300.0000
Benzoate	0.0000	0.0000	0.9338	-300.0000
2-Hydroxypentano	0.0000	0.0000	1.0000	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Prop)2-	0.0000	0.0000	0.9312	-300.0000
Na(Pent)2-	0.0000	0.0000	0.9312	-300.0000
Na(Pent) (aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxypentano	0.0000	0.0000	0.9312	-300.0000
Na(Lac)2-	0.0000	0.0000	0.9312	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)+	0.0000	0.0000	0.9312	-300.0000
Na(But)2-	0.0000	0.0000	0.9312	-300.0000
Suberic acid(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoic	0.0000	0.0000	1.0000	-300.0000
1-Pentanol (aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoic	0.0000	0.0000	0.9312	-300.0000
Ba(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoic	0.0000	0.0000	1.0000	-300.0000

	SpecE8_output_GSS_Decant	Pond	Jul y. txt	
1-Pentanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butanol(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)+	0.0000	0.0000	0.9312	-300.0000
Cu(CH3COO)3-	0.0000	0.0000	0.9312	-300.0000
Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)+	0.0000	0.0000	0.9312	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9312	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(CH3COO)3-	0.0000	0.0000	0.9312	-300.0000
Cu(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoi	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoa	0.0000	0.0000	0.9312	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)+	0.0000	0.0000	0.9312	-300.0000
Mg(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	0.9312	-300.0000
Mg(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Methionine(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Pent)+	0.0000	0.0000	0.9312	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoi	0.0000	0.0000	1.0000	-300.0000
1-Octanol(aq)	0.0000	0.0000	1.0000	-300.0000
Leucylglycine(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Leucine(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butylbenzene(a	0.0000	0.0000	1.0000	-300.0000
Ba(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptylbenzene(0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecano	0.0000	0.0000	0.9312	-300.0000
K(Prop)2-	0.0000	0.0000	0.9312	-300.0000
K(Pent)2-	0.0000	0.0000	0.9312	-300.0000
n-Octylbenzene(a	0.0000	0.0000	1.0000	-300.0000
K(Pent)(aq)	0.0000	0.0000	1.0000	-300.0000
K(Lac)2-	0.0000	0.0000	0.9312	-300.0000
Azelate	0.0000	0.0000	0.7511	-300.0000
o-Phthalate	0.0000	0.0000	0.7511	-300.0000
1-Octanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butanamine(aq)	0.0000	0.0000	1.0000	-300.0000
p-Toluidine(aq)	0.0000	0.0000	1.0000	-300.0000
(NH4)2Sb2S4(aq)	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states

log Q/K

log Q/K

Fluorapatite	6.0547s/sat	U02S04: 3H2O	-16.9565
Witherite	1.9760s/sat	CdF2	-17.1899
Malachite	1.3998s/sat	U02F2: 3H2O	-17.4547

SpecE8_output_GSS_Decant Pond July.txt			
Brochanti te	1. 3884s/sat	Larni te	-17. 4812
Tenori te	1. 3681s/sat	MnCl 2: 2H2O	-17. 6270
Bi xbyi te	0. 5891s/sat	UO2F2	-17. 6467
Tal c	0. 4616s/sat	Autuni te-H	-18. 0475
Quartz	0. 4527s/sat	CuCl 2	-18. 6326
Dol omi te	0. 4114s/sat	Na2UO4(al pha)	-19. 1563
Dol omi te-ord	0. 4114s/sat	MnCl 2: H2O	-19. 2114
Bari te	0. 3110s/sat	MgCl 2: 4H2O	-19. 2714
Tri dymi te	0. 2556s/sat	UO2S04	-20. 4433
Chal cedony	0. 1776s/sat	Li me	-20. 4701
Cri stobal i te(al p	-0. 1073	UO2. 3333(beta)	-21. 2558
I ce	-0. 1228	Ranki ni te	-22. 1481
Cal ci te	-0. 1612	Scacchi te	-22. 4912
Rhodochrosi te	-0. 2670	UO2(NO3)2: 6H2O	-23. 1265
Hausmanni te	-0. 2981	Tobermori te-14A	-23. 2499
Aragoni te	-0. 3057	Hydrophi l i te	-23. 3662
Coesi te	-0. 3682	K2UO4	-23. 4915
Cri stobal i te(bet	-0. 5628	Ni ngyoi te	-23. 6153
Di optase	-0. 6041	UOF0H: . 5H2O	-23. 9159
Si O2(am)	-0. 8683	UO2Cl	-24. 1663
Monohydrocal ci te	-0. 9842	UOF0H	-24. 3863
Pyrol usi te	-1. 0428	UO2(NO3)2: 3H2O	-24. 5959
Magnesi te	-1. 0819	MgCl 2: 2H2O	-24. 8075
Dol omi te-di s	-1. 1638	Tobermori te-11A	-25. 1512
Atacami te	-1. 8371	Anti gori te	-25. 2114
Gypsum	-2. 4365	UO2(NO3)2: 2H2O	-25. 8634
CaU04	-2. 4404	Merwi ni te	-26. 9112
Chrysoti l e	-2. 4423	Foshagi te	-27. 3448
Anhydri te	-2. 6524	UOF2: H2O	-27. 5295
Enstati te	-3. 0003	UOF2	-28. 1118
Rhodoni te	-3. 0880	MgCl 2: H2O	-28. 2176
Hydroxyl apati te	-3. 1069	UO2(P03)2	-28. 4008
Bassani te	-3. 2986	(U02)3(P04)2	-28. 4084
Fl uori te	-3. 3649	UO2Cl 2: 3H2O	-28. 4579
CaS04: 0. 5H2O(bet	-3. 4720	Tobermori te-9A	-28. 7244
CdSi O3	-3. 4787	(U02)2P207	-28. 8730
Sb203	-3. 4801	Ba2Si O4	-29. 2578
Azuri te	-3. 5906	UO2(NO3)2: H2O	-29. 5024
Schoepi te	-3. 7199	Afwi l l i te	-30. 3236
U03: 2H2O	-3. 7226	UO2Cl 2: H2O	-31. 2100
Di opsi de	-3. 8032	U(HP04)2: 4H2O	-32. 9886
U02(OH)2(beta)	-3. 8496	UP207	-32. 9933
U03: . 9H2O(al pha)	-3. 9137	UO2(NO3)2	-33. 0179
Schoepi te-dehy(.	-3. 9178	KMgCl 3: 2H2O	-33. 1710
Sepi ol i te	-3. 9650	Chl oromagnesi te	-34. 0577
Schoepi te-dehy(.	-3. 9954	UO2Cl 2	-35. 1379
Schoepi te-dehy(1	-4. 0043	As205	-35. 3509
Whi tl ocki te	-4. 0640	UO2S03	-36. 6375
Nesquehoni te	-4. 0952	Cd	-36. 7582
Bruci te	-4. 4148	(U02)2As207	-38. 6912
Sel l ai te	-4. 5085	BaO	-38. 7016
Wol l astoni te	-4. 9897	UF4: 2. 5H2O	-38. 7406
Cupri te	-4. 9946	UO2(As03)2	-39. 0729
Hunti te	-5. 0083	(U02)3(As04)2	-39. 1677
Okeni te	-5. 1017	Xonotl i te	-39. 3890
Mn(OH)2(am)	-5. 1163	BaSi F6	-39. 5524
Schoepi te-dehy(.	-5. 1268	Hatruri te	-40. 3493
Pseudowol l astoni	-5. 2362	KMgCl 3	-40. 5932
Tremol i te	-5. 2897	Ba2U207	-42. 7918
U02C03	-5. 5224	UF4	-43. 0148
Rutherfordi ne	-5. 5432	U(S04)2: 8H2O	-46. 6820
Schoepi te-dehy(.	-5. 6540	U(S04)2: 4H2O	-47. 7949
Nahcol i te	-5. 6583	U(S04)2	-47. 8773

SpecE8_output_GSS_Decant Pond July.txt			
Ni ter	-5.9668	Arsenol i te	-49.2152
Cd(OH)2	-6.1448	Cl audeti te	-49.2792
Tephroi te	-6.2428	U(CO3)2	-49.7485
Arti ni te	-6.4989	UOCl 2	-49.8935
UO3(gamma)	-6.6713	C	-50.1332
Sanborni te	-7.0316	UCI F3	-53.8119
UO3(beta)	-7.2827	Na3UO4	-54.5471
BaUO4	-7.5191	Na4Si O4	-54.9700
Montepo ni te	-7.5433	UOF4	-55.7630
UO3(al pha)	-7.6191	U5O12Cl	-56.4523
Forsteri te	-7.6808	Na	-56.8340
Chal canthi te	-7.6936	Na2O	-58.2460
Manganosi te	-7.7941	As	-58.6301
Cd(OH)Cl	-7.7945	(UO2)2Cl 3	-60.4397
Syl vi te	-7.9988	K	-60.7147
Arcani te	-8.0764	Mn	-60.8400
Mi rabi l i te	-8.1227	UOCl 3	-61.4804
Hal i te	-8.5256	UF5(beta)	-64.6562
UO2HP04: 4H2O	-8.5947	UF3	-64.7234
Nantoki te	-8.6021	UF5(al pha)	-65.0101
Gyrol i te	-8.8246	Ba3UO6	-65.1802
Monti cel l i te	-8.9497	UCI 2F2	-66.9242
Thenardi te	-9.1377	UOCl	-67.7537
UO2FOH: 2H2O	-9.1696	U2O3F6	-73.5592
Peri cl ase	-9.5485	S	-75.4469
UO2FOH: H2O	-9.5695	K2O	-75.5429
Ni trobari te	-9.7579	Covel l i te	-76.0017
Mg1. 25S04(OH)0. 5	-9.8048	UCI 3F	-79.8588
Cu	-9.9309	Na6Si 2O7	-79.8729
CdS04: 2. 667H2O	-9.9510	Chal coci te	-79.9940
UO2FOH	-10.0350	CdS	-84.3946
MgUO4	-10.0995	CdCl 2(NH3)2	-90.5493
CdS04: H2O	-10.1305	UCI 4	-90.6249
BaU2O7	-10.1786	U(SO3)2	-91.0717
Natron	-10.2680	UCI 3	-94.3816
Portl andi te	-10.2868	UF6	-94.9222
Na2U2O7	-10.4832	Al abandi te	-97.5044
Na2CO3: 7H2O	-10.6626	As4O6(mono)	-98.0748
Mg1. 5S04(OH)	-10.8309	U3O5F8	-98.2947
UO2. 25	-11.6565	As4O6(cubi)	-98.3116
UO2. 25(beta)	-11.7195	Mg	-99.3120
CdS04	-11.7269	U2F9	-104.4943
Thermonatri te	-11.8093	UCI 5	-110.4951
MnS04	-11.8901	U2O2Cl 5	-112.5860
Anthophyl l i te	-11.8998	UN1. 73(al pha)	-113.6605
Sal eei te	-11.9581	BaS	-114.9882
Na2CO3	-12.0582	UN1. 59(al pha)	-116.3777
UO2S04: H2O	-12.3420	Ca	-116.3870
MgS04	-12.4471	Ba	-120.5856
UO2Cl OH: 2H2O	-13.1328	P	-124.7877
Chal cocyani te	-13.4296	Si	-128.8429
Ca3(AsO4)2	-13.5119	UCI 6	-129.1989
UO2. 6667	-13.7527	UN	-129.2412
Uranoci rci te	-13.8988	CdCl 2(NH3)4	-175.8183
Torberni te	-14.0373	U	-175.9735
Hydromagnesi te	-14.1553	U4F17	-189.6476
BaCl 2: 2H2O	-14.3883	UAs	-192.3358
CdCl 2: H2O	-14.5041	US	-194.3875
Ba(OH)2: 8H2O	-14.8546	UC	-208.5342
Ba2Si 3O8	-14.8950	UH3(beta)	-209.0686
BaCl 2: H2O	-15.0365	US1. 9	-228.7876
(UO2)3(P04)2: 4H2	-15.1445	US2	-234.5942
Natrosil i te	-15.4595	UAs2	-248.3141

SpecE8_output_GSS_Decant Pond July.txt

Urani ni te	-15. 4773	UP	-253. 4965
CdCl 2	-15. 5579	UC1. 94(al pha)	-257. 7055
Hi l l ebrandi te	-15. 7749	CdCl 2(NH3)6	-261. 1641
Coffi ni te	-15. 8395	Sti bni te	-283. 4737
UP05	-15. 8768	US3	-306. 8374
Akermani te	-15. 9534	Orpi ment	-313. 6382
Na2Si O3	-16. 0157	UP2	-373. 1818
MgOHCl	-16. 0301	U2S3	-421. 9494
NaUO3	-16. 0940	o-Phthal i c_aci d	-433. 8985
Di cal ci um_si l i ca	-16. 1636	U2C3	-468. 7190
MnCl 2: 4H2O	-16. 3198	U3As4	-633. 5510
BaCl 2	-16. 4863	U3S5	-651. 9311
UO2S04: 3. 5H2O	-16. 8613	U3P4	-880. 1697
UO2S04: 2. 5H2O	-16. 8715		

Gases	fugaci ty	l og fug.
N2(g)	0. 1662	-0. 779
H2O(g)	0. 02011	-1. 697
CO2(g)	0. 0008846	-3. 053
NO2(g)	7. 243e-014	-13. 140
HF(g)	5. 240e-015	-14. 281
HCl (g)	3. 365e-019	-18. 473
O2(g)	1. 009e-023	-22. 996
NO(g)	5. 500e-028	-27. 260
H2(g)	1. 689e-031	-30. 772
CO(g)	4. 755e-038	-37. 323
Cl 2(g)	9. 527e-041	-40. 021
Si F4(g)	2. 527e-041	-40. 597
NH3(g)	2. 813e-044	-43. 551
SO2(g)	7. 821e-046	-45. 107
Cd(g)	2. 779e-051	-50. 556
UO2F2(g)	5. 011e-061	-60. 300
Cu(g)	1. 229e-063	-62. 910
Na(g)	2. 693e-071	-70. 570
UO3(g)	1. 052e-071	-70. 978
K(g)	3. 030e-072	-71. 519
UO2Cl 2(g)	2. 269e-072	-71. 644
UOF4(g)	1. 316e-076	-75. 881
UF5(g)	2. 274e-085	-84. 643
UF4(g)	2. 519e-088	-87. 599
F2(g)	2. 346e-096	-95. 630
UF6(g)	1. 348e-096	-95. 870
H2S(g)	5. 150e-101	-100. 288
CH4(g)	2. 483e-103	-102. 605
UO2(g)	3. 223e-114	-113. 492
UCl 4(g)	2. 814e-116	-115. 551
Mg(g)	4. 379e-120	-119. 359
UCl 5(g)	7. 856e-129	-128. 105
UF3(g)	3. 223e-133	-132. 492
UCl 6(g)	5. 153e-136	-135. 288
UCl 3(g)	1. 483e-141	-140. 829
Ca(g)	9. 505e-143	-142. 022
U2F10(g)	1. 852e-144	-143. 732
S2(g)	7. 788e-166	-165. 109
C(g)	3. 193e-170	-169. 496
C2H4(g)	1. 329e-174	-173. 876
UF2(g)	6. 948e-175	-174. 158
UCl 2(g)	2. 472e-184	-183. 607
UO(g)	1. 330e-188	-187. 876
Si (g)	1. 026e-201	-200. 989
UF(g)	4. 812e-210	-209. 318
U2Cl 8(g)	3. 761e-221	-220. 425

SpecE8_output_GSS_Decant Pond July.txt

UCI (g)	1.030e-224	-223.987
U2Cl10(g)	3.978e-230	-229.400
U(g)	1.333e-263	-262.875

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
Ba++	7.96e-007	7.96e-007	0.109			
Ca++	0.000541	0.000541	21.7			
Cd++	8.50e-009	8.50e-009	0.000955			
Cl -	0.000111	0.000111	3.94			
Cu++	2.01e-007	2.01e-007	0.0128			
F-	1.04e-005	1.04e-005	0.197			
H+	1.76e-005	1.76e-005	0.0177			
H2AsO4-	6.57e-008	6.57e-008	0.00926			
H2O	55.5	55.5	1.00e+006			
HCO3-	0.00160	0.00160	97.5			
HP04--	1.91e-007	1.91e-007	0.0183			
K+	0.000655	0.000655	25.6			
Mg++	0.000203	0.000203	4.92			
Mn++	3.75e-006	3.75e-006	0.206			
NH3(aq)	0.000234	0.000234	3.98			
NO3-	0.000956	0.000956	59.3			
Na+	0.00116	0.00116	26.6			
O2(aq)	0.000175	0.000175	5.61			
SO4--	0.000431	0.000431	41.3			
Sb(OH)3(aq)	5.01e-007	5.01e-007	0.0866			
Si O2(aq)	0.000221	0.000221	13.3			
UO2++	1.20e-008	1.20e-008	0.00324			

Elemental composition	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg
Antimony	5.015e-007	5.015e-007	0.06104		
Arsenic	6.572e-008	6.572e-008	0.004922		
Barium	7.960e-007	7.960e-007	0.1093		
Cadmium	8.497e-009	8.497e-009	0.0009549		
Calcium	0.0005406	0.0005406	21.66		
Carbon	0.001598	0.001598	19.18		
Chlorine	0.0001111	0.0001111	3.938		
Copper	2.015e-007	2.015e-007	0.01280		
Fluorine	1.037e-005	1.037e-005	0.1969		
Hydrogen	111.0	111.0	1.119e+005		
Magnesium	0.0002026	0.0002026	4.922		
Manganese	3.746e-006	3.746e-006	0.2057		
Nitrogen	0.001190	0.001190	16.66		
Oxygen	55.52	55.52	8.880e+005		
Phosphorus	1.908e-007	1.908e-007	0.005907		
Potassium	0.0006548	0.0006548	25.60		
Silicon	0.0002213	0.0002213	6.212		
Sodium	0.001157	0.001157	26.58		
Sulfur	0.0004305	0.0004305	13.80		
Uranium	1.200e-008	1.200e-008	0.002855		

SpecE8_output_GSS_Decant Pond June.txt

Temperature = 15.1 C Pressure = 1.013 bars
 pH = 8.000 log fO2 = -30.606
 Eh = 0.3360 volts pe = 5.8750
 Ionic strength = 0.004609
 Charge imbalance = -0.000231 eq/kg (-6.639% error)
 Activity of water = 0.999996
 Solvent mass = 1.000000 kg
 Solution mass = 1.000317 kg
 Solution density = 1.020 g/cm3
 Chlorinity = 0.000111 molal
 Dissolved solids = 316 mg/kg sol'n
 Hardness = 71.62 mg/kg sol'n as CaCO3
 carbonate = 71.62 mg/kg sol'n as CaCO3
 non-carbonate = 0.00 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 79.64 mg/kg sol'n as CaCO3
 Water type = Na-HCO3

Nernst redox couples	Eh (volts)	pe
e- + .25*O2(aq) + H+ = .5*H2O	0.3360	5.8750
8*e- + 9*H+ + NO3- = 3*H2O + NH3(aq)	0.5400	9.4422

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HCO3-	0.001562	95.26	0.9308	-2.8376
Na+	0.001277	29.35	0.9308	-2.9250
NO3-	0.001015	62.94	0.9294	-3.0252
K+	0.0007008	27.39	0.9294	-3.1862
Ca++	0.0004890	19.59	0.7581	-3.4309
SO4--	0.0004331	41.60	0.7495	-3.4886
SiO2(aq)	0.0002135	12.82	1.0000	-3.6706
Mg++	0.0001891	4.595	0.7661	-3.8390
N2(aq)	0.0001632	4.570	1.0000	-3.7873
Cl-	0.0001106	3.921	0.9294	-3.9879
CO2(aq)	3.861e-005	1.699	1.0000	-4.4133
CaSO4(aq)	1.482e-005	2.017	1.0000	-4.8291
F-	1.027e-005	0.1951	0.9301	-5.0198
MgSO4(aq)	9.432e-006	1.135	1.0000	-5.0254
Mn++	7.785e-006	0.4275	0.7581	-5.2290
CO3--	7.115e-006	0.4269	0.7517	-5.2717
CaHCO3+	6.612e-006	0.6682	0.9308	-5.2109
AlO2-	6.258e-006	0.3690	0.9308	-5.2347
CaCO3(aq)	3.470e-006	0.3472	1.0000	-5.4597
NaHCO3(aq)	2.996e-006	0.2516	1.0000	-5.5235
MgHCO3+	2.514e-006	0.2144	0.9308	-5.6308
HP04--	2.303e-006	0.2210	0.7495	-5.7629
HSiO3-	1.888e-006	0.1455	0.9308	-5.7553
KS04-	1.712e-006	0.2313	0.9308	-5.7977
Fe++	8.754e-007	0.04887	0.7581	-6.1781
Ba++	8.345e-007	0.1146	0.7539	-6.2013
MgCO3(aq)	6.715e-007	0.05660	1.0000	-6.1730
OH-	4.830e-007	0.008213	0.9301	-6.3475
Sb(OH)3(aq)	4.028e-007	0.06957	1.0000	-6.3949
MnSO4(aq)	3.918e-007	0.05914	1.0000	-6.4070
Cu++	3.703e-007	0.02352	0.7581	-6.5517
H2PO4-	3.255e-007	0.03156	0.9308	-6.5186
HAIO2(aq)	2.808e-007	0.01684	1.0000	-6.5516
NaHSiO3(aq)	1.160e-007	0.01161	1.0000	-6.9354
MgF+	3.398e-008	0.001471	0.9308	-7.4999

	SpecE8_output_GSS_Decant	Pond	June.txt	
NaCO3-	2.955e-008	0.002452	0.9308	-7.5605
HAsO4--	2.494e-008	0.003488	0.7495	-7.7283
NaCl (aq)	1.937e-008	0.001132	1.0000	-7.7129
Pb++	1.795e-008	0.003718	0.7517	-7.8699
CaF+	1.758e-008	0.001038	0.9308	-7.7863
MgCl +	1.242e-008	0.0007418	0.9308	-7.9371
Cd++	1.099e-008	0.001235	0.7539	-8.0816
H+	1.067e-008	1.075e-005	0.9370	-8.0000
CaCl +	8.286e-009	0.0006256	0.9308	-8.1128
UO2(CO3)3----	7.593e-009	0.003416	0.3152	-8.6210
UO2(CO3)2--	5.949e-009	0.002320	0.7495	-8.3508
Al (OH)2+	3.951e-009	0.0002409	0.9308	-8.4344
FeCO3+	2.748e-009	0.0003183	0.9308	-8.5921
KCl (aq)	1.766e-009	0.0001316	1.0000	-8.7530
H2AsO4-	1.240e-009	0.0001747	0.9308	-8.9377
BaCO3(aq)	1.208e-009	0.0002383	1.0000	-8.9180
NaAl O2(aq)	1.129e-009	9.248e-005	1.0000	-8.9475
NaF (aq)	1.054e-009	4.425e-005	1.0000	-8.9771
MnCl +	1.051e-009	9.494e-005	0.9308	-9.0097
HSO4-	2.609e-010	2.531e-005	0.9308	-9.6148
Cu+	1.529e-010	9.710e-006	0.9308	-9.8469
PO4---	1.247e-010	1.184e-005	0.5225	-10.1861
HF (aq)	1.185e-010	2.371e-006	1.0000	-9.9261
NaOH(aq)	9.068e-011	3.626e-006	1.0000	-10.0425
UO2CO3(aq)	5.659e-011	1.867e-005	1.0000	-10.2473
FeCl +	5.028e-011	4.589e-006	0.9308	-10.3298
Al OH++	4.175e-011	1.836e-006	0.7517	-10.5033
PbCl +	3.935e-011	9.545e-006	0.9308	-10.4362
Sb(OH)4-	2.427e-011	4.605e-006	0.9308	-10.6460
BaCl +	1.940e-011	3.351e-006	0.9308	-10.7434
AsO4---	6.682e-012	9.280e-007	0.5225	-11.4570
BaF+	3.847e-012	6.011e-007	0.9308	-11.4461
CdSO4(aq)	2.716e-012	5.661e-007	1.0000	-11.5660
MgP2O7--	1.033e-012	2.047e-007	0.7495	-12.1111
CaCl 2(aq)	1.023e-012	1.135e-007	1.0000	-11.9900
UO2OH+	8.362e-013	2.400e-007	0.9308	-12.1088
Sb(OH)2F(aq)	6.436e-013	1.124e-007	1.0000	-12.1914
H3PO4(aq)	4.068e-013	3.985e-008	1.0000	-12.3907
HN03(aq)	3.793e-013	2.390e-008	1.0000	-12.4210
HCl (aq)	2.150e-013	7.837e-009	1.0000	-12.6676
Sb(OH)2+	1.958e-013	3.048e-008	0.9308	-12.7394
CdCl 2(aq)	1.794e-013	3.287e-008	1.0000	-12.7463
Al +++	1.009e-013	2.722e-009	0.5549	-13.2519
Fe+++	6.235e-014	3.481e-009	0.5549	-13.4610
PbCl 2(aq)	1.335e-014	3.712e-009	1.0000	-13.8745
HP2O7---	1.211e-014	2.118e-009	0.5225	-14.1988
KHSO4(aq)	9.668e-015	1.316e-009	1.0000	-14.0147
FeF++	5.158e-015	3.859e-010	0.7517	-14.4115
UO2++	3.086e-015	8.330e-010	0.7517	-14.6345
UO2F+	2.646e-015	7.646e-010	0.9308	-14.6086
H3AsO4(aq)	1.842e-015	2.614e-010	1.0000	-14.7346
P2O7----	1.174e-015	2.042e-010	0.3152	-15.4315
FeSO4+	7.451e-016	1.131e-010	0.9308	-15.1590
UO2S04(aq)	6.905e-016	2.527e-010	1.0000	-15.1609
FeF2+	5.762e-016	5.406e-011	0.9308	-15.2706
H2P2O7--	3.759e-016	6.613e-011	0.7495	-15.5501
HF2-	2.637e-016	1.028e-011	0.9308	-15.6100
KP2O7---	8.970e-017	1.910e-011	0.5225	-16.3291
UO2F2(aq)	7.348e-017	2.263e-011	1.0000	-16.1339
FeCl 2(aq)	2.388e-017	3.026e-012	1.0000	-16.6219
CdCl 3-	4.041e-018	8.838e-013	0.9308	-17.4247
UO2(SO4)2--	2.003e-018	9.253e-013	0.7495	-17.8236
PbCl 3-	7.228e-019	2.266e-013	0.9308	-18.1721

	SpecE8_output_GSS_Decant	Pond	June.txt	
FeCl ++	4. 587e-019	4. 187e-014	0. 7517	-18. 4624
UO2Cl +	3. 332e-019	1. 018e-013	0. 9308	-18. 5085
UO2F3-	1. 304e-019	4. 261e-014	0. 9308	-18. 9160
(UO2)2(OH)2++	9. 619e-020	5. 520e-014	0. 7517	-19. 1408
UO2+	8. 349e-020	2. 254e-014	0. 9308	-19. 1095
(UO2)3(CO3)6(6-)	9. 620e-021	1. 125e-014	0. 0744	-21. 1451
(UO2)3(OH)5+	8. 489e-021	7. 597e-015	0. 9308	-20. 1023
H3P2O7-	6. 759e-022	1. 196e-016	0. 9308	-21. 2013
U(OH)4(aq)	1. 899e-022	5. 811e-017	1. 0000	-21. 7214
PbCl 4--	7. 486e-023	2. 612e-017	0. 7495	-22. 2509
HAsO2(aq)	3. 408e-023	3. 677e-018	1. 0000	-22. 4674
As(OH)3(aq)	2. 977e-023	3. 748e-018	1. 0000	-22. 5262
UO2F4--	9. 237e-024	3. 195e-018	0. 7495	-23. 1597
UO2Cl 2(aq)	1. 510e-024	5. 147e-019	1. 0000	-23. 8210
AsO2-	1. 350e-024	1. 443e-019	0. 9308	-23. 9007
H2AsO3-	1. 345e-024	1. 680e-019	0. 9308	-23. 9025
FeCl 4--	1. 030e-024	2. 036e-019	0. 7495	-24. 1122
Mn+++	1. 026e-025	5. 633e-021	0. 5293	-25. 2653
AsO2OH--	1. 083e-027	1. 342e-022	0. 7495	-27. 0904
H4P2O7(aq)	1. 748e-028	3. 111e-023	1. 0000	-27. 7573
Formate	3. 386e-031	1. 524e-026	0. 9301	-30. 5019
NO2-	2. 318e-031	1. 066e-026	0. 9294	-30. 6668
H2(aq)	8. 716e-032	1. 756e-028	1. 0000	-31. 0597
Ca(For)+	3. 631e-033	3. 089e-028	0. 9308	-32. 4711
Mg(For)+	1. 616e-033	1. 120e-028	0. 9308	-32. 8228
Na(For)(aq)	4. 370e-034	2. 971e-029	1. 0000	-33. 3595
O2(aq)	3. 821e-034	1. 222e-029	1. 0000	-33. 4178
K(For)(aq)	2. 144e-034	1. 803e-029	1. 0000	-33. 6688
Mn(For)+	1. 134e-034	1. 133e-029	0. 9308	-33. 9766
Formic acid(aq)	1. 828e-035	8. 409e-031	1. 0000	-34. 7381
Fe(For)+	1. 821e-035	1. 836e-030	0. 9308	-34. 7708
SO3--	1. 248e-035	9. 993e-031	0. 7517	-35. 0276
Cu(For)+	1. 086e-035	1. 179e-030	0. 9308	-34. 9952
Ba(For)+	4. 930e-036	8. 986e-031	0. 9308	-35. 3384
HNO2(aq)	4. 705e-036	2. 211e-031	1. 0000	-35. 3275
NH4+	2. 772e-036	5. 000e-032	0. 9286	-35. 5893
HSO3-	1. 488e-036	1. 206e-031	0. 9308	-35. 8585
Cl O-	6. 622e-037	3. 406e-032	0. 9308	-36. 2102
Pb(For)+	3. 523e-037	8. 884e-032	0. 9308	-36. 4842
Cd(For)+	2. 221e-037	3. 495e-032	0. 9308	-36. 6847
NH3(aq)	7. 163e-038	1. 220e-033	1. 0000	-37. 1449
CO(aq)	3. 323e-038	9. 306e-034	1. 0000	-37. 4784
HO2-	1. 593e-038	5. 257e-034	0. 9308	-37. 8289
Oxalate	1. 237e-038	1. 089e-033	0. 7495	-38. 0328
Si F6--	1. 314e-039	1. 867e-034	0. 7495	-39. 0065
U(CO3)5(6-)	1. 543e-040	8. 301e-035	0. 0744	-40. 9398
MnO4--	1. 310e-040	1. 558e-035	0. 7495	-40. 0078
UOH+++	5. 651e-042	1. 441e-036	0. 5293	-41. 5242
H-Oxalate	1. 691e-042	1. 505e-037	0. 9308	-41. 8030
SO2(aq)	8. 219e-043	5. 264e-038	1. 0000	-42. 0852
UF3+	6. 022e-043	1. 776e-037	0. 9308	-42. 2514
UF2++	3. 806e-043	1. 050e-037	0. 7517	-42. 5435
UF4(aq)	5. 288e-044	1. 660e-038	1. 0000	-43. 2767
MnO4-	9. 258e-045	1. 101e-039	0. 9301	-44. 0650
UF+++	7. 063e-045	1. 815e-039	0. 5293	-44. 4273
U(SO4)2(aq)	3. 189e-046	1. 371e-040	1. 0000	-45. 4964
USO4++	2. 581e-046	8. 621e-041	0. 7517	-45. 7121
HSO5-	5. 194e-047	5. 871e-042	0. 9308	-46. 3157
U++++	6. 283e-049	1. 495e-043	0. 3262	-48. 6883
Oxalic acid(aq)	2. 873e-049	2. 585e-044	1. 0000	-48. 5417
UCl +++	2. 676e-051	7. 316e-046	0. 5293	-50. 8488
S2O6--	2. 015e-060	3. 226e-055	0. 7495	-59. 8209
Cl O2-	6. 566e-062	4. 428e-057	0. 9308	-61. 2139

	SpecE8_output_GSS_Decant	Pond	June.txt	
Ca(For)2(aq)	8.903e-063	1.158e-057	1.0000	-62.0504
Mg(For)2(aq)	4.586e-063	5.241e-058	1.0000	-62.3386
S208--	6.074e-064	1.167e-058	0.7495	-63.3418
Mn(For)2(aq)	4.861e-064	7.045e-059	1.0000	-63.3133
Fe(For)2(aq)	1.044e-064	1.523e-059	1.0000	-63.9812
Na(For)2-	9.626e-065	1.088e-059	0.9308	-64.0477
Cu(For)2(aq)	8.379e-065	1.286e-059	1.0000	-64.0768
U+++	4.174e-065	9.933e-060	0.5293	-64.6557
K(For)2-	4.131e-065	5.333e-060	0.9308	-64.4151
Ba(For)2(aq)	9.714e-066	2.208e-060	1.0000	-65.0126
Cd(For)2(aq)	8.323e-066	1.685e-060	1.0000	-65.0797
Pb(For)2(aq)	1.865e-066	5.541e-061	1.0000	-65.7294
Formaldehyde(aq)	6.149e-068	1.846e-063	1.0000	-67.2112
N3-	2.171e-070	9.118e-066	0.9308	-69.6946
ClO3-	1.396e-072	1.164e-067	0.9301	-71.8867
HN3(aq)	1.301e-073	5.597e-069	1.0000	-72.8857
Cu(NH3)2++	1.016e-073	9.910e-069	0.7517	-73.1172
S205--	4.029e-077	5.806e-072	0.7495	-76.5200
Cd(NH3)2++	6.249e-078	9.150e-073	0.7517	-77.3281
Urea(aq)	5.440e-079	3.266e-074	1.0000	-78.2644
HCN(aq)	1.283e-080	3.465e-076	1.0000	-79.8919
CN-	4.166e-082	1.084e-077	0.9294	-81.4121
Methanol(aq)	1.225e-083	3.923e-079	1.0000	-82.9120
UO2ClO3+	1.068e-086	3.774e-081	0.9308	-86.0026
ClO4-	1.251e-087	1.244e-082	0.9301	-86.9342
HS-	6.639e-089	2.195e-084	0.9301	-88.2094
Glycolate	9.678e-090	7.260e-085	0.9308	-89.0454
H2S(aq)	8.568e-090	2.919e-085	1.0000	-89.0671
Ca(Glyc)+	1.778e-091	2.046e-086	0.9308	-90.7812
Mg(Glyc)+	3.260e-092	3.238e-087	0.9308	-91.5179
Na(Glyc)(aq)	1.267e-092	1.242e-087	1.0000	-91.8971
K(Glyc)(aq)	6.217e-093	7.094e-088	1.0000	-92.2064
Cu(Glyc)+	2.756e-093	3.819e-088	0.9308	-92.5908
Mn(Glyc)+	2.482e-093	3.226e-088	0.9308	-92.6363
Fe(Glyc)+	1.215e-093	1.590e-088	0.9308	-92.9467
Glycolic acid(aq)	6.317e-094	4.803e-089	1.0000	-93.1995
S--	4.783e-094	1.533e-089	0.7539	-93.4430
Methane(aq)	4.517e-094	7.244e-090	1.0000	-93.3452
Ba(Glyc)+	6.307e-095	1.339e-089	0.9308	-94.2314
Pb(Glyc)+	2.716e-095	7.664e-090	0.9308	-94.5972
S203--	1.960e-095	2.198e-090	0.7495	-94.8328
Cd(Glyc)+	6.918e-096	1.296e-090	0.9308	-95.1912
S204--	1.716e-096	2.198e-091	0.7539	-95.8882
Acetate	1.562e-099	9.221e-095	0.9314	-98.8371
MgCH3COO+	5.135e-102	4.279e-097	0.9308	-101.3206
CaCH3COO+	5.122e-102	5.076e-097	0.9308	-101.3217
NaCH3COO(aq)	1.421e-102	1.165e-097	1.0000	-101.8474
Acetic acid(aq)	8.388e-103	5.035e-098	1.0000	-102.0764
KCH3COO(aq)	4.980e-103	4.886e-098	1.0000	-102.3027
MnCH3COO+	1.680e-103	1.914e-098	0.9308	-102.8058
CuCH3COO+	9.271e-104	1.136e-098	0.9308	-103.0640
FeCH3COO+	2.245e-104	2.579e-099	0.9308	-103.6799
BaCH3COO+	9.090e-105	1.784e-099	0.9308	-104.0726
PbCH3COO+	5.527e-105	1.471e-099	0.9308	-104.2887
CdCH3COO+	1.282e-105	2.197e-100	0.9308	-104.9234
Cu(NH3)3++	6.515e-108	7.466e-103	0.7517	-107.3100
Malonate	2.561e-108	2.613e-103	0.7495	-107.7167
CuCH3COO(aq)	4.888e-109	5.990e-104	1.0000	-108.3109
AlCH3COO++	1.755e-110	1.510e-105	0.7517	-109.8796
H-Malonate	9.791e-111	1.009e-105	0.9308	-110.0403
AsH3(aq)	1.147e-111	8.934e-107	1.0000	-110.9406
Malonic acid(aq)	6.601e-116	6.867e-111	1.0000	-115.1804
Methanamine(aq)	2.905e-118	9.021e-114	1.0000	-117.5368

	SpecE8_output_GSS_Decant	Pond	June.txt	
Cu(Gl y)+	2. 658e-121	3. 656e-116	0. 9308	-120. 6067
Gl yci ne(aq)	1. 796e-121	1. 348e-116	1. 0000	-120. 7457
Mg(Gl y)+	8. 118e-124	7. 983e-119	0. 9308	-123. 1217
Mn(Gl y)+	5. 129e-125	6. 614e-120	0. 9308	-124. 3212
Fe(Gl y)+	2. 627e-125	3. 411e-120	0. 9308	-124. 6117
Ca(Gl y)+	1. 225e-125	1. 398e-120	0. 9308	-124. 9430
Pb(Gl y)+	7. 272e-126	2. 045e-120	0. 9308	-125. 1695
Cd(Gl y)+	8. 119e-127	1. 514e-121	0. 9308	-126. 1216
S306--	4. 043e-128	7. 768e-123	0. 7495	-127. 5185
Ba(Gl y)+	2. 613e-128	5. 522e-123	0. 9308	-127. 6140
SCN-	6. 456e-134	3. 749e-129	0. 9301	-133. 2215
Acetal dehyde(aq)	4. 268e-134	1. 879e-129	1. 0000	-133. 3698
NH4CH3COO(aq)	6. 128e-135	4. 722e-130	1. 0000	-134. 2127
Acetami de(aq)	5. 361e-135	3. 166e-130	1. 0000	-134. 2707
UO2SCN+	3. 699e-147	1. 213e-141	0. 9308	-146. 4631
Cd(NH3)4++	1. 110e-149	2. 004e-144	0. 7517	-149. 0786
Ethanol (aq)	2. 189e-154	1. 008e-149	1. 0000	-153. 6598
Ethyne(aq)	6. 190e-156	1. 611e-151	1. 0000	-155. 2083
Ethyl ene(aq)	4. 523e-159	1. 268e-154	1. 0000	-158. 3446
Lactate	1. 915e-160	1. 705e-155	0. 9308	-159. 7491
Ca(Lac)+	2. 072e-162	2. 675e-157	0. 9308	-161. 7148
Mg(Lac)+	6. 690e-163	7. 582e-158	0. 9308	-162. 2057
Na(Lac) (aq)	2. 546e-163	2. 852e-158	1. 0000	-162. 5941
K(Lac) (aq)	1. 249e-163	1. 601e-158	1. 0000	-162. 9034
Mn(Lac)+	3. 475e-164	5. 003e-159	0. 9308	-163. 4902
Cu(Lac)+	2. 519e-164	3. 844e-159	0. 9308	-163. 6299
Fe(Lac)+	2. 119e-164	3. 070e-159	0. 9308	-163. 7050
Lacti c_aci d(aq)	1. 320e-164	1. 189e-159	1. 0000	-163. 8793
Ba(Lac)+	5. 255e-166	1. 189e-160	0. 9308	-165. 3106
Pb(Lac)+	5. 197e-166	1. 539e-160	0. 9308	-165. 3154
Cd(Lac)+	9. 269e-167	1. 867e-161	0. 9308	-166. 0641
Ethane(aq)	4. 202e-169	1. 263e-164	1. 0000	-168. 3766
Propanoate	2. 128e-172	1. 555e-167	0. 9308	-171. 7031
Ca(Prop)+	3. 885e-175	4. 394e-170	0. 9308	-174. 4418
Na(Prop) (aq)	2. 790e-175	2. 679e-170	1. 0000	-174. 5544
Mg(Prop)+	1. 900e-175	1. 849e-170	0. 9308	-174. 7525
Propanoi c_aci d(a	1. 540e-175	1. 141e-170	1. 0000	-174. 8123
K(Prop) (aq)	1. 369e-175	1. 535e-170	1. 0000	-174. 8636
Mn(Prop)+	2. 662e-176	3. 406e-171	0. 9308	-175. 6060
Cu(Prop)+	1. 230e-176	1. 680e-171	0. 9308	-175. 9412
Fe(Prop)+	6. 887e-177	8. 876e-172	0. 9308	-176. 1931
Pb(Prop)+	6. 573e-178	1. 842e-172	0. 9308	-177. 2134
Ba(Prop)+	1. 680e-178	3. 534e-173	0. 9308	-177. 8058
Cd(Prop)+	1. 597e-178	2. 961e-173	0. 9308	-177. 8279
USCN+++	3. 252e-179	9. 628e-174	0. 5293	-178. 7641
Ca(Gl yc)2(aq)	3. 020e-179	5. 742e-174	1. 0000	-178. 5199
S406--	1. 510e-179	3. 386e-174	0. 7495	-178. 9461
Succi nate	1. 365e-179	1. 583e-174	0. 7495	-178. 9902
Mg(Gl yc)2(aq)	3. 245e-180	5. 657e-175	1. 0000	-179. 4888
Cu(Gl yc)2(aq)	1. 779e-180	3. 798e-175	1. 0000	-179. 7499
Fe(Gl yc)2(aq)	6. 281e-181	1. 293e-175	1. 0000	-180. 2020
Mn(Gl yc)2(aq)	1. 715e-181	3. 516e-176	1. 0000	-180. 7656
Na(Gl yc)2-	1. 242e-181	2. 149e-176	0. 9308	-180. 9371
K(Gl yc)2-	5. 422e-182	1. 025e-176	0. 9308	-181. 2970
H-Succi nate	4. 796e-182	5. 613e-177	0. 9308	-181. 3503
Pb(Gl yc)2(aq)	1. 454e-182	5. 193e-177	1. 0000	-181. 8375
Seri ne(aq)	5. 921e-183	6. 221e-178	1. 0000	-182. 2276
Ba(Gl yc)2(aq)	2. 689e-183	7. 727e-178	1. 0000	-182. 5704
Cd(Gl yc)2(aq)	1. 865e-183	4. 895e-178	1. 0000	-182. 7293
Succi ni c_aci d(aq)	7. 628e-186	9. 005e-181	1. 0000	-185. 1176
Ethanami ne(aq)	6. 202e-191	2. 795e-186	1. 0000	-190. 2075
Al ani ne(aq)	4. 846e-193	4. 316e-188	1. 0000	-192. 3147
Cu(Al a)+	1. 624e-193	2. 461e-188	0. 9308	-192. 8207

	SpecE8_output_GSS_Decant	Pond	June.txt	
Fe(Al a) +	8. 192e-197	1. 179e-191	0. 9308	-196. 1178
Mg(Al a) +	1. 594e-197	1. 790e-192	0. 9308	-196. 8288
Mn(Al a) +	1. 240e-197	1. 774e-192	0. 9308	-196. 9376
Ca(Al a) +	6. 444e-198	8. 256e-193	0. 9308	-197. 2220
Pb(Al a) +	1. 623e-198	4. 790e-193	0. 9308	-197. 8209
Cd(Al a) +	1. 238e-198	2. 481e-193	0. 9308	-197. 9385
Ca(CH3COO)2(aq)	1. 318e-199	2. 084e-194	1. 0000	-198. 8800
Mg(CH3COO)2(aq)	5. 337e-200	7. 597e-195	1. 0000	-199. 2727
Ba(Al a) +	5. 170e-201	1. 165e-195	0. 9308	-200. 3177
Cu(CH3COO)2(aq)	4. 131e-201	7. 502e-196	1. 0000	-200. 3839
Mn(CH3COO)2(aq)	1. 908e-201	3. 301e-196	1. 0000	-200. 7193
Na(CH3COO)2-	1. 152e-201	1. 624e-196	0. 9308	-200. 9698
Fe(CH3COO)2(aq)	6. 387e-202	1. 111e-196	1. 0000	-201. 1947
Acetone(aq)	4. 653e-202	2. 701e-197	1. 0000	-201. 3323
K(CH3COO)2-	2. 754e-202	4. 328e-197	0. 9308	-201. 5912
Pb(CH3COO)2(aq)	6. 905e-203	2. 245e-197	1. 0000	-202. 1608
Ba(CH3COO)2(aq)	4. 026e-203	1. 028e-197	1. 0000	-202. 3951
Cd(CH3COO)2(aq)	3. 607e-203	8. 311e-198	1. 0000	-202. 4429
Aspartic acid(aq)	6. 263e-204	8. 334e-199	1. 0000	-203. 2032
Propanal(aq)	2. 125e-206	1. 234e-201	1. 0000	-205. 6726
Al(CH3COO)2+	2. 333e-207	3. 384e-202	0. 9308	-206. 6632
Cu(CH3COO)2-	9. 465e-208	1. 719e-202	0. 9308	-207. 0550
S3--	2. 591e-224	2. 492e-219	0. 7495	-223. 7117
1-Propyne(aq)	1. 236e-224	4. 952e-220	1. 0000	-223. 9078
1-Propanol(aq)	3. 392e-227	2. 038e-222	1. 0000	-226. 4696
1-Propene(aq)	1. 182e-229	4. 974e-225	1. 0000	-228. 9272
2-Hydroxybutanoic acid(aq)	1. 091e-233	1. 124e-228	0. 9308	-232. 9934
NH4(CH3COO)2-	8. 429e-234	1. 147e-228	0. 9308	-233. 1054
Asparagine(aq)	4. 889e-236	6. 457e-231	1. 0000	-235. 3108
Cu(Gly)2(aq)	1. 153e-236	2. 440e-231	1. 0000	-235. 9381
2-Hydroxybutanoic acid(aq)	6. 674e-238	6. 945e-233	1. 0000	-237. 1756
Propane(aq)	2. 665e-242	1. 175e-237	1. 0000	-241. 5744
Mg(Gly)2(aq)	1. 520e-243	2. 620e-238	1. 0000	-242. 8182
Mn(Gly)2(aq)	7. 911e-245	1. 606e-239	1. 0000	-244. 1018
Diglycine(aq)	7. 553e-245	9. 976e-240	1. 0000	-244. 1219
Fe(Gly)2(aq)	7. 508e-245	1. 531e-239	1. 0000	-244. 1245
Pb(Gly)2(aq)	2. 541e-245	9. 025e-240	1. 0000	-244. 5950
Butanoate	9. 532e-246	8. 300e-241	0. 9308	-245. 0520
Cd(Gly)2(aq)	7. 515e-246	1. 957e-240	1. 0000	-245. 1241
Diketopiperazine	3. 646e-247	4. 159e-242	1. 0000	-246. 4382
Ca(Gly)2(aq)	9. 321e-248	1. 754e-242	1. 0000	-247. 0305
Na(But)(aq)	1. 198e-248	1. 319e-243	1. 0000	-247. 9214
Ca(But) +	1. 195e-248	1. 520e-243	0. 9308	-247. 9537
K(But)(aq)	5. 880e-249	7. 418e-244	1. 0000	-248. 2306
Butanoic acid(aq)	5. 593e-249	4. 926e-244	1. 0000	-248. 2524
Mg(But) +	5. 576e-249	6. 210e-244	0. 9308	-248. 2848
Mn(But) +	1. 015e-249	1. 442e-244	0. 9308	-249. 0245
Cu(But) +	4. 360e-250	6. 566e-245	0. 9308	-249. 3916
Fe(But) +	3. 178e-250	4. 541e-245	0. 9308	-249. 5290
Ba(Gly)2(aq)	2. 340e-250	6. 676e-245	1. 0000	-249. 6309
Pb(But) +	8. 388e-252	2. 468e-246	0. 9308	-251. 1075
Ba(But) +	5. 044e-252	1. 132e-246	0. 9308	-251. 3284
Gluconate	5. 016e-252	6. 524e-247	0. 7495	-251. 4249
Cd(But) +	3. 056e-252	6. 094e-247	0. 9308	-251. 5461
H-Gluconate	1. 044e-254	1. 368e-249	0. 9308	-254. 0126
Threonine(aq)	2. 741e-256	3. 263e-251	1. 0000	-255. 5622
Ethyl acetate(aq)	5. 479e-258	4. 825e-253	1. 0000	-257. 2613
Gluconic acid(aq)	2. 153e-258	2. 844e-253	1. 0000	-257. 6669
S506--	1. 729e-260	4. 430e-255	0. 7495	-259. 8875
1-Propanamine(aq)	3. 222e-263	1. 904e-258	1. 0000	-262. 4918
alpha-Aminobutyric acid(aq)	4. 664e-266	4. 808e-261	1. 0000	-265. 3312
Gluconic acid(aq)	2. 751e-275	4. 046e-270	1. 0000	-274. 5606
UO2(SCN)2(aq)	1. 268e-280	4. 897e-275	1. 0000	-279. 8968

	SpecE8_output_GSS_Decant	Pond	June.txt	
Butanal (aq)	3.504e-281	2.526e-276	1.0000	-280.4555
S4--	1.056e-289	1.354e-284	0.7495	-289.1016
1-Butyne(aq)	5.259e-298	2.844e-293	1.0000	-297.2791
Cu(CH3COO)3-	1.863e-298	4.481e-293	0.9308	-297.7611
Pb(CH3COO)3-	1.096e-299	4.212e-294	0.9308	-298.9912
Mn(CH3COO)3-	9.148e-300	2.122e-294	0.9308	-299.0699
1-Butanol (aq)	2.944e-301	2.181e-296	1.0000	-300.0000
Cd(CH3COO)3-	1.454e-301	4.208e-296	0.9308	-300.0000
1-Butene(aq)	3.250e-303	1.823e-298	1.0000	-300.0000
2-Hydroxypentano	9.287e-307	1.087e-301	0.9308	-300.0000
Glutamine(aq)	2.460e-309	3.595e-304	1.0000	-300.0000
2-Hydroxypentano	3.373e-311	3.983e-306	1.0000	-300.0000
U(SCN)2++	2.434e-311	8.617e-306	0.7517	-300.0000
n-Butane(aq)	1.479e-315	8.592e-311	1.0000	-300.0000
Alanylglycine(aq)	9.867e-317	1.442e-311	1.0000	-300.0000
Pentanoate	5.093e-319	5.148e-314	0.9308	-300.0000
Ca(Lac)2(aq)	4.511e-321	9.840e-316	1.0000	-300.0000
Mg(Lac)2(aq)	1.443e-321	2.920e-316	1.0000	-300.0000
Na(Pent)(aq)	6.719e-322	8.337e-317	1.0000	-300.0000
Ca(Pent)+	3.755e-322	5.300e-317	0.9308	-300.0000
K(Pent)(aq)	3.310e-322	4.640e-317	1.0000	-300.0000
Pentanoic acid(a	3.261e-322	3.329e-317	1.0000	-300.0000
Cu(Lac)2(aq)	2.075e-322	5.014e-317	1.0000	-300.0000
Fe(Lac)2(aq)	1.976e-322	4.623e-317	1.0000	-300.0000
Mg(Pent)+	1.680e-322	2.106e-317	0.9308	-300.0000
Mn(Lac)2(aq)	7.905e-323	1.842e-317	1.0000	-300.0000
Na(Lac)2-	5.435e-323	1.093e-317	0.9308	-300.0000
Mn(Pent)+	4.447e-323	6.937e-318	0.9308	-300.0000
Cu(Pent)+	1.976e-323	3.253e-318	0.9308	-300.0000
K(Lac)2-	1.976e-323	4.292e-318	0.9308	-300.0000
Fe(Pent)+	1.482e-323	2.326e-318	0.9308	-300.0000
Pb(Lac)2(aq)	4.941e-324	1.903e-318	1.0000	-300.0000
m-Toluate	0.0000	0.0000	0.9308	-300.0000
Valine(aq)	0.0000	0.0000	1.0000	-300.0000
Undecanoic acid(0.0000	0.0000	1.0000	-300.0000
Undecanoate	0.0000	0.0000	0.9308	-300.0000
H-Suberate	0.0000	0.0000	0.9308	-300.0000
UO2(SCN)3-	0.0000	0.0000	0.9308	-300.0000
H-Sebacate	0.0000	0.0000	0.9308	-300.0000
H-Pimelate	0.0000	0.0000	0.9308	-300.0000
2-Hexanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexyne(aq)	0.0000	0.0000	1.0000	-300.0000
H-Azelate	0.0000	0.0000	0.9308	-300.0000
Ca(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
H-Adipate	0.0000	0.0000	0.9308	-300.0000
Tyrosine(aq)	0.0000	0.0000	1.0000	-300.0000
Toluene(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Suberic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Suberate	0.0000	0.0000	0.7495	-300.0000
2-Heptanone(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacate	0.0000	0.0000	0.7495	-300.0000
Sb2S4--	0.0000	0.0000	0.7495	-300.0000
S5--	0.0000	0.0000	0.7495	-300.0000
2-Butanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexene(aq)	0.0000	0.0000	1.0000	-300.0000
Pimelic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Pimelate	0.0000	0.0000	0.7495	-300.0000
Phenylalanine(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol(aq)	0.0000	0.0000	1.0000	-300.0000
Pentanal(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000

	SpecE8_output_GSS_Decant	Pond	June.txt	
Fe(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Adipi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Pent)+	0.0000	0.0000	0.9308	-300.0000
Ca(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Adi pate	0.0000	0.0000	0.7495	-300.0000
1-Hexanol (aq)	0.0000	0.0000	1.0000	-300.0000
Pb(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9308	-300.0000
Fe(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanal (aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoate	0.0000	0.0000	0.9308	-300.0000
Nonanal (aq)	0.0000	0.0000	1.0000	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Prop)2-	0.0000	0.0000	0.9308	-300.0000
Na(Pent)2-	0.0000	0.0000	0.9308	-300.0000
Benzoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
Dodecanoate	0.0000	0.0000	0.9308	-300.0000
Benzoate	0.0000	0.0000	0.9334	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
Na(But)2-	0.0000	0.0000	0.9308	-300.0000
Decanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9308	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Decanal (aq)	0.0000	0.0000	1.0000	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)+	0.0000	0.0000	0.9308	-300.0000
2-Hydroxyoctanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanol (aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoa	0.0000	0.0000	0.9308	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9308	-300.0000
Methi oni ne(aq)	0.0000	0.0000	1.0000	-300.0000
Leucyl gl yci ne(aq)	0.0000	0.0000	1.0000	-300.0000

	SpecE8_output_GSS_Decant	Pond	June.txt	
Leuci ne(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoi	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoa	0.0000	0.0000	0.9308	-300.0000
K(Prop)2-	0.0000	0.0000	0.9308	-300.0000
Cd(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
K(Pent)2-	0.0000	0.0000	0.9308	-300.0000
Cd(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Pent)+	0.0000	0.0000	0.9308	-300.0000
Ba(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol (aq)	0.0000	0.0000	1.0000	-300.0000
K(But)2-	0.0000	0.0000	0.9308	-300.0000
I sol euci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9308	-300.0000
Azel ate	0.0000	0.0000	0.7495	-300.0000
2-Hydroxyheptano	0.0000	0.0000	0.9308	-300.0000
Hexanal (aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoi c_aci d(a	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9308	-300.0000
Heptanal (aq)	0.0000	0.0000	1.0000	-300.0000
Azel ai c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
HSb2S4-	0.0000	0.0000	0.9308	-300.0000
2-Hydroxydecanoi	0.0000	0.0000	1.0000	-300.0000
1-Octanol (aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoa	0.0000	0.0000	0.9308	-300.0000
p-Tol uate	0.0000	0.0000	0.9308	-300.0000
o-Tol uate	0.0000	0.0000	0.9308	-300.0000
Cd(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Propyl benzene(0.0000	0.0000	1.0000	-300.0000
m-Tol ui c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
n-Pentyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octyl benzene(a	0.0000	0.0000	1.0000	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexyl benzene(a	0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(a	0.0000	0.0000	1.0000	-300.0000
1-Octanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
o-Phthal ate	0.0000	0.0000	0.7495	-300.0000
1-Heptanami ne(aq	0.0000	0.0000	1.0000	-300.0000
o-Tol ui c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
1-Butanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
p-Tol ui c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
(NH4)2Sb2S4(aq)	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states

	log Q/K		log Q/K
Nontronite-Ca	24.7416s/sat	BaU207	-10.6158
Nontronite-Mg	24.6938s/sat	MgU04	-10.6863
Nontronite-K	24.6091s/sat	Na2U207	-10.7635
Nontronite-Na	24.3261s/sat	Portlandite	-10.7687
Nontronite-H	23.6719s/sat	Cotunnite	-10.8330
Hematite	20.1990s/sat	Mg1.5S04(OH)	-11.2407
Cronstedtite-7A	19.3646s/sat	AlF3	-11.2740

SpecE8_output_GSS_Decant Pond June.txt			
Magnetite	19. 1299s/sat	Torbernite	-11. 5875
Ferri te-Cu	18. 9411s/sat	Uranocircite	-11. 6012
Delafossite	15. 0009s/sat	MnSO4	-11. 7434
Clinoptilolite-h	12. 9113s/sat	CdSO4	-11. 7635
Clinoptilolite-C	12. 8547s/sat	UO2. 6667	-11. 7754
Andradite	12. 7477s/sat	Thermonatrite	-11. 8127
Clinoptilolite-K	12. 5588s/sat	Na2CO3	-12. 1046
Clinoptilolite-h	12. 3360s/sat	UO2SO4: H2O	-12. 3187
Muscovite	11. 1226s/sat	Uraninite	-12. 3322
Epidote-ord	11. 0277s/sat	Coffinite	-12. 5923
Epidote	11. 0269s/sat	(UO2)3(PO4)2: 4H2O	-12. 6577
Stilbite	10. 6439s/sat	PbF2	-12. 7171
Ferri te-Ca	10. 5340s/sat	MgSO4	-12. 7307
Ferri te-Mg	10. 5025s/sat	FeSO4	-12. 7412
Goethite	9. 6357s/sat	Przhevalskite	-13. 0273
Clinoptilolite-N	9. 2525s/sat	UP05	-13. 0752
Clinoptilolite-h	9. 2520s/sat	UO2ClOH: 2H2O	-13. 1324
Fluorapatite	9. 0603s/sat	Mordenite-dehy	-13. 2449
Illite	7. 8871s/sat	Chalcocyanite	-13. 3867
Mesolite	7. 4478s/sat	Ferri te-Dicalcium	-13. 4867
Paragonite	7. 2510s/sat	Anthophyllite	-13. 6354
Beddellite-Ca	7. 2285s/sat	Bassetite	-13. 9167
Beddellite-Mg	7. 1806s/sat	FeF2	-14. 1285
Beddellite-K	7. 0961s/sat	BaCl 2: 2H2O	-14. 3046
Beddellite-Na	6. 8130s/sat	CdCl 2: H2O	-14. 4177
Kaolinite	6. 5126s/sat	NaUO3	-14. 7132
Montmor-Mg	6. 4461s/sat	Ca3(AsO4)2	-14. 8798
Montmor-Ca	6. 4207s/sat	BaCl 2: H2O	-15. 0006
Montmor-K	6. 3616s/sat	Ba(OH)2: 8H2O	-15. 0512
Beddellite-H	6. 1592s/sat	Hydromagnesite	-15. 1926
Montmor-Na	6. 0742s/sat	Ba2Si 308	-15. 1954
Scolécite	6. 0607s/sat	CdCl 2	-15. 5091
Pyrophyllite	5. 9264s/sat	CaAl 204	-15. 5278
Jarosite	5. 7104s/sat	Natrosilite	-15. 5604
Smectite-low-Fe-	5. 3660s/sat	Autunite-H	-15. 6276
Celadonite	5. 2963s/sat	MnCl 2: 4H2O	-16. 0241
Smectite-high-Fe	4. 8945s/sat	Na2Si O3	-16. 2258
Maximiliani crocchi	4. 8082s/sat	Gehlenite	-16. 2564
K-Feldspar	4. 8067s/sat	MgOHCl	-16. 4524
Laumontite	4. 7655s/sat	BaCl 2	-16. 5067
Fe(OH)3	4. 3770s/sat	Hillebrandite	-16. 5459
Daphnite-14A	4. 3055s/sat	UO2SO4: 3. 5H2O	-16. 8052
Amesite-14A	4. 1026s/sat	UO2SO4: 2. 5H2O	-16. 8446
Margarite	4. 0879s/sat	UO2SO4: 3H2O	-16. 9250
Saniidine-high	3. 5417s/sat	Dicalciumsilicate	-16. 9853
Lawsonite	3. 1157s/sat	Akermanite	-17. 0287
Annite	3. 1014s/sat	UO2. 3333(beta)	-17. 0952
Diaspore	2. 9228s/sat	CdF2	-17. 2396
Mordenite	2. 5047s/sat	CaAl 407	-17. 3748
Boehmite	2. 5019s/sat	UO2F2: 3H2O	-17. 3800
Chamosite-7A	2. 4688s/sat	MnCl 2: 2H2O	-17. 4120
Gibbsite	2. 3751s/sat	UO2F2	-17. 6469
Ripidolite-14A	2. 3065s/sat	Pargassite	-17. 9934
Witherite	2. 0641s/sat	Ni ngoyite	-18. 0591
Brochantite	1. 8878s/sat	Larnite	-18. 3364
Albite-low	1. 8588s/sat	Al um-K	-18. 4202
Albite	1. 8587s/sat	CuCl 2	-18. 5388
Phlogopite	1. 8402s/sat	MnCl 2: H2O	-19. 0529
Malachite	1. 6640s/sat	MgCl 2: 4H2O	-19. 4288
Corundum	1. 6421s/sat	Na2UO4(al pha)	-19. 5652
Tenorite	1. 4015s/sat	UO2SO4	-20. 5765
Saponite-Ca	1. 2335s/sat	UOFOH: . 5H2O	-20. 7264
Analcite	1. 1884s/sat	Lime	-21. 1831

SpecE8_output_GSS_Decant Pond June.txt			
Saponi te-Mg	1. 1868s/sat	UOF0H	-21. 2189
Natrol i te	1. 1066s/sat	Scacchi te	-22. 4099
Saponi te-K	1. 1011s/sat	U02Cl	-22. 7001
Prehni te	0. 9323s/sat	U02(N03)2: 6H2O	-22. 8769
Daphni te-7A	0. 8413s/sat	Ranki ni te	-23. 2943
Saponi te-Na	0. 8181s/sat	Hydrophi l i te	-23. 6738
Kyani te	0. 8020s/sat	Lawrenci te	-23. 7427
Cerussi te	0. 5971s/sat	K2U04	-23. 9348
Quartz	0. 5692s/sat	Pb	-24. 0319
Andalusi te	0. 5055s/sat	Tobermori te-14A	-24. 2047
Cl i nozoi si te	0. 4746s/sat	UOF2: H2O	-24. 2177
Al bi te_hi gh	0. 4732s/sat	Cl i nopti l ol i te-d	-24. 3156
Cl i nochl ore-14A	0. 4689s/sat	U02(N03)2: 3H2O	-24. 4495
Bari te	0. 4649s/sat	Cl i nopti l ol i te-d	-24. 6367
Zoi si te	0. 4275s/sat	UOF2	-24. 8393
Hercyni te	0. 3836s/sat	MgCl 2: 2H2O	-25. 1328
Kal si l i te	0. 3812s/sat	U02(N03)2: 2H2O	-25. 7726
Tri dymi te	0. 3694s/sat	U02(P03)2	-26. 1925
Mi nnesotai te	0. 3298s/sat	(U02)3(P04)2	-26. 2777
Chal cedony	0. 2886s/sat	Tobermori te-11A	-26. 2974
Dol omi te-ord	0. 1955s/sat	Ettri ngi te	-26. 4377
Dol omi te	0. 1955s/sat	(U02)2P207	-26. 7018
Dawsoni te	0. 1816s/sat	U(HP04)2: 4H2O	-27. 2183
Saponi te-H	0. 1641s/sat	UP207	-27. 4111
Si l l i mani te	0. 1223s/sat	KAl (S04)2	-27. 5736
Wai raki te	0. 0357s/sat	Cl i nopti l ol i te-d	-27. 9270
Rhodochrosi te	0. 0052s/sat	U02Cl 2: 3H2O	-28. 5012
Cri stobal i te(al p	-0. 0041	Merwi ni te	-28. 5130
Tal c	-0. 0747	MgCl 2: H2O	-28. 6391
Strengi te	-0. 0793	Foshagi te	-28. 7218
I ce	-0. 1008	U02(N03)2: H2O	-29. 5197
Cal ci te	-0. 2647	Tobermori te-9A	-30. 0203
Coesi te	-0. 2671	Ba2Si 04	-30. 0617
Jadei te	-0. 3332	U02Cl 2: H2O	-31. 3683
Aragoni te	-0. 4095	Afwil l i te	-31. 5225
Di optase	-0. 4325	U02(N03)2	-33. 1303
Hydroxyl apati te	-0. 4644	KMgCl 3: 2H2O	-33. 4141
Cri stobal i te(bet	-0. 4762	U02S03	-33. 6252
Si 02(am)	-0. 7992	Cd	-34. 1320
Si deri te	-1. 0107	Chl oromagnesi te	-34. 6190
Monohydrocal ci te	-1. 0738	Fe	-34. 7530
Ri pi dol i te-7A	-1. 1521	UF4: 2. 5H2O	-35. 2230
Magnesi te	-1. 2308	U02Cl 2	-35. 3992
Atacami te	-1. 2610	As205	-36. 2408
Al amosi te	-1. 2894	Anti gori te	-37. 2881
Dol omi te-di s	-1. 4233	BaSi F6	-39. 1742
Greenal i te	-1. 4667	Ba0	-39. 6432
Anorthi te	-1. 5999	UF4	-39. 6481
Ferrosi l i te	-1. 6040	(U02)2As207	-39. 7201
Cupri te	-1. 6240	Al 2(S04)3: 6H2O	-39. 8477
Whi t l ocki te	-2. 2541	Mol ysi te	-39. 8603
Nephel i ne	-2. 4218	U02(As03)2	-40. 0993
Gypsum	-2. 4266	(U02)3(As04)2	-40. 2199
Anhydri te	-2. 7003	Ba2U207	-40. 5477
Bi xbyi te	-2. 7315	KMgCl 3	-41. 0321
CaU04	-2. 8073	Xonotl i te	-41. 2933
Lanarki te	-2. 8743	Hatruri te	-41. 9669
Al uni te	-2. 9179	Fe2(S04)3	-42. 0269
Rhodoni te	-2. 9818	U(S04)2: 8H2O	-43. 3206
Cl i nochl ore-7A	-2. 9869	Arsenol i te	-43. 4519
Azuri te	-3. 2169	Cl audeti te	-43. 5230
Chrysoti l e	-3. 2247	C	-43. 9540
Enstati te	-3. 2968	U(S04)2: 4H2O	-44. 5540

SpecE8_output_GSS_Decant Pond June.txt			
Fluorite	-3.3389	U(SO ₄) ₂	-44.7363
Bassanite	-3.3482	U(CO ₃) ₂	-46.9183
Anglesite	-3.4089	UOCl ₂	-46.9713
Sb ₂ O ₃	-3.4404	U ₅ O ₁₂ Cl	-48.8040
CaSO ₄ · 0.5H ₂ O (bet)	-3.5290	UClF ₃	-50.6442
CdSiO ₃	-3.5471	As	-51.2115
Hausmannite	-3.5576	Na ₃ UO ₄	-53.9264
Wustite	-3.6532	Ca ₄ Al ₂ Fe ₂ O ₁₀	-54.5139
Schoepite	-3.7786	Na ₄ SiO ₄	-55.9677
UO ₃ · 2H ₂ O	-3.7798	UOF ₄	-56.1672
UO ₂ (OH) ₂ (beta)	-3.9287	Na	-56.2553
Schoepite-dehyd.	-3.9930	Al ₂ (SO ₄) ₃	-58.1497
UO ₃ · 9H ₂ O (alpha)	-3.9935	Mn	-58.5025
Paralaurionite	-4.0042	Ca ₃ Al ₂ O ₆	-58.8647
Fayalite	-4.0183	(UO ₂) ₂ Cl ₃	-59.2710
Schoepite-dehyd.	-4.0773	Na ₂ O	-59.4133
Schoepite-dehyd(1)	-4.0903	K	-60.1918
Nesquehoniite	-4.2892	UF ₃	-60.2421
Dioptase	-4.3049	UOCl ₃	-60.2921
Cordierite-hydr	-4.3302	Pb ₂ Cl ₅ NH ₄	-60.4620
FeO	-4.3516	UF ₅ (beta)	-63.2730
Berlinite	-4.3637	Covelite	-63.3007
Pyrolusite	-4.4612	UF ₅ (alpha)	-63.6551
Sepiolite	-4.5390	UOCl	-63.6848
Sellaite	-4.5487	Chalcocite	-63.9112
PbCO ₃ · PbO	-4.5760	UCl ₂ F ₂	-63.9551
Fe(OH) ₂	-4.6612	S	-66.2158
Brucite	-4.8129	Ba ₃ UO ₆	-67.0174
Litharge	-4.9052	CdS	-71.9585
Massicot	-5.0976	Galena	-72.7066
Mn(OH) ₂ (am)	-5.1289	U ₂ O ₃ F ₆	-73.9621
Okenite	-5.2015	K ₂ O	-77.0032
Hedenbergite	-5.2292	UCl ₃ F	-77.0846
Schoepite-dehyd.	-5.2407	CdCl ₂ (NH ₃) ₂	-81.1722
NaFeO ₂	-5.2567	Na ₆ Si ₂ O ₇	-81.2996
Wollastonite	-5.2858	U(SO ₃) ₂	-81.8409
Analcite-dehyd	-5.4210	Troilite	-82.6013
UO ₂ CO ₃	-5.4603	Pyrrhotite	-82.7037
Rutherfordine	-5.4782	PbSO ₄ (NH ₃) ₂	-83.4469
Pseudowollastonite	-5.5429	Alabandite	-85.1749
Nahcolite	-5.5444	As ₄ O ₆ (mono)	-86.5612
Huntite	-5.6008	As ₄ O ₆ (cubi)	-86.7821
Ni ₂ S	-5.7803	UCl ₄	-88.0498
Schoepite-dehyd.	-5.7810	UCl ₃	-90.3736
Pb ₂ SiO ₄	-5.9288	UF ₆	-95.7164
Tephroite	-6.1398	Mg	-98.1874
Pb ₃ SO ₆	-6.1561	U ₃ O ₅ F ₈	-98.8347
Grossular	-6.2186	U ₂ F ₉	-99.6782
Cd(OH) ₂	-6.3482	BaS	-103.2237
Spinel	-6.3755	UN _{1.73} (alpha)	-105.2680
Cu	-6.6478	UN _{1.59} (alpha)	-108.0614
Tremolite	-6.7941	U ₂ O ₂ Cl ₅	-108.5178
Nantokite	-6.8159	UCl ₅	-109.8337
UO ₃ (gamma)	-6.8358	Ca	-115.5592
Cordierite-anhyd	-6.9147	P	-117.2064
Artinite	-6.9572	Ba	-119.6947
Sanbornite	-7.1232	Al	-119.9205
Melanterite	-7.2325	UN	-121.3085
UO ₂ HP0 ₄ · 4H ₂ O	-7.2806	Si	-124.3627
Chalcanthite	-7.3783	UCl ₆	-130.4253
UO ₃ (beta)	-7.4600	Pb(N ₃) ₂ (orth)	-137.9602
PbFCI	-7.6808	Pyrite	-138.0734
Ca-Al-Pyroxene	-7.7231	Pb(N ₃) ₂ (mono)	-138.4109

SpecE8_output_GSS_Decant Pond June.txt			
Mirabilite	-7.7355	Chalcopyrite	-139.7899
Cd(OH)Cl	-7.7917	CdCl ₂ (NH ₃) ₄	-157.3412
Montepionite	-7.8031	PbSO ₄ (NH ₃) ₄	-161.2503
UO ₃ (alpha)	-7.8064	U	-169.0862
BaUO ₄	-7.8532	US	-177.1208
Manganosite	-7.8881	UAs	-177.1946
Arcanite	-7.8913	U ₄ F ₁₇	-178.0789
Sylvite	-7.9034	UC	-195.1184
Forsterite	-8.4157	UH ₃ (beta)	-197.5607
Halite	-8.4666	US _{1.9}	-202.5438
Phosgenite	-8.6016	US ₂	-207.3899
Thenardite	-9.0205	Mayenite	-217.3524
UO ₂ F _{0H} : 2H ₂ O	-9.1288	UAs ₂	-225.6913
Gyrolite	-9.1708	CdCl ₂ (NH ₃) ₆	-233.5967
UO ₂ . 25	-9.3082	UP	-238.0663
FeF ₃	-9.3382	UC _{1.94} (alpha)	-238.5260
UO ₂ . 25(beta)	-9.3743	Stibnite	-246.3138
Pb ₄ SO ₇	-9.5074	Bornite	-267.2668
Nitrobarite	-9.5315	US ₃	-270.3336
UO ₂ F _{0H} : H ₂ O	-9.5611	Orpiment	-270.5038
Monticellite	-9.6710	Pb(SCN) ₂	-274.1957
Salteite	-9.8198	UP ₂	-350.0482
CdSO ₄ : 2.667H ₂ O	-9.8730	U ₂ S ₃	-377.3484
Natron	-10.0008	o-Phthalic acid	-388.5448
UO ₂ F _{0H}	-10.0610	U ₂ C ₃	-435.7527
Periclase	-10.0842	U ₃ S ₅	-580.0168
CdSO ₄ : H ₂ O	-10.0960	U ₃ As ₄	-580.6483
Mg _{1.25} SO ₄ (OH)0.5	-10.1124	U ₃ P ₄	-826.1780
Na ₂ CO ₃ : 7H ₂ O	-10.4842		

Gases	fugacity	log fug.
N ₂ (g)	0.2106	-0.676
H ₂ O(g)	0.01398	-1.854
CO ₂ (g)	0.0008301	-3.081
NO ₂ (g)	3.255e-012	-11.487
HF(g)	3.155e-015	-14.501
HCl(g)	1.825e-019	-18.739
H ₂ (g)	1.030e-028	-27.987
O ₂ (g)	2.479e-031	-30.606
NO(g)	4.923e-032	-31.308
CO(g)	2.791e-035	-34.554
NH ₃ (g)	6.947e-040	-39.158
SiF ₄ (g)	1.989e-041	-40.701
SO ₂ (g)	3.741e-043	-42.427
Cl ₂ (g)	1.009e-044	-43.996
Cd(g)	4.657e-049	-48.332
Pb(g)	2.341e-054	-53.631
Cu(g)	1.465e-061	-60.834
UO ₂ F ₂ (g)	4.231e-062	-61.374
Na(g)	4.183e-071	-70.378
K(g)	4.825e-072	-71.317
UO ₃ (g)	2.205e-073	-72.657
UO ₂ Cl ₂ (g)	1.332e-073	-72.875
UOF ₄ (g)	1.375e-077	-76.862
UF ₅ (g)	1.325e-084	-83.878
UF ₄ (g)	4.487e-086	-85.348
H ₂ S(g)	6.330e-089	-88.199
CH ₄ (g)	2.573e-091	-90.590
UF ₆ (g)	1.444e-097	-96.840
F ₂ (g)	1.556e-101	-100.808
UO ₂ (g)	3.084e-113	-112.511
UCl ₄ (g)	2.036e-114	-113.691

SpecE8_output_GSS_Decant Pond June.txt

Mg(g)	1.712e-119	-118.766
UCl 5(g)	9.956e-129	-128.002
UF3(g)	2.484e-130	-129.605
UCl 6(g)	1.597e-137	-136.797
UCl 3(g)	1.035e-138	-137.985
U2F10(g)	3.274e-142	-141.485
Ca(g)	1.454e-142	-141.838
S2(g)	7.877e-148	-147.104
C2H4(g)	7.328e-157	-156.135
C(g)	1.322e-166	-165.879
UF2(g)	2.794e-171	-170.554
Al (g)	2.611e-173	-172.583
UCl 2(g)	7.751e-181	-180.111
UO(g)	1.253e-185	-184.902
Si (g)	7.594e-199	-198.120
UF(g)	1.476e-205	-204.831
U2Cl 8(g)	4.995e-217	-216.301
UCl (g)	1.751e-220	-219.757
U2Cl 10(g)	3.375e-229	-228.472
U(g)	1.295e-258	-257.888

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
<hr/>						
Al +++	6.54e-006	6.54e-006	0.177			
Ba++	8.36e-007	8.36e-007	0.115			
Ca++	0.000514	0.000514	20.6			
Cd++	1.10e-008	1.10e-008	0.00124			
Cl -	0.000111	0.000111	3.92			
Cu++	3.70e-007	3.70e-007	0.0235			
F-	1.03e-005	1.03e-005	0.196			
Fe++	8.78e-007	8.78e-007	0.0490			
H+	-2.10e-006	-2.10e-006	-0.00212			
H2AsO4-	2.62e-008	2.62e-008	0.00369			
H2O	55.5	55.5	1.00e+006			
HC03-	0.00162	0.00162	99.0			
HP04--	2.63e-006	2.63e-006	0.252			
K+	0.000702	0.000702	27.5			
Mg++	0.000202	0.000202	4.90			
Mn++	8.18e-006	8.18e-006	0.449			
NH3(aq)	0.000326	0.000326	5.56			
NO3-	0.00102	0.00102	62.9			
Na+	0.00128	0.00128	29.4			
O2(aq)	0.000244	0.000244	7.82			
Pb++	1.80e-008	1.80e-008	0.00373			
S04--	0.000459	0.000459	44.1			
Sb(OH)3(aq)	4.03e-007	4.03e-007	0.0696			
Si O2(aq)	0.000215	0.000215	12.9			
U02++	1.36e-008	1.36e-008	0.00367			

Elemental composition	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg
<hr/>					
Al umi num	6.544e-006	6.544e-006	0.1765		
Anti mony	4.028e-007	4.028e-007	0.04903		
Arseni c	2.619e-008	2.619e-008	0.001961		
Bari um	8.357e-007	8.357e-007	0.1147		
Cadmi um	1.099e-008	1.099e-008	0.001236		
Cal ci um	0.0005140	0.0005140	20.59		
Carbon	0.001624	0.001624	19.50		
Chl ori ne	0.0001107	0.0001107	3.922		
Copper	3.705e-007	3.705e-007	0.02353		
Fl uori ne	1.033e-005	1.033e-005	0.1961		

	SpecE8_output_GSS_Decant Pond June.txt		
Hydrogen	111.0	111.0	1.119e+005
Iron	8.782e-007	8.782e-007	0.04903
Lead	1.799e-008	1.799e-008	0.003726
Magnesium	0.0002018	0.0002018	4.903
Manganese	8.178e-006	8.178e-006	0.4491
Nitrogen	0.001342	0.001342	18.79
Oxygen	55.52	55.52	8.880e+005
Phosphorus	2.629e-006	2.629e-006	0.08139
Potassium	0.0007025	0.0007025	27.46
Silicon	0.0002155	0.0002155	6.050
Sodium	0.001280	0.001280	29.42
Sulfur	0.0004595	0.0004595	14.73
Uranium	1.360e-008	1.360e-008	0.003236

SpecE8_output_GSS_Decant Pond May.txt

Temperature = 9.5 C Pressure = 1.013 bars
 pH = 7.600 log fO2 = -29.159
 Eh = 0.3970 volts pe = 7.0792
 Ionic strength = 0.004191
 Charge imbalance = -0.000351 eq/kg (-10.64% error)
 Activity of water = 0.999996
 Solvent mass = 1.000000 kg
 Solution mass = 1.000306 kg
 Solution density = 1.024 g/cm3
 Chlorinity = 0.000110 molal
 Dissolved solids = 306 mg/kg sol'n
 Hardness = 57.57 mg/kg sol'n as CaCO3
 carbonate = 57.57 mg/kg sol'n as CaCO3
 non-carbonate = 0.00 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 74.27 mg/kg sol'n as CaCO3
 Water type = Na-HCO3

Nernst redox couples	Eh (volts)	pe
$e^- + .25 \cdot O_2(aq) + H^+ = .5 \cdot H_2O$	0.3970	7.0792
$8 \cdot e^- + 9 \cdot H^+ + NO_3^- = 3 \cdot H_2O + NH_3(aq)$	0.5859	10.4477
$14 \cdot e^- + 16 \cdot H^+ + 2 \cdot SO_4^{--} = 8 \cdot H_2O + S^{2--}$	-0.2531	-4.5127

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HC03-	0.001469	89.62	0.9340	-2.8626
Na+	0.001314	30.21	0.9340	-2.9109
NO3-	0.001228	76.12	0.9328	-2.9410
K+	0.0006736	26.33	0.9328	-3.2018
Ca++	0.0003995	16.00	0.7681	-3.5131
SO4--	0.0003201	30.74	0.7602	-3.6137
N2(aq)	0.0002217	6.209	1.0000	-3.6542
Mg++	0.0001536	3.732	0.7755	-3.9240
SiO2(aq)	0.0001378	8.279	1.0000	-3.8607
Cl-	0.0001102	3.907	0.9328	-3.9880
CO2(aq)	0.0001026	4.514	1.0000	-3.9889
F-	1.024e-005	0.1946	0.9334	-5.0195
Mn++	9.510e-006	0.5223	0.7681	-5.1364
CaSO4(aq)	8.979e-006	1.222	1.0000	-5.0468
CaHCO3+	5.239e-006	0.5295	0.9340	-5.3104
MgSO4(aq)	5.040e-006	0.6065	1.0000	-5.2975
AlO2-	3.423e-006	0.2018	0.9340	-5.4952
NaHCO3(aq)	3.265e-006	0.2742	1.0000	-5.4861
HP04--	3.065e-006	0.2941	0.7602	-5.6326
CO3--	2.280e-006	0.1368	0.7623	-5.7600
MgHCO3+	1.971e-006	0.1681	0.9340	-5.7350
KS04-	1.230e-006	0.1663	0.9340	-5.9396
H2PO4-	1.163e-006	0.1128	0.9340	-5.9641
Fe++	1.025e-006	0.05722	0.7681	-6.1039
CaCO3(aq)	8.372e-007	0.08377	1.0000	-6.0772
Ba++	6.900e-007	0.09472	0.7642	-6.2779
S2--	6.095e-007	0.03908	0.7602	-6.3340
HAIO2(aq)	5.322e-007	0.03191	1.0000	-6.2740
Cu++	4.306e-007	0.02735	0.7681	-6.4805
HSiO3-	3.935e-007	0.03033	0.9340	-6.4347
Sb(OH)3(aq)	3.692e-007	0.06377	1.0000	-6.4327
MnSO4(aq)	3.442e-007	0.05197	1.0000	-6.4631
MgCO3(aq)	1.701e-007	0.01434	1.0000	-6.7693
OH-	1.186e-007	0.002017	0.9334	-6.9558

	SpecE8_output_GSS_Decant Pond May.txt			
NaHSi O3(aq)	2. 857e-008	0. 002858	1. 0000	-7. 5441
MgF+	2. 813e-008	0. 001218	0. 9340	-7. 5804
Al (OH) 2+	2. 740e-008	0. 001671	0. 9340	-7. 5918
H+	2. 673e-008	2. 693e-005	0. 9397	-7. 6000
FeCO3+	2. 497e-008	0. 002892	0. 9340	-7. 6322
HAsO4--	2. 304e-008	0. 003223	0. 7602	-7. 7566
NaCl (aq)	1. 944e-008	0. 001136	1. 0000	-7. 7113
CaF+	1. 422e-008	0. 0008398	0. 9340	-7. 8768
Pb++	1. 224e-008	0. 002535	0. 7623	-8. 0302
NaCO3-	1. 151e-008	0. 0009550	0. 9340	-7. 9686
MgCl +	1. 055e-008	0. 0006300	0. 9340	-8. 0066
Cd++	7. 736e-009	0. 0008693	0. 7642	-8. 2283
UO2(CO3) 2--	7. 049e-009	0. 002749	0. 7602	-8. 2709
CaCl +	6. 883e-009	0. 0005197	0. 9340	-8. 1919
UO2(CO3) 3----	4. 627e-009	0. 002082	0. 3337	-8. 8114
H2AsO4-	3. 049e-009	0. 0004296	0. 9340	-8. 5455
KCl (aq)	1. 526e-009	0. 0001138	1. 0000	-8. 8163
MnCl +	1. 145e-009	0. 0001035	0. 9340	-8. 9708
Al OH++	1. 072e-009	4. 713e-005	0. 7623	-9. 0878
NaF(aq)	1. 040e-009	4. 367e-005	1. 0000	-8. 9828
NaAl O2(aq)	6. 029e-010	4. 940e-005	1. 0000	-9. 2198
HSO4-	4. 254e-010	4. 128e-005	0. 9340	-9. 4009
BaCO3(aq)	2. 927e-010	5. 775e-005	1. 0000	-9. 5336
HF(aq)	2. 704e-010	5. 408e-006	1. 0000	-9. 5680
UO2CO3(aq)	2. 272e-010	7. 495e-005	1. 0000	-9. 6436
FeCl +	5. 938e-011	5. 419e-006	0. 9340	-10. 2560
PO4---	5. 645e-011	5. 359e-006	0. 5394	-10. 5164
PbCl +	2. 662e-011	6. 459e-006	0. 9340	-10. 6044
NaOH(aq)	2. 409e-011	9. 633e-007	1. 0000	-10. 6181
BaCl +	1. 509e-011	2. 606e-006	0. 9340	-10. 8510
Al +++	9. 562e-012	2. 579e-007	0. 5698	-11. 2637
Cu+	9. 174e-012	5. 828e-007	0. 9340	-11. 0671
Sb(OH) 4-	5. 658e-012	1. 073e-006	0. 9340	-11. 2770
H3PO4(aq)	3. 474e-012	3. 404e-007	1. 0000	-11. 4591
BaF+	3. 049e-012	4. 765e-007	0. 9340	-11. 5455
UO2OH+	2. 878e-012	8. 259e-007	0. 9340	-11. 5705
AsO4---	2. 054e-012	2. 853e-007	0. 5394	-11. 9554
Sb(OH) 2F(aq)	1. 994e-012	3. 484e-007	1. 0000	-11. 7002
CdSO4(aq)	1. 454e-012	3. 030e-007	1. 0000	-11. 8375
MgP2O7--	1. 141e-012	2. 261e-007	0. 7602	-12. 0619
HNO3(aq)	1. 023e-012	6. 444e-008	1. 0000	-11. 9902
Fe+++	9. 416e-013	5. 257e-008	0. 5698	-12. 2704
CaCl 2(aq)	9. 269e-013	1. 028e-007	1. 0000	-12. 0330
Sb(OH) 2+	5. 613e-013	8. 741e-008	0. 9340	-12. 2804
HCl (aq)	5. 505e-013	2. 007e-008	1. 0000	-12. 2592
CdCl 2(aq)	1. 241e-013	2. 274e-008	1. 0000	-12. 9063
FeF++	7. 213e-014	5. 397e-009	0. 7623	-13. 2598
HP2O7---	4. 747e-014	8. 302e-009	0. 5394	-13. 5917
UO2++	3. 780e-014	1. 020e-008	0. 7623	-13. 5404
UO2F+	3. 283e-014	9. 486e-009	0. 9340	-13. 5134
KHSO4(aq)	1. 432e-014	1. 950e-009	1. 0000	-13. 8440
H3AsO4(aq)	1. 080e-014	1. 533e-009	1. 0000	-13. 9664
PbCl 2(aq)	8. 907e-015	2. 476e-009	1. 0000	-14. 0503
FeF2+	7. 644e-015	7. 171e-010	0. 9340	-14. 1463
FeSO4+	7. 251e-015	1. 101e-009	0. 9340	-14. 1693
UO2SO4(aq)	5. 587e-015	2. 045e-009	1. 0000	-14. 2529
H2P2O7--	3. 772e-015	6. 636e-010	0. 7602	-14. 5424
P2O7----	1. 684e-015	2. 928e-010	0. 3337	-15. 2504
UO2F2(aq)	9. 245e-016	2. 847e-010	1. 0000	-15. 0341
HF2-	5. 622e-016	2. 192e-011	0. 9340	-15. 2797
KP2O7---	1. 219e-016	2. 596e-011	0. 5394	-16. 1821
FeCl 2(aq)	2. 800e-017	3. 548e-012	1. 0000	-16. 5528
UO2(SO4) 2--	1. 084e-017	5. 006e-012	0. 7602	-17. 0842

	SpecE8_output_GSS_Decant	Pond	May.txt	
FeCl ++	5. 333e-018	4. 868e-013	0. 7623	-17. 3909
UO2Cl +	3. 927e-018	1. 199e-012	0. 9340	-17. 4356
CdCl 3-	2. 584e-018	5. 651e-013	0. 9340	-17. 6173
(UO2)2(OH)2++	1. 672e-018	9. 598e-013	0. 7623	-17. 8946
UO2F3-	1. 621e-018	5. 300e-013	0. 9340	-17. 8198
PbCl 3-	4. 842e-019	1. 518e-013	0. 9340	-18. 3446
(UO2)3(OH)5+	7. 027e-020	6. 288e-014	0. 9340	-19. 1829
UO2+	5. 830e-020	1. 574e-014	0. 9340	-19. 2639
(UO2)3(CO3)6(6-)	4. 023e-020	4. 706e-014	0. 0846	-20. 4682
H3P2O7-	1. 735e-020	3. 070e-015	0. 9340	-19. 7902
UO2F4--	1. 180e-022	4. 080e-017	0. 7602	-22. 0473
PbCl 4--	5. 730e-023	1. 999e-017	0. 7602	-22. 3609
UO2Cl 2(aq)	1. 670e-023	5. 691e-018	1. 0000	-22. 7774
U(OH)4(aq)	1. 203e-023	3. 682e-018	1. 0000	-22. 9196
HAsO2(aq)	1. 013e-023	1. 093e-018	1. 0000	-22. 9942
As(OH)3(aq)	9. 004e-024	1. 134e-018	1. 0000	-23. 0455
Mn+++	1. 257e-024	6. 904e-020	0. 5458	-24. 1636
FeCl 4--	1. 174e-024	2. 320e-019	0. 7602	-24. 0494
H2AsO3-	1. 265e-025	1. 580e-020	0. 9340	-24. 9275
AsO2-	1. 249e-025	1. 335e-020	0. 9340	-24. 9330
H4P2O7(aq)	1. 054e-026	1. 876e-021	1. 0000	-25. 9770
AsO2OH--	3. 194e-029	3. 957e-024	0. 7602	-28. 6147
NO2-	8. 877e-031	4. 083e-026	0. 9328	-30. 0820
O2(aq)	1. 203e-032	3. 849e-028	1. 0000	-31. 9197
Formate	6. 846e-033	3. 081e-028	0. 9334	-32. 1945
H2(aq)	1. 644e-033	3. 314e-030	1. 0000	-32. 7840
Ca(For) +	6. 336e-035	5. 390e-030	0. 9340	-34. 2278
HN02(aq)	5. 093e-035	2. 394e-030	1. 0000	-34. 2930
Mg(For) +	3. 014e-035	2. 089e-030	0. 9340	-34. 5504
Na(For) (aq)	9. 346e-036	6. 354e-031	1. 0000	-35. 0294
K(For) (aq)	4. 128e-036	3. 472e-031	1. 0000	-35. 3842
Mn(For) +	3. 047e-036	3. 045e-031	0. 9340	-35. 5457
Cl O-	2. 121e-036	1. 091e-031	0. 9340	-35. 7031
Formi c_aci d(aq)	9. 486e-037	4. 365e-032	1. 0000	-36. 0229
Fe(For) +	4. 867e-037	4. 908e-032	0. 9340	-36. 3424
Cu(For) +	2. 869e-037	3. 114e-032	0. 9340	-36. 5719
SO3--	1. 831e-037	1. 465e-032	0. 7623	-36. 8553
Ba(For) +	8. 176e-038	1. 490e-032	0. 9340	-37. 1171
Si F6--	6. 561e-038	9. 319e-033	0. 7602	-37. 3021
NH4+	5. 552e-038	1. 001e-033	0. 9321	-37. 2861
HSO3-	5. 291e-038	4. 288e-033	0. 9340	-37. 3061
H02-	1. 166e-038	3. 848e-034	0. 9340	-37. 9628
Pb(For) +	4. 900e-039	1. 235e-033	0. 9340	-38. 3394
Cd(For) +	3. 461e-039	5. 447e-034	0. 9340	-38. 4904
CO(aq)	1. 522e-039	4. 262e-035	1. 0000	-38. 8175
NH3(aq)	3. 751e-040	6. 386e-036	1. 0000	-39. 4259
Oxal ate	2. 223e-040	1. 956e-035	0. 7602	-39. 7720
MnO4--	3. 464e-041	4. 119e-036	0. 7602	-40. 5795
UOH+++	6. 840e-042	1. 744e-036	0. 5458	-41. 4279
U(CO3)5(6-)	3. 145e-042	1. 692e-036	0. 0846	-42. 5752
UF3+	2. 906e-042	8. 571e-037	0. 9340	-41. 5664
UF2++	1. 849e-042	5. 101e-037	0. 7623	-41. 8511
UF4(aq)	2. 768e-043	8. 688e-038	1. 0000	-42. 5579
H-Oxal ate	7. 468e-044	6. 646e-039	0. 9340	-43. 1565
SO2(aq)	6. 591e-044	4. 222e-039	1. 0000	-43. 1810
UF+++	3. 380e-044	8. 685e-039	0. 5458	-43. 7340
MnO4-	1. 883e-044	2. 239e-039	0. 9334	-43. 7551
USO4++	8. 582e-046	2. 866e-040	0. 7623	-45. 1843
U(SO4)2(aq)	6. 714e-046	2. 887e-040	1. 0000	-45. 1730
HSO5-	1. 777e-046	2. 008e-041	0. 9340	-45. 7800
U++++	2. 767e-048	6. 583e-043	0. 3442	-48. 0212
Oxal i c_aci d(aq)	3. 179e-050	2. 862e-045	1. 0000	-49. 4977
UCl +++	1. 436e-050	3. 927e-045	0. 5458	-50. 1056

	SpecE8_output_GSS_Decant	Pond	May.txt	
Cl O2-	7. 979e-061	5. 380e-056	0. 9340	-60. 1277
S206--	7. 055e-062	1. 129e-056	0. 7602	-61. 2706
S208--	2. 501e-063	4. 804e-058	0. 7602	-62. 7209
U+++	4. 334e-066	1. 031e-060	0. 5458	-65. 6260
Ca(For) 2(aq)	3. 388e-066	4. 407e-061	1. 0000	-65. 4700
Mg(For) 2(aq)	2. 033e-066	2. 324e-061	1. 0000	-65. 6918
Mn(For) 2(aq)	2. 939e-067	4. 259e-062	1. 0000	-66. 5319
Fe(For) 2(aq)	6. 507e-068	9. 489e-063	1. 0000	-67. 1866
Cu(For) 2(aq)	5. 148e-068	7. 904e-063	1. 0000	-67. 2884
Na(For) 2-	4. 702e-068	5. 313e-063	0. 9340	-67. 3573
K(For) 2-	1. 726e-068	2. 228e-063	0. 9340	-67. 7927
Ba(For) 2(aq)	3. 263e-069	7. 418e-064	1. 0000	-68. 4863
Cd(For) 2(aq)	3. 045e-069	6. 162e-064	1. 0000	-68. 5165
Pb(For) 2(aq)	5. 311e-070	1. 578e-064	1. 0000	-69. 2748
Cl O3-	1. 227e-070	1. 023e-065	0. 9334	-69. 9412
Formal dehyde(aq)	6. 015e-071	1. 805e-066	1. 0000	-70. 2208
N3-	1. 603e-072	6. 734e-068	0. 9340	-71. 8247
HN3(aq)	2. 715e-075	1. 168e-070	1. 0000	-74. 5662
Cu(NH3) 2++	4. 576e-078	4. 465e-073	0. 7623	-77. 4574
S205--	5. 156e-080	7. 429e-075	0. 7602	-79. 4067
Cd(NH3) 2++	1. 476e-082	2. 162e-077	0. 7623	-81. 9488
Urea(aq)	4. 857e-083	2. 916e-078	1. 0000	-82. 3137
UO2Cl O3+	1. 213e-083	4. 285e-078	0. 9340	-82. 9459
HCN(aq)	2. 492e-084	6. 732e-080	1. 0000	-83. 6035
Cl O4-	7. 169e-085	7. 127e-080	0. 9334	-84. 1745
CN-	2. 177e-086	5. 663e-082	0. 9328	-85. 6923
Methanol (aq)	4. 973e-088	1. 593e-083	1. 0000	-87. 3034
Gl ycol ate	3. 608e-094	2. 707e-089	0. 9340	-93. 4724
HS-	1. 059e-094	3. 502e-090	0. 9334	-94. 0050
H2S(aq)	4. 268e-095	1. 454e-090	1. 0000	-94. 3697
Ca(Gl yc) +	5. 839e-096	6. 720e-091	0. 9340	-95. 2633
Mg(Gl yc) +	1. 049e-096	1. 041e-091	0. 9340	-96. 0090
Na(Gl yc) (aq)	5. 048e-097	4. 947e-092	1. 0000	-96. 2969
K(Gl yc) (aq)	2. 230e-097	2. 544e-092	1. 0000	-96. 6518
Cu(Gl yc) +	1. 415e-097	1. 960e-092	0. 9340	-96. 8789
Mn(Gl yc) +	1. 236e-097	1. 606e-092	0. 9340	-96. 9378
Fe(Gl yc) +	6. 146e-098	8. 042e-093	0. 9340	-97. 2410
Gl ycol i c_aci d(aq)	6. 077e-098	4. 620e-093	1. 0000	-97. 2163
Ba(Gl yc) +	2. 015e-099	4. 279e-094	0. 9340	-98. 7253
Methane(aq)	9. 588e-100	1. 538e-095	1. 0000	-99. 0183
Pb(Gl yc) +	7. 172e-100	2. 024e-094	0. 9340	-99. 1740
S--	2. 025e-100	6. 493e-096	0. 7642	-99. 8103
Cd(Gl yc) +	2. 011e-100	3. 769e-095	0. 9340	-99. 7262
S204--	6. 981e-101	8. 942e-096	0. 7642	-100. 2729
S203--	6. 579e-101	7. 375e-096	0. 7602	-100. 3009
Acetate	2. 933e-105	1. 731e-100	0. 9346	-104. 5620
MgCH3COO+	8. 779e-108	7. 315e-103	0. 9340	-107. 0862
CaCH3COO+	8. 139e-108	8. 065e-103	0. 9340	-107. 1191
Aceti c_aci d(aq)	4. 015e-108	2. 410e-103	1. 0000	-107. 3963
NaCH3COO(aq)	2. 831e-108	2. 321e-103	1. 0000	-107. 5481
KCH3COO(aq)	8. 869e-109	8. 701e-104	1. 0000	-108. 0521
MnCH3COO+	4. 140e-109	4. 718e-104	0. 9340	-108. 4126
CuCH3COO+	2. 323e-109	2. 847e-104	0. 9340	-108. 6635
FeCH3COO+	5. 316e-110	6. 105e-105	0. 9340	-109. 3041
BaCH3COO+	1. 388e-110	2. 724e-105	0. 9340	-109. 8874
PbCH3COO+	7. 385e-111	1. 966e-105	0. 9340	-110. 1613
CdCH3COO+	1. 881e-111	3. 224e-106	0. 9340	-110. 7552
Mal onate	3. 879e-114	3. 957e-109	0. 7602	-113. 5304
Al CH3COO++	3. 863e-114	3. 322e-109	0. 7623	-113. 5310
Cu(NH3) 3++	1. 821e-114	2. 087e-109	0. 7623	-113. 8576
CuCH3COO(aq)	5. 881e-116	7. 207e-111	1. 0000	-115. 2306
H-Mal onate	3. 710e-116	3. 823e-111	0. 9340	-115. 4602
AsH3(aq)	3. 432e-117	2. 675e-112	1. 0000	-116. 4644

	SpecE8_output_GSS_Decant Pond May.txt			
Maloni c_acid(aq)	6.428e-121	6.687e-116	1.0000	-120.1920
Methanamine(aq)	7.700e-125	2.391e-120	1.0000	-124.1135
Glucine(aq)	1.588e-127	1.192e-122	1.0000	-126.7992
Cu(Gl y)+	9.535e-128	1.312e-122	0.9340	-127.0503
Mg(Gl y)+	1.756e-130	1.727e-125	0.9340	-129.7851
Mn(Gl y)+	1.630e-131	2.102e-126	0.9340	-130.8175
Fe(Gl y)+	8.010e-132	1.040e-126	0.9340	-131.1260
Ca(Gl y)+	2.246e-132	2.562e-127	0.9340	-131.6783
Pb(Gl y)+	1.262e-132	3.548e-127	0.9340	-131.9287
S306--	2.745e-133	5.275e-128	0.7602	-132.6804
Cd(Gl y)+	1.533e-133	2.857e-128	0.9340	-132.8441
Ba(Gl y)+	4.595e-135	9.711e-130	0.9340	-134.3673
Acetaldehyde(aq)	4.036e-141	1.778e-136	1.0000	-140.3940
SCN-	1.231e-141	7.146e-137	0.9334	-140.9398
NH4CH3COO(aq)	2.324e-142	1.790e-137	1.0000	-141.6339
Acetamide(aq)	1.897e-142	1.120e-137	1.0000	-141.7219
UO2SCN+	8.646e-154	2.836e-148	0.9340	-153.0928
Cd(NH3)4++	8.357e-159	1.508e-153	0.7623	-158.1958
Ethanol(aq)	7.051e-163	3.247e-158	1.0000	-162.1518
Ethyne(aq)	1.925e-163	5.012e-159	1.0000	-162.7155
Ethylene(aq)	1.080e-167	3.028e-163	1.0000	-166.9667
Lactate	5.553e-169	4.945e-164	0.9340	-168.2851
Ca(Lac)+	5.282e-171	6.820e-166	0.9340	-170.3068
Mg(Lac)+	1.615e-171	1.831e-166	0.9340	-170.8213
Na(Lac)(aq)	7.942e-172	8.897e-167	1.0000	-171.1001
K(Lac)(aq)	3.509e-172	4.496e-167	1.0000	-171.4548
Mn(Lac)+	1.345e-172	1.936e-167	0.9340	-171.9009
Cu(Lac)+	9.973e-173	1.522e-167	0.9340	-172.0308
Lactic acid(aq)	9.798e-173	8.823e-168	1.0000	-172.0089
Fe(Lac)+	8.374e-173	1.213e-167	0.9340	-172.1067
Ba(Lac)+	1.276e-174	2.888e-169	0.9340	-173.9237
Pb(Lac)+	1.074e-174	3.180e-169	0.9340	-173.9987
Cd(Lac)+	2.094e-175	4.217e-170	0.9340	-174.7087
Ethane(aq)	5.700e-179	1.713e-174	1.0000	-178.2441
Propanoate	2.820e-182	2.060e-177	0.9340	-181.5794
Propanoic acid(aq)	5.183e-185	3.838e-180	1.0000	-184.2854
Ca(Prop)+	4.362e-185	4.934e-180	0.9340	-184.3900
Na(Prop)(aq)	3.977e-185	3.820e-180	1.0000	-184.4004
Mg(Prop)+	2.285e-185	2.224e-180	0.9340	-184.6708
K(Prop)(aq)	1.757e-185	1.970e-180	1.0000	-184.7553
Mn(Prop)+	4.675e-186	5.983e-181	0.9340	-185.3598
USCN+++	3.557e-186	1.053e-180	0.5458	-185.7118
Cu(Prop)+	2.190e-186	2.991e-181	0.9340	-185.6892
Fe(Prop)+	1.214e-186	1.565e-181	0.9340	-185.9454
Pb(Prop)+	6.223e-188	1.744e-182	0.9340	-187.2356
Ca(Gl yc)2(aq)	4.073e-188	7.744e-183	1.0000	-187.3900
S406--	3.256e-188	7.300e-183	0.7602	-187.6064
Ba(Prop)+	1.753e-188	3.688e-183	0.9340	-187.7858
Cd(Prop)+	1.663e-188	3.083e-183	0.9340	-187.8088
Mg(Gl yc)2(aq)	4.510e-189	7.863e-184	1.0000	-188.3458
Cu(Gl yc)2(aq)	4.015e-189	8.574e-184	1.0000	-188.3964
Succinate	1.548e-189	1.796e-184	0.7602	-188.9293
Fe(Gl yc)2(aq)	1.403e-189	2.888e-184	1.0000	-188.8531
Mn(Gl yc)2(aq)	3.515e-190	7.205e-185	1.0000	-189.4541
Na(Gl yc)2-	2.094e-190	3.623e-185	0.9340	-189.7087
K(Gl yc)2-	7.899e-191	1.494e-185	0.9340	-190.1320
Pb(Gl yc)2(aq)	1.491e-191	5.325e-186	1.0000	-190.8265
H-Succinate	1.417e-191	1.659e-186	0.9340	-190.8782
Serine(aq)	8.086e-192	8.495e-187	1.0000	-191.0923
Ba(Gl yc)2(aq)	3.253e-192	9.348e-187	1.0000	-191.4877
Cd(Gl yc)2(aq)	2.292e-192	6.014e-187	1.0000	-191.6398
Succinic acid(aq)	5.917e-195	6.985e-190	1.0000	-194.2279
Ethanamine(aq)	1.200e-201	5.408e-197	1.0000	-200.9208

	SpecE8_output_GSS_Decant Pond May.txt			
Al ani ne(aq)	3. 325e-203	2. 961e-198	1. 0000	-202. 4782
Cu(Al a) +	4. 520e-204	6. 852e-199	0. 9340	-203. 3745
Fe(Al a) +	2. 040e-207	2. 935e-202	0. 9340	-206. 7200
Mn(Al a) +	2. 888e-208	4. 129e-203	0. 9340	-207. 5691
Mg(Al a) +	2. 456e-208	2. 759e-203	0. 9340	-207. 6395
Ca(Al a) +	8. 981e-209	1. 151e-203	0. 9340	-208. 0763
Pb(Al a) +	2. 105e-209	6. 213e-204	0. 9340	-208. 7065
Cd(Al a) +	1. 806e-209	3. 619e-204	0. 9340	-208. 7730
Ca(CH3COO)2(aq)	4. 394e-211	6. 948e-206	1. 0000	-210. 3571
Mg(CH3COO)2(aq)	2. 037e-211	2. 899e-206	1. 0000	-210. 6911
Ba(Al a) +	6. 780e-212	1. 528e-206	0. 9340	-211. 1984
Cu(CH3COO)2(aq)	2. 294e-212	4. 166e-207	1. 0000	-211. 6393
Mn(CH3COO)2(aq)	9. 922e-213	1. 716e-207	1. 0000	-212. 0034
Na(CH3COO)2-	4. 892e-213	6. 900e-208	0. 9340	-212. 3401
Acetone(aq)	3. 667e-213	2. 129e-208	1. 0000	-212. 4357
Fe(CH3COO)2(aq)	3. 398e-213	5. 908e-208	1. 0000	-212. 4688
K(CH3COO)2-	9. 991e-214	1. 570e-208	0. 9340	-213. 0300
Asparti c_aci d(aq)	9. 492e-214	1. 263e-208	1. 0000	-213. 0227
Pb(CH3COO)2(aq)	1. 654e-214	5. 379e-209	1. 0000	-213. 7814
Ba(CH3COO)2(aq)	1. 165e-214	2. 973e-209	1. 0000	-213. 9339
Cd(CH3COO)2(aq)	1. 136e-214	2. 618e-209	1. 0000	-213. 9445
Al (CH3COO)2+	1. 280e-216	1. 856e-211	0. 9340	-215. 9224
Propanal (aq)	1. 441e-217	8. 364e-213	1. 0000	-216. 8415
Cu(CH3COO)2-	2. 516e-220	4. 569e-215	0. 9340	-219. 6289
1-Propyne(aq)	3. 246e-236	1. 300e-231	1. 0000	-235. 4887
S3--	4. 451e-238	4. 281e-233	0. 7602	-237. 4706
1-Propanol (aq)	7. 748e-240	4. 655e-235	1. 0000	-239. 1108
1-Propene(aq)	2. 159e-242	9. 083e-238	1. 0000	-241. 6657
2-Hydroxybutanoa	2. 173e-246	2. 240e-241	0. 9340	-245. 6926
NH4(CH3COO)2-	6. 643e-247	9. 040e-242	0. 9340	-246. 2073
Asparagi ne(aq)	5. 308e-248	7. 011e-243	1. 0000	-247. 2751
Cu(Gl y)2(aq)	1. 330e-249	2. 814e-244	1. 0000	-248. 8762
2-Hydroxybutanoi	3. 392e-250	3. 531e-245	1. 0000	-249. 4695
Propane(aq)	2. 492e-256	1. 099e-251	1. 0000	-255. 6034
Mg(Gl y)2(aq)	9. 046e-257	1. 559e-251	1. 0000	-256. 0436
Di gl yci ne(aq)	5. 573e-257	7. 361e-252	1. 0000	-256. 2539
Mn(Gl y)2(aq)	6. 512e-258	1. 322e-252	1. 0000	-257. 1863
Fe(Gl y)2(aq)	6. 075e-258	1. 239e-252	1. 0000	-257. 2164
Pb(Gl y)2(aq)	1. 045e-258	3. 712e-253	1. 0000	-257. 9809
Cd(Gl y)2(aq)	3. 868e-259	1. 008e-253	1. 0000	-258. 4125
Di ketopi perazi ne	2. 065e-259	2. 355e-254	1. 0000	-258. 6852
Butanoate	8. 722e-260	7. 594e-255	0. 9340	-259. 0890
Ca(Gl y)2(aq)	3. 883e-261	7. 305e-256	1. 0000	-260. 4109
Butanoi c_aci d(aq)	1. 279e-262	1. 127e-257	1. 0000	-261. 8930
Na(But) (aq)	1. 185e-262	1. 304e-257	1. 0000	-261. 9264
Ca(But) +	9. 246e-263	1. 176e-257	0. 9340	-262. 0637
K(But) (aq)	5. 233e-263	6. 602e-258	1. 0000	-262. 2812
Mg(But) +	4. 616e-263	5. 141e-258	0. 9340	-262. 3654
Mn(But) +	1. 234e-263	1. 752e-258	0. 9340	-262. 9384
Ba(Gl y)2(aq)	8. 744e-264	2. 495e-258	1. 0000	-263. 0583
Cu(But) +	5. 362e-264	8. 075e-259	0. 9340	-263. 3003
Fe(But) +	3. 890e-264	5. 559e-259	0. 9340	-263. 4397
Pb(But) +	5. 377e-266	1. 582e-260	0. 9340	-265. 2991
Gl utarate	4. 038e-266	5. 252e-261	0. 7602	-265. 5129
Ba(But) +	3. 625e-266	8. 133e-261	0. 9340	-265. 4703
Cd(But) +	2. 172e-266	4. 331e-261	0. 9340	-265. 6928
H-Gl utarate	2. 145e-268	2. 812e-263	0. 9340	-267. 6982
Threoni ne(aq)	2. 561e-269	3. 050e-264	1. 0000	-268. 5916
Gl utari c_aci d(aq)	1. 131e-271	1. 494e-266	1. 0000	-270. 9466
Ethyl acetate(aq)	8. 605e-272	7. 580e-267	1. 0000	-271. 0652
S506--	5. 017e-273	1. 286e-267	0. 7602	-272. 4186
1-Propanami ne(aq)	4. 451e-278	2. 630e-273	1. 0000	-277. 3516
a-Ami nobutyri c_a	2. 219e-280	2. 288e-275	1. 0000	-279. 6538

	SpecE8_output_GSS_Decant Pond May. txt			
Glutamic acid(aq)	2.843e-289	4.182e-284	1.0000	-288.5462
UO2(SCN)2(aq)	5.421e-295	2.093e-289	1.0000	-294.2659
Butanal(aq)	1.521e-296	1.096e-291	1.0000	-295.8179
S4--	3.691e-307	4.733e-302	0.7602	-300.0000
1-Butyne(aq)	9.427e-314	5.097e-309	1.0000	-300.0000
Cu(CH3COO)3-	2.536e-315	6.101e-310	0.9340	-300.0000
Mn(CH3COO)3-	1.090e-316	2.529e-311	0.9340	-300.0000
Pb(CH3COO)3-	5.705e-317	2.192e-311	0.9340	-300.0000
1-Butanol(aq)	4.502e-318	3.336e-313	1.0000	-300.0000
Cd(CH3COO)3-	1.047e-318	3.031e-313	0.9340	-300.0000
1-Butene(aq)	4.018e-320	2.254e-315	1.0000	-300.0000
2-Hydroxypentano	9.881e-324	1.157e-318	0.9340	-300.0000
m-Toluate	0.0000	0.0000	0.9340	-300.0000
Valine(aq)	0.0000	0.0000	1.0000	-300.0000
Undecanoic acid(0.0000	0.0000	1.0000	-300.0000
Undecanoate	0.0000	0.0000	0.9340	-300.0000
UO2(SCN)3-	0.0000	0.0000	0.9340	-300.0000
H-Suberate	0.0000	0.0000	0.9340	-300.0000
Alanylglycine(aq)	0.0000	0.0000	1.0000	-300.0000
H-Sebacate	0.0000	0.0000	0.9340	-300.0000
H-Pimelate	0.0000	0.0000	0.9340	-300.0000
2-Hexanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexyne(aq)	0.0000	0.0000	1.0000	-300.0000
H-Azelate	0.0000	0.0000	0.9340	-300.0000
Ca(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
H-Adipate	0.0000	0.0000	0.9340	-300.0000
Tyrosine(aq)	0.0000	0.0000	1.0000	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
Toluene(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Suberic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Suberate	0.0000	0.0000	0.7602	-300.0000
Ca(Pent)+	0.0000	0.0000	0.9340	-300.0000
2-Heptanone(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Sb2S4--	0.0000	0.0000	0.7602	-300.0000
Glutamine(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
S5--	0.0000	0.0000	0.7602	-300.0000
2-Butanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexene(aq)	0.0000	0.0000	1.0000	-300.0000
Pimelic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Pimelate	0.0000	0.0000	0.7602	-300.0000
Phenylalanine(aq)	0.0000	0.0000	1.0000	-300.0000
Pentanoic acid(a	0.0000	0.0000	1.0000	-300.0000
Pentanal(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Adipic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Pent)+	0.0000	0.0000	0.9340	-300.0000
Pb(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Pent)+	0.0000	0.0000	0.9340	-300.0000
Pb(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Adipate	0.0000	0.0000	0.7602	-300.0000
1-Hexanol(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9340	-300.0000

	SpecE8_output_GSS_Decant Pond May. txt			
Fe(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanal (aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoate	0.0000	0.0000	0.9340	-300.0000
Nonanal (aq)	0.0000	0.0000	1.0000	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Prop)2-	0.0000	0.0000	0.9340	-300.0000
Na(Pent)2-	0.0000	0.0000	0.9340	-300.0000
Benzoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Pent) (aq)	0.0000	0.0000	1.0000	-300.0000
Na(Lac)2-	0.0000	0.0000	0.9340	-300.0000
Dodecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
Dodecanoate	0.0000	0.0000	0.9340	-300.0000
Benzoate	0.0000	0.0000	0.9364	-300.0000
Pentanoate	0.0000	0.0000	0.9340	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol (aq)	0.0000	0.0000	1.0000	-300.0000
Na(But)2-	0.0000	0.0000	0.9340	-300.0000
Decanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9340	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Decanal (aq)	0.0000	0.0000	1.0000	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxypentano	0.0000	0.0000	1.0000	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)+	0.0000	0.0000	0.9340	-300.0000
Mn(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)+	0.0000	0.0000	0.9340	-300.0000
Ba(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)+	0.0000	0.0000	0.9340	-300.0000
2-Hydroxyoctanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanol (aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoa	0.0000	0.0000	0.9340	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacate	0.0000	0.0000	0.7602	-300.0000
Mg(Pent)+	0.0000	0.0000	0.9340	-300.0000
Mg(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9340	-300.0000
Methi oni ne(aq)	0.0000	0.0000	1.0000	-300.0000
Leucyl gl yci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Leuci ne(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoi	0.0000	0.0000	1.0000	-300.0000

	SpecE8_output_GSS_Decant		Pond May. txt	
U(SCN)2++	0.0000	0.0000	0.7623	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoa	0.0000	0.0000	0.9340	-300.0000
K(Prop)2-	0.0000	0.0000	0.9340	-300.0000
Cd(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
K(Pent)2-	0.0000	0.0000	0.9340	-300.0000
K(Pent)(aq)	0.0000	0.0000	1.0000	-300.0000
K(Lac)2-	0.0000	0.0000	0.9340	-300.0000
Cd(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Pent)+	0.0000	0.0000	0.9340	-300.0000
Ba(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol(aq)	0.0000	0.0000	1.0000	-300.0000
K(But)2-	0.0000	0.0000	0.9340	-300.0000
I sol euci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9340	-300.0000
Azel ate	0.0000	0.0000	0.7602	-300.0000
2-Hydroxyheptano	0.0000	0.0000	0.9340	-300.0000
Hexanal(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoi c_aci d(a	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9340	-300.0000
Heptanal(aq)	0.0000	0.0000	1.0000	-300.0000
Azel ai c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
HSb2S4-	0.0000	0.0000	0.9340	-300.0000
2-Hydroxydecano	0.0000	0.0000	1.0000	-300.0000
1-Octanol(aq)	0.0000	0.0000	1.0000	-300.0000
p-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecano	0.0000	0.0000	0.9340	-300.0000
o-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
o-Phthal ate	0.0000	0.0000	0.7602	-300.0000
Cd(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Pentyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octyl benzene(a	0.0000	0.0000	1.0000	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexyl benzene(a	0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
Cd(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(a	0.0000	0.0000	1.0000	-300.0000
n-Butane(aq)	0.0000	0.0000	1.0000	-300.0000
m-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
n-Propyl benzene(0.0000	0.0000	1.0000	-300.0000
1-Octanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
o-Tol uate	0.0000	0.0000	0.9340	-300.0000
1-Heptanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
p-Tol uate	0.0000	0.0000	0.9340	-300.0000
1-Butanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
(NH4)2Sb2S4(aq)	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states

log Q/K

log Q/K

Nontronite-Ca	24.0610s/sat	U02S04: H2O	-11.4745
Nontronite-Mg	24.0018s/sat	Periclase	-11.5055
Nontronite-K	23.9782s/sat	Na2U207	-11.5658
Nontronite-Na	23.6867s/sat	BaU207	-11.6171
Nontronite-H	23.1641s/sat	Gyrolite	-11.7632

SpecE8_output_GSS_Decant Pond May.txt			
Hemati te	19. 7244s/sat	Przhevalski te	-11. 7686
Magneti te	17. 6186s/sat	MgUO4	-11. 9889
Ferri te-Cu	17. 4432s/sat	MnSO4	-11. 9908
Cronstedti te-7A	16. 8508s/sat	Portlandi te	-12. 1114
Delafossi te	13. 3051s/sat	CdSO4	-12. 2077
Muscovi te	11. 6819s/sat	Thermonatri te	-12. 3064
Clinoptilolite-K	11. 5219s/sat	Monticelli te	-12. 3126
Clinoptilolite-h	11. 3112s/sat	Mg1. 5SO4(OH)	-12. 3271
Clinoptilolite-C	11. 2223s/sat	UO2. 6667	-12. 4774
Clinoptilolite-h	11. 1664s/sat	UO2Cl OH: 2H2O	-12. 5632
Stilbite	9. 9090s/sat	UP05	-12. 5666
Goethi te	9. 4029s/sat	Basseti te	-12. 5815
Epidote-ord	8. 9314s/sat	Na2CO3	-12. 6407
Epidote	8. 9302s/sat	PbF2	-12. 8823
Ferri te-Ca	8. 6944s/sat	FeSO4	-13. 0359
Ferri te-Mg	8. 6061s/sat	Pb4SO7	-13. 1486
Andradi te	8. 4866s/sat	MgSO4	-13. 2479
Clinoptilolite-N	8. 0805s/sat	Autunite-H	-13. 2709
Clinoptilolite-h	8. 0785s/sat	Urani ni te	-13. 5459
Illite	8. 0043s/sat	Chalcocyanite	-13. 6864
Paragoni te	7. 7237s/sat	Coffinite	-13. 8779
Beidellite-Ca	7. 6983s/sat	Mordenite-dehy	-14. 0939
Beidellite-Mg	7. 6389s/sat	FeF2	-14. 2240
Beidellite-K	7. 6155s/sat	BaCl 2: 2H2O	-14. 3087
Fluorapatite	7. 5759s/sat	CdCl 2: H2O	-14. 5815
Beidellite-Na	7. 3238s/sat	BaCl 2: H2O	-15. 0521
Kaolinite	7. 1973s/sat	CdCl 2	-15. 7091
Mesolite	7. 0041s/sat	Tremolite	-15. 8439
Beidellite-H	6. 8019s/sat	NaUO3	-15. 8577
Pyrophyllite	6. 4386s/sat	UO2SO4: 3. 5H2O	-15. 9348
Montmor-Mg	6. 2567s/sat	MnCl 2: 4H2O	-15. 9710
Montmor-Ca	6. 2407s/sat	UO2SO4: 2. 5H2O	-16. 0021
Montmor-K	6. 2332s/sat	UO2SO4: 3H2O	-16. 0786
Jarosite	5. 9532s/sat	Ba(OH) 2: 8H2O	-16. 1594
Montmor-Na	5. 9366s/sat	UO2F2: 3H2O	-16. 3316
Scolecite	5. 4960s/sat	CaAl 2O4	-16. 3873
Maximum_Microclite	4. 5705s/sat	Ca3(AsO4) 2	-16. 4735
K-Feldspar	4. 5684s/sat	BaCl 2	-16. 6139
Margarite	4. 1511s/sat	Al um-K	-16. 6297
Laumontite	4. 0984s/sat	UO2F2	-16. 6676
Fe(OH) 3	4. 0677s/sat	Natrosilite	-16. 8718
Smectite-low-Fe-	4. 0091s/sat	Ferri te-Dicalcium	-16. 9597
Celadonite	3. 8127s/sat	MgOHCl	-17. 3554
Diaspore	3. 3185s/sat	MnCl 2: 2H2O	-17. 4369
Sanidine_high	3. 2644s/sat	Na2Si O3	-17. 4718
Smectite-high-Fe	3. 0260s/sat	CdF2	-17. 5383
Boehmite	2. 8876s/sat	CaAl 4O7	-17. 6780
Gibbsite	2. 8007s/sat	Ba2Si 3O8	-17. 8141
Lawsonite	2. 5870s/sat	Ni ngoyite	-18. 1233
Corundum	2. 3175s/sat	Gehlenite	-18. 3635
Mordenite	2. 0194s/sat	CuCl 2	-18. 6291
Witherite	1. 6268s/sat	UO2. 3333(beta)	-18. 6336
Albite_low	1. 5496s/sat	MnCl 2: H2O	-19. 1330
Albite	1. 5495s/sat	Hydromagnesite	-19. 1736
Kyanite	1. 3831s/sat	Hillebrandite	-19. 2131
Chamosite-7A	1. 1198s/sat	MgCl 2: 4H2O	-19. 6637
Andalusite	1. 0714s/sat	Dicalcium_silica	-19. 7022
Alcime	0. 9446s/sat	UO2SO4	-19. 8883
Strengite	0. 7263s/sat	Na2UO4(al pha)	-20. 6659
Natrolite	0. 7143s/sat	Akermanite	-21. 0592
Dawsonite	0. 6850s/sat	Larnite	-21. 0869
Sillimani te	0. 6766s/sat	UOFOH: . 5H2O	-21. 4934
Amesite-14A	0. 5198s/sat	UO2(NO3) 2: 6H2O	-21. 5385

SpecE8_output_GSS_Decant Pond May.txt			
Quartz	0. 5111s/sat	UOF0H	-22. 0073
Tenori te	0. 4417s/sat	Scacchi te	-22. 5648
Bari te	0. 3740s/sat	Li me	-22. 7566
Mal achi te	0. 3128s/sat	U02Cl	-22. 9236
Tri dymi te	0. 3109s/sat	Anthophyll i te	-22. 9670
Chal cedony	0. 2250s/sat	U02(N03)2: 3H2O	-23. 2130
Kal si l i te	0. 1931s/sat	Lawrenci te	-23. 9555
Al bi te_hi gh	0. 1248s/sat	Hydrophi l i te	-24. 0314
Cerussi te	0. 0690s/sat	(U02)3(P04)2	-24. 0352
Hercyni te	0. 0040s/sat	U02(P03)2	-24. 2117
Cri stobal i te(al p	-0. 0756	UOF2: H2O	-24. 4568
I ce	-0. 0801	(U02)2P207	-24. 5886
Daphni te-14A	-0. 1214	U02(N03)2: 2H2O	-24. 5913
Anni te	-0. 1563	K2U04	-25. 1056
Coesi te	-0. 3408	UOF2	-25. 1163
Rhodochrosi te	-0. 3971	Pargasi te	-25. 5128
Al uni te	-0. 4092	MgCl 2: 2H2O	-25. 5346
Cri stobal i te(bet	-0. 5645	U(HP04)2: 4H2O	-26. 2050
Jadei te	-0. 5830	KAl (S04)2	-26. 3129
Wai raki te	-0. 8187	UP207	-26. 5844
Cal ci te	-0. 8567	Cl i nopti l ol i te-d	-26. 5887
Prehni te	-0. 8921	Pb	-26. 7311
Si O2(am)	-0. 9059	Cl i nopti l ol i te-d	-26. 8482
Brochanti te	-0. 9421	Ranki ni te	-27. 3705
Aragoni te	-1. 0019	U02Cl 2: 3H2O	-27. 5716
Dol omi te-ord	-1. 0227	U02(N03)2: H2O	-28. 4507
Dol omi te	-1. 0228	MgCl 2: H2O	-29. 1366
Cl i nozoi si te	-1. 0796	Ettri ngi te	-29. 5737
Zoi si te	-1. 1280	Cl i nopti l ol i te-d	-30. 0045
Di optase	-1. 4290	Tobermori te-14A	-30. 5170
Si deri te	-1. 4705	U02Cl 2: H2O	-30. 5511
Monohydrocal ci te	-1. 6548	U02(N03)2	-32. 1551
Magnesi te	-1. 8935	Tobermori te-11A	-32. 7957
Phl ogopi te	-2. 0781	Ba2Si O4	-32. 8349
Anorthi te	-2. 3337	KMgCl 3: 2H2O	-33. 7781
Al amosi te	-2. 4805	Merwi ni te	-33. 9304
Gypsum	-2. 6257	Foshagi te	-34. 0648
Nephel i ne	-2. 6762	U02S03	-34. 3220
Dol omi te-di s	-2. 6852	UF4: 2. 5H2O	-34. 4447
Saponi te-Ca	-2. 6969	U02Cl 2	-34. 6812
Ferrosi l i te	-2. 7113	As205	-34. 7713
Hydroxyl apati te	-2. 7175	Chl oromagnesi te	-35. 2559
Saponi te-Mg	-2. 7544	Afwi l l i te	-35. 6535
Saponi te-K	-2. 7797	Tobermori te-9A	-36. 6671
Ri pi dol i te-14A	-2. 7966	Al 2(S04)3: 6H2O	-36. 9451
Mi nnesotai te	-2. 9269	Cd	-37. 0931
Anhydri te	-2. 9562	BaSi F6	-37. 4342
Berl i ni te	-2. 9702	Fe	-37. 5511
Saponi te-Na	-3. 0713	(U02)2As207	-38. 0505
Sb203	-3. 2858	(U02)3(As04)2	-38. 4038
Fl uori te	-3. 3627	U02(As03)2	-38. 5966
Whi tlocki te	-3. 4211	UF4	-39. 0200
Saponi te-H	-3. 5935	Mol ysi te	-39. 1837
Bassani te	-3. 6061	Fe2(S04)3	-40. 8566
Bi xbyi te	-3. 6161	BaO	-41. 4826
Angl esi te	-3. 6356	KMgCl 3	-41. 5896
Daphni te-7A	-3. 6391	U(S04)2: 8H2O	-43. 0272
U03: 2H2O	-3. 6685	Ba2U207	-44. 2948
Schoepi te	-3. 6689	U(S04)2: 4H2O	-44. 3756
CaS04: 0. 5H2O(bet	-3. 7944	Arsenol i te	-44. 3856
U02(OH)2(beta)	-3. 8386	Cl audeti te	-44. 4636
Atacami te	-3. 8555	U(S04)2	-44. 6522
CaU04	-3. 8647	Hatruri te	-46. 3403

SpecE8_output_GSS_Decant Pond May.txt			
Schoepi te-dehy(. UO3: . 9H2O(al pha)	-3. 8987	C	-46. 4605
Tal c	-3. 9074	UOCl 2	-47. 5982
Schoepi te-dehy(. UO3: . 9H2O(al pha)	-3. 9662	U(CO3)2	-47. 7078
Schoepi te-dehy(. UO3: . 9H2O(al pha)	-3. 9932	Xonotli te	-49. 4047
Schoepi te-dehy(1 Rhodoni te	-4. 0106	UCI F3	-50. 2154
Rhodoni te	-4. 0964	U5O12Cl	-51. 2334
Lanarki te	-4. 1830	As	-53. 9343
Pyrol usi te	-4. 3976	UOF4	-54. 7846
Paral auri oni te	-4. 5288	Al 2(SO4)3	-55. 7908
Greenal i te	-4. 5881	Na3UO4	-56. 6928
Wusti te	-4. 6295	Na	-58. 3728
Enstati te	-4. 6534	(UO2)2Cl 3	-58. 8125
Sell ai te	-4. 6641	Na4Si O4	-58. 8631
Cupri te	-4. 7682	UOCl 3	-59. 9838
CdSi O3	-4. 8478	Ca4Al 2Fe2O10	-60. 5776
UO2CO3	-4. 8596	UF3	-61. 1693
Rutherfordi ne	-4. 8721	Na2O	-61. 4427
Nesquehoni te	-5. 0212	Mn	-61. 7406
Azuri te	-5. 0803	UF5(beta)	-61. 9597
Cl i nochl ore-14A	-5. 0954	Pb2Cl 5NH4	-62. 2028
Schoepi te-dehy(. UO3: . 9H2O(al pha)	-5. 1883	UF5(al pha)	-62. 3700
Hausmanni te	-5. 4276	K	-62. 3809
FeO	-5. 4556	Ca3Al 2O6	-62. 9272
Nahcol i te	-5. 4894	UCI 2F2	-63. 7250
Ni ter	-5. 5785	UOCl	-65. 8346
Fe(OH)2	-5. 7284	Covel l i te	-69. 0549
Schoepi te-dehy(. UO3: . 9H2O(al pha)	-5. 7414	S	-69. 9185
Anal ci me-dehy(. UO3: . 9H2O(al pha)	-5. 8091	Ba3UO6	-71. 3863
Cordi eri te_hydr UO2HPO4: 4H2O	-5. 9918	U2O3F6	-71. 6004
UO2HPO4: 4H2O	-6. 0036	Chal coci te	-71. 8049
Bruci te	-6. 0971	UCI 3F	-77. 0497
Li tharge	-6. 1024	CdS	-78. 0417
Fayal i te	-6. 1887	Gal ena	-78. 7573
Mn(OH)2(am)	-6. 1908	K2O	-79. 3570
NaFeO2	-6. 2404	U(SO3)2	-84. 7232
Massi cot	-6. 3009	CdCl 2(NH3)2	-85. 6482
Ri pi dol i te-7A	-6. 3055	Na6Si 2O7	-85. 6599
PbCO3. PbO	-6. 3161	PbSO4(NH3)2	-88. 1848
Okeni te	-6. 5149	UCI 4	-88. 2143
Wol l astoni te	-6. 6168	As4O6(mono)	-88. 4418
UO3(gamma)	-6. 8306	As4O6(cubi)	-88. 6466
Pseudowol l astoni	-6. 8845	Troi l i te	-88. 7415
Di opsi de	-6. 9016	Pyrrhoti te	-88. 8459
Chrysoti l e	-7. 0139	Al abandi te	-91. 3540
Spi nel	-7. 1008	UCI 3	-91. 7745
Mel anteri te	-7. 2350	UF6	-93. 9116
Chal canthi te	-7. 4122	U3O5F8	-95. 6270
UO3(beta)	-7. 4677	U2F9	-97. 6707
Mi rabi l i te	-7. 5725	Anti gori te	-97. 8227
Cd(OH)2	-7. 6130	Mg	-102. 4786
Hedenbergi te	-7. 6327	U2O2Cl 5	-108. 8785
PbFCI	-7. 7071	UCI 5	-109. 2428
UO3(al pha)	-7. 8241	BaS	-109. 8190
Syl vi te	-7. 8489	UN1. 73(al pha)	-110. 3087
Nantoki te	-7. 8824	UN1. 59(al pha)	-113. 1774
Arcani te	-7. 9461	Ca	-120. 1222
FeF3	-8. 1732	P	-121. 1463
Hunti te	-8. 2235	Al	-123. 6654
Tephroi te	-8. 3048	Ba	-124. 2316
Pb2Si O4	-8. 3227	UN	-126. 8316
Hal i te	-8. 4329	UCI 6	-129. 0533
Cd(OH)Cl	-8. 4438	Si	-129. 1162
Sanborni te	-8. 4613	Pb(N3)2(orth)	-142. 1078

SpecE8_output_GSS_Decant Pond May.txt			
Ca-Al_Pyroxene	-8.4900	Pb(N3)2(mono)	-142.5660
U02F0H: 2H2O	-8.5189	Pyri te	-147.6645
Cl i nochl ore-7A	-8.5997	Chal copyri te	-151.5676
Sal eei te	-8.6297	CdCl 2(NH3)4	-166.3220
Pb3S06	-8.6313	PbS04(NH3)4	-170.5482
Cordi eri te_anhyd	-8.6374	U4F17	-174.7960
Arti ni te	-8.8169	U	-175.6177
Cu	-8.8863	UAs	-185.6128
BaU04	-8.9149	US	-186.2053
U02F0H: H2O	-8.9823	UC	-203.8066
Manganosi te	-9.0285	UH3(beta)	-206.2714
Thenardi te	-9.1122	US1.9	-214.2909
Monteponi te	-9.1230	US2	-219.4698
Phosgeni te	-9.1512	Mayeni te	-231.4603
Ni trobari te	-9.2798	UAs2	-236.7694
Grossul ar	-9.3527	CdCl 2(NH3)6	-247.0924
Al F3	-9.3831	UP	-247.5751
U02F0H	-9.5152	UC1.94(al pha)	-249.6152
Sepi ol i te	-9.7066	Sti bni te	-261.4143
(U02)3(P04)2: 4H2	-10.0708	US3	-286.0477
U02.25	-10.1867	Orpi ment	-286.4530
Torberni te	-10.1937	Pb(SCN)2	-289.7630
CdS04: 2.667H2O	-10.2062	Borni te	-294.8740
Natron	-10.2284	UP2	-363.3746
U02.25(beta)	-10.2558	U2S3	-398.3827
Uranoci rci te	-10.2634	o-Phthal i c_aci d	-408.1520
CdS04: H2O	-10.4703	U2C3	-455.6804
Na2C03: 7H2O	-10.8072	U3As4	-608.5652
Mg1.25S04(OH)0.5	-10.8753	U3S5	-613.0206
Cotunni te	-10.8845	U3P4	-858.5242
Forsteri te	-11.0962		

Gases	fugaci ty	I og fug.
N2(g)	0.2577	-0.589
H2O(g)	0.009706	-2.013
CO2(g)	0.001832	-2.737
NO2(g)	2.591e-012	-11.587
HF(g)	4.813e-015	-14.318
HCl(g)	2.506e-019	-18.601
O2(g)	6.928e-030	-29.159
H2(g)	1.855e-030	-29.732
NO(g)	1.530e-031	-30.815
CO(g)	1.140e-036	-35.943
Si F4(g)	4.390e-040	-39.358
NH3(g)	2.707e-042	-41.568
Cl 2(g)	2.312e-043	-42.636
S02(g)	2.374e-044	-43.625
Cd(g)	2.006e-052	-51.698
Pb(g)	9.291e-058	-57.032
U02F2(g)	3.398e-062	-61.469
Cu(g)	5.202e-065	-64.284
Na(g)	1.297e-073	-72.887
U02Cl 2(g)	7.440e-074	-73.128
K(g)	1.480e-074	-73.830
U03(g)	6.799e-075	-74.168
U0F4(g)	8.824e-077	-76.054
UF5(g)	6.555e-084	-83.183
UF4(g)	1.455e-086	-85.837
H2S(g)	2.677e-094	-93.572
UF6(g)	6.135e-096	-95.212
CH4(g)	4.794e-097	-96.319
F2(g)	2.212e-101	-100.655

SpecE8_output_GSS_Decant Pond May.txt

UCI 4(g)	2.677e-115	-114.572
UO2(g)	1.285e-116	-115.891
Mg(g)	2.534e-124	-123.596
UCI 5(g)	1.070e-128	-127.971
UF3(g)	7.441e-133	-132.128
UCI 6(g)	1.961e-136	-135.708
U2F10(g)	4.184e-140	-139.378
UCI 3(g)	2.806e-141	-140.552
Ca(g)	8.886e-148	-147.051
S2(g)	1.079e-155	-154.967
C2H4(g)	1.495e-165	-164.825
C(g)	1.109e-171	-170.955
UF2(g)	9.334e-176	-175.030
Al(g)	2.991e-178	-177.524
UCI 2(g)	2.018e-185	-184.695
UO(g)	1.516e-191	-190.819
Si(g)	3.238e-205	-204.490
UF(g)	8.035e-212	-211.095
U2Cl8(g)	2.195e-218	-217.659
UCI(g)	5.286e-227	-226.277
U2Cl10(g)	2.063e-228	-227.685
U(g)	4.774e-267	-266.321

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg

Al+++	3.98e-006	3.98e-006	0.107			
Ba++	6.90e-007	6.90e-007	0.0948			
Ca++	0.000415	0.000415	16.6			
Cd++	7.74e-009	7.74e-009	0.000870			
Cl-	0.000110	0.000110	3.91			
Cu++	4.31e-007	4.31e-007	0.0274			
F-	1.03e-005	1.03e-005	0.195			
Fe++	1.05e-006	1.05e-006	0.0586			
H+	8.45e-005	8.45e-005	0.0851			
H2AsO4-	2.61e-008	2.61e-008	0.00368			
H2O	55.5	55.5	1.00e+006			
HC03-	0.00159	0.00159	96.7			
HP04--	4.23e-006	4.23e-006	0.406			
K+	0.000675	0.000675	26.4			
Mg++	0.000161	0.000161	3.91			
Mn++	9.85e-006	9.85e-006	0.541			
NH3(aq)	0.000443	0.000443	7.55			
NO3-	0.00123	0.00123	76.1			
Na+	0.00132	0.00132	30.3			
O2(aq)	0.000333	0.000333	10.6			
Pb++	1.23e-008	1.23e-008	0.00254			
S2--	6.10e-007	6.10e-007	0.0391			
S04--	0.000336	0.000336	32.2			
Sb(OH)3(aq)	3.69e-007	3.69e-007	0.0638			
SiO2(aq)	0.000138	0.000138	8.30			
UO2++	1.19e-008	1.19e-008	0.00321			

Elemental composition	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg

Aluminum	3.984e-006	3.984e-006	0.1075		
Antimony	3.692e-007	3.692e-007	0.04494		
Arsenic	2.609e-008	2.609e-008	0.001954		
Barium	6.903e-007	6.903e-007	0.09477		
Cadmium	7.737e-009	7.737e-009	0.0008695		
Calcium	0.0004145	0.0004145	16.61		
Carbon	0.001586	0.001586	19.04		

	SpecE8_output_GSS_Decant Pond May. txt		
Chl ori ne	0.0001103	0.0001103	3.908
Copper	4.306e-007	4.306e-007	0.02736
Fl uori ne	1.029e-005	1.029e-005	0.1954
Hydrogen	111.0	111.0	1.119e+005
Iron	1.050e-006	1.050e-006	0.05862
Lead	1.226e-008	1.226e-008	0.002540
Magnesi um	0.0001608	0.0001608	3.908
Manganese	9.855e-006	9.855e-006	0.5412
Ni trogen	0.001671	0.001671	23.40
Oxygen	55.52	55.52	8.880e+005
Phosphorus	4.228e-006	4.228e-006	0.1309
Potassi um	0.0006749	0.0006749	26.38
Si l i con	0.0001383	0.0001383	3.882
Sodi um	0.001318	0.001318	30.29
Sul fur	0.0003369	0.0003369	10.80
Urani um	1.191e-008	1.191e-008	0.002833

SpecE8_output_GSS_Mine Water 9 East April 2003.txt

Temperature = 25.0 C Pressure = 1.013 bars
 pH = 7.200 log fO2 = -30.382
 Eh = 0.3538 volts pe = 5.9809
 Ionic strength = 0.002470
 Charge imbalance = 0.000089 eq/kg (5.785% error)
 Activity of water = 1.000000
 Solvent mass = 1.000000 kg
 Solution mass = 1.000136 kg
 Solution density = 1.013 g/cm3
 Chlorinity = 0.000000 molal
 Dissolved solids = 136 mg/kg sol'n
 Elect. conductivity = 151.53 uS/cm (or umho/cm)
 Hardness = 81.16 mg/kg sol'n as CaCO3
 carbonate = 60.86 mg/kg sol'n as CaCO3
 non-carbonate = 20.30 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 60.86 mg/kg sol'n as CaCO3
 Water type = Ca-HCO3

Nernst redox couples

Eh (volts)

pe

e- + .25*O2(aq) + H+ = .5*H2O	0.3538	5.9809
8*e- + 9*H+ + NO3- = 3*H2O + NH3(aq)	0.5662	9.5714

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HC03-	0.001205	73.50	0.9468	-2.9428
Ca++	0.0006018	24.11	0.8082	-3.3131
Mg++	0.0001886	4.584	0.8132	-3.8142
CO2(aq)	0.0001592	7.004	1.0000	-3.7981
SO4--	0.0001330	12.77	0.8029	-3.9716
SiO2(aq)	7.546e-005	4.533	1.0000	-4.1223
NO3-	2.742e-005	1.700	0.9460	-4.5861
CaSO4(aq)	6.706e-006	0.9128	1.0000	-5.1736
CaHCO3+	6.525e-006	0.6595	0.9468	-5.2092
MgSO4(aq)	4.226e-006	0.5086	1.0000	-5.3740
N2(aq)	2.763e-006	0.07738	1.0000	-5.5587
MgHCO3+	2.007e-006	0.1712	0.9468	-5.7213
CO3--	1.054e-006	0.06326	0.8043	-6.0716
CuCO3(aq)	9.471e-007	0.1170	1.0000	-6.0236
CaCO3(aq)	8.758e-007	0.08764	1.0000	-6.0576
OH-	1.694e-007	0.002880	0.9464	-6.7951
Cu++	1.532e-007	0.009733	0.8082	-6.9073
HSiO3-	1.409e-007	0.01086	0.9468	-6.8748
MgCO3(aq)	1.239e-007	0.01045	1.0000	-6.9069
CuOH+	1.069e-007	0.008610	0.9468	-6.9948
Mn++	1.027e-007	0.005642	0.8082	-7.0809
CaNO3+	6.678e-008	0.006816	0.9468	-7.1991
H+	6.638e-008	6.690e-005	0.9505	-7.2000
HAsO4--	5.043e-008	0.007055	0.8029	-7.3927
H2AsO4-	1.547e-008	0.002179	0.9468	-7.8344
CuSO4(aq)	3.028e-009	0.0004832	1.0000	-8.5188
MnCO3(aq)	2.331e-009	0.0002679	1.0000	-8.6325
MnSO4(aq)	1.997e-009	0.0003016	1.0000	-8.6996
Cu(CO3)2--	1.685e-009	0.0003093	0.8029	-8.8686
CaOH+	1.150e-009	6.564e-005	0.9468	-8.9631
MnHCO3+	7.615e-010	8.828e-005	0.9468	-9.1421
HSO4-	6.781e-010	6.581e-005	0.9468	-9.1925
Cu+	7.349e-011	4.669e-006	0.9468	-10.1575
MnOH+	3.572e-011	2.569e-006	0.9468	-10.4709

SpecE8_output_GSS_Mine Water 9 East April 2003. txt

MnNO3+	3. 604e-012	4. 214e-007	0. 9468	-11. 4669
AsO4---	2. 629e-012	3. 652e-007	0. 6101	-11. 7948
CuCO3(OH)2--	2. 519e-012	3. 969e-007	0. 8029	-11. 6941
H2SiO4--	2. 588e-013	2. 435e-008	0. 8029	-12. 6823
H3AsO4(aq)	1. 640e-013	2. 328e-008	1. 0000	-12. 7852
HNO3(aq)	8. 155e-014	5. 138e-009	1. 0000	-13. 0886
Mn(OH)2(aq)	1. 316e-015	1. 170e-010	1. 0000	-14. 8809
H6(H2SiO4)4--	2. 323e-016	8. 884e-011	0. 8029	-15. 7292
Mn(NO3)2(aq)	2. 223e-016	3. 978e-011	1. 0000	-15. 6530
Mn2(OH)3+	3. 648e-017	5. 868e-012	0. 9468	-16. 4617
Mn2OH+++	4. 894e-018	6. 208e-013	0. 6147	-17. 5217
CuO2--	3. 454e-018	3. 300e-013	0. 8029	-17. 5570
H2SO4(aq)	4. 051e-020	3. 972e-015	1. 0000	-19. 3925
HAsO2(aq)	2. 240e-020	2. 417e-015	1. 0000	-19. 6498
Mn(OH)3-	2. 066e-020	2. 189e-015	0. 9468	-19. 7087
As(OH)3(aq)	1. 887e-020	2. 376e-015	1. 0000	-19. 7242
AsO2-	2. 022e-022	2. 162e-017	0. 9468	-21. 7179
H2AsO3-	1. 971e-022	2. 462e-017	0. 9468	-21. 7290
H4(H2SiO4)4----	5. 654e-024	2. 151e-018	0. 4153	-23. 6292
AsO2OH--	3. 542e-026	4. 388e-021	0. 8029	-25. 5461
Mg4(OH)4++++	1. 467e-026	2. 424e-021	0. 4236	-26. 2066
Mn+++	3. 383e-027	1. 858e-022	0. 6147	-26. 6821
Mn(OH)4--	3. 269e-027	4. 020e-022	0. 8029	-26. 5809
Formate	8. 240e-030	3. 709e-025	0. 9464	-29. 1080
H2(aq)	3. 417e-030	6. 887e-027	1. 0000	-29. 4664
Ca(For)+	1. 078e-031	9. 175e-027	0. 9468	-30. 9910
Mg(For)+	3. 402e-032	2. 358e-027	0. 9468	-31. 4921
NO2-	1. 316e-032	6. 054e-028	0. 9460	-31. 9048
Formic_acid(aq)	2. 786e-033	1. 282e-028	1. 0000	-32. 5550
O2(aq)	5. 243e-034	1. 677e-029	1. 0000	-33. 2804
SO3--	1. 377e-034	1. 102e-029	0. 8043	-33. 9558
HSO3-	1. 184e-034	9. 598e-030	0. 9468	-33. 9504
NH4+	1. 037e-034	1. 871e-030	0. 9455	-34. 0084
Cu(For)+	9. 736e-035	1. 057e-029	0. 9468	-34. 0354
Mn(For)+	3. 425e-035	3. 423e-030	0. 9468	-34. 4891
CO(aq)	6. 269e-036	1. 756e-031	1. 0000	-35. 2028
HNO2(aq)	1. 306e-036	6. 137e-032	1. 0000	-35. 8842
NH3(aq)	8. 924e-037	1. 520e-032	1. 0000	-36. 0494
Oxalate	2. 428e-037	2. 137e-032	0. 8029	-36. 7100
CuNO2+	1. 705e-037	1. 867e-032	0. 9468	-36. 7921
HO2-	1. 913e-038	6. 312e-034	0. 9468	-37. 7421
CuNH3++	1. 506e-039	1. 214e-034	0. 8043	-38. 9167
H2SO3(aq)	7. 202e-040	5. 910e-035	1. 0000	-39. 1426
SO2(aq)	5. 127e-040	3. 284e-035	1. 0000	-39. 2902
H-Oxalate	2. 346e-040	2. 088e-035	0. 9468	-39. 6534
MnO4--	1. 316e-044	1. 566e-039	0. 8029	-43. 9759
HSO5-	8. 423e-046	9. 522e-041	0. 9468	-45. 0983
Oxalic_acid(aq)	2. 611e-046	2. 351e-041	1. 0000	-45. 5831
NH4SO4-	5. 530e-047	6. 309e-042	0. 9468	-46. 2810
MnO4-	4. 428e-048	5. 266e-043	0. 9464	-47. 3777
S2O6--	9. 386e-058	1. 503e-052	0. 8029	-57. 1229
Ca(For)2(aq)	5. 903e-060	7. 680e-055	1. 0000	-59. 2289
Mg(For)2(aq)	1. 862e-060	2. 129e-055	1. 0000	-59. 7300
S2O8--	4. 093e-062	7. 862e-057	0. 8029	-61. 4833
Cu(For)2(aq)	1. 502e-062	2. 307e-057	1. 0000	-61. 8233
Mn(For)2(aq)	3. 111e-063	4. 510e-058	1. 0000	-62. 5071
Formaldehyde(aq)	3. 674e-064	1. 103e-059	1. 0000	-63. 4349
Cu(NO2)2(aq)	2. 056e-068	3. 199e-063	1. 0000	-67. 6869
N3-	3. 038e-071	1. 276e-066	0. 9468	-70. 5412
Cu(NH3)2++	3. 465e-072	3. 381e-067	0. 8043	-71. 5550
S2O5--	2. 383e-073	3. 435e-068	0. 8029	-72. 7181
HN3(aq)	9. 096e-074	3. 914e-069	1. 0000	-73. 0411
Urea(aq)	2. 455e-076	1. 474e-071	1. 0000	-75. 6099

SpecE8_output_GSS_Mine Water 9 East April 2003.txt

HCN(aq)	4. 299e-077	1. 162e-072	1. 0000	-76. 3667
Methanol (aq)	7. 468e-079	2. 392e-074	1. 0000	-78. 1268
CN-	4. 184e-079	1. 088e-074	0. 9460	-78. 4026
HS-	1. 248e-083	4. 128e-079	0. 9464	-82. 9276
H2S(aq)	7. 246e-084	2. 469e-079	1. 0000	-83. 1399
Gl ycol ate	4. 726e-085	3. 546e-080	0. 9468	-84. 3493
Ca(Gl yc)+	1. 027e-086	1. 182e-081	0. 9468	-86. 0124
Mg(Gl yc)+	1. 549e-087	1. 538e-082	0. 9468	-86. 8338
Gl ycol i c_aci d(aq)	1. 924e-088	1. 463e-083	1. 0000	-87. 7157
Methane(aq)	1. 892e-088	3. 034e-084	1. 0000	-87. 7232
Cu(Gl yc)+	4. 544e-089	6. 296e-084	0. 9468	-88. 3664
S--	2. 699e-089	8. 653e-085	0. 8056	-88. 6627
S2O3--	5. 938e-090	6. 657e-085	0. 8029	-89. 3217
Mn(Gl yc)+	1. 498e-090	1. 946e-085	0. 9468	-89. 8484
S2O4--	1. 641e-091	2. 102e-086	0. 8056	-90. 8788
Acetate	5. 616e-094	3. 315e-089	0. 9472	-93. 2742
HS2O3-	3. 281e-096	3. 711e-091	0. 9468	-95. 5078
CaCH3COO+	2. 330e-096	2. 310e-091	0. 9468	-95. 6563
Aceti c_aci d(aq)	1. 919e-096	1. 152e-091	1. 0000	-95. 7170
MgCH3COO+	1. 639e-096	1. 366e-091	0. 9468	-95. 8092
CuCH3COO+	1. 187e-098	1. 454e-093	0. 9468	-97. 9494
MnCH3COO+	7. 682e-100	8. 755e-095	0. 9468	-99. 1383
Mal onate	8. 666e-103	8. 842e-098	0. 8029	-102. 1575
CuCH3COO(aq)	7. 908e-104	9. 693e-099	1. 0000	-103. 1019
H-Mal onate	2. 300e-104	2. 370e-099	0. 9468	-103. 6620
ASH3(aq)	2. 167e-104	1. 689e-099	1. 0000	-103. 6642
Cu(NH3)3++	2. 046e-105	2. 345e-100	0. 8043	-104. 7837
Mal oni c_aci d(aq)	9. 756e-109	1. 015e-103	1. 0000	-108. 0107
Methanami ne(aq)	1. 514e-112	4. 701e-108	1. 0000	-111. 8199
Gl yci ne(aq)	2. 594e-115	1. 947e-110	1. 0000	-114. 5860
Cu(Gl y)+	3. 331e-116	4. 583e-111	0. 9468	-115. 5012
Mg(Gl y)+	3. 131e-118	3. 079e-113	0. 9468	-117. 5281
Ca(Gl y)+	7. 882e-120	8. 995e-115	0. 9468	-119. 1271
S3O6--	6. 865e-121	1. 319e-115	0. 8029	-120. 2587
Mn(Gl y)+	2. 745e-121	3. 540e-116	0. 9468	-120. 5852
Acetal dehyde(aq)	3. 533e-126	1. 556e-121	1. 0000	-125. 4519
SCN-	8. 148e-127	4. 732e-122	0. 9464	-126. 1129
NH4CH3COO(aq)	8. 511e-128	6. 559e-123	1. 0000	-127. 0700
Acetami de(aq)	8. 477e-128	5. 007e-123	1. 0000	-127. 0717
Ethanol (aq)	2. 595e-145	1. 195e-140	1. 0000	-144. 5858
Ethyne(aq)	3. 397e-147	8. 845e-143	1. 0000	-146. 4689
Ethyl ene(aq)	8. 881e-150	2. 491e-145	1. 0000	-149. 0515
Lactate	1. 880e-151	1. 674e-146	0. 9468	-150. 7497
Ca(Lac)+	2. 403e-153	3. 103e-148	0. 9468	-152. 6429
Mg(Lac)+	6. 758e-154	7. 661e-149	0. 9468	-153. 1939
Lacti c_aci d(aq)	8. 189e-155	7. 375e-150	1. 0000	-154. 0868
Cu(Lac)+	8. 442e-156	1. 288e-150	0. 9468	-155. 0973
Mn(Lac)+	4. 201e-157	6. 048e-152	0. 9468	-156. 4004
Ethane(aq)	4. 926e-159	1. 481e-154	1. 0000	-158. 3075
Propanoate	1. 820e-162	1. 329e-157	0. 9468	-161. 7638
Propanoi c_aci d(a	8. 423e-165	6. 239e-160	1. 0000	-164. 0746
Ca(Prop)+	4. 167e-165	4. 714e-160	0. 9468	-164. 4039
Mg(Prop)+	1. 440e-165	1. 402e-160	0. 9468	-164. 8654
Cu(Prop)+	3. 679e-167	5. 026e-162	0. 9468	-166. 4580
S4O6--	4. 046e-168	9. 072e-163	0. 8029	-167. 4883
Mn(Prop)+	2. 829e-168	3. 621e-163	0. 9468	-167. 5721
Succi nate	9. 884e-170	1. 147e-164	0. 8029	-169. 1004
Ca(Gl yc)2(aq)	7. 560e-170	1. 437e-164	1. 0000	-169. 1215
Mg(Gl yc)2(aq)	6. 128e-171	1. 069e-165	1. 0000	-170. 2127
H-Succi nate	2. 254e-171	2. 638e-166	0. 9468	-170. 6709
Cu(Gl yc)2(aq)	1. 132e-171	2. 418e-166	1. 0000	-170. 9462
Seri ne(aq)	2. 335e-173	2. 453e-168	1. 0000	-172. 6317
Mn(Gl yc)2(aq)	4. 471e-174	9. 166e-169	1. 0000	-173. 3496

SpecE8_output_GSS_Mine Water 9 East April 2003.txt

Succinic acid(aq)	2.171e-174	2.563e-169	1.0000	-173.6634
HAsS2(aq)	1.977e-179	2.769e-174	1.0000	-178.7039
Ethanamine(aq)	7.231e-181	3.260e-176	1.0000	-180.1408
Alanine(aq)	1.413e-182	1.259e-177	1.0000	-181.8498
Cu(Ala)+	4.094e-184	6.207e-179	0.9468	-183.4116
Mg(Ala)+	1.429e-187	1.606e-182	0.9468	-186.8687
Ca(Ala)+	8.637e-188	1.107e-182	0.9468	-187.0874
Ca(CH3COO)2(aq)	1.869e-188	2.955e-183	1.0000	-187.7284
Mg(CH3COO)2(aq)	4.774e-189	6.797e-184	1.0000	-188.3211
Mn(Ala)+	1.473e-189	2.107e-184	0.9468	-188.8555
Acetone(aq)	6.884e-190	3.998e-185	1.0000	-189.1621
Cu(CH3COO)2(aq)	1.501e-190	2.726e-185	1.0000	-189.8236
Mn(CH3COO)2(aq)	2.694e-192	4.661e-187	1.0000	-191.5695
Aspartic acid(aq)	1.027e-192	1.367e-187	1.0000	-191.9885
Propanal(aq)	4.058e-194	2.357e-189	1.0000	-193.3916
Cu(CH3COO)2-S3--	4.153e-197	7.542e-192	0.9468	-196.4054
1-Propyne(aq)	1.673e-210	1.609e-205	0.8029	-209.8718
1-Propanol(aq)	1.183e-211	4.741e-207	1.0000	-210.9268
1-Propene(aq)	9.463e-214	5.686e-209	1.0000	-213.0240
2-Hydroxybutanoic acid(aq)	4.786e-216	2.014e-211	1.0000	-215.3200
NH4(CH3COO)2-	2.663e-220	2.745e-215	0.9468	-219.5984
2-Hydroxybutanoic acid(aq)	3.528e-221	4.802e-216	0.9468	-220.4762
Asparagine(aq)	1.031e-223	1.073e-218	1.0000	-222.9868
Cu(Gly)2(aq)	5.844e-224	7.720e-219	1.0000	-223.2333
Propane(aq)	3.929e-226	8.314e-221	1.0000	-225.4058
Butanoate	7.684e-228	3.388e-223	1.0000	-227.1144
Mg(Gly)2(aq)	2.002e-231	1.744e-226	0.9468	-230.7222
Diglycine(aq)	2.077e-232	3.580e-227	1.0000	-231.6827
Butanoic acid(aq)	1.745e-232	2.305e-227	1.0000	-231.7582
Ca(But)+	7.696e-234	6.780e-229	1.0000	-233.1137
Diketopiperazine	3.158e-234	4.015e-229	0.9468	-233.5244
Mg(But)+	1.324e-234	1.510e-229	1.0000	-233.8782
Mn(Gly)2(aq)	1.042e-234	1.161e-229	0.9468	-234.0057
Cu(But)+	1.662e-235	3.374e-230	1.0000	-234.7794
Ca(Gly)2(aq)	3.196e-236	4.814e-231	0.9468	-235.5192
Mn(But)+	2.943e-236	5.537e-231	1.0000	-235.5313
Glytarate	2.638e-237	3.746e-232	0.9468	-236.6025
H-Glytarate	8.524e-238	1.109e-232	0.8029	-237.1647
Threonine(aq)	1.191e-239	1.562e-234	0.9468	-238.9478
Glytartic acid(aq)	2.701e-242	3.217e-237	1.0000	-241.5684
Ethyl acetate(aq)	1.555e-242	2.054e-237	1.0000	-241.8084
S506--	1.433e-242	1.262e-237	1.0000	-241.8438
1-Propanamine(aq)	3.140e-244	8.049e-239	0.8029	-243.5983
alpha-Aminobutyric acid(aq)	8.739e-249	5.165e-244	1.0000	-248.0586
Glytamic acid(aq)	3.322e-251	3.425e-246	1.0000	-250.4786
Butanal(aq)	1.133e-259	1.666e-254	1.0000	-258.9459
S4--	1.870e-264	1.348e-259	1.0000	-263.7282
1-Butyne(aq)	2.291e-271	2.938e-266	0.8029	-270.7353
Cu(CH3COO)3-	1.259e-280	6.809e-276	1.0000	-279.9000
1-Butanol(aq)	1.537e-282	3.699e-277	0.9468	-281.8370
1-Butene(aq)	2.130e-283	1.579e-278	1.0000	-282.6716
Mn(CH3COO)3-	3.338e-285	1.873e-280	1.0000	-284.4765
2-Hydroxypentanoic acid(aq)	3.290e-285	7.635e-280	0.9468	-284.5065
2-Hydroxypentanoic acid(aq)	5.562e-289	6.514e-284	0.9468	-288.2785
Glytamine(aq)	1.298e-292	1.533e-287	1.0000	-291.8867
n-Butane(aq)	7.359e-293	1.075e-287	1.0000	-292.1332
Alanine(aq)	1.054e-296	6.128e-292	1.0000	-295.9770
Alanine(aq)	4.285e-300	6.261e-295	1.0000	-299.3681
Pentanoate	2.651e-300	2.681e-295	0.9468	-299.6003
Pentanoic acid(aq)	1.109e-302	1.132e-297	1.0000	-300.0000
Ca(Lac)2(aq)	4.648e-303	1.014e-297	1.0000	-300.0000
Ca(Pent)+	2.445e-303	3.451e-298	0.9468	-300.0000
Mg(Lac)2(aq)	1.193e-303	2.415e-298	1.0000	-300.0000

SpecE8_output_GSS_Mine Water 9 East April 2003.txt

Mg(Pent)+	7.896e-304	9.902e-299	0.9468	-300.0000
Cu(Lac)2(aq)	5.541e-305	1.339e-299	1.0000	-300.0000
Cu(Pent)+	4.112e-305	6.771e-300	0.9468	-300.0000
Mn(Pent)+	2.636e-306	4.113e-301	0.9468	-300.0000
Mn(Lac)2(aq)	8.317e-307	1.938e-301	1.0000	-300.0000
Adipate	3.489e-308	5.028e-303	0.8029	-300.0000
H-Adipate	4.794e-310	6.957e-305	0.9468	-300.0000
Adipic_acid(aq)	7.407e-313	1.082e-307	1.0000	-300.0000
1-Butanamine(aq)	6.095e-318	4.457e-313	1.0000	-300.0000
Valine(aq)	6.168e-320	7.225e-315	1.0000	-300.0000
Benzoic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Benzoate	0.0000	0.0000	0.9483	-300.0000
2-Hydroxyheptano	0.0000	0.0000	0.9468	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
Azelate	0.0000	0.0000	0.8029	-300.0000
2-Hydroxydecano	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Azelic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecano	0.0000	0.0000	0.9468	-300.0000
Mn(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hexanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octanol(aq)	0.0000	0.0000	1.0000	-300.0000
Methionine(aq)	0.0000	0.0000	1.0000	-300.0000
Leucylglycine(aq)	0.0000	0.0000	1.0000	-300.0000
2-Heptanone(aq)	0.0000	0.0000	1.0000	-300.0000
Leucine(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Isoleucine(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9468	-300.0000
Cu(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanal(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
2-Butanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9468	-300.0000
Heptanal(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(o-Phthalate)(0.0000	0.0000	1.0000	-300.0000
1-Hexyne(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
p-Toluidic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
p-Toluate	0.0000	0.0000	0.9468	-300.0000
o-Toluate	0.0000	0.0000	0.9468	-300.0000
o-Phthalate	0.0000	0.0000	0.8029	-300.0000
Ca(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Pentylbenzene(0.0000	0.0000	1.0000	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octylbenzene(a	0.0000	0.0000	1.0000	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexylbenzene(a	0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_Mine Water 9 East April 2003.txt

n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(a	0.0000	0.0000	1.0000	-300.0000
m-Tol u i c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
H-Suberate	0.0000	0.0000	0.9468	-300.0000
H-Sebacate	0.0000	0.0000	0.9468	-300.0000
Undecanoate	0.0000	0.0000	0.9468	-300.0000
H-Pi mel ate	0.0000	0.0000	0.9468	-300.0000
Tyrosi ne(aq)	0.0000	0.0000	1.0000	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol (aq)	0.0000	0.0000	1.0000	-300.0000
Tol uene(aq)	0.0000	0.0000	1.0000	-300.0000
Suberi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Suberate	0.0000	0.0000	0.8029	-300.0000
H-Azel ate	0.0000	0.0000	0.9468	-300.0000
Sebacate	0.0000	0.0000	0.8029	-300.0000
H(o-Phthal ate) -	0.0000	0.0000	0.9468	-300.0000
S5--	0.0000	0.0000	0.8029	-300.0000
2-Hydroxyoctanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoa	0.0000	0.0000	0.9468	-300.0000
Pi mel i c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Sebaci c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Pi mel ate	0.0000	0.0000	0.8029	-300.0000
Ca(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexanol (aq)	0.0000	0.0000	1.0000	-300.0000
Phenyl al ani ne(aq	0.0000	0.0000	1.0000	-300.0000
Phenol (aq)	0.0000	0.0000	1.0000	-300.0000
Pentanal (aq)	0.0000	0.0000	1.0000	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoi c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9468	-300.0000
Undecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9468	-300.0000
Octanal (aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoi c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
m-Tol uate	0.0000	0.0000	0.9468	-300.0000
Nonanoate	0.0000	0.0000	0.9468	-300.0000
Nonanal (aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanol (aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoate	0.0000	0.0000	0.9468	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoi c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoa	0.0000	0.0000	0.9468	-300.0000
Decanoate	0.0000	0.0000	0.9468	-300.0000
n-Propyl benzene(0.0000	0.0000	1.0000	-300.0000
Decanal (aq)	0.0000	0.0000	1.0000	-300.0000
o-Phthal i c_aci d(0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	1.0000	-300.0000
o-Tol u i c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
1-Pentanami ne(aq	0.0000	0.0000	1.0000	-300.0000
1-Hexanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanami ne(aq	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states
log Q/K

log Q/K

SpecE8_output_GSS_Mine Water 9 East April 2003.txt

Quartz	-0.1230	Mg1.25SO4(OH)0.5	-10.3993
Ice	-0.1387	Manganosite	-10.6049
Tenorite	-0.1633	Periclase	-10.7496
Tridymite	-0.2945	Forsterite	-10.8132
Chalcedony	-0.3942	Mn(OH)3	-11.4233
Cristobalite(alp)	-0.6735	Portlandite	-11.4683
Calcite	-0.9046	Mg1.5SO4(OH)	-11.7479
Coesite	-0.9330	Monticellite	-12.0347
Aragonite	-1.0490	Tephroite	-12.5622
Malachite	-1.0972	MgSO4	-12.6638
Cristobalite(bet)	-1.1170	Sepiolite	-12.8344
Dolomite	-1.1264	Hausmannite	-13.0048
Dolomite-ord	-1.1264	Gyrolite	-13.1030
SiO2(am)	-1.4087	MnSO4	-13.7086
Monohydrocalcite	-1.7383	Chalcocyanite	-13.8027
Magnesite	-1.8506	Ca3(AsO4)2	-14.6239
Rhodochrosite	-2.6309	Hydromagnesite	-18.4960
Dolomite-diss	-2.6708	Hillebrandite	-18.7674
Diopside	-2.7069	Tremolite	-19.1121
Gypsum	-2.8024	Diocalciumsilica	-19.1209
Chrysocolia	-2.8438	Larnite	-20.4149
Anhydrite	-2.9783	Akermanite	-20.8039
Bassanite	-3.6232	Lime	-21.4892
CaSO4:0.5H2O(bet)	-3.7913	Anthophyllite	-25.6741
Brochantite	-3.8369	Rankinite	-26.8916
Cuprite	-4.0120	As2O5	-32.2288
Lansfordite	-4.3979	Merwinite	-32.9120
Nesquehoniite	-4.5525	Tobermorite-14A	-33.1437
Antlerite	-4.6236	Foshagite	-33.9402
Enstatite	-4.8634	Tobermorite-11A	-34.9113
Manganite	-4.9175	Afwillite	-35.0290
Brucite	-5.7122	Arsenolite	-38.0216
Epsomite	-5.8234	Claudetite	-38.0934
Talc	-5.8700	Tobermorite-9A	-38.3790
Hexahydrite	-6.0589	C	-41.0359
Pentahydrite	-6.3985	Todorokite	-43.0443
Rhodnite	-6.5333	Hatrurite	-44.2671
Starkeyite	-6.7858	As	-46.6766
Wollastonite	-6.7959	Xonotlite	-50.0389
Azurite	-6.9681	Birnessite	-52.9858
Pseudowollastonite	-7.0351	Mn	-58.9911
Cu	-7.3788	Covelite	-59.8039
Kieserite	-7.5187	Chalcocite	-61.3085
Diopside	-7.5361	S	-61.6698
Okenite	-7.5393	As4O6(mono)	-75.6787
Chrysotile	-7.6125	As4O6(cubi)	-75.9526
Pyrolusite	-7.8845	Alabandite	-82.4141
Mn(OH)2(am)	-7.9911	Mg	-95.3104
Huntite	-8.0279	Realgar	-96.0268
Chalcanthite	-8.2573	Antigorite	-109.2324
Artinite	-8.6271	Ca	-112.1193
Lammerite	-9.1447	Si	-119.7478
Bixbyite	-9.1987	Orpiment	-248.8250
MnO2(gamma)	-9.4023	o-Phthalic acid	-364.5975

Gases	fugacity	log fug.
H2O(g)	0.02598	-1.585
CO2(g)	0.004686	-2.329
N2(g)	0.004244	-2.372
NO2(g)	1.468e-012	-11.833
H2(g)	4.351e-027	-26.361

SpecE8_output_GSS_Mine Water 9 East April 2003.txt

O2(g)	4.148e-031	-30.382
NO(g)	2.883e-032	-31.540
CO(g)	6.368e-033	-32.196
NH3(g)	1.425e-038	-37.846
SO2(g)	3.466e-040	-39.460
Cu(g)	2.959e-060	-59.529
H2S(g)	7.052e-083	-82.152
CH4(g)	1.340e-085	-84.873
Mg(g)	9.477e-116	-115.023
S2(g)	5.857e-138	-137.232
Ca(g)	4.460e-138	-137.351
C2H4(g)	1.871e-147	-146.728
C(g)	2.319e-159	-158.635
Si(g)	1.612e-191	-190.793

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
<hr/>						
Ca++	0.000616	0.000616	24.7			
Cu++	1.21e-006	1.21e-006	0.0770			
H+	0.000155	0.000155	0.156			
H2AsO4-	6.59e-008	6.59e-008	0.00929			
H2O	55.5	55.5	1.00e+006			
HC03-	0.00138	0.00138	83.9			
Mg++	0.000195	0.000195	4.74			
Mn++	1.08e-007	1.08e-007	0.00592			
NH3(aq)	5.53e-006	5.53e-006	0.0941			
NO3-	2.75e-005	2.75e-005	1.70			
O2(aq)	3.88e-006	3.88e-006	0.124			
SO4--	0.000144	0.000144	13.8			
SiO2(aq)	7.56e-005	7.56e-005	4.54			

Elemental composition	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg
<hr/>					
Arsenic	6.590e-008	6.590e-008	0.004936		
Calcium	0.0006159	0.0006159	24.68		
Carbon	0.001376	0.001376	16.52		
Copper	1.212e-006	1.212e-006	0.07701		
Hydrogen	111.0	111.0	1.119e+005		
Magnesium	0.0001950	0.0001950	4.739		
Manganese	1.078e-007	1.078e-007	0.005924		
Nitrogen	3.301e-005	3.301e-005	0.4623		
Oxygen	55.51	55.51	8.881e+005		
Silicon	7.560e-005	7.560e-005	2.123		
Sulfur	0.0001439	0.0001439	4.614		

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

Temperature = 5.3 C Pressure = 1.013 bars
 pH = 7.800 log fO2 = -28.736
 Eh = 0.4050 volts pe = 7.3321
 Ionic strength = 0.001989
 Charge imbalance = -0.000109 eq/kg (-8.459% error)
 Activity of water = 1.000000
 Solvent mass = 1.000000 kg
 Solution mass = 1.000122 kg
 Solution density = 1.026 g/cm3
 Chlorinity = 0.000000 molal
 Dissolved solids = 122 mg/kg sol'n
 Hardness = 53.40 mg/kg sol'n as CaCO3
 carbonate = 39.68 mg/kg sol'n as CaCO3
 non-carbonate = 13.73 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 39.68 mg/kg sol'n as CaCO3
 Water type = Ca-HCO3

Nernst redox couples	Eh (volts)	pe
e- + .25*O2(aq) + H+ = .5*H2O	0.4050	7.3321
8*e- + 9*H+ + NO3- = 3*H2O + NH3(aq)	0.5819	10.5353

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HCO3-	0.0007850	47.89	0.9531	-3.1260
Ca++	0.0004042	16.20	0.8290	-3.4749
NO3-	0.0002239	13.88	0.9524	-3.6710
SO4--	0.0001648	15.83	0.8247	-3.8667
SiO2(aq)	0.0001535	9.221	1.0000	-3.8139
Na+	0.0001270	2.919	0.9531	-3.9172
Mg++	0.0001172	2.849	0.8330	-4.0103
N2(aq)	5.858e-005	1.641	1.0000	-4.2323
K+	4.980e-005	1.947	0.9524	-4.3240
CO2(aq)	3.893e-005	1.713	1.0000	-4.4097
Mn++	5.420e-006	0.2977	0.8290	-5.3475
CaSO4(aq)	5.402e-006	0.7354	1.0000	-5.2674
CaHCO3+	3.123e-006	0.3157	0.9531	-5.5263
MgSO4(aq)	2.074e-006	0.2497	1.0000	-5.6831
AlO2-	1.909e-006	0.1126	0.9531	-5.7400
CO3--	1.612e-006	0.09670	0.8258	-5.8759
Cu++	1.288e-006	0.08184	0.8290	-5.9715
HP04--	1.266e-006	0.1215	0.8247	-5.9813
MgHCO3+	8.793e-007	0.07502	0.9531	-6.0767
CaCO3(aq)	6.468e-007	0.06473	1.0000	-6.1892
HSiO3-	5.797e-007	0.04469	0.9531	-6.2577
Ba++	4.611e-007	0.06331	0.8268	-6.4188
H2PO4-	3.385e-007	0.03282	0.9531	-6.4913
HAlO2(aq)	2.465e-007	0.01479	1.0000	-6.6081
NaHCO3(aq)	1.913e-007	0.01607	1.0000	-6.7183
OH-	1.258e-007	0.002139	0.9528	-6.9214
MnSO4(aq)	1.139e-007	0.01720	1.0000	-6.9435
MgCO3(aq)	1.030e-007	0.008683	1.0000	-6.9872
Sb(OH)3(aq)	6.403e-008	0.01106	1.0000	-7.1936
KS04-	5.099e-008	0.006892	0.9531	-7.3133
Pb++	3.762e-008	0.007795	0.8258	-7.5077
HAsO4--	2.381e-008	0.003331	0.8247	-7.7070
H+	1.658e-008	1.671e-005	0.9560	-7.8000
Al(OH)2+	1.057e-008	0.0006444	0.9531	-7.9969
Cd++	7.541e-009	0.0008476	0.8268	-8.2051

	SpecE8_output_GSS_Mine Water	Adit	Ditch	May.txt
NaHSi O3(aq)	4. 701e-009	0. 0004705	1. 0000	-8. 3278
H2AsO4-	2. 205e-009	0. 0003107	0. 9531	-8. 6775
NaCO3-	9. 531e-010	7. 910e-005	0. 9531	-9. 0417
UO2(CO3)2--	3. 752e-010	0. 0001463	0. 8247	-9. 5094
Al OH++	3. 354e-010	1. 475e-005	0. 8258	-9. 5576
UO2(CO3)3----	2. 204e-010	9. 918e-005	0. 4623	-9. 9919
BaCO3(aq)	1. 488e-010	2. 935e-005	1. 0000	-9. 8275
HSO4-	1. 326e-010	1. 287e-005	0. 9531	-9. 8984
NaAl O2(aq)	3. 242e-011	2. 657e-006	1. 0000	-10. 4892
PO4---	2. 973e-011	2. 824e-006	0. 6480	-10. 7152
UO2CO3(aq)	1. 811e-011	5. 977e-006	1. 0000	-10. 7420
Cu+	1. 388e-011	8. 818e-007	0. 9531	-10. 8785
NaOH(aq)	2. 675e-012	1. 070e-007	1. 0000	-11. 5726
AsO4---	2. 663e-012	3. 700e-007	0. 6480	-11. 7630
Al +++	2. 381e-012	6. 423e-008	0. 6663	-11. 7996
Sb(OH)4-	1. 067e-012	2. 025e-007	0. 9531	-11. 9927
CdSO4(aq)	8. 574e-013	1. 787e-007	1. 0000	-12. 0668
H3PO4(aq)	6. 268e-013	6. 141e-008	1. 0000	-12. 2029
UO2OH+	3. 491e-013	1. 002e-007	0. 9531	-12. 4779
MgP2O7--	1. 384e-013	2. 744e-008	0. 8247	-12. 9425
HNO3(aq)	1. 096e-013	6. 908e-009	1. 0000	-12. 9601
Sb(OH)2+	7. 168e-014	1. 116e-008	0. 9531	-13. 1655
H3AsO4(aq)	4. 823e-015	6. 846e-010	1. 0000	-14. 3166
HP2O7---	4. 637e-015	8. 111e-010	0. 6480	-14. 5222
UO2++	3. 612e-015	9. 752e-010	0. 8258	-14. 5254
KHSO4(aq)	3. 282e-016	4. 468e-011	1. 0000	-15. 4839
UO2SO4(aq)	2. 912e-016	1. 066e-010	1. 0000	-15. 5359
H2P2O7--	2. 558e-016	4. 500e-011	0. 8247	-15. 6758
P2O7----	2. 136e-016	3. 714e-011	0. 4623	-16. 0055
KP2O7---	1. 308e-018	2. 786e-013	0. 6480	-18. 0719
UO2(SO4)2--	2. 709e-019	1. 252e-013	0. 8247	-18. 6509
(UO2)2(OH)2++	3. 200e-020	1. 837e-014	0. 8258	-19. 5780
UO2+	3. 056e-021	8. 252e-016	0. 9531	-20. 5357
H3P2O7-	7. 986e-022	1. 413e-016	0. 9531	-21. 1185
(UO2)3(OH)5+	3. 929e-022	3. 517e-016	0. 9531	-21. 4266
(UO2)3(CO3)6(6-)	8. 271e-024	9. 677e-018	0. 1762	-23. 8365
HAsO2(aq)	9. 906e-025	1. 069e-019	1. 0000	-24. 0041
As(OH)3(aq)	8. 966e-025	1. 129e-019	1. 0000	-24. 0474
Mn+++	8. 163e-025	4. 484e-020	0. 6518	-24. 2741
U(OH)4(aq)	4. 860e-025	1. 487e-019	1. 0000	-24. 3133
H2AsO3-	1. 608e-026	2. 008e-021	0. 9531	-25. 8146
AsO2-	1. 558e-026	1. 666e-021	0. 9531	-25. 8282
H4P2O7(aq)	2. 953e-028	5. 254e-023	1. 0000	-27. 5298
AsO2OH--	5. 035e-030	6. 239e-025	0. 8247	-29. 3817
NO2-	1. 082e-030	4. 976e-026	0. 9524	-29. 9870
O2(aq)	3. 519e-032	1. 126e-027	1. 0000	-31. 4536
Formate	4. 088e-034	1. 840e-029	0. 9528	-33. 4095
H2(aq)	1. 665e-034	3. 357e-031	1. 0000	-33. 7785
HN02(aq)	4. 361e-035	2. 050e-030	1. 0000	-34. 3604
Ca(For)+	4. 296e-036	3. 655e-031	0. 9531	-35. 3878
Mg(For)+	1. 625e-036	1. 127e-031	0. 9531	-35. 8099
Mn(For)+	1. 191e-037	1. 191e-032	0. 9531	-36. 9448
Cu(For)+	6. 032e-038	6. 548e-033	0. 9531	-37. 2404
Na(For)(aq)	5. 723e-038	3. 891e-033	1. 0000	-37. 2424
Formic acid(aq)	3. 718e-038	1. 711e-033	1. 0000	-37. 4297
K(For)(aq)	1. 880e-038	1. 581e-033	1. 0000	-37. 7259
H02-	1. 286e-038	4. 244e-034	0. 9531	-37. 9116
SO3--	9. 905e-039	7. 930e-034	0. 8258	-38. 0873
Ba(For)+	3. 491e-039	6. 365e-034	0. 9531	-38. 4779
HSO3-	1. 865e-039	1. 512e-034	0. 9531	-38. 7502
NH4+	1. 209e-039	2. 181e-035	0. 9521	-38. 9388
Pb(For)+	9. 763e-040	2. 462e-034	0. 9531	-39. 0313
Cd(For)+	2. 332e-040	3. 670e-035	0. 9531	-39. 6532

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

MnO4--	1. 294e-040	1. 539e-035	0. 8247	-39. 9718
CO(aq)	5. 443e-041	1. 524e-036	1. 0000	-40. 2642
NH3(aq)	9. 492e-042	1. 616e-037	1. 0000	-41. 0226
Oxal ate	6. 528e-042	5. 746e-037	0. 8247	-41. 2689
MnO4-	7. 522e-044	8. 945e-039	0. 9528	-43. 1447
UOH+++	6. 804e-044	1. 735e-038	0. 6518	-43. 3531
U(CO3)5(6-)	5. 444e-045	2. 929e-039	0. 1762	-45. 0181
H-Oxal ate	1. 443e-045	1. 285e-040	0. 9531	-44. 8616
S02(aq)	1. 381e-045	8. 848e-041	1. 0000	-44. 8597
HS05-	4. 381e-047	4. 953e-042	0. 9531	-46. 3793
US04++	4. 451e-048	1. 487e-042	0. 8258	-47. 4347
U(S04)2(aq)	1. 839e-048	7. 910e-043	1. 0000	-47. 7354
U++++	2. 073e-050	4. 933e-045	0. 4696	-50. 0118
Oxal i c_aci d(aq)	3. 956e-052	3. 562e-047	1. 0000	-51. 4027
S206--	5. 023e-064	8. 042e-059	0. 8247	-63. 3828
S208--	1. 456e-064	2. 798e-059	0. 8247	-63. 9204
Ca(For)2(aq)	1. 516e-068	1. 972e-063	1. 0000	-67. 8194
U+++	9. 487e-069	2. 258e-063	0. 6518	-68. 2088
Mg(For)2(aq)	7. 741e-069	8. 850e-064	1. 0000	-68. 1112
Mn(For)2(aq)	7. 773e-070	1. 127e-064	1. 0000	-69. 1094
Cu(For)2(aq)	7. 514e-070	1. 154e-064	1. 0000	-69. 1241
Na(For)2-	1. 900e-071	2. 147e-066	0. 9531	-70. 7422
Cd(For)2(aq)	1. 434e-071	2. 904e-066	1. 0000	-70. 8433
Ba(For)2(aq)	8. 666e-072	1. 970e-066	1. 0000	-71. 0622
Pb(For)2(aq)	6. 639e-072	1. 973e-066	1. 0000	-71. 1779
K(For)2-	4. 972e-072	6. 420e-067	0. 9531	-71. 3243
Formal dehyde(aq)	2. 379e-073	7. 141e-069	1. 0000	-72. 6237
N3-	1. 548e-074	6. 502e-070	0. 9531	-73. 8312
HN3(aq)	1. 838e-077	7. 909e-073	1. 0000	-76. 7356
Cu(NH3)2++	1. 138e-080	1. 111e-075	0. 8258	-80. 0270
S205--	6. 224e-083	8. 970e-078	0. 8247	-82. 2896
Cd(NH3)2++	1. 071e-085	1. 569e-080	0. 8258	-85. 0532
Urea(aq)	1. 376e-086	8. 261e-082	1. 0000	-85. 8615
HCN(aq)	1. 907e-087	5. 152e-083	1. 0000	-86. 7197
CN-	1. 893e-089	4. 925e-085	0. 9524	-88. 7440
Methanol (aq)	3. 687e-091	1. 181e-086	1. 0000	-90. 4333
Gl ycol ate	1. 401e-097	1. 051e-092	0. 9531	-96. 8745
HS-	1. 700e-098	5. 623e-094	0. 9528	-97. 7905
H2S(aq)	5. 257e-099	1. 792e-094	1. 0000	-98. 2792
Ca(Gl yc)+	2. 622e-099	3. 018e-094	0. 9531	-98. 6022
Mg(Gl yc)+	3. 496e-100	3. 473e-095	0. 9531	-99. 4773
Cu(Gl yc)+	2. 021e-100	2. 800e-095	0. 9531	-99. 7154
Mn(Gl yc)+	3. 160e-101	4. 107e-096	0. 9531	-100. 5211
Na(Gl yc)(aq)	2. 031e-101	1. 990e-096	1. 0000	-100. 6924
Gl ycol i c_aci d(aq)	1. 554e-101	1. 182e-096	1. 0000	-100. 8084
K(Gl yc)(aq)	6. 670e-102	7. 612e-097	1. 0000	-101. 1759
Pb(Gl yc)+	9. 532e-103	2. 690e-097	0. 9531	-102. 0417
Ba(Gl yc)+	5. 805e-103	1. 233e-097	0. 9531	-102. 2570
Methane(aq)	1. 600e-103	2. 566e-099	1. 0000	-102. 7959
Cd(Gl yc)+	8. 912e-104	1. 670e-098	0. 9531	-103. 0709
S--	3. 577e-104	1. 147e-099	0. 8268	-103. 5290
S204--	1. 284e-104	1. 645e-099	0. 8268	-103. 9740
S203--	3. 872e-105	4. 341e-100	0. 8247	-104. 4958
Acetate	2. 483e-109	1. 466e-104	0. 9534	-108. 6257
CaCH3COO+	7. 722e-112	7. 653e-107	0. 9531	-111. 1331
MgCH3COO+	6. 656e-112	5. 547e-107	0. 9531	-111. 1976
Aceti c_aci d(aq)	2. 216e-112	1. 331e-107	1. 0000	-111. 6544
CuCH3COO+	7. 060e-113	8. 653e-108	0. 9531	-112. 1721
NaCH3COO(aq)	2. 469e-113	2. 025e-108	1. 0000	-112. 6075
MnCH3COO+	2. 278e-113	2. 596e-108	0. 9531	-112. 6633
KCH3COO(aq)	5. 719e-114	5. 613e-109	1. 0000	-113. 2426
PbCH3COO+	2. 147e-114	5. 717e-109	0. 9531	-113. 6890
BaCH3COO+	8. 361e-115	1. 642e-109	0. 9531	-114. 0986

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

CdCH3COO+	1. 821e-115	3. 121e-110	0. 9531	-114. 7607
Malonate	1. 493e-118	1. 523e-113	0. 8247	-117. 9097
Cu(NH3)3++	1. 304e-118	1. 495e-113	0. 8258	-117. 9677
AlCH3COO++	1. 057e-118	9. 088e-114	0. 8258	-118. 0592
CuCH3COO(aq)	8. 265e-120	1. 013e-114	1. 0000	-119. 0828
H-Malonate	9. 538e-121	9. 829e-116	0. 9531	-120. 0414
AsH3(aq)	4. 705e-121	3. 667e-116	1. 0000	-120. 3274
Malonic acid(aq)	1. 085e-125	1. 129e-120	1. 0000	-124. 9644
Methanamine(aq)	1. 716e-129	5. 328e-125	1. 0000	-128. 7655
Cu(Gly)+	4. 310e-132	5. 930e-127	0. 9531	-131. 3864
Glycine(aq)	1. 590e-132	1. 194e-127	1. 0000	-131. 7985
Mg(Gly)+	1. 797e-135	1. 768e-130	0. 9531	-134. 7662
Mn(Gly)+	1. 229e-136	1. 585e-131	0. 9531	-135. 9314
Pb(Gly)+	5. 056e-137	1. 422e-131	0. 9531	-136. 3171
Ca(Gly)+	2. 677e-137	3. 055e-132	0. 9531	-136. 5933
Cd(Gly)+	2. 029e-138	3. 783e-133	0. 9531	-137. 7136
S3O6--	1. 867e-138	3. 587e-133	0. 8247	-137. 8127
Ba(Gly)+	3. 474e-140	7. 342e-135	0. 9531	-139. 4801
Acetaldehyde(aq)	2. 334e-146	1. 028e-141	1. 0000	-145. 6320
SCN-	1. 676e-147	9. 732e-143	0. 9528	-146. 7968
NH4CH3COO(aq)	4. 500e-148	3. 469e-143	1. 0000	-147. 3468
Acetamide(aq)	3. 471e-148	2. 050e-143	1. 0000	-147. 4596
UO2SCN+	1. 212e-160	3. 975e-155	0. 9531	-159. 9375
Cd(NH3)4++	4. 340e-165	7. 835e-160	0. 8258	-164. 4456
Ethanol(aq)	6. 548e-169	3. 016e-164	1. 0000	-168. 1839
Ethyne(aq)	4. 677e-169	1. 218e-164	1. 0000	-168. 3301
Ethylene(aq)	7. 943e-174	2. 228e-169	1. 0000	-173. 1000
Lactate	2. 659e-175	2. 368e-170	0. 9531	-174. 5961
Ca(Lac)+	2. 926e-177	3. 778e-172	0. 9531	-176. 5547
Mg(Lac)+	6. 471e-178	7. 336e-173	0. 9531	-177. 2099
Cu(Lac)+	1. 747e-178	2. 666e-173	0. 9531	-177. 7785
Mn(Lac)+	4. 248e-179	6. 117e-174	0. 9531	-178. 3927
Na(Lac)(aq)	3. 967e-179	4. 445e-174	1. 0000	-178. 4015
Lactic acid(aq)	3. 070e-179	2. 765e-174	1. 0000	-178. 5128
K(Lac)(aq)	1. 303e-179	1. 670e-174	1. 0000	-178. 8849
Pb(Lac)+	1. 771e-180	5. 246e-175	0. 9531	-179. 7727
Ba(Lac)+	4. 458e-181	1. 009e-175	0. 9531	-180. 3717
Cd(Lac)+	1. 145e-181	2. 306e-176	0. 9531	-180. 9622
Ethane(aq)	1. 005e-185	3. 023e-181	1. 0000	-184. 9976
Propanoate	2. 728e-189	1. 994e-184	0. 9531	-188. 5849
Ca(Prop)+	4. 745e-192	5. 369e-187	0. 9531	-191. 3446
Propanoic acid(aq)	3. 260e-192	2. 414e-187	1. 0000	-191. 4868
Mg(Prop)+	1. 980e-192	1. 927e-187	0. 9531	-191. 7243
Cu(Prop)+	7. 662e-193	1. 047e-187	0. 9531	-192. 1366
Na(Prop)(aq)	4. 017e-193	3. 858e-188	1. 0000	-192. 3961
Mn(Prop)+	2. 969e-193	3. 800e-188	0. 9531	-192. 5482
K(Prop)(aq)	1. 319e-193	1. 480e-188	1. 0000	-192. 8796
USCN+++	5. 122e-194	1. 516e-188	0. 6518	-193. 4765
Pb(Prop)+	2. 079e-194	5. 827e-189	0. 9531	-193. 7029
Ca(Glyc)2(aq)	8. 016e-195	1. 524e-189	1. 0000	-194. 0960
Cu(Glyc)2(aq)	2. 652e-195	5. 664e-190	1. 0000	-194. 5765
Cd(Prop)+	1. 851e-195	3. 433e-190	0. 9531	-194. 7535
Ba(Prop)+	1. 180e-195	2. 483e-190	0. 9531	-194. 9489
Mg(Glyc)2(aq)	6. 861e-196	1. 196e-190	1. 0000	-195. 1636
S4O6--	3. 116e-196	6. 987e-191	0. 8247	-195. 5901
Succinate	7. 129e-197	8. 274e-192	0. 8247	-196. 2307
Mn(Glyc)2(aq)	3. 950e-197	8. 097e-192	1. 0000	-196. 4034
Pb(Glyc)2(aq)	8. 301e-198	2. 965e-192	1. 0000	-197. 0809
Na(Glyc)2-	3. 645e-198	6. 308e-193	0. 9531	-197. 4592
K(Glyc)2-	9. 887e-199	1. 870e-193	0. 9531	-198. 0258
Serine(aq)	4. 580e-199	4. 812e-194	1. 0000	-198. 3392
Cd(Glyc)2(aq)	4. 555e-199	1. 196e-193	1. 0000	-198. 3415
H-Succinate	4. 498e-199	5. 266e-194	0. 9531	-198. 3678

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

Ba(Gl yc)2(aq)	3. 849e-199	1. 106e-193	1. 0000	-198. 4147
Succi ni c_aci d(aq)	1. 255e-202	1. 481e-197	1. 0000	-201. 9015
Ethanami ne(aq)	3. 138e-209	1. 414e-204	1. 0000	-208. 5034
Al ani ne(aq)	4. 100e-211	3. 653e-206	1. 0000	-210. 3872
Cu(Al a)+	2. 521e-211	3. 821e-206	0. 9531	-210. 6194
Mg(Al a)+	2. 901e-216	3. 260e-211	0. 9531	-215. 5583
Mn(Al a)+	2. 567e-216	3. 672e-211	0. 9531	-215. 6114
Ca(Al a)+	1. 299e-216	1. 665e-211	0. 9531	-215. 9071
Pb(Al a)+	1. 010e-216	2. 983e-211	0. 9531	-216. 0164
Cd(Al a)+	2. 939e-217	5. 892e-212	0. 9531	-216. 5527
Ca(CH3COO)2(aq)	4. 027e-219	6. 368e-214	1. 0000	-218. 3950
Mg(CH3COO)2(aq)	1. 567e-219	2. 231e-214	1. 0000	-218. 8050
Cu(CH3COO)2(aq)	7. 033e-220	1. 277e-214	1. 0000	-219. 1529
Ba(Al a)+	6. 129e-220	1. 381e-214	0. 9531	-219. 2335
Mn(CH3COO)2(aq)	5. 299e-221	9. 168e-216	1. 0000	-220. 2758
Acetone(aq)	2. 760e-221	1. 603e-216	1. 0000	-220. 5592
Pb(CH3COO)2(aq)	4. 103e-222	1. 335e-216	1. 0000	-221. 3869
Na(CH3COO)2-	4. 026e-222	5. 678e-217	0. 9531	-221. 4160
Asparti c_aci d(aq)	3. 818e-222	5. 082e-217	1. 0000	-221. 4181
Cd(CH3COO)2(aq)	1. 082e-222	2. 495e-217	1. 0000	-221. 9656
Ba(CH3COO)2(aq)	6. 250e-223	1. 596e-217	1. 0000	-222. 2041
K(CH3COO)2-	5. 857e-223	9. 205e-218	0. 9531	-222. 2532
Al (CH3COO)2+	4. 002e-225	5. 805e-220	0. 9531	-224. 4186
Propanal (aq)	9. 631e-226	5. 593e-221	1. 0000	-225. 0163
Cu(CH3COO)2-	3. 434e-228	6. 237e-223	0. 9531	-227. 4850
1-Propyne(aq)	1. 036e-244	4. 151e-240	1. 0000	-243. 9846
S3--	7. 536e-248	7. 249e-243	0. 8247	-247. 2066
1-Propanol (aq)	8. 255e-249	4. 960e-244	1. 0000	-248. 0833
1-Propene(aq)	1. 933e-251	8. 134e-247	1. 0000	-250. 7137
2-Hydroxybutanoa	1. 164e-255	1. 200e-250	0. 9531	-254. 9549
NH4(CH3COO)2-	1. 192e-256	1. 622e-251	0. 9531	-255. 9447
Asparagi ne(aq)	6. 907e-258	9. 124e-253	1. 0000	-257. 1607
Cu(Gl y)2(aq)	9. 102e-259	1. 926e-253	1. 0000	-258. 0409
2-Hydroxybutanoi	1. 187e-259	1. 235e-254	1. 0000	-258. 9257
Propane(aq)	4. 932e-266	2. 175e-261	1. 0000	-265. 3070
Mg(Gl y)2(aq)	1. 241e-266	2. 140e-261	1. 0000	-265. 9061
Di gl yci ne(aq)	5. 363e-267	7. 085e-262	1. 0000	-266. 2706
Mn(Gl y)2(aq)	6. 280e-268	1. 275e-262	1. 0000	-267. 2020
Pb(Gl y)2(aq)	4. 987e-268	1. 772e-262	1. 0000	-267. 3022
Cd(Gl y)2(aq)	6. 819e-269	1. 776e-263	1. 0000	-268. 1663
Di ketopi perazi ne	1. 614e-269	1. 842e-264	1. 0000	-268. 7920
Butanoate	9. 473e-270	8. 250e-265	0. 9531	-269. 0444
Ca(Gl y)2(aq)	5. 338e-271	1. 004e-265	1. 0000	-270. 2726
Ca(But)+	1. 128e-272	1. 435e-267	0. 9531	-271. 9684
Butanoi c_aci d(aq)	8. 927e-273	7. 865e-268	1. 0000	-272. 0493
Mg(But)+	4. 483e-273	4. 993e-268	0. 9531	-272. 3693
Cu(But)+	2. 109e-273	3. 176e-268	0. 9531	-272. 6968
Na(But) (aq)	1. 349e-273	1. 485e-268	1. 0000	-272. 8701
Mn(But)+	8. 817e-274	1. 252e-268	0. 9531	-273. 0755
Ba(Gl y)2(aq)	7. 228e-274	2. 063e-268	1. 0000	-273. 1410
K(But) (aq)	4. 431e-274	5. 591e-269	1. 0000	-273. 3535
Pb(But)+	1. 988e-275	5. 850e-270	0. 9531	-274. 7224
Ba(But)+	2. 736e-276	6. 140e-271	0. 9531	-275. 5837
Cd(But)+	2. 692e-276	5. 370e-271	0. 9531	-275. 5908
Gl utarate	2. 132e-276	2. 773e-271	0. 8247	-275. 7550
H-Gl utarate	7. 684e-279	1. 007e-273	0. 9531	-278. 1353
Threoni ne(aq)	1. 618e-279	1. 927e-274	1. 0000	-278. 7910
Ethyl acetate(aq)	4. 480e-282	3. 946e-277	1. 0000	-281. 3488
Gl utari c_aci d(aq)	2. 649e-282	3. 499e-277	1. 0000	-281. 5769
S506--	3. 438e-284	8. 811e-279	0. 8247	-283. 5474
1-Propanami ne(aq)	1. 342e-288	7. 930e-284	1. 0000	-287. 8723
a-Ami nobutyri c_a	3. 084e-291	3. 180e-286	1. 0000	-290. 5109
Gl utami c_aci d(aq)	1. 271e-300	1. 870e-295	1. 0000	-299. 8957

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

Butanal (aq)	1.077e-307	7.764e-303	1.0000	-300.0000
UO2(SCN)2(aq)	1.038e-307	4.008e-302	1.0000	-300.0000
S4--	6.199e-320	7.950e-315	0.8247	-300.0000
Cd(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
K(Prop)2-	0.0000	0.0000	0.9531	-300.0000
K(Pent)2-	0.0000	0.0000	0.9531	-300.0000
K(Pent)(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
K(Lac)2-	0.0000	0.0000	0.9531	-300.0000
Alanyl glycine(aq)	0.0000	0.0000	1.0000	-300.0000
K(But)2-	0.0000	0.0000	0.9531	-300.0000
2-Hexanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexyne(aq)	0.0000	0.0000	1.0000	-300.0000
Isoleucine(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9531	-300.0000
Hexanal (aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9531	-300.0000
Heptanal (aq)	0.0000	0.0000	1.0000	-300.0000
2-Heptanone(aq)	0.0000	0.0000	1.0000	-300.0000
HSb2S4-	0.0000	0.0000	0.9531	-300.0000
Ca(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
p-Toluate	0.0000	0.0000	0.9531	-300.0000
o-Toluidic acid(aq)	0.0000	0.0000	1.0000	-300.0000
o-Toluate	0.0000	0.0000	0.9531	-300.0000
Ca(Pent)+	0.0000	0.0000	0.9531	-300.0000
n-Propyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
2-Butanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butyne(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butane(aq)	0.0000	0.0000	1.0000	-300.0000
m-Toluidic acid(aq)	0.0000	0.0000	1.0000	-300.0000
m-Toluate	0.0000	0.0000	0.9531	-300.0000
Valine(aq)	0.0000	0.0000	1.0000	-300.0000
Undecanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Undecanoate	0.0000	0.0000	0.9531	-300.0000
UO2(SCN)3-	0.0000	0.0000	0.9531	-300.0000
Adipic acid(aq)	0.0000	0.0000	1.0000	-300.0000
U(SCN)2++	0.0000	0.0000	0.8258	-300.0000
Tyrosine(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Adipate	0.0000	0.0000	0.8247	-300.0000
1-Hexanol (aq)	0.0000	0.0000	1.0000	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
Toluene(aq)	0.0000	0.0000	1.0000	-300.0000
Suberic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Suberate	0.0000	0.0000	0.8247	-300.0000
Ca(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
H-Suberate	0.0000	0.0000	0.9531	-300.0000
Sebacate	0.0000	0.0000	0.8247	-300.0000
Sb2S4--	0.0000	0.0000	0.8247	-300.0000
H-Sebacate	0.0000	0.0000	0.9531	-300.0000
H-Pimelate	0.0000	0.0000	0.9531	-300.0000
S5--	0.0000	0.0000	0.8247	-300.0000

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

H-Azel ate	0.0000	0.0000	0.9531	-300.0000
H-Adi pate	0.0000	0.0000	0.9531	-300.0000
1-Hexanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butene(aq)	0.0000	0.0000	1.0000	-300.0000
Pi mel i c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Gl utami ne(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol (aq)	0.0000	0.0000	1.0000	-300.0000
Pentanoi c_aci d(a	0.0000	0.0000	1.0000	-300.0000
Pentanal (aq)	0.0000	0.0000	1.0000	-300.0000
Benzoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Pent)+	0.0000	0.0000	0.9531	-300.0000
Benzoate	0.0000	0.0000	0.9543	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(CH3COO)3-	0.0000	0.0000	0.9531	-300.0000
Pb(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
Pentanoate	0.0000	0.0000	0.9531	-300.0000
Dodecanoate	0.0000	0.0000	0.9531	-300.0000
2-Hydroxypentano	0.0000	0.0000	1.0000	-300.0000
Phenyl al ani ne(aq)	0.0000	0.0000	1.0000	-300.0000
Pi mel ate	0.0000	0.0000	0.8247	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9531	-300.0000
Ba(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanal (aq)	0.0000	0.0000	1.0000	-300.0000
Decanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9531	-300.0000
2-Hydroxypentano	0.0000	0.0000	0.9531	-300.0000
Nonanoate	0.0000	0.0000	0.9531	-300.0000
Nonanal (aq)	0.0000	0.0000	1.0000	-300.0000
Decanal (aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)+	0.0000	0.0000	0.9531	-300.0000
2-Hydroxyoctanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanol (aq)	0.0000	0.0000	1.0000	-300.0000
Na(Prop)2-	0.0000	0.0000	0.9531	-300.0000
Na(Pent)2-	0.0000	0.0000	0.9531	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Pent) (aq)	0.0000	0.0000	1.0000	-300.0000
Na(Lac)2-	0.0000	0.0000	0.9531	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoa	0.0000	0.0000	0.9531	-300.0000
Cu(Pent)+	0.0000	0.0000	0.9531	-300.0000
Sebaci c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Na(But)2-	0.0000	0.0000	0.9531	-300.0000
Cu(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butanol (aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9531	-300.0000

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)+	0.0000	0.0000	0.9531	-300.0000
Cu(CH3COO)3-	0.0000	0.0000	0.9531	-300.0000
2-Hydroxyhexanoic	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoic	0.0000	0.0000	0.9531	-300.0000
Mn(CH3COO)3-	0.0000	0.0000	0.9531	-300.0000
Cu(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptanoic	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)+	0.0000	0.0000	0.9531	-300.0000
Cd(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Azelate	0.0000	0.0000	0.8247	-300.0000
2-Hydroxyheptanoic	0.0000	0.0000	0.9531	-300.0000
Mg(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Pent)+	0.0000	0.0000	0.9531	-300.0000
Azelate	0.0000	0.0000	1.0000	-300.0000
Cd(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoic	0.0000	0.0000	1.0000	-300.0000
1-Octanol(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
Methionine(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoic	0.0000	0.0000	0.9531	-300.0000
Leucylglycine(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octyl benzene(a	0.0000	0.0000	1.0000	-300.0000
Leucine(aq)	0.0000	0.0000	1.0000	-300.0000
n-Pentyl benzene(0.0000	0.0000	1.0000	-300.0000
Cd(CH3COO)3-	0.0000	0.0000	0.9531	-300.0000
o-Phthalate	0.0000	0.0000	0.8247	-300.0000
1-Octanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butanamine(aq)	0.0000	0.0000	1.0000	-300.0000
p-Toluate	0.0000	0.0000	1.0000	-300.0000
(NH4)2Sb2S4(aq)	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states

log Q/K

log Q/K

Stilbite	10.5604s/sat	Salinite	-11.2705
Muscovite	10.4610s/sat	Thenardite	-11.3703
Beidellite-Ca	7.6939s/sat	Mg1.25S04(OH)0.5	-11.3920
Beidellite-Mg	7.6053s/sat	U02.25	-11.4240
Illite	7.4140s/sat	U02.25(beta)	-11.4955
Beidellite-K	7.2668s/sat	Periclase	-11.6121
Kaolinite	7.0250s/sat	Portlandite	-12.0343
Beidellite-Na	6.9988s/sat	Montcellite	-12.0471
Beidellite-H	6.7466s/sat	Natron	-12.1750
Mesolite	6.5548s/sat	Torbernite	-12.2021
Pyrophyllite	6.5291s/sat	CdS04	-12.5691
Paragonite	6.5271s/sat	MnS04	-12.6189

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

Montmor-Ca	6. 4899s/sat	U02S04: H2O	-12. 8062
Montmor-Mg	6. 4784s/sat	MgU04	-12. 8179
Montmor-K	6. 1398s/sat	Na2CO3: 7H2O	-12. 8327
Montmor-Na	5. 8662s/sat	Mg1. 5S04(OH)	-12. 8454
Scoloci te	5. 6727s/sat	Uranocircite	-12. 8490
Laumontite	4. 3936s/sat	BaU2O7	-13. 0904
Maximucroci	3. 8311s/sat	(U02)3(P04)2: 4H2O	-13. 4441
K-Feldspar	3. 8283s/sat	Chalcocyanite	-13. 6196
Margarite	3. 5161s/sat	Przhevalskite	-13. 6870
Celadonite	3. 3019s/sat	MgSO4	-13. 8233
Diaspore	3. 0755s/sat	Tremolite	-14. 0514
Boehmite	2. 6370s/sat	Mordenite-dehydrate	-14. 1277
Lawsonite	2. 5935s/sat	UP05	-14. 2789
Gibbsite	2. 5820s/sat	U02. 6667	-14. 4214
Sandineigh	2. 4941s/sat	Thermonatrite	-14. 4581
Mordenite	2. 2695s/sat	Na2CO3	-14. 8252
Corundum	1. 7463s/sat	Na2U2O7	-14. 8342
Brochantite	1. 5763s/sat	Uraninite	-14. 9571
Witherite	1. 4749s/sat	Coffinite	-15. 1458
Malachite	1. 4574s/sat	Autunite-H	-16. 0046
Tenorite	1. 1693s/sat	Ba(OH)2: 8H2O	-16. 0911
Kyanite	0. 9317s/sat	Ca3(AsO4)2	-16. 2386
Albite-low	0. 8464s/sat	CaAl2O4	-17. 0363
Albite	0. 8464s/sat	U02S04: 3. 5H2O	-17. 2504
Quartz	0. 6643s/sat	U02S04: 2. 5H2O	-17. 3385
Andalusite	0. 6080s/sat	Ba2Si3O8	-17. 3769
Cerussite	0. 5749s/sat	U02S04: 3H2O	-17. 4124
Tridymite	0. 4656s/sat	NaUO3	-17. 8966
Chalcedony	0. 3739s/sat	Natrosilite	-18. 5105
Sillimanite	0. 2041s/sat	Gehlenite	-18. 6620
Analcite	0. 1438s/sat	Alum-K	-18. 7361
Bari te	0. 0717s/sat	Hillebrandite	-18. 8456
Cristobalite(alp)	0. 0672s/sat	CaAl4O7	-19. 0023
Ice	-0. 0649	Na2SiO3	-19. 2581
Amesite-14A	-0. 0745	Dicalciumsilica	-19. 3731
Coesite	-0. 1996	Hydromagnesite	-20. 2824
Cristobalite(bet)	-0. 4351	Ningyote	-20. 4534
Diopside	-0. 5313	Akermanite	-20. 5537
Prehnite	-0. 5722	Larnite	-20. 7840
Albite-high	-0. 6091	U02. 3333(beta)	-20. 9962
Wairakit e	-0. 6696	U02S04	-21. 3416
Rhodochrosite	-0. 7249	Anthophyllite	-21. 6402
Dawsonite	-0. 7358	Lime	-22. 8599
SiO2(am)	-0. 7915	Na2UO4(alpha)	-23. 3525
Kalsilite	-0. 9049	U02(N03)2: 6H2O	-23. 9223
Calcite	-0. 9467	U02(N03)2: 3H2O	-25. 6758
Aragonite	-1. 0922	Pargosite	-26. 0882
Clinozoisite	-1. 1006	U02(P03)2	-26. 4209
Natrolite	-1. 1119	Rankinite	-26. 7782
Zoisite	-1. 1500	Pb	-26. 8153
Dolomite-ord	-1. 3521	U02(N03)2: 2H2O	-27. 0968
Dolomite	-1. 3522	(U02)2P2O7	-27. 5095
Jadeite	-1. 4384	(U02)3(P04)2	-27. 6703
Alamosite	-1. 5385	K2UO4	-28. 0348
Monohydrocalcite	-1. 7382	Tobermorite-14A	-28. 6287
Saponite-Ca	-2. 0227	U(HP04)2: 4H2O	-28. 8182
Saponite-Mg	-2. 1088	KAl(SO4)2	-28. 8290
Magnesite	-2. 1615	UP2O7	-29. 3424
Saponite-K	-2. 4498	Ettringite	-29. 9290
Anorthite	-2. 4871	U02(N03)2: H2O	-31. 0468
Phlogopite	-2. 6811	Tobermorite-11A	-31. 0494
Saponite-Na	-2. 7176	Ba2SiO4	-32. 9161
Gypsum	-2. 8304	Foshagite	-33. 1374

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

Saponi te-H	-2. 9701	Merwi ni te	-33. 3820
Lanarki te	-3. 0361	UO2(NO3)2	-34. 8239
Dol omi te-di s	-3. 0488	Tobermori te-9A	-35. 0360
Al uni te	-3. 0543	Afwil li te	-35. 1053
Bi xbyi te	-3. 1781	As2O5	-35. 5210
Tal c	-3. 1999	UO2SO3	-36. 5064
Anhydri te	-3. 2046	Cd	-37. 8931
Hydroxyl apati te	-3. 2546	Al 2(SO4)3: 6H2O	-39. 3105
Angl esi te	-3. 3152	UO2(AsO3)2	-40. 1357
Azuri te	-3. 6128	(UO2)2As2O7	-40. 2743
Nephel i ne	-3. 7339	(UO2)3(AsO4)2	-41. 3282
Bassani te	-3. 8562	BaO	-41. 9772
Whi t l ocki te	-3. 8930	U(SO4)2: 8H2O	-45. 6205
Berl i ni te	-3. 9040	Hatruri te	-46. 1794
Cupri te	-3. 9173	Arsenol i te	-46. 3077
Pyrol usi te	-3. 9599	Cl audeti te	-46. 3910
Rhodoni te	-4. 0279	Ba2UO7	-46. 6185
CaSO4: 0. 5H2O(bet	-4. 0503	U(SO4)2: 4H2O	-47. 0553
CaUO4	-4. 3792	U(SO4)2	-47. 4018
UO3: 2H2O	-4. 3974	Xonotl i te	-47. 7879
Schoepi te	-4. 3991	C	-48. 4694
CdSi O3	-4. 5027	U(CO3)2	-50. 2997
Enstati te	-4. 5106	As	-56. 1301
UO2(OH)2(beta)	-4. 5836	Al 2(SO4)3	-58. 5785
Sb2O3	-4. 6275	Na	-60. 3584
Schoepi te-dehy(.	-4. 6403	Na3UO4	-61. 0112
UO3: . 9H2O(al pha)	-4. 6578	Na4Si O4	-62. 9503
Schoepi te-dehy(.	-4. 7451	Mn	-63. 1790
Schoepi te-dehy(1	-4. 7661	Ca3Al 2O6	-63. 8337
Cl i nochl ore-14A	-4. 9311	Na2O	-64. 0412
Hausmanni te	-5. 0187	K	-64. 5151
PbCO3. PbO	-5. 0860	Covel l i te	-71. 8399
Nesquehoni te	-5. 3600	Ba3UO6	-72. 7915
Li tharge	-5. 3660	S	-72. 9308
Massi cot	-5. 5692	Chal coci te	-74. 4310
UO2CO3	-5. 9617	CdS	-81. 3983
Schoepi te-dehy(.	-5. 9651	Gal ena	-81. 5789
Rutherfordi ne	-5. 9682	K2O	-82. 3937
Okeni te	-6. 0198	U(SO3)2	-89. 0858
Bruci te	-6. 0968	Anti gori te	-90. 1764
Mn(OH)2(am)	-6. 2818	PbSO4(NH3)2	-91. 0205
Wol l astoni te	-6. 3313	Na6Si 2O7	-91. 7047
Di opsi de	-6. 4024	As4O6(mono)	-92. 2962
Cordi eri te_hydr	-6. 4972	As4O6(cubi)	-92. 4887
Schoepi te-dehy(.	-6. 5279	Al abandi te	-95. 2083
Chrysoti l e	-6. 5655	Mg	-104. 4778
Pseudowol l astoni	-6. 6073	BaS	-113. 7961
Pb2Si O4	-6. 6488	UN1. 73(al pha)	-114. 6016
Nahcol i te	-6. 7063	UN1. 59(al pha)	-117. 4815
Anal ci me-dehy	-6. 7225	Ca	-122. 2118
Pb3SO6	-6. 7238	P	-124. 8173
Chal canthi te	-7. 1416	Ba	-126. 4839
UO2HP04: 4H2O	-7. 2971	Al	-126. 6184
Ni ter	-7. 3246	UN	-131. 2706
Cd(OH)2	-7. 4413	Si	-131. 9304
UO3(gamma)	-7. 6418	Pb(N3)2(orth)	-145. 3842
Spi nel	-7. 7675	Pb(N3)2(mono)	-145. 8484
Sanborni te	-8. 1659	PbSO4(NH3)4	-176. 5780
UO3(beta)	-8. 2892	U	-180. 4839
Tephroi te	-8. 3168	UAs	-192. 0199
Cl i nochl ore-7A	-8. 4738	US	-193. 1836
Sepi ol i te	-8. 5381	UC	-210. 4076
UO3(al pha)	-8. 6534	UH3(beta)	-212. 3172

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

Cu	-8.8053	US1.9	-223.4556
Ca-Al_Pyroxene	-8.8680	US2	-228.9066
Grossular	-8.9291	Mayenite	-236.7296
Montepontite	-8.9936	UAs2	-245.3229
Artinite	-9.0171	UP	-255.3584
Huntite	-9.1507	UC1.94(alpha)	-258.1395
Manganosite	-9.1789	Stibnite	-273.9683
Cordierite_anhyd	-9.1902	US3	-298.4432
BaUO4	-9.6159	Orpiment	-299.4082
Mirabilite	-9.6413	Pb(SCN)2	-300.9248
Arcanite	-10.3597	UP2	-374.7331
Pb4SO7	-10.4802	U2S3	-414.6957
CdSO4: 2.667H2O	-10.4830	o-Phthalic_acid	-424.2301
Nitrobarite	-10.7513	U2C3	-470.9262
CdSO4: H2O	-10.7779	U3As4	-629.9350
Gyrolite	-10.9404	U3S5	-638.6823
Forsterite	-10.9839	U3P4	-885.4510

Gases	fugacity	log fug.
N2(g)	0.06239	-1.205
H2O(g)	0.007285	-2.138
CO2(g)	0.0005978	-3.223
NO2(g)	1.736e-013	-12.761
O2(g)	1.836e-029	-28.736
H2(g)	1.805e-031	-30.743
NO(g)	7.712e-032	-31.113
CO(g)	3.704e-038	-37.431
NH3(g)	5.428e-044	-43.265
SO2(g)	4.124e-046	-45.385
Cd(g)	1.526e-053	-52.816
Pb(g)	2.150e-058	-57.668
Cu(g)	7.024e-066	-65.153
Na(g)	6.588e-076	-75.181
UO3(g)	6.820e-077	-76.166
K(g)	6.030e-077	-76.220
H2S(g)	2.889e-098	-97.539
CH4(g)	7.177e-101	-100.144
UO2(g)	1.002e-119	-118.999
Mg(g)	9.521e-127	-126.021
Ca(g)	2.212e-150	-149.655
S2(g)	4.463e-162	-161.350
C2H4(g)	9.668e-172	-171.015
C(g)	1.047e-175	-174.980
Al(g)	3.807e-182	-181.419
UO(g)	2.760e-196	-195.559
Si(g)	2.672e-209	-208.573
U(g)	2.106e-273	-272.677

Original	basis	total	moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
Al+++		2.17e-006	2.17e-006	0.0585				
Ba++		4.61e-007	4.61e-007	0.0633				
Ca++		0.000413	0.000413	16.6				
Cd++		7.54e-009	7.54e-009	0.000848				
Cu++		1.29e-006	1.29e-006	0.0818				
H+		2.77e-005	2.77e-005	0.0280				
H2AsO4-		2.60e-008	2.60e-008	0.00367				
H2O		55.5	55.5	1.00e+006				
HC03-		0.000830	0.000830	50.7				
HP04--		1.60e-006	1.60e-006	0.154				
K+		4.98e-005	4.98e-005	1.95				

SpecE8_output_GSS_Mine Water Adit Ditch May.txt

Mg++	0.000120	0.000120	2.92
Mn++	5.53e-006	5.53e-006	0.304
NH3(aq)	0.000117	0.000117	2.00
NO3-	0.000224	0.000224	13.9
Na+	0.000127	0.000127	2.92
O2(aq)	8.79e-005	8.79e-005	2.81
Pb++	3.76e-008	3.76e-008	0.00779
SO4--	0.000172	0.000172	16.6
Sb(OH)3(aq)	6.40e-008	6.40e-008	0.0111
SiO2(aq)	0.000154	0.000154	9.26
UO2++	6.14e-010	6.14e-010	0.000166

Elemental composition	In fluid		Sorbed	
	total moles	moles	moles	mg/kg
<hr/>				
Aluminum	2.167e-006	2.167e-006		0.05846
Antimony	6.403e-008	6.403e-008		0.007795
Arsenic	2.601e-008	2.601e-008		0.001949
Barium	4.612e-007	4.612e-007		0.06333
Cadmium	7.542e-009	7.542e-009		0.0008477
Calcium	0.0004133	0.0004133		16.56
Carbon	0.0008305	0.0008305		9.973
Copper	1.288e-006	1.288e-006		0.08184
Hydrogen	111.0	111.0	1.119e+005	
Lead	3.762e-008	3.762e-008		0.007795
Magnesium	0.0001203	0.0001203		2.923
Manganese	5.534e-006	5.534e-006		0.3040
Nitrogen	0.0003411	0.0003411		4.777
Oxygen	55.51	55.51	8.881e+005	
Phosphorus	1.605e-006	1.605e-006		0.04969
Potassium	4.985e-005	4.985e-005		1.949
Silicon	0.0001541	0.0001541		4.327
Sodium	0.0001272	0.0001272		2.923
Sulfur	0.0001724	0.0001724		5.529
Uranium	6.141e-010	6.141e-010		0.0001462

SpecE8_output_GSS_Mine Water Adit Pipe May.txt

Temperature = 6.0 C Pressure = 1.013 bars
 pH = 7.800 log fO2 = -29.280
 Eh = 0.3950 volts pe = 7.1308
 Ionic strength = 0.003519
 Charge imbalance = -0.000033 eq/kg (-1.414% error)
 Activity of water = 1.000000
 Solvent mass = 1.000000 kg
 Solution mass = 1.000204 kg
 Solution density = 1.026 g/cm3
 Chlorinity = 0.000000 molal
 Dissolved solids = 204 mg/kg sol'n
 Hardness = 106.00 mg/kg sol'n as CaCO3
 carbonate = 73.41 mg/kg sol'n as CaCO3
 non-carbonate = 32.59 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 73.41 mg/kg sol'n as CaCO3
 Water type = Ca-HCO3

Nernst redox couples	Eh (volts)	pe
$e^- + .25*O_2(aq) + H^+ = .5*H_2O$	0.3950	7.1308
$8*e^- + 9*H^+ + NO_3^- = 3*H_2O + NH_3(aq)$	0.5795	10.4608

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HCO3-	0.001445	88.14	0.9392	-2.8675
Ca++	0.0007574	30.35	0.7841	-3.2263
NO3-	0.0005562	34.48	0.9381	-3.2825
Mg++	0.0002726	6.623	0.7906	-3.6666
Na+	0.0001692	3.889	0.9392	-3.7989
SO4--	0.0001594	15.31	0.7773	-3.9068
SiO2(aq)	0.0001245	7.476	1.0000	-3.9050
N2(aq)	8.590e-005	2.406	1.0000	-4.0660
CO2(aq)	6.927e-005	3.048	1.0000	-4.1594
K+	4.983e-005	1.948	0.9381	-4.3303
CaHCO3+	1.014e-005	1.025	0.9392	-5.0211
CaSO4(aq)	8.752e-006	1.191	1.0000	-5.0579
MgSO4(aq)	4.257e-006	0.5123	1.0000	-5.3709
MgHCO3+	3.557e-006	0.3034	0.9392	-5.4762
CO3--	3.171e-006	0.1902	0.7791	-5.6073
AlO2-	2.887e-006	0.1703	0.9392	-5.5668
Mn++	2.841e-006	0.1560	0.7841	-5.6521
HP04--	2.411e-006	0.2314	0.7773	-5.7272
CaCO3(aq)	2.159e-006	0.2161	1.0000	-5.6657
Fe++	8.303e-007	0.04636	0.7841	-6.1864
Cu++	6.291e-007	0.03997	0.7841	-6.3069
H2PO4-	6.107e-007	0.05922	0.9392	-6.2414
Ba++	4.967e-007	0.06820	0.7808	-6.4114
HSiO3-	4.916e-007	0.03789	0.9392	-6.3356
NaHCO3(aq)	4.483e-007	0.03765	1.0000	-6.3485
MgCO3(aq)	4.246e-007	0.03579	1.0000	-6.3721
HAIO2(aq)	3.501e-007	0.02100	1.0000	-6.4559
OH-	1.372e-007	0.002333	0.9386	-6.8901
Sb(OH)3(aq)	1.201e-007	0.02075	1.0000	-6.9204
MnSO4(aq)	5.184e-008	0.007826	1.0000	-7.2853
KS04-	4.646e-008	0.006279	0.9392	-7.3601
FeCO3+	4.259e-008	0.004933	0.9392	-7.3980
HAsO4--	2.392e-008	0.003347	0.7773	-7.7306
Cd++	1.908e-008	0.002144	0.7808	-7.8269
H+	1.679e-008	1.692e-005	0.9441	-7.8000

	SpecE8_output_GSS_Mi ne Water	Adi t	Pi pe	May. txt
Al (OH) 2+	1. 440e-008	0. 0008781	0. 9392	-7. 8689
Pb++	9. 411e-009	0. 001950	0. 7791	-8. 1348
NaHSi O3(aq)	5. 057e-009	0. 0005060	1. 0000	-8. 2961
UO2(CO3) 3----	2. 881e-009	0. 001296	0. 3647	-8. 9785
UO2(CO3) 2--	2. 386e-009	0. 0009303	0. 7773	-8. 7318
NaCO3-	2. 309e-009	0. 0001916	0. 9392	-8. 6638
H2AsO4-	2. 102e-009	0. 0002961	0. 9392	-8. 7047
Al OH++	4. 502e-010	1. 980e-005	0. 7791	-9. 4550
BaCO3(aq)	2. 853e-010	5. 630e-005	1. 0000	-9. 5446
HSO4-	1. 250e-010	1. 213e-005	0. 9392	-9. 9303
NaAl O2(aq)	6. 392e-011	5. 238e-006	1. 0000	-10. 1944
PO4---	6. 235e-011	5. 920e-006	0. 5671	-10. 4515
UO2CO3(aq)	5. 790e-011	1. 910e-005	1. 0000	-10. 2373
Cu+	1. 065e-011	6. 765e-007	0. 9392	-11. 0000
NaOH(aq)	3. 745e-012	1. 498e-007	1. 0000	-11. 4265
Al +++	3. 191e-012	8. 608e-008	0. 5942	-11. 7221
AsO4---	2. 956e-012	4. 105e-007	0. 5671	-11. 7757
Sb(OH) 4-	2. 173e-012	4. 124e-007	0. 9392	-11. 6901
CdSO4(aq)	1. 867e-012	3. 892e-007	1. 0000	-11. 7288
H3PO4(aq)	1. 122e-012	1. 099e-007	1. 0000	-11. 9500
MgP2O7--	1. 089e-012	2. 158e-007	0. 7773	-12. 0725
Fe+++	7. 381e-013	4. 121e-008	0. 5942	-12. 3579
UO2OH+	6. 444e-013	1. 849e-007	0. 9392	-12. 2181
HNO3(aq)	2. 728e-013	1. 718e-008	1. 0000	-12. 5642
Sb(OH) 2+	1. 321e-013	2. 057e-008	0. 9392	-12. 9065
HP2O7---	1. 730e-014	3. 026e-009	0. 5671	-14. 0083
UO2++	6. 605e-015	1. 783e-009	0. 7791	-14. 2886
H3AsO4(aq)	4. 566e-015	6. 480e-010	1. 0000	-14. 3404
FeSO4+	2. 713e-015	4. 121e-010	0. 9392	-14. 5938
P2O7----	8. 939e-016	1. 555e-010	0. 3647	-15. 4868
H2P2O7--	8. 876e-016	1. 561e-010	0. 7773	-15. 1612
UO2SO4(aq)	4. 668e-016	1. 709e-010	1. 0000	-15. 3308
KHSO4(aq)	3. 031e-016	4. 126e-011	1. 0000	-15. 5184
KP2O7---	4. 888e-018	1. 041e-012	0. 5671	-17. 5572
UO2(SO4) 2--	4. 256e-019	1. 967e-013	0. 7773	-18. 4804
(UO2) 2(OH) 2++	1. 061e-019	6. 088e-014	0. 7791	-19. 0828
UO2+	8. 630e-021	2. 330e-015	0. 9392	-20. 0913
H3P2O7-	2. 644e-021	4. 678e-016	0. 9392	-20. 6050
(UO2) 3(CO3) 6(6-)	2. 610e-021	3. 053e-015	0. 1033	-21. 5692
(UO2) 3(OH) 5+	2. 321e-021	2. 077e-015	0. 9392	-20. 6615
HAsO2(aq)	2. 130e-024	2. 298e-019	1. 0000	-23. 6717
U(OH) 4(aq)	2. 030e-024	6. 210e-019	1. 0000	-23. 6926
As(OH) 3(aq)	1. 920e-024	2. 418e-019	1. 0000	-23. 7166
Mn+++	3. 094e-025	1. 700e-020	0. 5728	-24. 7514
H2AsO3-	3. 627e-026	4. 531e-021	0. 9392	-25. 4677
AsO2-	3. 530e-026	3. 773e-021	0. 9392	-25. 4795
H4P2O7(aq)	9. 736e-028	1. 732e-022	1. 0000	-27. 0116
AsO2OH--	1. 230e-029	1. 524e-024	0. 7773	-29. 0196
NO2-	5. 510e-031	2. 534e-026	0. 9381	-30. 2866
O2(aq)	9. 867e-033	3. 157e-028	1. 0000	-32. 0058
Formate	1. 940e-033	8. 729e-029	0. 9386	-32. 7398
H2(aq)	4. 371e-034	8. 809e-031	1. 0000	-33. 3594
Ca(For) +	3. 584e-035	3. 049e-030	0. 9392	-34. 4729
HN02(aq)	2. 153e-035	1. 012e-030	1. 0000	-34. 6670
Mg(For) +	1. 670e-035	1. 157e-030	0. 9392	-34. 8046
Na(For) (aq)	3. 499e-037	2. 379e-032	1. 0000	-36. 4561
Mn(For) +	2. 768e-037	2. 766e-032	0. 9392	-36. 5851
Formi c_ aci d(aq)	1. 731e-037	7. 966e-033	1. 0000	-36. 7617
Cu(For) +	1. 300e-037	1. 411e-032	0. 9392	-36. 9133
Fe(For) +	1. 227e-037	1. 237e-032	0. 9392	-36. 9384
K(For) (aq)	8. 674e-038	7. 295e-033	1. 0000	-37. 0618
SO3--	2. 497e-038	1. 998e-033	0. 7791	-37. 7111
Ba(For) +	1. 687e-038	3. 076e-033	0. 9392	-37. 8001

	SpecE8_output_GSS_Mi ne Water	Adi t	Pi pe	May. txt
H02-	8. 157e-039	2. 692e-034	0. 9392	-38. 1157
NH4+	5. 489e-039	9. 899e-035	0. 9375	-38. 2885
HS03-	4. 521e-039	3. 665e-034	0. 9392	-38. 3720
Cd(For)+	2. 608e-039	4. 105e-034	0. 9392	-38. 6109
Pb(For)+	1. 092e-039	2. 754e-034	0. 9392	-38. 9890
CO(aq)	2. 577e-040	7. 217e-036	1. 0000	-39. 5889
Oxal ate	5. 919e-041	5. 209e-036	0. 7773	-40. 3371
NH3(aq)	4. 516e-041	7. 689e-037	1. 0000	-40. 3453
Mn04--	2. 318e-041	2. 757e-036	0. 7773	-40. 7442
U(CO3)5(6-)	7. 387e-043	3. 974e-037	0. 1033	-43. 1173
UOH+++	3. 141e-043	8. 009e-038	0. 5728	-42. 7449
H-Oxal ate	1. 255e-044	1. 117e-039	0. 9392	-43. 9284
Mn04-	9. 037e-045	1. 075e-039	0. 9386	-44. 0715
S02(aq)	3. 348e-045	2. 145e-040	1. 0000	-44. 4752
HS05-	2. 528e-047	2. 858e-042	0. 9392	-46. 6244
US04++	1. 650e-047	5. 512e-042	0. 7791	-46. 8909
U(S04)2(aq)	6. 014e-048	2. 587e-042	1. 0000	-47. 2208
U++++	9. 945e-050	2. 367e-044	0. 3745	-49. 4289
Oxal i c_aci d(aq)	3. 391e-051	3. 052e-046	1. 0000	-50. 4697
S206--	1. 274e-063	2. 039e-058	0. 7773	-63. 0043
S208--	8. 538e-065	1. 640e-059	0. 7773	-64. 1780
Ca(For)2(aq)	5. 758e-067	7. 491e-062	1. 0000	-66. 2397
Mg(For)2(aq)	3. 575e-067	4. 087e-062	1. 0000	-66. 4467
U+++	7. 605e-068	1. 810e-062	0. 5728	-67. 3609
Mn(For)2(aq)	8. 183e-069	1. 186e-063	1. 0000	-68. 0871
Cu(For)2(aq)	7. 302e-069	1. 121e-063	1. 0000	-68. 1365
Fe(For)2(aq)	5. 140e-069	7. 497e-064	1. 0000	-68. 2890
Cd(For)2(aq)	7. 225e-070	1. 462e-064	1. 0000	-69. 1412
Na(For)2-	5. 405e-070	6. 108e-065	0. 9392	-69. 2944
Ba(For)2(aq)	1. 929e-070	4. 384e-065	1. 0000	-69. 7148
K(For)2-	1. 076e-070	1. 390e-065	0. 9392	-69. 9953
Pb(For)2(aq)	3. 413e-071	1. 014e-065	1. 0000	-70. 4668
Formal dehyde(aq)	2. 909e-072	8. 733e-068	1. 0000	-71. 5362
N3-	6. 511e-074	2. 735e-069	0. 9392	-73. 2136
HN3(aq)	7. 502e-077	3. 227e-072	1. 0000	-76. 1248
Cu(NH3)2++	1. 200e-079	1. 171e-074	0. 7791	-79. 0291
S205--	3. 761e-082	5. 420e-077	0. 7773	-81. 5341
Cd(NH3)2++	5. 970e-084	8. 742e-079	0. 7791	-83. 3325
Urea(aq)	5. 385e-085	3. 234e-080	1. 0000	-84. 2688
HCN(aq)	4. 433e-086	1. 198e-081	1. 0000	-85. 3533
CN-	4. 741e-088	1. 233e-083	0. 9381	-87. 3519
Methanol (aq)	1. 054e-089	3. 378e-085	1. 0000	-88. 9770
Gl ycol ate	7. 400e-096	5. 552e-091	0. 9392	-95. 1580
HS-	5. 650e-097	1. 868e-092	0. 9386	-96. 2755
Ca(Gl yc)+	2. 428e-097	2. 794e-092	0. 9392	-96. 6420
H2S(aq)	1. 665e-097	5. 672e-093	1. 0000	-96. 7787
Mg(Gl yc)+	4. 038e-098	4. 011e-093	0. 9392	-97. 4211
Cu(Gl yc)+	4. 810e-099	6. 664e-094	0. 9392	-98. 3451
Na(Gl yc)(aq)	1. 380e-099	1. 352e-094	1. 0000	-98. 8602
Fe(Gl yc)+	1. 143e-099	1. 496e-094	0. 9392	-98. 9691
Mn(Gl yc)+	8. 168e-100	1. 061e-094	0. 9392	-99. 1152
Gl ycol i c_aci d(aq)	8. 055e-100	6. 125e-095	1. 0000	-99. 0939
K(Gl yc)(aq)	3. 421e-100	3. 904e-095	1. 0000	-99. 4659
Ba(Gl yc)+	3. 103e-101	6. 589e-096	0. 9392	-100. 5355
Pb(Gl yc)+	1. 182e-101	3. 335e-096	0. 9392	-100. 9547
Cd(Gl yc)+	1. 108e-101	2. 076e-096	0. 9392	-100. 9828
Methane(aq)	1. 033e-101	1. 657e-097	1. 0000	-100. 9857
S--	1. 314e-102	4. 211e-098	0. 7808	-101. 9890
S204--	1. 884e-103	2. 413e-098	0. 7808	-102. 8325
S203--	1. 203e-103	1. 349e-098	0. 7773	-103. 0292
Acetate	2. 981e-107	1. 760e-102	0. 9397	-106. 5526
MgCH3COO+	1. 734e-109	1. 445e-104	0. 9392	-108. 7883
CaCH3COO+	1. 635e-109	1. 620e-104	0. 9392	-108. 8138

SpecE8_output_GSS_Mine Water Adit Pipe May. txt				
Acetic_acid(aq)	2. 615e-110	1. 570e-105	1. 0000	-109. 5825
CuCH3COO+	3. 837e-111	4. 702e-106	0. 9392	-110. 4433
NaCH3COO(aq)	3. 817e-111	3. 131e-106	1. 0000	-110. 4183
MnCH3COO+	1. 341e-111	1. 529e-106	0. 9392	-110. 8997
KCH3COO(aq)	6. 681e-112	6. 556e-107	1. 0000	-111. 1751
FeCH3COO+	4. 674e-112	5. 370e-107	0. 9392	-111. 3575
BaCH3COO+	1. 024e-112	2. 010e-107	0. 9392	-112. 0171
PbCH3COO+	6. 047e-113	1. 610e-107	0. 9392	-112. 2457
CdCH3COO+	5. 141e-113	8. 813e-108	0. 9392	-112. 3162
Malonate	3. 478e-116	3. 548e-111	0. 7773	-115. 5681
AlCH3COO++	1. 535e-116	1. 320e-111	0. 7791	-115. 9223
Cu(NH3)3++	6. 390e-117	7. 324e-112	0. 7791	-116. 3030
CuCH3COO(aq)	7. 317e-118	8. 968e-113	1. 0000	-117. 1357
H-Malonate	2. 126e-118	2. 191e-113	0. 9392	-117. 6996
AsH3(aq)	1. 731e-119	1. 349e-114	1. 0000	-118. 7618
Malonic_acid(aq)	2. 374e-123	2. 470e-118	1. 0000	-122. 6244
Methanamine(aq)	2. 260e-127	7. 018e-123	1. 0000	-126. 6459
Cu(Gly)+	4. 680e-130	6. 439e-125	0. 9392	-129. 3570
Glycine(aq)	3. 613e-130	2. 712e-125	1. 0000	-129. 4422
Mg(Gly)+	9. 526e-133	9. 368e-128	0. 9392	-132. 0483
Fe(Gly)+	1. 957e-134	2. 542e-129	0. 9392	-133. 7356
Mn(Gly)+	1. 468e-134	1. 894e-129	0. 9392	-133. 8604
Ca(Gly)+	1. 169e-134	1. 334e-129	0. 9392	-133. 9595
Pb(Gly)+	2. 886e-135	8. 115e-130	0. 9392	-134. 5670
Cd(Gly)+	1. 163e-135	2. 168e-130	0. 9392	-134. 9616
S3O6--	5. 976e-137	1. 148e-131	0. 7773	-136. 3330
Ba(Gly)+	8. 775e-138	1. 855e-132	0. 9392	-137. 0840
Acetaldehyde(aq)	7. 184e-144	3. 164e-139	1. 0000	-143. 1436
SCN-	4. 827e-145	2. 803e-140	0. 9386	-144. 3439
NH4CH3COO(aq)	2. 376e-145	1. 831e-140	1. 0000	-144. 6242
Acetamide(aq)	1. 852e-145	1. 094e-140	1. 0000	-144. 7323
UO2SCN+	6. 024e-158	1. 976e-152	0. 9392	-157. 2474
Cd(NH3)4++	5. 363e-162	9. 680e-157	0. 7791	-161. 3790
Ethanol(aq)	4. 852e-166	2. 235e-161	1. 0000	-165. 3141
Ethyne(aq)	1. 694e-166	4. 410e-162	1. 0000	-165. 7711
Ethylene(aq)	6. 148e-171	1. 724e-166	1. 0000	-170. 2113
Lactate	3. 651e-172	3. 251e-167	0. 9392	-171. 4648
Ca(Lac)+	7. 039e-174	9. 089e-169	0. 9392	-173. 1797
Mg(Lac)+	1. 952e-174	2. 212e-169	0. 9392	-173. 7368
Cu(Lac)+	1. 082e-175	1. 651e-170	0. 9392	-174. 9931
Na(Lac)(aq)	6. 996e-176	7. 839e-171	1. 0000	-175. 1551
Fe(Lac)+	5. 010e-176	7. 259e-171	0. 9392	-175. 3274
Lactic_acid(aq)	4. 140e-176	3. 728e-171	1. 0000	-175. 3830
Mn(Lac)+	2. 852e-176	4. 107e-171	0. 9392	-175. 5721
K(Lac)(aq)	1. 735e-176	2. 223e-171	1. 0000	-175. 7607
Ba(Lac)+	6. 212e-178	1. 406e-172	0. 9392	-177. 2340
Pb(Lac)+	5. 698e-178	1. 688e-172	0. 9392	-177. 2715
Cd(Lac)+	3. 697e-178	7. 447e-173	0. 9392	-177. 4594
Ethane(aq)	1. 737e-182	5. 223e-178	1. 0000	-181. 7601
Propanoate	8. 636e-186	6. 309e-181	0. 9392	-185. 0909
Ca(Prop)+	2. 646e-188	2. 994e-183	0. 9392	-187. 6046
Mg(Prop)+	1. 359e-188	1. 323e-183	0. 9392	-187. 8940
Propanoic_acid(aq)	1. 015e-188	7. 515e-184	1. 0000	-187. 9937
Na(Prop)(aq)	1. 633e-189	1. 568e-184	1. 0000	-188. 7870
Cu(Prop)+	1. 096e-189	1. 497e-184	0. 9392	-188. 9875
Mn(Prop)+	4. 600e-190	5. 887e-185	0. 9392	-189. 3645
K(Prop)(aq)	4. 048e-190	4. 540e-185	1. 0000	-189. 3927
Fe(Prop)+	3. 335e-190	4. 298e-185	0. 9392	-189. 5042
USCN+++	6. 104e-191	1. 807e-185	0. 5728	-190. 4564
Ca(Glyc)2(aq)	3. 748e-191	7. 126e-186	1. 0000	-190. 4262
Pb(Prop)+	1. 542e-191	4. 320e-186	0. 9392	-190. 8392
Cd(Prop)+	1. 376e-191	2. 552e-186	0. 9392	-190. 8886
Mg(Glyc)2(aq)	3. 972e-192	6. 925e-187	1. 0000	-191. 4010

	SpecE8_output_GSS_Mi	ne Water	Adi t	Pi pe	May. txt
Ba(Prop)+	3. 826e-192	8. 047e-187	0. 9392	-191. 4446	
Cu(Gl yc)2(aq)	3. 155e-192	6. 739e-187	1. 0000	-191. 5010	
Fe(Gl yc)2(aq)	6. 043e-193	1. 244e-187	1. 0000	-192. 2188	
Succinate	4. 341e-193	5. 038e-188	0. 7773	-192. 4718	
S4O6--	1. 171e-193	2. 625e-188	0. 7773	-193. 0410	
Mn(Gl yc)2(aq)	5. 152e-194	1. 056e-188	1. 0000	-193. 2880	
Na(Gl yc)2-	1. 281e-194	2. 218e-189	0. 9392	-193. 9195	
Pb(Gl yc)2(aq)	5. 241e-195	1. 872e-189	1. 0000	-194. 2806	
Cd(Gl yc)2(aq)	2. 846e-195	7. 470e-190	1. 0000	-194. 5457	
K(Gl yc)2-	2. 641e-195	4. 994e-190	0. 9392	-194. 6056	
H-Succinate	2. 605e-195	3. 050e-190	0. 9392	-194. 6114	
Serine(aq)	1. 190e-195	1. 250e-190	1. 0000	-194. 9246	
Ba(Gl yc)2(aq)	1. 052e-195	3. 022e-190	1. 0000	-194. 9781	
Succinic_acid(aq)	7. 109e-199	8. 393e-194	1. 0000	-198. 1482	
Ethanediamine(aq)	1. 084e-205	4. 886e-201	1. 0000	-204. 9650	
Alanine(aq)	2. 421e-207	2. 157e-202	1. 0000	-206. 6159	
Cu(Al a)+	7. 111e-208	1. 078e-202	0. 9392	-207. 1753	
Fe(Al a)+	1. 651e-211	2. 375e-206	0. 9392	-210. 8096	
Mg(Al a)+	4. 045e-212	4. 545e-207	0. 9392	-211. 4203	
Ca(Al a)+	1. 479e-212	1. 895e-207	0. 9392	-211. 8574	
Mn(Al a)+	8. 039e-213	1. 149e-207	0. 9392	-212. 1221	
Cd(Al a)+	4. 380e-213	8. 781e-208	0. 9392	-212. 3857	
Pb(Al a)+	1. 507e-213	4. 448e-208	0. 9392	-212. 8492	
Ca(CH3COO)2(aq)	9. 764e-215	1. 544e-209	1. 0000	-214. 0104	
Mg(CH3COO)2(aq)	4. 630e-215	6. 592e-210	1. 0000	-214. 3344	
Cu(CH3COO)2(aq)	4. 342e-216	7. 885e-211	1. 0000	-215. 3623	
Ba(Al a)+	4. 046e-216	9. 119e-211	0. 9392	-215. 4202	
Mn(CH3COO)2(aq)	3. 570e-217	6. 177e-212	1. 0000	-216. 4473	
Fe(CH3COO)2(aq)	3. 452e-217	6. 004e-212	1. 0000	-216. 4619	
Acetone(aq)	2. 185e-217	1. 269e-212	1. 0000	-216. 6605	
Na(CH3COO)2-	7. 318e-218	1. 032e-212	0. 9392	-217. 1628	
Aspartic_acid(aq)	4. 130e-218	5. 496e-213	1. 0000	-217. 3840	
Cd(CH3COO)2(aq)	3. 488e-218	8. 039e-213	1. 0000	-217. 4574	
Pb(CH3COO)2(aq)	1. 355e-218	4. 406e-213	1. 0000	-217. 8682	
Ba(CH3COO)2(aq)	8. 899e-219	2. 273e-213	1. 0000	-218. 0506	
K(CH3COO)2-	8. 102e-219	1. 273e-213	0. 9392	-218. 1187	
Al (CH3COO)2+	6. 315e-221	9. 160e-216	0. 9392	-220. 2269	
Propanal (aq)	7. 798e-222	4. 528e-217	1. 0000	-221. 1080	
Cu(CH3COO)2-	3. 556e-224	6. 457e-219	0. 9392	-223. 4763	
1-Propyne(aq)	9. 638e-241	3. 861e-236	1. 0000	-240. 0160	
S3--	4. 377e-244	4. 210e-239	0. 7773	-243. 4682	
1-Propanol (aq)	1. 611e-244	9. 681e-240	1. 0000	-243. 7928	
1-Propene(aq)	3. 898e-247	1. 640e-242	1. 0000	-246. 4091	
2-Hydroxybutanoic acid	4. 230e-251	4. 360e-246	0. 9392	-250. 4009	
NH4(CH3COO)2-	7. 423e-252	1. 010e-246	0. 9392	-251. 1567	
Asparagine(aq)	3. 394e-253	4. 483e-248	1. 0000	-252. 4693	
Cu(Gl y)2(aq)	2. 240e-254	4. 740e-249	1. 0000	-253. 6498	
2-Hydroxybutanone	4. 237e-255	4. 410e-250	1. 0000	-254. 3729	
Propane(aq)	2. 255e-261	9. 940e-257	1. 0000	-260. 6469	
Mg(Gl y)2(aq)	1. 525e-261	2. 630e-256	1. 0000	-260. 8166	
Diglycine(aq)	2. 789e-262	3. 684e-257	1. 0000	-261. 5546	
Fe(Gl y)2(aq)	4. 508e-263	9. 193e-258	1. 0000	-262. 3460	
Mn(Gl y)2(aq)	1. 755e-263	3. 564e-258	1. 0000	-262. 7556	
Cd(Gl y)2(aq)	9. 075e-264	2. 364e-258	1. 0000	-263. 0422	
Pb(Gl y)2(aq)	6. 749e-264	2. 397e-258	1. 0000	-263. 1708	
Diketopiperazine	8. 729e-265	9. 958e-260	1. 0000	-264. 0590	
Butanoate	7. 932e-265	6. 907e-260	0. 9392	-264. 1279	
Ca(Gl y)2(aq)	5. 559e-266	1. 046e-260	1. 0000	-265. 2550	
Ca(But)+	1. 664e-267	2. 116e-262	0. 9392	-266. 8060	
Mg(But)+	8. 142e-268	9. 069e-263	0. 9392	-267. 1165	
Butanoic_acid(aq)	7. 367e-268	6. 489e-263	1. 0000	-267. 1327	
Na(But)(aq)	1. 449e-268	1. 595e-263	1. 0000	-267. 8389	
Cu(But)+	7. 977e-269	1. 201e-263	0. 9392	-268. 1254	

	SpecE8_output_GSS_Mi ne Water	Adi t	Pi pe	May. txt
Ba(Gl y)2(aq)	4. 399e-269	1. 255e-263	1. 0000	-268. 3566
Mn(But)+	3. 611e-269	5. 129e-264	0. 9392	-268. 4696
K(But)(aq)	3. 593e-269	4. 534e-264	1. 0000	-268. 4445
Fe(But)+	3. 186e-269	4. 554e-264	0. 9392	-268. 5240
Cd(But)+	5. 301e-271	1. 057e-265	0. 9392	-270. 3029
Pb(But)+	3. 909e-271	1. 150e-265	0. 9392	-270. 4351
Gl utarate	3. 420e-271	4. 449e-266	0. 7773	-270. 5753
Ba(But)+	2. 346e-271	5. 264e-266	0. 9392	-270. 6569
H-Gl utarate	1. 176e-273	1. 542e-268	0. 9392	-272. 9568
Threoni ne(aq)	1. 113e-274	1. 326e-269	1. 0000	-273. 9534
Gl utari c_aci d(aq)	3. 983e-277	5. 261e-272	1. 0000	-276. 3998
Ethyl acetate(aq)	3. 905e-277	3. 440e-272	1. 0000	-276. 4083
S506--	1. 721e-280	4. 412e-275	0. 7773	-279. 8735
1-Propanami ne(aq)	1. 220e-283	7. 211e-279	1. 0000	-282. 9136
a-Ami nobutyri c_a	4. 814e-286	4. 963e-281	1. 0000	-285. 3175
Gl utami c_aci d(aq)	3. 644e-295	5. 361e-290	1. 0000	-294. 4384
Butanal (aq)	2. 332e-302	1. 681e-297	1. 0000	-300. 0000
UO2(SCN)2(aq)	1. 453e-302	5. 611e-297	1. 0000	-300. 0000
S4--	4. 512e-315	5. 787e-310	0. 7773	-300. 0000
1-Butyne(aq)	8. 257e-320	4. 465e-315	1. 0000	-300. 0000
Cu(CH3COO)3-	5. 830e-321	1. 403e-315	0. 9392	-300. 0000
Pb(CH3COO)3-	5. 435e-323	2. 088e-317	0. 9392	-300. 0000
Mn(CH3COO)3-	4. 941e-323	1. 146e-317	0. 9392	-300. 0000
1-Butanol (aq)	4. 941e-324	3. 661e-319	1. 0000	-300. 0000
Cd(CH3COO)3-	4. 941e-324	1. 430e-318	0. 9392	-300. 0000
Heptanoate	0. 0000	0. 0000	0. 9392	-300. 0000
Heptanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
HSb2S4-	0. 0000	0. 0000	0. 9392	-300. 0000
Cd(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Cd(Al a)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Al anyl gl yci ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
p-Tol ui c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
p-Tol uate	0. 0000	0. 0000	0. 9392	-300. 0000
2-Hexanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexyne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
o-Tol uate	0. 0000	0. 0000	0. 9392	-300. 0000
o-Phthal ate	0. 0000	0. 0000	0. 7773	-300. 0000
n-Propyl benzene(0. 0000	0. 0000	1. 0000	-300. 0000
n-Pentane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Octyl benzene(a	0. 0000	0. 0000	1. 0000	-300. 0000
n-Octane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Hexane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Heptane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Butane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
m-Tol ui c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
m-Tol uate	0. 0000	0. 0000	0. 9392	-300. 0000
2-Heptanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Val i ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Undecanoi c_aci d(0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Pent)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Undecanoate	0. 0000	0. 0000	0. 9392	-300. 0000
Ca(Pent)+	0. 0000	0. 0000	0. 9392	-300. 0000
Ca(Lac)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
H-Suberate	0. 0000	0. 0000	0. 9392	-300. 0000
UO2(SCN)3-	0. 0000	0. 0000	0. 9392	-300. 0000
H-Sebacate	0. 0000	0. 0000	0. 9392	-300. 0000
2-Butanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
H-Pi mel ate	0. 0000	0. 0000	0. 9392	-300. 0000
U(SCN)2++	0. 0000	0. 0000	0. 7791	-300. 0000
H-Azel ate	0. 0000	0. 0000	0. 9392	-300. 0000
Tryptophan(aq)	0. 0000	0. 0000	1. 0000	-300. 0000

	SpecE8_output_GSS_Mi ne Water	Adi t	Pi pe	May. txt
Tol uene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
H-Adi pate	0. 0000	0. 0000	0. 9392	-300. 0000
Suberi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Suberate	0. 0000	0. 0000	0. 7773	-300. 0000
Sebaci c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Sebacate	0. 0000	0. 0000	0. 7773	-300. 0000
Sb2S4--	0. 0000	0. 0000	0. 7773	-300. 0000
Adi pi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Gl utami ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
S5--	0. 0000	0. 0000	0. 7773	-300. 0000
Ca(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Adi pate	0. 0000	0. 0000	0. 7773	-300. 0000
1-Hexanol (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Al a)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pi mel i c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Fe(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pi mel ate	0. 0000	0. 0000	0. 7773	-300. 0000
Phenyl al ani ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Fe(Pent)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pentanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Fe(Pent)+	0. 0000	0. 0000	0. 9392	-300. 0000
Fe(Lac)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pb(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pb(Pent)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pb(Pent)+	0. 0000	0. 0000	0. 9392	-300. 0000
Pb(Lac)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexanami ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Butene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pb(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pb(Al a)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Fe(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Fe(Al a)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Benzoi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Octanoi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Octanoate	0. 0000	0. 0000	0. 9392	-300. 0000
Octanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Benzoate	0. 0000	0. 0000	0. 9412	-300. 0000
2-Pentanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Pentyne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Nonanoi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Nonanoate	0. 0000	0. 0000	0. 9392	-300. 0000
Nonanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ethyl benzene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Benzene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Octanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Na(Prop)2-	0. 0000	0. 0000	0. 9392	-300. 0000
Pentanoate	0. 0000	0. 0000	0. 9392	-300. 0000
Pentanoi c_aci d(a	0. 0000	0. 0000	1. 0000	-300. 0000
Phenol (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Na(Pent)2-	0. 0000	0. 0000	0. 9392	-300. 0000
Dodecanoi c_aci d(0. 0000	0. 0000	1. 0000	-300. 0000
Na(Pent) (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Na(Lac)2-	0. 0000	0. 0000	0. 9392	-300. 0000
Dodecanoate	0. 0000	0. 0000	0. 9392	-300. 0000
2-Hydroxypentano	0. 0000	0. 0000	1. 0000	-300. 0000
1-Pentene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Heptyne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ba(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Decanoi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Na(But)2-	0. 0000	0. 0000	0. 9392	-300. 0000
Decanoate	0. 0000	0. 0000	0. 9392	-300. 0000
2-Hydroxypentano	0. 0000	0. 0000	0. 9392	-300. 0000
Decanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000

SpecE8_output_GSS_Mine Water Adit Pipe May. txt

Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)+	0.0000	0.0000	0.9392	-300.0000
2-Hydroxyoctanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanol(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoa	0.0000	0.0000	0.9392	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)+	0.0000	0.0000	0.9392	-300.0000
Cu(Pent)+	0.0000	0.0000	0.9392	-300.0000
Mn(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9392	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoi	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Tyrosine(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)+	0.0000	0.0000	0.9392	-300.0000
Mg(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoa	0.0000	0.0000	0.9392	-300.0000
Cu(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Methionine(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Leucylglycine(aq)	0.0000	0.0000	1.0000	-300.0000
Leucine(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Azelate	0.0000	0.0000	0.7773	-300.0000
2-Hydroxyheptano	0.0000	0.0000	0.9392	-300.0000
Cd(Pent)+	0.0000	0.0000	0.9392	-300.0000
Azelai c_acid(aq)	0.0000	0.0000	1.0000	-300.0000
K(Prop)2-	0.0000	0.0000	0.9392	-300.0000
Cd(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoi	0.0000	0.0000	1.0000	-300.0000
1-Octanol(aq)	0.0000	0.0000	1.0000	-300.0000
K(Pent)2-	0.0000	0.0000	0.9392	-300.0000
K(Pent)(aq)	0.0000	0.0000	1.0000	-300.0000
K(Lac)2-	0.0000	0.0000	0.9392	-300.0000
n-Butyl benzene(a	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecano	0.0000	0.0000	0.9392	-300.0000
n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
K(But)2-	0.0000	0.0000	0.9392	-300.0000
n-Hexyl benzene(a	0.0000	0.0000	1.0000	-300.0000
Isoleucine(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9392	-300.0000
n-Pentyl benzene(0.0000	0.0000	1.0000	-300.0000
Hexanal(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_Mine Water Adit Pipe May. txt

Heptanoic acid(a	0.0000	0.0000	1.0000	-300.0000
1-Octanamine(aq)	0.0000	0.0000	1.0000	-300.0000
o-Toluidic acid(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butanamine(aq)	0.0000	0.0000	1.0000	-300.0000
(NH4)2Sb2S4(aq)	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states

log Q/K

log Q/K

Nontronite-Ca	25.0427s/sat	UO3(alpha)	-8.3706
Nontronite-Mg	24.9716s/sat	Monteponite	-8.5601
Nontronite-K	24.5663s/sat	Pb3S06	-8.6065
Nontronite-Na	24.3425s/sat	Cu	-8.7528
Nontronite-H	24.0497s/sat	Tephroite	-8.9353
Hematite	20.4604s/sat	BaU04	-9.3011
Ferriite-Cu	18.5653s/sat	Manganosite	-9.4198
Magnetite	18.4744s/sat	Mirabilite	-9.4813
Cronstedtite-7A	17.9267s/sat	Salteite	-9.8912
Delafossite	14.0424s/sat	Nitrobarite	-9.9915
Clinoptilolite-h	12.3715s/sat	CdS04: 2.667H2O	-10.1361
Clinoptilolite-C	12.2620s/sat	Forsterite	-10.2820
Andradite	10.7947s/sat	CdS04: H2O	-10.4253
Muscovite	10.5320s/sat	Arcanite	-10.4284
Stilbite	10.4023s/sat	Gyrolite	-10.6804
Epidote-ord	10.1308s/sat	UO2.25	-10.8904
Epidote	10.1295s/sat	Mg1.25S04(OH)0.5	-10.9546
Ferriite-Ca	9.8103s/sat	UO2.25(beta)	-10.9614
Goethite	9.7727s/sat	Thenardite	-11.1754
Ferriite-Mg	9.6584s/sat	Periclase	-11.1893
Clinoptilolite-K	8.4743s/sat	Monticellite	-11.4458
Clinoptilolite-h	8.0349s/sat	Torbernite	-11.5090
Beidellite-Ca	7.6552s/sat	Natron	-11.7039
Beidellite-Mg	7.5839s/sat	Portlandite	-11.7178
Illite	7.4350s/sat	Uranocircite	-11.8268
Beidellite-K	7.1789s/sat	MgU04	-12.1325
Kaolinite	7.0715s/sat	(UO2)3(P04)2: 4H2	-12.2027
Beidellite-Na	6.9548s/sat	CdS04	-12.2064
Mesolite	6.7642s/sat	Mg1.5S04(OH)	-12.3070
Paragonite	6.7399s/sat	Na2C03: 7H2O	-12.3466
Beidellite-H	6.6629s/sat	Tremolite	-12.3726
Montmor-Mg	6.4556s/sat	BaU207	-12.5034
Montmor-Ca	6.4501s/sat	UO2S04: H2O	-12.5921
Pyrophyllite	6.3612s/sat	MnS04	-12.9331
Montmor-K	6.0505s/sat	Pb4S07	-12.9596
Scolecite	5.9021s/sat	Przhevalskite	-13.3000
Montmor-Na	5.8209s/sat	MgS04	-13.4756
Clinoptilolite-N	5.7142s/sat	FeS04	-13.5631
Clinoptilolite-h	5.7111s/sat	UP05	-13.6010
Jarosite	4.7575s/sat	UO2.6667	-13.6474
Laumontite	4.5183s/sat	Thermonatrite	-13.9487
Fe(OH)3	4.3894s/sat	Chalcocyanite	-13.9597
Smectite-low-Fe-	4.2255s/sat	Na2U207	-14.0325
Margarite	4.1596s/sat	Bassettite	-14.1918
Maximilianiite	3.6325s/sat	Na2C03	-14.3097
K-Feldspar	3.6299s/sat	Uraniinite	-14.3327
Smectite-high-Fe	3.4151s/sat	Mordenite-dehy	-14.3708
Celadonite	3.3806s/sat	Coffinite	-14.6309
Diaspore	3.2107s/sat	Autunite-H	-15.0106
Lawsonite	2.9374s/sat	Ca3(As04)2	-15.4689
Boehmite	2.7736s/sat	Ferriite-Dicalcium	-15.6348
Gibbsite	2.7126s/sat	Ba(OH)2: 8H2O	-16.0472
Sandine-high	2.3013s/sat	CaAl204	-16.4043

SpecE8_output_GSS_Mi ne Water Adi t Pi pe May. txt

Corundum	2. 0324s/sat	U02S04: 3. 5H2O	-17. 0391
Mordeni te	1. 9731s/sat	U02S04: 2. 5H2O	-17. 1233
Amesi te-14A	1. 8412s/sat	U02S04: 3H2O	-17. 1976
Wi theri te	1. 7309s/sat	NaU03	-17. 3036
Chamosi te-7A	1. 1799s/sat	Hydromagnesi te	-17. 3749
Kyani te	1. 1130s/sat	Ba2Si 308	-17. 5943
Mal achi te	1. 0846s/sat	Gehl eni te	-17. 8474
Tenori te	0. 8681s/sat	CaAl 407	-18. 0639
Andal usi te	0. 7917s/sat	Hi l l ebrandi te	-18. 3351
Al bi te_low	0. 7874s/sat	Natrosi l i te	-18. 4339
Al bi te	0. 7874s/sat	Al um-K	-18. 7561
Strengi te	0. 6743s/sat	Di cal ci um_si l i ca	-18. 8555
Quartz	0. 5530s/sat	Na2Si 03	-19. 0713
Daphni te-14A	0. 4760s/sat	Ni ngyoi te	-19. 0990
Si l l i mani te	0. 3895s/sat	Anthophyl l i te	-19. 7305
Tri dymi te	0. 3539s/sat	Akermani te	-19. 7490
Brochanti te	0. 2821s/sat	U02. 3333(beta)	-19. 9887
Chal cedony	0. 2634s/sat	Larni te	-20. 2614
Cerussi te	0. 1974s/sat	U02S04	-21. 1047
Anal ci me	0. 1801s/sat	Li me	-22. 5094
Bari te	0. 0216s/sat	Na2U04(al pha)	-22. 7871
Prehni te	-0. 0387	U02(N03)2: 6H2O	-22. 9201
Cri stobal i te(al p	-0. 0421	Pargasi te	-23. 9668
I ce	-0. 0677	U02(N03)2: 3H2O	-24. 6588
Hercyni te	-0. 1768	U02(P03)2	-25. 6526
Dol omi te-ord	-0. 2138	Ranki ni te	-26. 0615
Dol omi te	-0. 2139	U02(N03)2: 2H2O	-26. 0718
Dawsoni te	-0. 2410	Cl i nopti l ol i te-d	-26. 3793
Coesi te	-0. 3087	(U02)3(P04)2	-26. 3802
Cl i nozoi si te	-0. 4135	(U02)2P207	-26. 4805
Cal ci te	-0. 4275	Pb	-27. 0209
Zoi si te	-0. 4628	K2U04	-27. 7164
Wai raki te	-0. 5176	U(HP04)2: 4H2O	-27. 7417
Cri stobal i te(bet	-0. 5419	Tobermori te-14A	-27. 8232
Aragoni te	-0. 5729	Ettringi te	-28. 1840
Al bi te_hi gh	-0. 6624	UP207	-28. 2387
Anni te	-0. 6779	KAl (S04)2	-28. 7724
Rhodochrosi te	-0. 7611	U02(N03)2: H2O	-30. 0046
Kal si l i te	-0. 8712	Cl i nopti l ol i te-d	-30. 2167
Si 02(am)	-0. 8955	Tobermori te-11A	-30. 2174
Natrol i te	-0. 9034	Foshagi te	-32. 2352
Di optase	-0. 9471	Merwi ni te	-32. 2543
Hydroxyl apati te	-1. 1593	Ba2Si 04	-32. 8711
Saponi te-Ca	-1. 1598	Cl i nopti l ol i te-d	-32. 9452
Monohydrocal ci te	-1. 2201	U02(N03)2	-33. 7681
Saponi te-Mg	-1. 2287	Tobermori te-9A	-34. 1824
Ri pi dol i te-14A	-1. 3066	Afwil l i te	-34. 3803
Jadei te	-1. 3863	As205	-35. 5596
Si deri te	-1. 4259	U02S03	-35. 8998
Magnesi te	-1. 5371	Cd	-37. 0527
Saponi te-K	-1. 6361	Fe	-38. 0312
Phl ogopi te	-1. 6971	Al 2(S04)3: 6H2O	-39. 1760
Saponi te-Na	-1. 8601	(U02)2As207	-39. 7819
Dol omi te-di s	-1. 9041	U02(As03)2	-39. 8990
Anorthi te	-2. 1132	(U02)3(As04)2	-40. 5773
Saponi te-H	-2. 1523	Ba0	-41. 8282
Al amosi te	-2. 2519	Fe2(S04)3	-42. 4304
Mi nesotai te	-2. 3453	U(S04)2: 8H2O	-45. 0996
Tal c	-2. 4657	Hatruri te	-45. 2991
Ferrosi l i te	-2. 5575	Ba2U207	-45. 5832
Whi tlocki te	-2. 5851	Arsenol i te	-45. 6614
Gypsum	-2. 6240	Cl audeti te	-45. 7437
Al uni te	-2. 7944	U(S04)2: 4H2O	-46. 5184

SpecE8_output_GSS_Mine Water Adit Pipe May.txt

Anhydrite	-2.9900	Xonotlite	-46.6010
Clinochlore-14A	-3.0168	U(SO ₄) ₂	-46.8520
Daphnite-7A	-3.0758	C	-47.4561
Berlinite	-3.5259	U(CO ₃) ₂	-49.1106
Nepheline	-3.5613	As	-55.2266
Bassanite	-3.6413	Al ₂ (SO ₄) ₃	-58.3649
CaUO ₄	-3.8246	Ca ₄ Al ₂ Fe ₂ O ₁₀	-59.1310
CaSO ₄ ·0.5H ₂ O(bet)	-3.8343	Na	-59.9023
Anglesite	-3.9922	Na ₃ UO ₄	-60.0562
Bixbyite	-4.0310	Na ₄ SiO ₄	-62.3974
Azurite	-4.0395	Ca ₃ Al ₂ O ₆	-62.4908
Greenalite	-4.0787	Mn	-62.9456
Sb ₂ O ₃	-4.1148	Na ₂ O	-63.6192
UO ₃ ·2H ₂ O	-4.1335	K	-64.1775
Schoepite	-4.1349	Covelite	-70.7152
Cuprite	-4.1741	S	-71.8310
CdSiO ₃	-4.1923	Ba ₃ UO ₆	-72.2318
Enstatite	-4.2172	Chalcocite	-73.2684
UO ₂ (OH) ₂ (beta)	-4.3167	Antigorite	-75.0871
Lanarkite	-4.3221	CdS	-79.5439
Schoepite-dehydrate	-4.3740	Galena	-80.7365
UO ₃ ·9H ₂ O(al pha)	-4.3897	K ₂ O	-82.1822
Rhodonite	-4.3925	U(SO ₃) ₂	-87.7685
Schoepite-dehydrate	-4.4768	PbSO ₄ (NH ₃) ₂	-90.3497
Schoepite-dehydrate(1)	-4.4971	Troilite	-90.9049
Wustite	-4.5400	Na ₆ Si ₂ O ₇	-90.9328
Pyrolusite	-4.5436	As ₄ O ₆ (mono)	-91.0017
Nesquehonite	-4.7212	Pyrrhotite	-91.0106
Ripidolite-7A	-4.8476	As ₄ O ₆ (cubi)	-91.1965
FeO	-5.3785	Alabandite	-93.9872
UO ₂ CO ₃	-5.4563	Mg	-103.4671
Rutherfordine	-5.4640	BaS	-112.2267
Chrysotile	-5.6066	UN1.73(al pha)	-112.9211
Fe(OH) ₂	-5.6283	UN1.59(al pha)	-115.8029
Corderite-hydr	-5.6512	Ca	-121.2558
Schoepite-dehydrate	-5.6921	P	-123.4337
Brucite	-5.6942	Al	-125.6256
Diopside	-5.9275	Ba	-125.7723
Fayalite	-5.9319	UN	-129.5825
Okenite	-5.9466	Si	-131.0195
Litharge	-5.9581	Pb(N ₃) ₂ (orth)	-144.8163
PbCO ₃ ·PbO	-6.0532	Pb(N ₃) ₂ (mono)	-145.2794
Hausmannite	-6.1352	Pyrite	-151.5805
Wollastonite	-6.1365	Chalcopyrite	-155.3169
Massicot	-6.1604	PbSO ₄ (NH ₃) ₄	-174.5523
Schoepite-dehydrate	-6.2531	U	-178.7457
Nahcolite	-6.3395	UAs	-189.5013
Pseudowollastonite	-6.4109	US	-190.5149
Mn(OH) ₂ (am)	-6.5337	UC	-207.7075
Clinochlore-7A	-6.5523	UH ₃ (beta)	-210.0086
Analcite-dehydrate	-6.6650	US1.9	-219.8957
Huntite	-6.7519	US ₂	-225.2422
NaFeO ₂	-6.7917	Mayenite	-230.5117
UO ₂ HP0 ₄ ·4H ₂ O	-6.8136	UAs ₂	-241.9100
Niter	-6.9625	UP	-252.3783
Hedenbergite	-6.9766	UC1.94(al pha)	-254.4803
Cd(OH) ₂	-7.0157	Stibnite	-269.0514
Spinel	-7.0605	US ₃	-293.6891
UO ₃ (gamma)	-7.3624	Orpiment	-294.3908
Chalcanthite	-7.5197	Pb(SCN) ₂	-296.6520
Melanterite	-7.5785	Bornite	-301.5776
Sepiolite	-7.6417	UP ₂	-370.3874
Pb ₂ SiO ₄	-7.9534	U ₂ Si ₃	-408.3817

SpecE8_output_GSS_Mine Water Adit Pipe May. txt

Artinite	-8.0026	o-Phthalic_acid	-416.3540
UO3(beta)	-8.0079	U2C3	-464.5062
Grossular	-8.0841	U3As4	-621.4843
Sanbornite	-8.3303	U3S5	-628.7206
Corderite_anhyd	-8.3353	U3P4	-875.1448
Ca-Al_Pyroxene	-8.3694		

Gases	fugacity	log fug.
N2(g)	0.09304	-1.031
H2O(g)	0.007689	-2.114
CO2(g)	0.001095	-2.961
NO2(g)	6.132e-013	-12.212
O2(g)	5.246e-030	-29.280
H2(g)	4.775e-031	-30.321
NO(g)	5.481e-032	-31.261
CO(g)	1.786e-037	-36.748
NH3(g)	2.698e-043	-42.569
SO2(g)	1.036e-045	-44.985
Cd(g)	1.213e-052	-51.916
Pb(g)	1.700e-058	-57.769
Cu(g)	1.196e-065	-64.922
Na(g)	2.153e-075	-74.667
UO3(g)	2.169e-076	-75.664
K(g)	1.466e-076	-75.834
H2S(g)	9.378e-097	-96.028
CH4(g)	4.733e-099	-98.325
UO2(g)	8.792e-119	-118.056
Mg(g)	1.174e-125	-124.930
Ca(g)	2.499e-149	-148.602
S2(g)	8.252e-160	-159.083
C2H4(g)	7.670e-169	-168.115
C(g)	2.584e-174	-173.588
Al(g)	5.635e-181	-180.249
UO(g)	8.375e-195	-194.077
Si(g)	3.771e-208	-207.424
U(g)	2.195e-271	-270.659

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
Al+++	3.25e-006	3.25e-006	0.0877			
Ba++	4.97e-007	4.97e-007	0.0682			
Ca++	0.000778	0.000778	31.2			
Cd++	1.91e-008	1.91e-008	0.00214			
Cu++	6.29e-007	6.29e-007	0.0400			
Fe++	8.73e-007	8.73e-007	0.0487			
H+	5.08e-005	5.08e-005	0.0512			
H2AsO4-	2.60e-008	2.60e-008	0.00367			
H2O	55.5	55.5	1.00e+006			
HCO3-	0.00153	0.00153	93.6			
HP04--	3.02e-006	3.02e-006	0.290			
K+	4.99e-005	4.99e-005	1.95			
Mg++	0.000281	0.000281	6.82			
Mn++	2.89e-006	2.89e-006	0.159			
NH3(aq)	0.000172	0.000172	2.93			
NO3-	0.000556	0.000556	34.5			
Na+	0.000170	0.000170	3.90			
O2(aq)	0.000129	0.000129	4.12			
Pb++	9.41e-009	9.41e-009	0.00195			
SO4--	0.000173	0.000173	16.6			
Sb(OH)3(aq)	1.20e-007	1.20e-007	0.0207			
SiO2(aq)	0.000125	0.000125	7.51			

U02++ SpecE8_output_GSS_Mine Water Adit Pipe May. txt
5.33e-009 5.33e-009 0.00144

Elemental composition	total moles		In fluid		Sorbed	
			moles	mg/kg	moles	mg/kg
Aluminium	3.252e-006		3.252e-006	0.08773		
Antimony	1.201e-007		1.201e-007	0.01462		
Arsenic	2.603e-008		2.603e-008	0.001950		
Barium	4.970e-007		4.970e-007	0.06824		
Cadmium	1.908e-008		1.908e-008	0.002145		
Calcium	0.0007785		0.0007785	31.19		
Carbon	0.001534		0.001534	18.42		
Copper	6.291e-007		6.291e-007	0.03997		
Hydrogen	111.0		111.0	1.119e+005		
Iron	8.729e-007		8.729e-007	0.04874		
Lead	9.411e-009		9.411e-009	0.001950		
Magnesium	0.0002808		0.0002808	6.824		
Manganese	2.893e-006		2.893e-006	0.1589		
Nitrogen	0.0007280		0.0007280	10.19		
Oxygen	55.52		55.52	8.880e+005		
Phosphorus	3.022e-006		3.022e-006	0.09358		
Potassium	4.987e-005		4.987e-005	1.950		
Silicon	0.0001250		0.0001250	3.509		
Sodium	0.0001696		0.0001696	3.899		
Sulfur	0.0001725		0.0001725	5.532		
Uranium	5.325e-009		5.325e-009	0.001267		

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt

Temperature = 25.0 C Pressure = 1.013 bars
 pH = 7.300 log fO2 = -30.442
 Eh = 0.3470 volts pe = 5.8659
 Ionic strength = 0.002847
 Charge imbalance = 0.000124 eq/kg (6.787% error)
 Activity of water = 1.000000
 Solvent mass = 1.000000 kg
 Solution mass = 1.000163 kg
 Solution density = 1.013 g/cm3
 Chlorinity = 0.000000 molal
 Dissolved solids = 163 mg/kg sol'n
 Elect. conductivity = 172.28 uS/cm (or umho/cm)
 Hardness = 96.28 mg/kg sol'n as CaCO3
 carbonate = 80.74 mg/kg sol'n as CaCO3
 non-carbonate = 15.54 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 80.74 mg/kg sol'n as CaCO3
 Water type = Ca-HCO3

Nernst redox couples

Eh (volts)

pe

e- + .25*O2(aq) + H+ = .5*H2O	0.3470	5.8659
8*e- + 9*H+ + NO3- = 3*H2O + NH3(aq)	0.5596	9.4597

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
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HCO3-	0.001596	97.35	0.9432	-2.8224
Ca++	0.0006747	27.04	0.7968	-3.2695
Mg++	0.0002651	6.443	0.8024	-3.6721
CO2(aq)	0.0001668	7.341	1.0000	-3.7777
SiO2(aq)	0.0001295	7.781	1.0000	-3.8876
SO4--	7.112e-005	6.830	0.7909	-4.2499
NO3-	2.883e-005	1.787	0.9423	-4.5660
CaHCO3+	9.553e-006	0.9656	0.9432	-5.0452
CaSO4(aq)	3.905e-006	0.5316	1.0000	-5.4083
MgHCO3+	3.686e-006	0.3144	0.9432	-5.4588
MgSO4(aq)	3.088e-006	0.3716	1.0000	-5.5103
N2(aq)	2.390e-006	0.06693	1.0000	-5.6217
CO3--	1.778e-006	0.1067	0.7924	-5.8512
CaCO3(aq)	1.608e-006	0.1609	1.0000	-5.7936
Mn++	8.329e-007	0.04575	0.7968	-6.1781
CuCO3(aq)	5.360e-007	0.06621	1.0000	-6.2709
HSiO3-	3.057e-007	0.02356	0.9432	-6.5401
MgCO3(aq)	2.855e-007	0.02407	1.0000	-6.5444
OH-	2.140e-007	0.003640	0.9428	-6.6951
Zn++	2.114e-007	0.01382	0.7968	-6.7735
CaNO3+	7.760e-008	0.007920	0.9432	-7.1355
Cu++	5.293e-008	0.003363	0.7968	-7.3750
H+	5.290e-008	5.331e-005	0.9474	-7.3000
CuOH+	4.602e-008	0.003706	0.9432	-7.3625
MnCO3(aq)	3.096e-008	0.003558	1.0000	-7.5093
MnSO4(aq)	8.413e-009	0.001270	1.0000	-8.0751
MnHCO3+	8.063e-009	0.0009349	0.9432	-8.1189
ZnHCO3+	7.071e-009	0.0008937	0.9432	-8.1759
ZnOH+	3.907e-009	0.0003219	0.9432	-8.4335
ZnSO4(aq)	1.918e-009	0.0003096	1.0000	-8.7172
ZnCO3(aq)	1.885e-009	0.0002363	1.0000	-8.7247
Cu(CO3)2--	1.608e-009	0.0002952	0.7909	-8.8955
CaOH+	1.606e-009	9.169e-005	0.9432	-8.8195
CuSO4(aq)	5.434e-010	8.672e-005	1.0000	-9.2649

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt

MnOH+	3. 609e-010	2. 596e-005	0. 9432	-9. 4681
Zn(OH) 2(aq)	3. 150e-010	3. 131e-005	1. 0000	-9. 5017
HSO4-	2. 848e-010	2. 764e-005	0. 9432	-9. 5708
Cu+	3. 274e-011	2. 080e-006	0. 9432	-10. 5103
MnNO3+	3. 029e-011	3. 542e-006	0. 9432	-10. 5441
CuCO3(OH) 2--	2. 294e-012	3. 613e-007	0. 7909	-11. 7414
H2Si O4--	7. 149e-013	6. 726e-008	0. 7909	-12. 2476
HN03(aq)	6. 784e-014	4. 274e-009	1. 0000	-13. 1685
Zn(OH) 3-	2. 065e-014	2. 404e-009	0. 9432	-13. 7104
Mn(OH) 2(aq)	1. 667e-014	1. 483e-009	1. 0000	-13. 7781
Mn2(OH) 3+	4. 670e-015	7. 512e-010	0. 9432	-14. 3561
H6(H2Si O4) 4--	3. 245e-015	1. 241e-009	0. 7909	-14. 5906
Mn(NO3) 2(aq)	1. 949e-015	3. 488e-010	1. 0000	-14. 7101
Mn2OH+++	4. 070e-016	5. 163e-011	0. 5948	-15. 6161
CuO2--	3. 001e-018	2. 866e-013	0. 7909	-17. 6247
Mn(OH) 3-	3. 308e-019	3. 504e-014	0. 9432	-18. 5059
Zn(OH) 4--	8. 379e-020	1. 118e-014	0. 7909	-19. 1787
H2SO4(aq)	1. 346e-020	1. 320e-015	1. 0000	-19. 8708
H4(H2Si O4) 4----	1. 310e-022	4. 982e-017	0. 3910	-22. 2906
Mg4(OH) 4++++	1. 444e-025	2. 386e-020	0. 3999	-25. 2385
Mn(OH) 4--	6. 665e-026	8. 195e-021	0. 7909	-25. 2781
Mn+++	2. 145e-026	1. 178e-021	0. 5948	-25. 8942
Formate	1. 169e-029	5. 263e-025	0. 9428	-28. 9577
H2(aq)	3. 660e-030	7. 378e-027	1. 0000	-29. 4365
Ca(For) +	1. 692e-031	1. 439e-026	0. 9432	-30. 7971
Mg(For) +	6. 694e-032	4. 640e-027	0. 9432	-31. 1997
NO2-	1. 395e-032	6. 417e-028	0. 9423	-31. 8812
Formic_acid(aq)	3. 128e-033	1. 440e-028	1. 0000	-32. 5047
O2(aq)	4. 568e-034	1. 462e-029	1. 0000	-33. 3402
Mn(For) +	3. 886e-034	3. 883e-029	0. 9432	-33. 4359
Zn(For) +	1. 160e-034	1. 280e-029	0. 9432	-33. 9610
NH4+	8. 529e-035	1. 538e-030	0. 9418	-34. 0951
SO3--	7. 886e-035	6. 313e-030	0. 7924	-34. 2042
HSO3-	5. 328e-035	4. 319e-030	0. 9432	-34. 2988
Cu(For) +	4. 706e-035	5. 108e-030	0. 9432	-34. 3527
CO(aq)	7. 039e-036	1. 971e-031	1. 0000	-35. 1525
HN02(aq)	1. 095e-036	5. 147e-032	1. 0000	-35. 9606
NH3(aq)	9. 202e-037	1. 567e-032	1. 0000	-36. 0361
Oxal ate	4. 598e-037	4. 047e-032	0. 7909	-36. 4393
CuNO2+	6. 155e-038	6. 742e-033	0. 9432	-37. 2361
HO2-	2. 256e-038	7. 445e-034	0. 9432	-37. 6720
CuNH3++	5. 370e-040	4. 326e-035	0. 7924	-39. 3711
H-Oxal ate	3. 489e-040	3. 106e-035	0. 9432	-39. 4827
H2SO3(aq)	2. 565e-040	2. 105e-035	1. 0000	-39. 5910
SO2(aq)	1. 826e-040	1. 169e-035	1. 0000	-39. 7386
Zn(NH3) ++	2. 209e-041	1. 820e-036	0. 7924	-40. 7569
MnO4--	2. 339e-043	2. 781e-038	0. 7909	-42. 7329
HSO5-	3. 303e-046	3. 734e-041	0. 9432	-45. 5065
Oxalic_acid(aq)	3. 073e-046	2. 767e-041	1. 0000	-45. 5124
MnO4-	5. 969e-047	7. 099e-042	0. 9428	-46. 2497
NH4SO4-	2. 395e-047	2. 733e-042	0. 9432	-46. 6460
S2O6--	1. 788e-058	2. 862e-053	0. 7909	-57. 8496
Ca(For) 2(aq)	1. 304e-059	1. 697e-054	1. 0000	-58. 8847
Mg(For) 2(aq)	5. 161e-060	5. 900e-055	1. 0000	-59. 2873
Mn(For) 2(aq)	4. 970e-062	7. 204e-057	1. 0000	-61. 3036
Zn(For) 2(aq)	1. 704e-062	2. 648e-057	1. 0000	-61. 7686
Cu(For) 2(aq)	1. 023e-062	1. 570e-057	1. 0000	-61. 9903
S2O8--	6. 793e-063	1. 305e-057	0. 7909	-62. 2698
Formal dehyde(aq)	4. 419e-064	1. 327e-059	1. 0000	-63. 3547
Cu(NO2) 2(aq)	7. 810e-069	1. 215e-063	1. 0000	-68. 1073
N3-	3. 196e-071	1. 343e-066	0. 9432	-70. 5208
Cu(NH3) 2++	1. 274e-072	1. 243e-067	0. 7924	-71. 9960
HN3(aq)	7. 573e-074	3. 258e-069	1. 0000	-73. 1207

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt

S205--	4. 863e-074	7. 008e-069	0. 7909	-73. 4150
Zn(NH3)2++	3. 268e-075	3. 250e-070	0. 7924	-74. 5867
Urea(aq)	2. 736e-076	1. 643e-071	1. 0000	-75. 5628
HCN(aq)	4. 977e-077	1. 345e-072	1. 0000	-76. 3030
ZnN3+	1. 490e-077	1. 600e-072	0. 9432	-76. 8523
Methanol (aq)	9. 623e-079	3. 083e-074	1. 0000	-78. 0167
CN-	6. 122e-079	1. 593e-074	0. 9423	-78. 2389
HS-	6. 907e-084	2. 284e-079	0. 9428	-83. 1863
H2S(aq)	3. 172e-084	1. 081e-079	1. 0000	-83. 4986
Glycolate	8. 066e-085	6. 052e-080	0. 9432	-84. 1187
Ca(Glyc)+	1. 937e-086	2. 230e-081	0. 9432	-85. 7383
Mg(Glyc)+	3. 666e-087	3. 642e-082	0. 9432	-86. 4612
Methane(aq)	2. 612e-088	4. 189e-084	1. 0000	-87. 5831
Glycolic_acid(aq)	2. 599e-088	1. 977e-083	1. 0000	-87. 5851
Zn(Glyc)+	3. 259e-089	4. 577e-084	0. 9432	-88. 5122
Cu(Glyc)+	2. 642e-089	3. 661e-084	0. 9432	-88. 6035
Mn(Glyc)+	2. 044e-089	2. 656e-084	0. 9432	-88. 7150
S--	1. 900e-089	6. 092e-085	0. 7939	-88. 8214
S203--	1. 390e-090	1. 559e-085	0. 7909	-89. 9587
S204--	3. 585e-092	4. 593e-087	0. 7939	-91. 5457
Acetate	1. 027e-093	6. 061e-089	0. 9437	-93. 0137
Zn(For)3-	7. 366e-094	1. 476e-088	0. 9432	-93. 1581
CaCH3COO+	4. 710e-096	4. 668e-091	0. 9432	-95. 3523
MgCH3COO+	4. 156e-096	3. 463e-091	0. 9432	-95. 4067
Acetic_acid(aq)	2. 776e-096	1. 667e-091	1. 0000	-95. 5565
HS203-	6. 033e-097	6. 825e-092	0. 9432	-96. 2448
MnCH3COO+	1. 123e-098	1. 280e-093	0. 9432	-97. 9750
CuCH3COO+	7. 391e-099	9. 059e-094	0. 9432	-98. 1567
ZnCH3COO+	6. 974e-099	8. 676e-094	0. 9432	-98. 1819
Malonate	2. 115e-102	2. 158e-097	0. 7909	-101. 7766
CuCH3COO(aq)	6. 394e-104	7. 837e-099	1. 0000	-103. 1942
H-Malonate	4. 408e-104	4. 542e-099	0. 9432	-103. 3811
Cu(NH3)3++	7. 756e-106	8. 890e-101	0. 7924	-105. 2114
Malonic_acid(aq)	1. 480e-108	1. 540e-103	1. 0000	-107. 8298
Zn(NH3)3++	4. 836e-109	5. 632e-104	0. 7924	-108. 4165
Methanamine(aq)	2. 012e-112	6. 247e-108	1. 0000	-111. 6964
Glycine(aq)	3. 614e-115	2. 712e-110	1. 0000	-114. 4421
Cu(Gly)+	1. 997e-116	2. 747e-111	0. 9432	-115. 7250
Mg(Gly)+	7. 642e-118	7. 516e-113	0. 9432	-117. 1422
Zn(Gly)+	5. 150e-119	7. 181e-114	0. 9432	-118. 3136
Ca(Gly)+	1. 534e-119	1. 750e-114	0. 9432	-118. 8397
Mn(Gly)+	3. 863e-120	4. 982e-115	0. 9432	-119. 4385
S306--	5. 344e-122	1. 027e-116	0. 7909	-121. 3740
Zn(For)4--	1. 718e-122	4. 217e-117	0. 7909	-121. 8668
Acetaldehyde(aq)	5. 477e-126	2. 412e-121	1. 0000	-125. 2615
SCN-	4. 873e-127	2. 830e-122	0. 9428	-126. 3378
NH4CH3COO(aq)	1. 270e-127	9. 787e-123	1. 0000	-126. 8962
Acetamide(aq)	1. 265e-127	7. 471e-123	1. 0000	-126. 8979
Zn(NH3)4++	3. 605e-143	4. 812e-138	0. 7924	-142. 5441
Ethanol (aq)	4. 310e-145	1. 985e-140	1. 0000	-144. 3656
Ethyne(aq)	5. 266e-147	1. 371e-142	1. 0000	-146. 2785
Zn(N3)2(aq)	2. 400e-147	3. 586e-142	1. 0000	-146. 6198
Ethylene(aq)	1. 475e-149	4. 137e-145	1. 0000	-148. 8313
Lactate	4. 134e-151	3. 682e-146	0. 9432	-150. 4090
Ca(Lac)+	5. 844e-153	7. 546e-148	0. 9432	-152. 2587
Mg(Lac)+	2. 062e-153	2. 337e-148	0. 9432	-152. 7112
Lactic_acid(aq)	1. 425e-154	1. 284e-149	1. 0000	-153. 8461
Zn(Lac)+	1. 103e-155	1. 704e-150	0. 9432	-154. 9828
Mn(Lac)+	7. 387e-156	1. 064e-150	0. 9432	-155. 1569
Cu(Lac)+	6. 325e-156	9. 651e-151	0. 9432	-155. 2243
Ethane(aq)	8. 763e-159	2. 635e-154	1. 0000	-158. 0573
Propanoate	4. 288e-162	3. 133e-157	0. 9432	-161. 9331
Propanoic_acid(a	1. 571e-164	1. 163e-159	1. 0000	-163. 8039

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt

Ca(Prop)+	1.085e-164	1.228e-159	0.9432	-163.9898
Mg(Prop)+	4.706e-165	4.581e-160	0.9432	-164.3528
Mn(Prop)+	5.329e-167	6.821e-162	0.9432	-166.2987
Cu(Prop)+	2.953e-167	4.034e-162	0.9432	-166.5551
Zn(Prop)+	1.262e-167	1.748e-162	0.9432	-166.9242
Succinate	3.108e-169	3.607e-164	0.7909	-168.6094
Ca(Glyc)2(aq)	2.417e-169	4.595e-164	1.0000	-168.6168
S4O6--	1.287e-169	2.886e-164	0.7909	-168.9923
Mg(Glyc)2(aq)	2.458e-170	4.285e-165	1.0000	-169.6095
H-Succinate	5.566e-171	6.515e-166	0.9432	-170.2799
Zn(Glyc)2(aq)	1.656e-171	3.568e-166	1.0000	-170.7809
Cu(Glyc)2(aq)	1.115e-171	2.382e-166	1.0000	-170.9527
Mn(Glyc)2(aq)	1.034e-172	2.119e-167	1.0000	-171.9856
Serine(aq)	3.912e-173	4.111e-168	1.0000	-172.4076
Succinylcinate(aq)	4.243e-174	5.009e-169	1.0000	-173.3724
Ethanimide(aq)	1.238e-180	5.582e-176	1.0000	-179.9072
Alanine(aq)	2.536e-182	2.259e-177	1.0000	-181.5958
Cu(Ala)+	3.163e-184	4.796e-179	0.9432	-183.5252
Mg(Ala)+	4.495e-187	5.052e-182	0.9432	-186.3726
Zn(Ala)+	3.479e-187	5.338e-182	0.9432	-186.4840
Ca(Ala)+	2.166e-187	2.775e-182	0.9432	-186.6898
Ca(CH3COO)2(aq)	6.856e-188	1.084e-182	1.0000	-187.1639
Mn(Ala)+	2.671e-188	3.820e-183	0.9432	-187.5986
Mg(CH3COO)2(aq)	2.197e-188	3.128e-183	1.0000	-187.6582
Acetone(aq)	1.375e-189	7.986e-185	1.0000	-188.8616
Zn(CH3COO)2(aq)	4.482e-190	8.222e-185	1.0000	-189.3485
Cu(CH3COO)2(aq)	1.697e-190	3.081e-185	1.0000	-189.7704
Mn(CH3COO)2(aq)	7.149e-191	1.237e-185	1.0000	-190.1458
Aspartylcinate(aq)	1.932e-192	2.571e-187	1.0000	-191.7140
Propanal(aq)	8.107e-194	4.708e-189	1.0000	-193.0911
Cu(CH3COO)2-	6.140e-197	1.115e-191	0.9432	-196.2372
1-Propyne(aq)	2.364e-211	9.470e-207	1.0000	-210.6263
S3--	1.969e-211	1.894e-206	0.7909	-210.8077
1-Propanol(aq)	2.025e-213	1.217e-208	1.0000	-212.6936
1-Propene(aq)	1.024e-215	4.309e-211	1.0000	-214.9896
2-Hydroxybutanoate	7.549e-220	7.781e-215	0.9432	-219.1475
NH4(CH3COO)2-	9.626e-221	1.310e-215	0.9432	-220.0419
2-Hydroxybutanoic	2.313e-223	2.407e-218	1.0000	-222.6359
Asparagine(aq)	1.134e-223	1.498e-218	1.0000	-222.9455
Cu(Gly)2(aq)	4.115e-226	8.708e-221	1.0000	-225.3857
Propane(aq)	1.762e-227	7.767e-223	1.0000	-226.7541
Butanoate	6.080e-231	5.295e-226	0.9432	-230.2415
Zn(Gly)2(aq)	1.569e-231	3.349e-226	1.0000	-230.8044
Mg(Gly)2(aq)	8.855e-232	1.527e-226	1.0000	-231.0528
Dilysine(aq)	3.385e-232	4.472e-227	1.0000	-231.4704
Butanoylcinate(aq)	1.849e-233	1.629e-228	1.0000	-232.7330
Ca(But)+	1.060e-233	1.348e-228	0.9432	-233.0001
Mg(But)+	4.390e-234	4.890e-229	0.9432	-233.3829
Mn(Gly)2(aq)	4.086e-234	8.295e-229	1.0000	-233.3888
Diketopiperazine	2.568e-234	2.930e-229	1.0000	-233.5904
Ca(Gly)2(aq)	1.000e-235	1.882e-230	1.0000	-234.9999
Mn(But)+	6.404e-236	9.094e-231	0.9432	-235.2189
Cu(But)+	3.306e-236	4.979e-231	0.9432	-235.5061
Zn(But)+	2.823e-236	4.304e-231	0.9432	-235.5747
Gluconate	3.454e-237	4.494e-232	0.7909	-236.5635
H-Gluconate	3.791e-239	4.970e-234	0.9432	-238.4466
Threonine(aq)	5.833e-242	6.947e-237	1.0000	-241.2341
Gluconylcinate(aq)	3.916e-242	5.172e-237	1.0000	-241.4072
Ethyl acetate(aq)	3.443e-242	3.033e-237	1.0000	-241.4631
S5O6--	4.083e-246	1.046e-240	0.7909	-245.4909
1-Propanamine(aq)	1.928e-248	1.140e-243	1.0000	-247.7148
alpha-Aminobutyric acid	7.684e-251	7.922e-246	1.0000	-250.1144
Gluconylcinate(aq)	2.746e-259	4.039e-254	1.0000	-258.5613

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt

Zn(SCN)2(aq)	2.697e-259	4.895e-254	1.0000	-258.5692
Butanal(aq)	4.813e-264	3.470e-259	1.0000	-263.3176
S4--	1.102e-272	1.413e-267	0.7909	-272.0598
1-Butyne(aq)	3.241e-280	1.753e-275	1.0000	-279.4894
Cu(CH3COO)3-	3.177e-282	7.646e-277	0.9432	-281.5233
Zn(CH3COO)3-	2.575e-282	6.245e-277	0.9432	-281.6145
1-Butanol(aq)	5.874e-283	4.354e-278	1.0000	-282.2310
Mn(CH3COO)3-	1.596e-283	3.704e-278	0.9432	-282.8223
1-Butene(aq)	9.206e-285	5.164e-280	1.0000	-284.0359
2-Hydroxypentano	2.032e-288	2.379e-283	0.9432	-287.7175
2-Hydroxypentano	3.752e-292	4.432e-287	1.0000	-291.4257
Glutamine(aq)	1.840e-292	2.688e-287	1.0000	-291.7352
n-Butane(aq)	3.115e-296	1.810e-291	1.0000	-295.5065
Alanylglycine(aq)	1.071e-299	1.565e-294	1.0000	-298.9701
Pentanoate	1.037e-299	1.049e-294	0.9432	-299.0094
Pentanoic acid(a	3.434e-302	3.507e-297	1.0000	-300.0000
Ca(Lac)2(aq)	2.467e-302	5.383e-297	1.0000	-300.0000
Ca(Pent)+	1.057e-302	1.493e-297	0.9432	-300.0000
Mg(Lac)2(aq)	7.946e-303	1.608e-297	1.0000	-300.0000
Mg(Pent)+	4.285e-303	5.374e-298	0.9432	-300.0000
Zn(CN)4--	1.193e-303	2.022e-298	0.7909	-300.0000
Zn(Lac)2(aq)	1.441e-304	3.508e-299	1.0000	-300.0000
Cu(Lac)2(aq)	9.063e-305	2.190e-299	1.0000	-300.0000
Mn(Pent)+	8.245e-305	1.287e-299	0.9432	-300.0000
Cu(Pent)+	5.482e-305	9.025e-300	0.9432	-300.0000
Zn(Pent)+	3.988e-305	6.640e-300	0.9432	-300.0000
Mn(Lac)2(aq)	3.193e-305	7.441e-300	1.0000	-300.0000
Adipate	1.822e-307	2.626e-302	0.7909	-300.0000
H-Adipate	1.966e-309	2.853e-304	0.9432	-300.0000
Adipic acid(aq)	2.404e-312	3.513e-307	1.0000	-300.0000
1-Butanamine(aq)	1.733e-317	1.267e-312	1.0000	-300.0000
Valine(aq)	1.838e-319	2.153e-314	1.0000	-300.0000
Decanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9432	-300.0000
Decanal(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	0.9432	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecano	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Benzoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecano	0.0000	0.0000	0.9432	-300.0000
Benzoate	0.0000	0.0000	0.9450	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Azelate	0.0000	0.0000	0.7909	-300.0000
Azelic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hexanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octanol(aq)	0.0000	0.0000	1.0000	-300.0000
2-Heptanone(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Methionine(aq)	0.0000	0.0000	1.0000	-300.0000
2-Butanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt

Leucyl glyci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Leuci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
p-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
I sol euci ne(aq)	0.0000	0.0000	1.0000	-300.0000
o-Tol uate	0.0000	0.0000	0.9432	-300.0000
o-Phthal i c_aci d(0.0000	0.0000	1.0000	-300.0000
Hexanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9432	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanal (aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoi c_aci d(a	0.0000	0.0000	1.0000	-300.0000
n-Hexyl benzene(a	0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9432	-300.0000
Heptanal (aq)	0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(a	0.0000	0.0000	1.0000	-300.0000
m-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
m-Tol uate	0.0000	0.0000	0.9432	-300.0000
1-Hexyne(aq)	0.0000	0.0000	1.0000	-300.0000
Zn(SCN)4--	0.0000	0.0000	0.7909	-300.0000
Zn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Zn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(o-Phthal ate)(0.0000	0.0000	1.0000	-300.0000
Ca(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol (aq)	0.0000	0.0000	1.0000	-300.0000
Zn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Zn(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
H-Suberate	0.0000	0.0000	0.9432	-300.0000
H-Sebacate	0.0000	0.0000	0.9432	-300.0000
Undecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
H-Pi mel ate	0.0000	0.0000	0.9432	-300.0000
Undecanoate	0.0000	0.0000	0.9432	-300.0000
Tyrosi ne(aq)	0.0000	0.0000	1.0000	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
Tol uene(aq)	0.0000	0.0000	1.0000	-300.0000
Suberi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
H-Azel ate	0.0000	0.0000	0.9432	-300.0000
2-Hydroxyoctanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Suberate	0.0000	0.0000	0.7909	-300.0000
Sebaci c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacate	0.0000	0.0000	0.7909	-300.0000
H(o-Phthal ate)-	0.0000	0.0000	0.9432	-300.0000
2-Hydroxyoctanoa	0.0000	0.0000	0.9432	-300.0000
S5--	0.0000	0.0000	0.7909	-300.0000
2-Hydroxynonanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexanol (aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9432	-300.0000
Pi mel i c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Pi mel ate	0.0000	0.0000	0.7909	-300.0000
Phenyl al ani ne(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol (aq)	0.0000	0.0000	1.0000	-300.0000
Pentanal (aq)	0.0000	0.0000	1.0000	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoi	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt

1-Pentanol (aq)	0.0000	0.0000	1.0000	-300.0000
Octanoi c_acid(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9432	-300.0000
Ca(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanal (aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoi c_acid(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoa	0.0000	0.0000	0.9432	-300.0000
Nonanoate	0.0000	0.0000	0.9432	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
Nonanal (aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoi c_acid(0.0000	0.0000	1.0000	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octyl benzene(a	0.0000	0.0000	1.0000	-300.0000
Dodecanoate	0.0000	0.0000	0.9432	-300.0000
n-Pentyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Propyl benzene(0.0000	0.0000	1.0000	-300.0000
o-Phthalate	0.0000	0.0000	0.7909	-300.0000
2-Hydroxyheptano	0.0000	0.0000	1.0000	-300.0000
1-Pentanami ne(aq	0.0000	0.0000	1.0000	-300.0000
o-Tolui c_acid(aq	0.0000	0.0000	1.0000	-300.0000
p-Toluate	0.0000	0.0000	0.9432	-300.0000
1-Hexanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanami ne(aq	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states
log Q/K

log Q/K

Quartz	0.1117s/sat	Chalcanthite	-9.0034
Tridymite	-0.0598	ZnSO4: 7H2O	-9.1551
Ice	-0.1387	ZnSO4: 6H2O	-9.3388
Chalcedony	-0.1595	Manganosite	-9.5021
Tenorite	-0.4310	Hausmannite	-9.7263
Cristobalite(alp	-0.4388	Forsterite	-9.8945
Dolomite	-0.5000	Sepiolite	-10.0583
Dolomite-ord	-0.5000	Tephroite	-10.1219
Calcite	-0.6406	Mn(OH)3	-10.3354
Coesite	-0.6983	Periclase	-10.4075
Aragonite	-0.7850	Mg1.25SO4(OH)0.5	-10.4501
Cristobalite(bet	-0.8823	ZnSO4: H2O	-10.4851
SiO2(am)	-1.1740	Zn2SO4(OH)2	-10.7785
Monohydrocalcite	-1.4743	Monticellite	-11.2145
Magnesite	-1.4881	Portlandite	-11.2247
Rhodochrosite	-1.5077	Hydrozincite	-11.4199
Malachite	-1.6122	Mg1.5SO4(OH)	-11.7132
Dolomite-dis	-2.0444	Gyrolite	-11.9119
Zn2SiO4	-2.1042	MgSO4	-12.8001
ZnCO3: H2O	-2.4357	MnSO4	-13.0841
Dioptase	-2.7399	Chalcocyanite	-14.5488
Smithsonite	-2.7592	ZnSO4	-14.5686
Chrysocolia	-2.8768	Tremolite	-15.0375
Gypsum	-3.0371	Hydromagnesite	-16.7041
Anhydrite	-3.2130	Hillebrandite	-18.0457
Zincite	-3.3822	Dicalciumsilica	-18.3992
Manganite	-3.8296	Zn(NO3)2: 6H2O	-19.3157
Zn(OH)2(epsilon)	-3.8360	Akermanite	-19.5055
Bassanite	-3.8579	Larnite	-19.6932
Talc	-3.9053	Lime	-21.2456
CaSO4: 0.5H2O(bet	-4.0260	Anthophyllite	-21.4025
Lansfordite	-4.0354	Rankinite	-25.6917
Zn(OH)2(gamma)	-4.0567	Zn5(NO3)2(OH)8	-27.2669
Zn(OH)2(beta)	-4.1076	Tobermorite-14A	-30.5180
Nesquehoni te	-4.1900	Merwinite	-31.3700

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt

Enstatite	-4.2867	Foshagite	-32.2620
Cuprite	-4.5175	Tobermorite-11A	-32.2856
Rhodonite	-5.1958	Zn30(SO4)2	-33.3391
Brucite	-5.3701	Afwillite	-33.8291
Brochantite	-5.3860	Todorokite	-35.4741
Antlerite	-5.9050	Tobermorite-9A	-35.7533
Epsomite	-5.9597	C	-40.9557
Chrysotile	-6.1171	Hatrurite	-43.3018
Hexahydrite	-6.1952	Zn	-44.3069
Wollastonite	-6.3177	Birnessite	-44.3427
Dioptase	-6.4812	Xonotlite	-47.1697
Pentahydrite	-6.5348	Mn	-57.8584
Pseudowollastonite	-6.5569	Covelite	-60.4303
Huntite	-6.6765	S	-62.0584
Pyrolusite	-6.8116	Chalcocite	-62.1727
Okenite	-6.8264	Sphalerite	-71.2198
Mn(OH)2(am)	-6.8883	Wurtzite	-73.5192
Starkeyite	-6.9221	Alabandite	-81.6700
Bixbyite	-7.0229	Antigorite	-84.8362
Cu	-7.6166	Mg	-94.9385
Kieserite	-7.6550	Ca	-111.8459
Azurite	-7.7304	Si	-119.4533
Artinite	-7.9227	o-Phthalic_acid	-363.9855
MnO2(gamma)	-8.3294		

Gases	fugacity	log fug.
H2O(g)	0.02598	-1.585
CO2(g)	0.004911	-2.309
N2(g)	0.003670	-2.435
NO2(g)	1.264e-012	-11.898
H2(g)	4.661e-027	-26.331
O2(g)	3.615e-031	-30.442
NO(g)	2.503e-032	-31.602
CO(g)	7.150e-033	-32.146
NH3(g)	1.470e-038	-37.833
SO2(g)	1.234e-040	-39.909
Cu(g)	1.711e-060	-59.767
Zn(g)	1.210e-061	-60.917
H2S(g)	3.087e-083	-82.510
CH4(g)	1.850e-085	-84.733
Mg(g)	2.232e-115	-114.651
Ca(g)	8.372e-138	-137.077
S2(g)	9.783e-139	-138.010
C2H4(g)	3.107e-147	-146.508
C(g)	2.790e-159	-158.554
Si(g)	3.175e-191	-190.498

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
Ca++	0.000690	0.000690	27.6			
Cu++	6.37e-007	6.37e-007	0.0405			
H+	0.000161	0.000161	0.162			
H2O	55.5	55.5	1.00e+006			
HC03-	0.00178	0.00178	109.			
Mg++	0.000272	0.000272	6.61			
Mn++	8.81e-007	8.81e-007	0.0484			
NH3(aq)	4.78e-006	4.78e-006	0.0814			
NO3-	2.89e-005	2.89e-005	1.79			
O2(aq)	3.24e-006	3.24e-006	0.104			
SO4--	7.81e-005	7.81e-005	7.50			
SiO2(aq)	0.000130	0.000130	7.80			

SpecE8_output_GSS_Mine Water UQ-1 April 2003.txt
 Zn++ 2.27e-007 2.27e-007 0.0148

Elemental composition	total moles		In fluid		Sorbed	
			moles	mg/kg	moles	mg/kg
Calci um	0.0006899		0.0006899	27.64		
Carbon	0.001780		0.001780	21.38		
Copper	6.371e-007		6.371e-007	0.04048		
Hydrogen	111.0		111.0	1.119e+005		
Magnesi um	0.0002722		0.0002722	6.615		
Manganese	8.807e-007		8.807e-007	0.04838		
Ni trogen	3.368e-005		3.368e-005	0.4717		
Oxygen	55.51		55.51	8.881e+005		
Si l i con	0.0001298		0.0001298	3.646		
Sul fur	7.812e-005		7.812e-005	2.505		
Zi nc	2.265e-007		2.265e-007	0.01481		

SpecE8_output_GSS_MW-01-15 July.txt

Temperature = 22.2 C Pressure = 1.013 bars
 pH = 7.300 log fO2 = -24.587
 Eh = 0.4420 volts pe = 7.5432
 Ionic strength = 0.004698
 Charge imbalance = -0.000095 eq/kg (-2.69% error)
 Activity of water = 0.999996
 Solvent mass = 1.000000 kg
 Solution mass = 1.000320 kg
 Solution density = 1.015 g/cm3
 Chlorinity = 0.000111 molal
 Dissolved solids = 320 mg/kg sol'n
 Elect. conductivity = 358.14 uS/cm (or umho/cm)
 Hardness = 81.80 mg/kg sol'n as CaCO3
 carbonate = 81.80 mg/kg sol'n as CaCO3
 non-carbonate = 0.00 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 83.83 mg/kg sol'n as CaCO3
 Water type = Na-HCO3

Nernst redox couples

Eh (volts)

pe

$e^- + .25 \cdot O_2(aq) + H^+ = .5 \cdot H_2O$	0.4420	7.5432
$8 \cdot e^- + 9 \cdot H^+ + NO_3^- = 3 \cdot H_2O + NH_3(aq)$	0.6033	10.2957

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
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HC03-	0.001658	101.2	0.9295	-2.8121
Na+	0.001326	30.48	0.9295	-2.9092
NO3-	0.000922	61.50	0.9281	-3.0358
Ca++	0.0005886	23.58	0.7542	-3.3527
K+	0.0005534	21.63	0.9281	-3.2894
SO4--	0.0003922	37.66	0.7455	-3.5340
SiO2(aq)	0.0002013	12.09	1.0000	-3.6962
Mg++	0.0001897	4.609	0.7624	-3.8398
CO2(aq)	0.0001822	8.014	1.0000	-3.7396
N2(aq)	0.0001189	3.329	1.0000	-3.9249
Cl-	0.0001112	3.940	0.9281	-3.9865
CaSO4(aq)	1.662e-005	2.263	1.0000	-4.7792
F-	1.032e-005	0.1960	0.9288	-5.0185
MgSO4(aq)	1.019e-005	1.226	1.0000	-4.9918
Mn++	8.399e-006	0.4613	0.7542	-5.1983
CaHCO3+	8.360e-006	0.8449	0.9295	-5.1095
NaHCO3(aq)	2.875e-006	0.2414	1.0000	-5.5414
MgHCO3+	2.664e-006	0.2272	0.9295	-5.6062
CO3--	1.783e-006	0.1070	0.7477	-5.8751
KS04-	1.234e-006	0.1668	0.9295	-5.9403
CaCO3(aq)	1.195e-006	0.1195	1.0000	-5.9228
Fe++	1.059e-006	0.05911	0.7542	-6.0977
Ba++	7.604e-007	0.1044	0.7499	-6.2439
HSiO3-	4.568e-007	0.03521	0.9295	-6.3720
MnSO4(aq)	4.088e-007	0.06171	1.0000	-6.3885
Sb(OH)3(aq)	2.105e-007	0.03635	1.0000	-6.6768
HP04--	1.909e-007	0.01832	0.7455	-6.8467
MgCO3(aq)	1.801e-007	0.01518	1.0000	-6.7444
OH-	1.709e-007	0.002906	0.9288	-6.7993
H2P04-	1.273e-007	0.01234	0.9295	-6.9269
Cu++	1.086e-007	0.006897	0.7542	-7.0868
H+	5.355e-008	5.396e-005	0.9359	-7.3000
MgF+	3.418e-008	0.001479	0.9295	-7.4980
NaHSiO3(aq)	2.498e-008	0.002499	1.0000	-7.6024

	SpecE8_output_GSS_MW-01-15	July.txt	
CaF+	2.196e-008	0.001297	0.9295 -7.6901
NaCl (aq)	2.101e-008	0.001227	1.0000 -7.6776
MgCl +	1.212e-008	0.0007240	0.9295 -7.9482
CaCl +	1.001e-008	0.0007558	0.9295 -8.0313
UO2(CO3)2--	6.828e-009	0.002663	0.7455 -8.2932
NaCO3-	6.275e-009	0.0005206	0.9295 -8.2342
KCl (aq)	1.603e-009	0.0001194	1.0000 -8.7952
HSO4-	1.419e-009	0.0001377	0.9295 -8.8798
MnCl +	1.335e-009	0.0001207	0.9295 -8.9061
UO2(CO3)3----	1.218e-009	0.0005480	0.3085 -9.4251
NaF (aq)	1.167e-009	4.900e-005	1.0000 -8.9328
HF (aq)	6.766e-010	1.353e-005	1.0000 -9.1697
BaCO3(aq)	3.180e-010	6.274e-005	1.0000 -9.4975
UO2CO3(aq)	2.308e-010	7.613e-005	1.0000 -9.6369
FeCl +	6.163e-011	5.625e-006	0.9295 -10.2419
NaOH(aq)	3.196e-011	1.278e-006	1.0000 -10.4955
BaCl +	1.948e-011	3.364e-006	0.9295 -10.7423
Sb(OH)4-	4.280e-012	8.119e-007	0.9295 -11.4003
UO2OH+	4.111e-012	1.180e-006	0.9295 -11.4178
BaF+	3.776e-012	5.901e-007	0.9295 -11.4547
PO4---	2.431e-012	2.308e-007	0.5161 -11.9015
HN03(aq)	2.171e-012	1.368e-007	1.0000 -11.6633
Cu+	1.228e-012	7.804e-008	0.9295 -11.9424
Sb(OH)2F(aq)	1.181e-012	2.063e-007	1.0000 -11.9277
CaCl 2(aq)	1.131e-012	1.255e-007	1.0000 -11.9466
HCl (aq)	1.062e-012	3.873e-008	1.0000 -11.9737
H3PO4(aq)	8.551e-013	8.377e-008	1.0000 -12.0680
Sb(OH)2+	3.917e-013	6.099e-008	0.9295 -12.4388
UO2++	4.946e-014	1.335e-008	0.7477 -13.4320
KHSO4(aq)	4.504e-014	6.131e-009	1.0000 -13.3464
UO2F+	4.266e-014	1.233e-008	0.9295 -13.4017
UO2SO4(aq)	1.191e-014	4.358e-009	1.0000 -13.9242
MgP2O7--	1.013e-014	2.007e-009	0.7455 -14.1221
HF2-	1.629e-015	6.350e-011	0.9295 -14.8199
UO2F2(aq)	1.182e-015	3.638e-010	1.0000 -14.9276
HP2O7---	5.049e-016	8.830e-011	0.5161 -15.5840
H2P2O7--	7.674e-017	1.350e-011	0.7455 -16.2425
UO2(SO4)2--	3.587e-017	1.657e-011	0.7455 -16.5729
FeCl 2(aq)	3.000e-017	3.802e-012	1.0000 -16.5228
P2O7----	1.040e-017	1.809e-012	0.3085 -17.4937
UO2Cl +	5.719e-018	1.746e-012	0.9295 -17.2744
UO2F3-	2.116e-018	6.916e-013	0.9295 -17.7063
(UO2)2(OH)2++	1.440e-018	8.266e-013	0.7477 -17.9678
KP2O7---	6.594e-019	1.404e-013	0.5161 -18.4681
UO2+	3.255e-020	8.787e-015	0.9295 -19.5192
(UO2)3(OH)5+	2.942e-020	2.633e-014	0.9295 -19.5631
(UO2)3(CO3)6(6-)	4.524e-021	5.292e-015	0.0709 -21.4940
H3P2O7-	6.897e-022	1.220e-016	0.9295 -21.1931
UO2F4--	1.461e-022	5.055e-017	0.7455 -21.9628
UO2Cl 2(aq)	2.802e-023	9.551e-018	1.0000 -22.5525
Mn+++	8.971e-024	4.927e-019	0.5231 -23.3286
FeCl 4--	1.351e-024	2.670e-019	0.7455 -23.9969
U(OH)4(aq)	1.039e-024	3.179e-019	1.0000 -23.9834
NO2-	2.025e-027	9.315e-023	0.9281 -26.7259
H4P2O7(aq)	9.721e-028	1.729e-022	1.0000 -27.0123
O2(aq)	3.486e-028	1.115e-023	1.0000 -27.4577
HN02(aq)	1.782e-031	8.376e-027	1.0000 -30.7491
Formate	4.934e-033	2.221e-028	0.9288 -32.3388
H2(aq)	1.411e-033	2.843e-030	1.0000 -32.8505
Cl 0-	1.238e-033	6.365e-029	0.9295 -32.9392
Ca(For)+	6.097e-035	5.187e-030	0.9295 -34.2466
Mg(For)+	2.068e-035	1.433e-030	0.9295 -34.7161
H02-	1.154e-035	3.807e-031	0.9295 -34.9696

	SpecE8_output_GSS_MW-01-15 July.txt			
Na(For) (aq)	6.481e-036	4.406e-031	1.0000	-35.1883
K(For) (aq)	2.523e-036	2.122e-031	1.0000	-35.5980
Mn(For)+	1.648e-036	1.646e-031	0.9295	-35.8149
Formic acid(aq)	1.317e-036	6.059e-032	1.0000	-35.8805
MnO4--	9.528e-037	1.133e-031	0.7455	-36.1485
SiF6--	3.694e-037	5.247e-032	0.7455	-36.5600
Fe(For)+	2.834e-037	2.858e-032	0.9295	-36.5793
SO3--	1.709e-037	1.368e-032	0.7477	-36.8936
HSO3-	1.092e-037	8.853e-033	0.9295	-36.9934
Ba(For)+	6.760e-038	1.232e-032	0.9295	-37.2018
Cu(For)+	4.123e-038	4.474e-033	0.9295	-37.4166
MnO4-	7.791e-039	9.264e-034	0.9288	-38.1405
NH4+	7.374e-039	1.330e-034	0.9273	-38.1650
CO(aq)	2.822e-039	7.903e-035	1.0000	-38.5494
Oxalate	2.026e-040	1.783e-035	0.7455	-39.8209
NH3(aq)	6.358e-041	1.083e-036	1.0000	-40.1966
UOH+++	2.947e-042	7.514e-037	0.5231	-41.8120
UF3+	9.565e-043	2.821e-037	0.9295	-42.0511
HSO5-	8.903e-043	1.006e-037	0.9295	-42.0822
UF2++	5.872e-043	1.620e-037	0.7477	-42.3575
SO2(aq)	3.508e-043	2.247e-038	1.0000	-42.4549
U(CO3)5(6-)	1.853e-043	9.968e-038	0.0709	-43.8816
H-Oxalate	1.458e-043	1.297e-038	0.9295	-42.8681
UF4(aq)	7.843e-044	2.462e-038	1.0000	-43.1055
UF+++	1.081e-044	2.777e-039	0.5231	-44.2476
U(SO4)2(aq)	5.590e-046	2.404e-040	1.0000	-45.2526
USO4++	3.991e-046	1.333e-040	0.7477	-45.5252
U++++	1.032e-048	2.456e-043	0.3196	-48.4817
Oxalic acid(aq)	1.259e-049	1.134e-044	1.0000	-48.8998
UCl+++	3.574e-051	9.770e-046	0.5231	-50.7283
ClO2-	1.881e-055	1.268e-050	0.9295	-54.7573
S2O8--	8.068e-059	1.550e-053	0.7455	-58.2208
S2O6--	1.466e-060	2.347e-055	0.7455	-59.9613
ClO3-	2.799e-063	2.335e-058	0.9288	-62.5851
U+++	4.904e-066	1.167e-060	0.5231	-65.5909
Ca(For)2(aq)	2.021e-066	2.628e-061	1.0000	-65.6945
Mg(For)2(aq)	7.156e-067	8.179e-062	1.0000	-66.1454
Mn(For)2(aq)	9.210e-068	1.335e-062	1.0000	-67.0357
Fe(For)2(aq)	2.025e-068	2.953e-063	1.0000	-67.6936
Na(For)2-	1.809e-068	2.044e-063	0.9295	-67.7742
K(For)2-	6.539e-069	8.441e-064	0.9295	-68.2163
Cu(For)2(aq)	3.977e-069	6.106e-064	1.0000	-68.4004
Ba(For)2(aq)	1.966e-069	4.468e-064	1.0000	-68.7065
Formaldehyde(aq)	7.136e-071	2.142e-066	1.0000	-70.1466
N3-	6.595e-071	2.771e-066	0.9295	-70.2125
HN3(aq)	1.712e-073	7.365e-069	1.0000	-72.7665
ClO4-	1.989e-075	1.978e-070	0.9288	-74.7334
UO2ClO3+	3.268e-076	1.155e-070	0.9295	-75.5174
S2O5--	2.113e-079	3.045e-074	0.7455	-78.8026
Cu(NH3)2++	1.520e-080	1.483e-075	0.7477	-79.9445
Urea(aq)	1.561e-084	9.371e-080	1.0000	-83.8066
HCN(aq)	1.242e-084	3.356e-080	1.0000	-83.9058
CN-	1.271e-086	3.306e-082	0.9281	-85.9282
Methanol(aq)	8.791e-089	2.816e-084	1.0000	-88.0560
Glycolate	7.590e-095	5.694e-090	0.9295	-94.1515
HS-	2.020e-096	6.678e-092	0.9288	-95.7268
Ca(Glyc)+	1.576e-096	1.814e-091	0.9295	-95.8341
H2S(aq)	1.021e-096	3.477e-092	1.0000	-95.9911
Mg(Glyc)+	2.451e-097	2.434e-092	0.9295	-96.6424
Na(Glyc)(aq)	1.005e-097	9.846e-093	1.0000	-96.9980
K(Glyc)(aq)	3.911e-098	4.463e-093	1.0000	-97.4077
Glycolic acid(aq)	2.433e-098	1.850e-093	1.0000	-97.6139
Mn(Glyc)+	1.941e-098	2.523e-093	0.9295	-97.7436

	SpecE8_output_GSS_MW-01-15 July.txt			
Fe(Gl yc) +	9. 907e-099	1. 296e-093	0. 9295	-98. 0358
Cu(Gl yc) +	5. 309e-099	7. 355e-094	0. 9295	-98. 3068
Ba(Gl yc) +	4. 436e-100	9. 418e-095	0. 9295	-99. 3847
S2O4--	7. 693e-101	9. 854e-096	0. 7499	-100. 2389
Methane(aq)	1. 523e-101	2. 443e-097	1. 0000	-100. 8172
S--	4. 733e-102	1. 517e-097	0. 7499	-101. 4499
S2O3--	2. 357e-102	2. 642e-097	0. 7455	-101. 7553
Acetate	5. 989e-107	3. 535e-102	0. 9302	-106. 2540
CaCH3COO+	2. 330e-109	2. 309e-104	0. 9295	-108. 6644
MgCH3COO+	1. 767e-109	1. 473e-104	0. 9295	-108. 7844
Aceti c_aci d(aq)	1. 601e-109	9. 610e-105	1. 0000	-108. 7957
NaCH3COO(aq)	5. 554e-110	4. 555e-105	1. 0000	-109. 2554
KCH3COO(aq)	1. 558e-110	1. 529e-105	1. 0000	-109. 8073
MnCH3COO+	6. 557e-111	7. 472e-106	0. 9295	-110. 2150
FeCH3COO+	9. 787e-112	1. 124e-106	0. 9295	-111. 0411
CuCH3COO+	9. 066e-112	1. 111e-106	0. 9295	-111. 0743
BaCH3COO+	3. 333e-112	6. 543e-107	0. 9295	-111. 5089
Mal onate	1. 236e-115	1. 261e-110	0. 7455	-115. 0355
H-Mal onate	2. 435e-117	2. 508e-112	0. 9295	-116. 6453
Cu(NH3)3++	6. 997e-118	8. 019e-113	0. 7477	-117. 2813
CuCH3COO(aq)	1. 418e-118	1. 738e-113	1. 0000	-117. 8483
Mal oni c_aci d(aq)	8. 113e-122	8. 440e-117	1. 0000	-121. 0908
Methanami ne(aq)	1. 416e-126	4. 397e-122	1. 0000	-125. 8489
Gl yci ne(aq)	3. 074e-129	2. 307e-124	1. 0000	-128. 5123
Cu(Gl y) +	3. 182e-130	4. 377e-125	0. 9295	-129. 5291
Mg(Gl y) +	3. 976e-132	3. 910e-127	0. 9295	-131. 4323
S3O6--	3. 148e-133	6. 049e-128	0. 7455	-132. 6295
Mn(Gl y) +	2. 784e-133	3. 590e-128	0. 9295	-132. 5871
Fe(Gl y) +	1. 595e-133	2. 071e-128	0. 9295	-132. 8291
Ca(Gl y) +	8. 868e-134	1. 012e-128	0. 9295	-133. 0839
Ba(Gl y) +	1. 528e-136	3. 229e-131	0. 9295	-135. 8477
Acetal dehyde(aq)	1. 239e-142	5. 456e-138	1. 0000	-141. 9069
SCN-	9. 791e-144	5. 685e-139	0. 9288	-143. 0413
NH4CH3COO(aq)	6. 311e-145	4. 863e-140	1. 0000	-144. 1999
Acetami de(aq)	5. 970e-145	3. 525e-140	1. 0000	-144. 2241
UO2SCN+	9. 150e-156	3. 001e-150	0. 9295	-155. 0703
Ethyne(aq)	6. 987e-164	1. 819e-159	1. 0000	-163. 1557
Ethanol (aq)	5. 005e-165	2. 305e-160	1. 0000	-164. 3006
Ethyl ene(aq)	1. 493e-169	4. 187e-165	1. 0000	-168. 8259
Lactate	4. 893e-171	4. 357e-166	0. 9295	-170. 3422
Ca(Lac) +	6. 020e-173	7. 772e-168	0. 9295	-172. 2522
Mg(Lac) +	1. 717e-173	1. 946e-168	0. 9295	-172. 7971
Na(Lac) (aq)	6. 546e-174	7. 333e-169	1. 0000	-173. 1840
K(Lac) (aq)	2. 549e-174	3. 266e-169	1. 0000	-173. 5937
Lacti c_aci d(aq)	1. 674e-174	1. 508e-169	1. 0000	-173. 7761
Mn(Lac) +	8. 889e-175	1. 280e-169	0. 9295	-174. 0829
Fe(Lac) +	5. 623e-175	8. 146e-170	0. 9295	-174. 2818
Cu(Lac) +	1. 603e-175	2. 446e-170	0. 9295	-174. 8267
Ba(Lac) +	1. 244e-176	2. 814e-171	0. 9295	-175. 9371
Ethane(aq)	5. 864e-182	1. 763e-177	1. 0000	-181. 2318
Propanoate	2. 998e-185	2. 190e-180	0. 9295	-184. 5549
Propanoi c_aci d(a	1. 084e-187	8. 027e-183	1. 0000	-186. 9650
Ca(Prop) +	6. 503e-188	7. 356e-183	0. 9295	-187. 2186
Na(Prop) (aq)	3. 953e-188	3. 796e-183	1. 0000	-187. 4031
Mg(Prop) +	2. 419e-188	2. 355e-183	0. 9295	-187. 6481
K(Prop) (aq)	1. 539e-188	1. 726e-183	1. 0000	-187. 8127
USCN+++	6. 061e-189	1. 794e-183	0. 5231	-188. 4989
Mn(Prop) +	3. 785e-189	4. 843e-184	0. 9295	-188. 4537
Ca(Gl yc) 2(aq)	1. 917e-189	3. 644e-184	1. 0000	-188. 7174
Fe(Prop) +	1. 037e-189	1. 336e-184	0. 9295	-189. 0161
S4O6--	6. 950e-190	1. 558e-184	0. 7455	-189. 2856
Cu(Prop) +	4. 399e-190	6. 007e-185	0. 9295	-189. 3884
Mg(Gl yc) 2(aq)	1. 643e-190	2. 865e-185	1. 0000	-189. 7843

	SpecE8_output_GSS_MW-01-15 July.txt			
Fe(Gl yc)2(aq)	3. 360e-191	6. 917e-186	1. 0000	-190. 4737
Ba(Prop)+	2. 360e-191	4. 964e-186	0. 9295	-190. 6588
Cu(Gl yc)2(aq)	2. 259e-191	4. 825e-186	1. 0000	-190. 6460
Mn(Gl yc)2(aq)	9. 594e-192	1. 966e-186	1. 0000	-191. 0180
Na(Gl yc)2-	6. 739e-192	1. 166e-186	0. 9295	-191. 2032
K(Gl yc)2-	2. 447e-192	4. 629e-187	0. 9295	-191. 6431
Succi nate	2. 244e-192	2. 604e-187	0. 7455	-191. 7765
Ba(Gl yc)2(aq)	1. 492e-193	4. 287e-188	1. 0000	-192. 8262
Seri ne(aq)	6. 754e-194	7. 096e-189	1. 0000	-193. 1704
H-Succi nate	3. 871e-194	4. 531e-189	0. 9295	-193. 4439
Succi ni c_aci d(aq)	2. 965e-197	3. 500e-192	1. 0000	-196. 5280
Ethanami ne(aq)	1. 064e-204	4. 795e-200	1. 0000	-203. 9731
Al ani ne(aq)	2. 713e-206	2. 417e-201	1. 0000	-205. 5665
Cu(Al a)+	6. 386e-208	9. 680e-203	0. 9295	-207. 2265
Fe(Al a)+	1. 535e-210	2. 209e-205	0. 9295	-209. 8456
Mg(Al a)+	2. 844e-211	3. 195e-206	0. 9295	-210. 5779
Mn(Al a)+	2. 372e-211	3. 391e-206	0. 9295	-210. 6567
Ca(Al a)+	1. 571e-211	2. 013e-206	0. 9295	-210. 8355
Ca(CH3COO)2(aq)	2. 063e-214	3. 262e-209	1. 0000	-213. 6855
Ba(Al a)+	1. 042e-214	2. 349e-209	0. 9295	-214. 0138
Mg(CH3COO)2(aq)	5. 865e-215	8. 349e-210	1. 0000	-214. 2317
Acetone(aq)	4. 051e-216	2. 352e-211	1. 0000	-215. 3924
Mn(CH3COO)2(aq)	2. 549e-216	4. 410e-211	1. 0000	-215. 5936
Asparti c_aci d(aq)	2. 043e-216	2. 719e-211	1. 0000	-215. 6896
Na(CH3COO)2-	1. 506e-216	2. 124e-211	0. 9295	-215. 8540
Cu(CH3COO)2(aq)	1. 301e-216	2. 363e-211	1. 0000	-215. 8857
Fe(CH3COO)2(aq)	8. 807e-217	1. 531e-211	1. 0000	-216. 0552
K(CH3COO)2-	3. 038e-217	4. 773e-212	0. 9295	-216. 5492
Ba(CH3COO)2(aq)	5. 736e-218	1. 465e-212	1. 0000	-217. 2414
Propanal (aq)	2. 221e-220	1. 290e-215	1. 0000	-219. 6534
Cu(CH3COO)2-	8. 814e-225	1. 601e-219	0. 9295	-224. 0865
1-Propyne(aq)	4. 120e-238	1. 650e-233	1. 0000	-237. 3851
S3--	2. 643e-242	2. 542e-237	0. 7455	-241. 7054
1-Propanol (aq)	2. 829e-243	1. 700e-238	1. 0000	-242. 5483
1-Propene(aq)	1. 298e-245	5. 461e-241	1. 0000	-244. 8867
2-Hydroxybutanoa	1. 057e-249	1. 090e-244	0. 9295	-249. 0076
NH4(CH3COO)2-	3. 004e-251	4. 088e-246	0. 9295	-250. 5540
Asparagi ne(aq)	9. 670e-252	1. 277e-246	1. 0000	-251. 0146
2-Hydroxybutanoi	3. 218e-253	3. 349e-248	1. 0000	-252. 4924
Cu(Gl y)2(aq)	5. 350e-254	1. 132e-248	1. 0000	-253. 2717
Mg(Gl y)2(aq)	3. 485e-260	6. 007e-255	1. 0000	-259. 4578
Di gl yci ne(aq)	2. 393e-260	3. 161e-255	1. 0000	-259. 6210
Propane(aq)	1. 403e-260	6. 187e-256	1. 0000	-259. 8528
Fe(Gl y)2(aq)	2. 226e-261	4. 540e-256	1. 0000	-260. 6524
Mn(Gl y)2(aq)	2. 164e-261	4. 393e-256	1. 0000	-260. 6647
Di ketopi perazi ne	1. 592e-262	1. 816e-257	1. 0000	-261. 7980
Butanoate	5. 047e-264	4. 395e-259	0. 9295	-263. 3287
Ca(Gl y)2(aq)	3. 982e-264	7. 492e-259	1. 0000	-263. 3999
Butanoi c_aci d(aq)	1. 509e-266	1. 329e-261	1. 0000	-265. 8213
Ba(Gl y)2(aq)	8. 735e-267	2. 493e-261	1. 0000	-266. 0587
Ca(But)+	7. 564e-267	9. 617e-262	0. 9295	-266. 1530
Na(But) (aq)	6. 367e-267	7. 007e-262	1. 0000	-266. 1960
Mg(But)+	2. 687e-267	2. 993e-262	0. 9295	-266. 6024
K(But) (aq)	2. 479e-267	3. 128e-262	1. 0000	-266. 6057
Mn(But)+	5. 429e-268	7. 708e-263	0. 9295	-267. 2971
Fe(But)+	1. 791e-268	2. 559e-263	0. 9295	-267. 7787
Cu(But)+	5. 873e-269	8. 845e-264	0. 9295	-268. 2629
Gl utarate	2. 999e-270	3. 900e-265	0. 7455	-269. 6506
Ba(But)+	2. 680e-270	6. 013e-265	0. 9295	-269. 6036
H-Gl utarate	3. 139e-272	4. 115e-267	0. 9295	-271. 5349
Threoni ne(aq)	1. 191e-272	1. 418e-267	1. 0000	-271. 9241
Gl utari c_aci d(aq)	3. 220e-275	4. 253e-270	1. 0000	-274. 4921
Ethyl acetate(aq)	2. 341e-275	2. 062e-270	1. 0000	-274. 6305

	SpecE8_output_GSS_MW-01-15 July.txt			
S506--	1. 321e-275	3. 385e-270	0. 7455	-275. 0067
1-Propanamine(aq)	2. 000e-282	1. 182e-277	1. 0000	-281. 6990
a-Ami nobutyri c_a	9. 783e-285	1. 009e-279	1. 0000	-284. 0095
Gl utami c_aci d(aq)	3. 425e-293	5. 038e-288	1. 0000	-292. 4653
UO2(SCN)2(aq)	5. 042e-299	1. 947e-293	1. 0000	-298. 2974
Butanal (aq)	1. 512e-300	1. 090e-295	1. 0000	-299. 8206
S4--	1. 089e-312	1. 396e-307	0. 7455	-300. 0000
1-Butyne(aq)	6. 690e-317	3. 618e-312	1. 0000	-300. 0000
Cu(CH3COO)3-	1. 665e-321	4. 006e-316	0. 9295	-300. 0000
Mn(CH3COO)3-	3. 755e-322	8. 711e-317	0. 9295	-300. 0000
1-Butanol (aq)	9. 387e-323	6. 956e-318	1. 0000	-300. 0000
Ba(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Cu(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
K(Prop)2-	0. 0000	0. 0000	0. 9295	-300. 0000
Cu(Al a)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
K(Pent)2-	0. 0000	0. 0000	0. 9295	-300. 0000
K(Pent) (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ba(Al a)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
K(Lac)2-	0. 0000	0. 0000	0. 9295	-300. 0000
2-Hexanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexyne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
K(But)2-	0. 0000	0. 0000	0. 9295	-300. 0000
Azel ate	0. 0000	0. 0000	0. 7455	-300. 0000
I sol euci ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Hexanoi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Hexanoate	0. 0000	0. 0000	0. 9295	-300. 0000
Azel ai c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Heptanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Hexanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Heptanoi c_aci d(a	0. 0000	0. 0000	1. 0000	-300. 0000
Heptanoate	0. 0000	0. 0000	0. 9295	-300. 0000
Heptanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
p-Tol ui c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
o-Tol ui c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
HSb2S4-	0. 0000	0. 0000	0. 9295	-300. 0000
o-Tol uate	0. 0000	0. 0000	0. 9295	-300. 0000
o-Phthal ate	0. 0000	0. 0000	0. 7455	-300. 0000
n-Propyl benzene(0. 0000	0. 0000	1. 0000	-300. 0000
n-Pentyl benzene(0. 0000	0. 0000	1. 0000	-300. 0000
2-Butanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Octyl benzene(a	0. 0000	0. 0000	1. 0000	-300. 0000
n-Octane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Hexyl benzene(a	0. 0000	0. 0000	1. 0000	-300. 0000
n-Heptyl benzene(0. 0000	0. 0000	1. 0000	-300. 0000
n-Heptane(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
n-Butyl benzene(a	0. 0000	0. 0000	1. 0000	-300. 0000
Al anyl gl yci ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
m-Tol ui c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
m-Tol uate	0. 0000	0. 0000	0. 9295	-300. 0000
Val i ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Undecanoi c_aci d(0. 0000	0. 0000	1. 0000	-300. 0000
Undecanoate	0. 0000	0. 0000	0. 9295	-300. 0000
Ca(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Pent)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Adi pi c_aci d(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Pent)+	0. 0000	0. 0000	0. 9295	-300. 0000
Ca(Lac)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Adi pate	0. 0000	0. 0000	0. 7455	-300. 0000
1-Hexanol (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
UO2(SCN)3-	0. 0000	0. 0000	0. 9295	-300. 0000
H-Suberate	0. 0000	0. 0000	0. 9295	-300. 0000
U(SCN)2++	0. 0000	0. 0000	0. 7477	-300. 0000

SpecE8_output_GSS_MW-01-15 July.txt

H-Sebacate	0.0000	0.0000	0.9295	-300.0000
H-Pi mel ate	0.0000	0.0000	0.9295	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
Tol uene(aq)	0.0000	0.0000	1.0000	-300.0000
Suberi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Suberate	0.0000	0.0000	0.7455	-300.0000
H-Azel ate	0.0000	0.0000	0.9295	-300.0000
H-Adi pate	0.0000	0.0000	0.9295	-300.0000
Sebaci c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacate	0.0000	0.0000	0.7455	-300.0000
Sb2S4--	0.0000	0.0000	0.7455	-300.0000
1-Hexanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butene(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
S5--	0.0000	0.0000	0.7455	-300.0000
Gl utami ne(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Pi mel i c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Pi mel ate	0.0000	0.0000	0.7455	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
Phenyl al ani ne(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol (aq)	0.0000	0.0000	1.0000	-300.0000
Pentanoi c_aci d(a	0.0000	0.0000	1.0000	-300.0000
Pentanoate	0.0000	0.0000	0.9295	-300.0000
Fe(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pentanal (aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxypentano	0.0000	0.0000	1.0000	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Pent)+	0.0000	0.0000	0.9295	-300.0000
Octanoate	0.0000	0.0000	0.9295	-300.0000
Octanal (aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoate	0.0000	0.0000	0.9295	-300.0000
2-Hydroxypentano	0.0000	0.0000	0.9295	-300.0000
Nonanal (aq)	0.0000	0.0000	1.0000	-300.0000
Benzoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanol (aq)	0.0000	0.0000	1.0000	-300.0000
Na(Prop)2-	0.0000	0.0000	0.9295	-300.0000
Na(Pent)2-	0.0000	0.0000	0.9295	-300.0000
Benzoate	0.0000	0.0000	0.9322	-300.0000
Na(Pent) (aq)	0.0000	0.0000	1.0000	-300.0000
Na(Lac)2-	0.0000	0.0000	0.9295	-300.0000
Tyrosi ne(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoa	0.0000	0.0000	0.9295	-300.0000
Fe(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Na(But)2-	0.0000	0.0000	0.9295	-300.0000
2-Hydroxynonanoi	0.0000	0.0000	1.0000	-300.0000
1-Pentanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9295	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_MW-01-15 July.txt

Dodecanoate	0.0000	0.0000	0.9295	-300.0000
Ba(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoi	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)+	0.0000	0.0000	0.9295	-300.0000
Mn(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9295	-300.0000
Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoa	0.0000	0.0000	0.9295	-300.0000
Decanal(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)+	0.0000	0.0000	0.9295	-300.0000
Mn(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	0.9295	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)+	0.0000	0.0000	0.9295	-300.0000
Mg(Pent)+	0.0000	0.0000	0.9295	-300.0000
Mg(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butane(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoic	0.0000	0.0000	1.0000	-300.0000
1-Octanol(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoic	0.0000	0.0000	0.9295	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
Methionine(aq)	0.0000	0.0000	1.0000	-300.0000
Leucylglycine(aq)	0.0000	0.0000	1.0000	-300.0000
Leucine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanamine(aq)	0.0000	0.0000	1.0000	-300.0000
p-Toluate	0.0000	0.0000	0.9295	-300.0000
1-Butanamine(aq)	0.0000	0.0000	1.0000	-300.0000
(NH4)2Sb2S4(aq)	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states

log Q/K

log Q/K

Fluorapatite	4.4547s/sat	UO2F2	-16.2942
Witherite	1.2729s/sat	BaCl2	-16.5006
Quartz	0.3897s/sat	MgOHCl	-16.6410
Baryte	0.2517s/sat	Natrosilite	-16.7756
Tridymite	0.1935s/sat	MnCl2·2H2O	-17.2331
Chalcedony	0.1158s/sat	Na2SiO3	-17.2704
Ice	-0.1278	Ba2Si3O8	-17.7791
Cristobalite(alp)	-0.1674	Hydromagnesite	-18.0045
Tenoriite	-0.2538	Hillebrandite	-18.3446
Coesite	-0.4278	Dicalciumsilicate	-18.7223
Rhodochrosite	-0.5504	MnCl2·H2O	-18.8051
Cristobalite(bet)	-0.6194	CuCl2	-18.8686
Calcite	-0.7543	UO2SO4	-19.0692
Dolomite	-0.8253	MgCl2·4H2O	-19.2401
Dolomite-ord	-0.8253	Akermanite	-19.8965
Aragonite	-0.8987	Larnite	-20.0326

SpecE8_output_GSS_MW-01-15 July.txt

Si O2(am)	-0.9211	UO2.3333(beta)	-20.0910
Malachi te	-1.1531	Na2UO4(al pha)	-20.3766
Si deri te	-1.4690	Ni ngyoi te	-21.1976
Monohydrocal ci te	-1.5807	Anthophyl li te	-21.6380
Magnesi te	-1.7176	Li me	-21.6653
Bi xbyi te	-2.2272	UO2(NO3)2: 6H2O	-21.7848
Di optase	-2.2938	Scacchi te	-22.0679
Dol omi te-di s	-2.3910	UOFOH: . 5H2O	-22.3350
Gypsum	-2.3961	UOFOH	-22.8005
Anhydri te	-2.5994	UO2Cl	-23.0213
Ferrosi li te	-2.7104	UO2(NO3)2: 3H2O	-23.2317
Pyrol usi te	-2.8799	Hydrophi li te	-23.2499
Bassani te	-3.2454	Lawrenci te	-23.3025
Mi nnesotai te	-3.3193	UO2(NO3)2: 2H2O	-24.4872
Fl uori te	-3.3208	MgCl 2: 2H2O	-24.7399
CaSO4: 0. 5H2O(bet	-3.4171	K2UO4	-24.9758
Schoepi te	-3.7538	UOF2: H2O	-25.2756
CaUO4	-3.7567	UO2(P03)2	-25.5274
UO3: 2H2O	-3.7567	(UO2)3(P04)2	-25.6110
Tal c	-3.7783	UOF2	-25.8492
Brochanti te	-3.7876	Ranki ni te	-26.0237
UO2(OH)2(beta)	-3.8789	(UO2)2P207	-26.0379
UO3: . 9H2O(al pha)	-3.9434	UO2Cl 2: 3H2O	-27.0967
Schoepi te-dehy(.	-3.9480	UO2(NO3)2: H2O	-28.1034
Schoepi te-dehy(.	-4.0246	MgCl 2: H2O	-28.1290
Schoepi te-dehy(1	-4.0326	UP207	-29.2314
Hausmanni te	-4.0858	U(HP04)2: 4H2O	-29.2675
Rhodoni te	-4.1070	UO2Cl 2: H2O	-29.8236
Sb203	-4.2838	Tobermori te-14A	-30.0174
Enstati te	-4.3756	UO2(NO3)2	-31.5982
Sel lai te	-4.5025	Tobermori te-11A	-31.8766
Nesquehoni te	-4.7243	Ba2Si O4	-31.9576
Greenal i te	-4.7811	Merwi ni te	-32.0898
UO2C03	-4.8489	Foshagi te	-32.5575
Rutherfordi ne	-4.8701	KMgCl 3: 2H2O	-33.1885
Whi tlocki te	-5.1054	UO2Cl 2	-33.7288
Schoepi te-dehy(.	-5.1491	Chl oromagnesi te	-33.9389
Fe0	-5.2121	Afwil li te	-34.1880
Hydroxyl apati te	-5.3279	UO2S03	-34.3260
Fe(OH)2	-5.5669	UF4: 2. 5H2O	-35.1321
Nahcol i te	-5.5804	Tobermori te-9A	-35.4173
Schoepi te-dehy(.	-5.6734	BaSi F6	-36.8928
Bruci te	-5.7304	Fe	-37.4435
Fayal i te	-6.0349	UF4	-39.3737
Ni ter	-6.0523	BaO	-39.9161
Mn(OH)2(am)	-6.0722	KMgCl 3	-40.5681
Wol l astoni te	-6.3127	U(SO4)2: 8H2O	-43.0557
Okeni te	-6.5054	Hatruri te	-44.0852
Di opsi de	-6.5211	U(SO4)2: 4H2O	-44.1418
Chrysoti le	-6.5536	U(SO4)2	-44.2019
Pseudowol l astoni	-6.5569	Ba2U207	-44.5887
UO3(gamma)	-6.6821	Xonotli te	-47.2993
Atacami te	-7.0223	U(CO3)2	-47.3224
Azuri te	-7.0482	C	-47.5000
UO2HP04: 4H2O	-7.2276	UOCl 2	-47.5549
Mel anteri te	-7.2530	UCl F3	-50.1276
UO3(beta)	-7.2908	UOF4	-52.9249
Cupri te	-7.3348	U5O12Cl	-53.5951
Hedenbergi te	-7.4048	Na3UO4	-55.8125
Hunti te	-7.4994	Na	-56.7817
UO3(al pha)	-7.6251	Na4Si O4	-57.3170
Chal canthi te	-7.9858	(UO2)2Cl 3	-57.8778
Mi rabi li te	-8.0789	UOCl 3	-58.8764

SpecE8_output_GSS_MW-01-15 July.txt

Syl vi te	-8. 0864	Na2O	-59. 2564
Tephroi te	-8. 2194	Mn	-60. 6862
Arcani te	-8. 2577	UF5(beta)	-60. 6958
Arti ni te	-8. 4696	K	-60. 7858
Hal i te	-8. 4699	UF5(al pha)	-61. 0437
Sanborni te	-8. 5051	UF3	-61. 2138
U02F0H: 2H2O	-8. 5262	UCI 2F2	-63. 1970
Manganosi te	-8. 7320	UOCl	-65. 5539
BaUO4	-8. 9111	Ba3UO6	-69. 0843
U02F0H: H2O	-8. 9188	U2O3F6	-69. 3692
Thenardi te	-9. 0336	S	-71. 1977
Nantoki te	-9. 0945	Covel l i te	-72. 4994
U02F0H	-9. 3768	UCI 3F	-76. 0894
Sepi ol i te	-9. 6744	K2O	-76. 7607
Ni trobari te	-9. 7821	Chal coci te	-77. 2196
Mg1. 25S04(OH)0. 5	-10. 0986	Na6Si 207	-83. 4274
Forsteri te	-10. 3629	U(SO3)2	-85. 4773
Sal eei te	-10. 5181	UCI 4	-86. 8123
Cu	-10. 6376	UF6	-90. 6009
Peri cl ase	-10. 8344	Troi l i te	-90. 7117
Natron	-10. 8550	UCI 3	-90. 7695
U02. 25	-10. 9973	Pyrrhoti te	-90. 8117
U02S04: H2O	-11. 0019	Anti gori te	-91. 1631
U02. 25(beta)	-11. 0596	U3O5F8	-92. 7224
Na2C03: 7H2O	-11. 2313	Al abandi te	-93. 2655
MgUO4	-11. 4207	U2F9	-96. 9070
Mg1. 5S04(OH)	-11. 4530	Mg	-99. 3397
MnS04	-11. 4883	UCI 5	-106. 3783
Portl andi te	-11. 5319	U2O2Cl 5	-107. 6342
Gyrol i te	-11. 5617	UN1. 73(al pha)	-110. 6359
Monti cel l i te	-11. 5824	BaS	-111. 0869
BaU2O7	-11. 5956	UN1. 59(al pha)	-113. 3393
Na2U2O7	-11. 7777	Ca	-116. 2982
Thermonatri te	-12. 3372	Ba	-120. 5791
MgS04	-12. 3956	P	-120. 6842
FeS04	-12. 4013	UCI 6	-124. 7834
(UO2)3(P04)2: 4H2	-12. 4256	UN	-126. 1350
U02Cl OH: 2H2O	-12. 4805	Si	-126. 6145
Uranoci rci te	-12. 5094	Pyri te	-151. 4714
Na2C03	-12. 5768	Chal copyri te	-157. 2407
Torberni te	-12. 8778	U	-172. 6566
U02. 6667	-13. 1946	U4F17	-174. 7821
Chal cocyani te	-13. 6621	US	-187. 0692
FeF2	-13. 8313	UC	-202. 6594
UP05	-14. 0213	UH3(beta)	-204. 3249
BaCl 2: 2H2O	-14. 4252	US1. 9	-217. 7894
Basseti te	-14. 5164	US2	-223. 1791
Urani ni te	-14. 5859	UP	-246. 2837
Tremol i te	-14. 9731	UC1. 94(al pha)	-249. 3459
Coffi ni te	-15. 0078	Sti bni te	-268. 7128
BaCl 2: H2O	-15. 0630	US3	-291. 1878
Autuni te-H	-15. 2668	Borni te	-311. 2840
U02S04: 3. 5H2O	-15. 5292	UP2	-361. 8918
U02S04: 2. 5H2O	-15. 5328	U2S3	-403. 2440
U02S04: 3H2O	-15. 6189	o-Phthal i c_aci d	-413. 5692
MnCl 2: 4H2O	-15. 9436	U2C3	-454. 3266
U02F2: 3H2O	-16. 1192	U3S5	-621. 8341
Ba(OH)2: 8H2O	-16. 2318	U3P4	-854. 4535
NaUO3	-16. 2612		

Gases	fugaci ty	l og fug.
N2(g)	0. 1724	-0. 763

SpecE8_output_GSS_MW-01-15 July.txt

H2O(g)	0.02175	-1.663
CO2(g)	0.004857	-2.314
NO2(g)	1.096e-012	-11.960
HF(g)	2.925e-014	-13.534
HCl(g)	1.921e-018	-17.716
O2(g)	2.586e-025	-24.587
NO(g)	1.045e-028	-27.981
H2(g)	1.750e-030	-29.757
CO(g)	2.691e-036	-35.570
SiF4(g)	1.531e-038	-37.815
Cl2(g)	4.159e-040	-39.381
NH3(g)	8.807e-043	-42.055
SO2(g)	2.105e-043	-42.677
UO2F2(g)	1.923e-059	-58.716
Cu(g)	4.395e-064	-63.357
UO2Cl2(g)	9.420e-071	-70.026
Na(g)	3.679e-071	-70.434
UO3(g)	2.176e-071	-70.662
K(g)	3.014e-072	-71.521
UOF4(g)	1.208e-073	-72.918
UF5(g)	2.820e-081	-80.550
UF4(g)	1.919e-084	-83.717
UF6(g)	3.083e-092	-91.511
F2(g)	1.860e-095	-94.731
H2S(g)	9.132e-096	-95.039
CH4(g)	1.004e-098	-97.998
UCl4(g)	2.608e-112	-111.584
UO2(g)	7.356e-113	-112.133
Mg(g)	5.345e-120	-119.272
UCl5(g)	1.356e-124	-123.868
UF3(g)	2.300e-129	-128.638
UCl6(g)	1.543e-131	-130.812
U2F10(g)	1.996e-136	-135.700
UCl3(g)	1.083e-137	-136.966
Ca(g)	1.602e-142	-141.795
S2(g)	3.080e-157	-156.511
C(g)	4.893e-167	-166.310
C2H4(g)	2.892e-167	-166.539
UF2(g)	4.461e-171	-170.351
UCl2(g)	1.676e-180	-179.776
UO(g)	4.664e-186	-185.331
Si(g)	3.858e-199	-198.414
UF(g)	2.561e-206	-205.592
U2Cl8(g)	2.642e-213	-212.578
UCl(g)	6.228e-221	-220.206
U2Cl10(g)	8.280e-222	-221.082
U(g)	7.106e-260	-259.148

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
<hr/>						
Ba++	7.61e-007	7.61e-007	0.104			
Ca++	0.000615	0.000615	24.6			
Cl-	0.000111	0.000111	3.94			
Cu++	1.09e-007	1.09e-007	0.00690			
F-	1.04e-005	1.04e-005	0.197			
Fe++	1.06e-006	1.06e-006	0.0591			
H+	0.000178	0.000178	0.180			
H2O	55.5	55.5	1.00e+006			
HCO3-	0.00186	0.00186	113.			
HP04--	3.18e-007	3.18e-007	0.0305			
K+	0.000555	0.000555	21.7			
Mg++	0.000203	0.000203	4.93			

SpecE8_output_GSS_MW-01-15 July.txt

Mn++	8.81e-006	8.81e-006	0.484
NH3(aq)	0.000238	0.000238	4.05
NO3-	0.000992	0.000992	61.5
Na+	0.00133	0.00133	30.5
O2(aq)	0.000178	0.000178	5.70
SO4--	0.000421	0.000421	40.4
Sb(OH)3(aq)	2.10e-007	2.10e-007	0.0364
SiO2(aq)	0.000202	0.000202	12.1
UO2++	8.28e-009	8.28e-009	0.00224

Elemental composition	In fluid		Sorbed	
	total moles	moles	moles	mg/kg
Antimony	2.105e-007	2.105e-007	0.02562	
Barium	7.608e-007	7.608e-007	0.1044	
Calcium	0.0006148	0.0006148	24.63	
Carbon	0.001858	0.001858	22.30	
Chlorine	0.0001112	0.0001112	3.941	
Copper	1.086e-007	1.086e-007	0.006897	
Fluorine	1.038e-005	1.038e-005	0.1971	
Hydrogen	111.0	111.0	1.119e+005	
Iron	1.059e-006	1.059e-006	0.05912	
Magnesium	0.0002028	0.0002028	4.926	
Manganese	8.809e-006	8.809e-006	0.4838	
Nitrogen	0.001230	0.001230	17.22	
Oxygen	55.52	55.52	8.880e+005	
Phosphorus	3.182e-007	3.182e-007	0.009853	
Potassium	0.0005546	0.0005546	21.68	
Silicon	0.0002018	0.0002018	5.665	
Sodium	0.001329	0.001329	30.54	
Sulfur	0.0004207	0.0004207	13.48	
Uranium	8.281e-009	8.281e-009	0.001971	

SpecE8_output_GSS_MW-01-15 June.txt

Temperature = 4.1 C Pressure = 1.013 bars
 pH = 7.500 log fO2 = -34.124
 Eh = 0.3510 volts pe = 6.3808
 Ionic strength = 0.004546
 Charge imbalance = -0.000016 eq/kg (-0.4658% error)
 Activity of water = 0.999996
 Solvent mass = 1.000000 kg
 Solution mass = 1.000307 kg
 Solution density = 1.027 g/cm3
 Chlorinity = 0.000110 molal
 Dissolved solids = 307 mg/kg sol'n
 Hardness = 75.96 mg/kg sol'n as CaCO3
 carbonate = 75.61 mg/kg sol'n as CaCO3
 non-carbonate = 0.35 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 75.61 mg/kg sol'n as CaCO3
 Water type = Na-HCO3

Nernst redox couples	Eh (volts)	pe
$e^- + .25 \cdot O_2(aq) + H^+ = .5 \cdot H_2O$	0.3510	6.3808
$8 \cdot e^- + 9 \cdot H^+ + NO_3^- = 3 \cdot H_2O + NH_3(aq)$	0.5785	10.5161
$14 \cdot e^- + 16 \cdot H^+ + 2 \cdot SO_4^{--} = 8 \cdot H_2O + S^{2--}$	-0.2387	-4.3391

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HC03-	0.001495	91.19	0.9320	-2.8560
Na+	0.001352	31.06	0.9320	-2.8997
NO3-	0.001008	62.49	0.9306	-3.0277
K+	0.0005716	22.34	0.9306	-3.2741
Ca++	0.0005365	21.50	0.7619	-3.3885
SO4--	0.0003931	37.75	0.7535	-3.5284
Mg++	0.0001910	4.640	0.7697	-3.8327
SiO2(aq)	0.0001844	11.08	1.0000	-3.7342
CO2(aq)	0.0001488	6.548	1.0000	-3.8273
N2(aq)	0.0001340	3.754	1.0000	-3.8727
Cl-	0.0001098	3.893	0.9306	-3.9905
F-	1.529e-005	0.2905	0.9313	-4.8464
CaSO4(aq)	1.432e-005	1.949	1.0000	-4.8440
Mn++	7.946e-006	0.4364	0.7619	-5.2180
CaHC03+	7.307e-006	0.7385	0.9320	-5.1669
MgSO4(aq)	6.614e-006	0.7959	1.0000	-5.1795
NaHC03(aq)	3.798e-006	0.3189	1.0000	-5.4205
MgHC03+	2.537e-006	0.2164	0.9320	-5.6263
CO3--	1.589e-006	0.09530	0.7557	-5.9207
KS04-	1.277e-006	0.1725	0.9320	-5.9244
Ba++	7.586e-007	0.1041	0.7578	-6.2405
CaCO3(aq)	6.970e-007	0.06974	1.0000	-6.1568
S2--	6.074e-007	0.03894	0.7535	-6.3394
HP04--	5.220e-007	0.05009	0.7535	-6.4052
HSiO3-	3.416e-007	0.02632	0.9320	-6.4971
MnSO4(aq)	3.313e-007	0.05001	1.0000	-6.4798
H2PO4-	2.640e-007	0.02560	0.9320	-6.6090
Sb(OH)3(aq)	2.320e-007	0.04006	1.0000	-6.6346
MgCO3(aq)	1.386e-007	0.01168	1.0000	-6.8584
Cu++	1.379e-007	0.008761	0.7619	-6.9785
OH-	5.803e-008	0.0009866	0.9313	-7.2673
MgF+	5.289e-008	0.002290	0.9320	-7.3072
H+	3.371e-008	3.397e-005	0.9381	-7.5000
NaHSiO3(aq)	2.904e-008	0.002905	1.0000	-7.5370

	SpecE8_output_GSS_MW-01-15 June.txt			
CaF+	2. 800e-008	0. 001654	0. 9320	-7. 5834
NaCl (aq)	1. 936e-008	0. 001131	1. 0000	-7. 7132
MgCl +	1. 360e-008	0. 0008122	0. 9320	-7. 8972
NaCO3-	9. 432e-009	0. 0007826	0. 9320	-8. 0560
CaCl +	9. 285e-009	0. 0007011	0. 9320	-8. 0628
UO2(CO3)2--	4. 031e-009	0. 001572	0. 7535	-8. 5175
UO2(CO3)3----	3. 131e-009	0. 001409	0. 3220	-8. 9964
NaF (aq)	1. 525e-009	6. 402e-005	1. 0000	-8. 8167
KCl (aq)	1. 154e-009	8. 600e-005	1. 0000	-8. 9378
MnCl +	8. 440e-010	7. 627e-005	0. 9320	-9. 1042
HSO4-	5. 738e-010	5. 568e-005	0. 9320	-9. 2719
HF (aq)	4. 633e-010	9. 266e-006	1. 0000	-9. 3342
UO2CO3(aq)	1. 999e-010	6. 596e-005	1. 0000	-9. 6992
BaCO3(aq)	1. 978e-010	3. 901e-005	1. 0000	-9. 7038
BaCl +	1. 538e-011	2. 656e-006	0. 9320	-10. 8437
NaOH(aq)	1. 271e-011	5. 081e-007	1. 0000	-10. 8959
Cu+	1. 196e-011	7. 600e-007	0. 9320	-10. 9527
PO4---	6. 659e-012	6. 323e-007	0. 5288	-11. 4533
BaF+	4. 741e-012	7. 409e-007	0. 9320	-11. 3547
Sb(OH)2F(aq)	3. 159e-012	5. 518e-007	1. 0000	-11. 5005
UO2OH+	2. 020e-012	5. 795e-007	0. 9320	-11. 7253
Sb(OH)4-	1. 794e-012	3. 404e-007	0. 9320	-11. 7768
CaCl 2(aq)	1. 357e-012	1. 506e-007	1. 0000	-11. 8674
H3PO4(aq)	9. 443e-013	9. 251e-008	1. 0000	-12. 0249
HN03(aq)	9. 387e-013	5. 913e-008	1. 0000	-12. 0275
HCl (aq)	7. 053e-013	2. 571e-008	1. 0000	-12. 1516
Sb(OH)2+	5. 560e-013	8. 658e-008	0. 9320	-12. 2855
UO2F+	6. 240e-014	1. 803e-008	0. 9320	-13. 2354
UO2++	4. 815e-014	1. 300e-008	0. 7557	-13. 4390
MgP2O7--	3. 046e-014	6. 038e-009	0. 7535	-13. 6391
KHSO4(aq)	1. 539e-014	2. 095e-009	1. 0000	-13. 8129
UO2SO4(aq)	7. 532e-015	2. 756e-009	1. 0000	-14. 1231
UO2F2(aq)	2. 652e-015	8. 166e-010	1. 0000	-14. 5765
HP2O7---	1. 580e-015	2. 764e-010	0. 5288	-15. 0779
HF2-	1. 343e-015	5. 239e-011	0. 9320	-14. 9024
H2P2O7--	1. 549e-016	2. 725e-011	0. 7535	-15. 9328
P2O7----	4. 202e-017	7. 307e-012	0. 3220	-16. 8687
UO2(SO4)2--	1. 641e-017	7. 583e-012	0. 7535	-16. 9077
UO2F3-	6. 854e-018	2. 241e-012	0. 9320	-17. 1946
UO2Cl +	4. 738e-018	1. 447e-012	0. 9320	-17. 3550
KP2O7---	2. 445e-018	5. 208e-013	0. 5288	-17. 8884
(UO2)2(OH)2++	1. 216e-018	6. 978e-013	0. 7557	-18. 0367
UO2+	3. 337e-019	9. 009e-014	0. 9320	-18. 5072
(UO2)3(CO3)6(6-)	2. 204e-020	2. 578e-014	0. 0781	-20. 7643
(UO2)3(OH)5+	1. 919e-020	1. 717e-014	0. 9320	-19. 7476
H3P2O7-	9. 054e-022	1. 602e-016	0. 9320	-21. 0738
UO2F4--	7. 818e-022	2. 704e-016	0. 7535	-21. 2298
U(OH)4(aq)	5. 049e-022	1. 545e-016	1. 0000	-21. 2968
UO2Cl 2(aq)	1. 861e-023	6. 343e-018	1. 0000	-22. 7303
Mn+++	1. 360e-025	7. 469e-021	0. 5355	-25. 1377
H4P2O7(aq)	6. 429e-028	1. 144e-022	1. 0000	-27. 1918
Formate	2. 407e-031	1. 083e-026	0. 9313	-30. 6494
H2(aq)	5. 011e-032	1. 010e-028	1. 0000	-31. 3001
Ca(For)+	3. 120e-033	2. 654e-028	0. 9320	-32. 5364
Mg(For)+	1. 480e-033	1. 026e-028	0. 9320	-32. 8603
Na(For) (aq)	3. 448e-034	2. 344e-029	1. 0000	-33. 4624
K(For) (aq)	1. 210e-034	1. 018e-029	1. 0000	-33. 9171
Mn(For)+	9. 620e-035	9. 613e-030	0. 9320	-34. 0474
NO2-	7. 017e-035	3. 227e-030	0. 9306	-34. 1851
Formic acid(aq)	4. 296e-035	1. 976e-030	1. 0000	-34. 3670
NH4+	2. 398e-035	4. 325e-031	0. 9299	-34. 6517
SO3--	7. 165e-036	5. 735e-031	0. 7557	-35. 2665
Si F6--	4. 921e-036	6. 989e-031	0. 7535	-35. 4308

	SpecE8_output_GSS_MW-01-15 June.txt			
Cu(For) +	3. 581e-036	3. 886e-031	0. 9320	-35. 4766
Ba(For) +	3. 091e-036	5. 634e-031	0. 9320	-35. 5406
HSO3-	2. 503e-036	2. 028e-031	0. 9320	-35. 6322
O2(aq)	1. 483e-037	4. 743e-033	1. 0000	-36. 8289
NH3(aq)	8. 412e-038	1. 432e-033	1. 0000	-37. 0751
CO(aq)	6. 136e-038	1. 718e-033	1. 0000	-37. 2121
Oxal ate	7. 566e-039	6. 657e-034	0. 7535	-38. 2440
HN02(aq)	5. 645e-039	2. 653e-034	1. 0000	-38. 2483
Cl O-	4. 224e-039	2. 172e-034	0. 9320	-38. 4049
UF3+	1. 946e-039	5. 738e-034	0. 9320	-38. 7416
UF2++	8. 459e-040	2. 334e-034	0. 7557	-39. 1943
UOH+++	7. 124e-040	1. 816e-034	0. 5355	-39. 4186
UF4(aq)	3. 024e-040	9. 495e-035	1. 0000	-39. 5193
U(CO3)5(6-)	1. 496e-040	8. 045e-035	0. 0781	-40. 9326
H02-	1. 061e-041	3. 501e-037	0. 9320	-41. 0049
UF+++	1. 046e-041	2. 689e-036	0. 5355	-41. 2515
S02(aq)	3. 542e-042	2. 268e-037	1. 0000	-41. 4508
H-Oxal ate	3. 105e-042	2. 763e-037	0. 9320	-41. 5386
US04++	1. 976e-043	6. 601e-038	0. 7557	-42. 8258
U(S04)2(aq)	1. 571e-043	6. 754e-038	1. 0000	-42. 8039
U++++	5. 465e-046	1. 300e-040	0. 3329	-45. 7400
Mn04--	3. 810e-047	4. 530e-042	0. 7535	-46. 5420
UCl +++	3. 339e-048	9. 129e-043	0. 5355	-47. 7476
Oxal i c_aci d(aq)	1. 662e-048	1. 496e-043	1. 0000	-47. 7794
HS05-	3. 147e-049	3. 558e-044	0. 9320	-48. 5326
Mn04-	1. 979e-051	2. 353e-046	0. 9313	-50. 7345
S206--	2. 738e-060	4. 383e-055	0. 7535	-59. 6854
Ca(For)2(aq)	6. 304e-063	8. 200e-058	1. 0000	-62. 2004
Mg(For)2(aq)	4. 112e-063	4. 700e-058	1. 0000	-62. 3860
U+++	1. 557e-063	3. 704e-058	0. 5355	-63. 0791
Mn(For)2(aq)	3. 618e-064	5. 244e-059	1. 0000	-63. 4415
Na(For)2-	6. 929e-065	7. 829e-060	0. 9320	-64. 1899
Cu(For)2(aq)	2. 590e-065	3. 976e-060	1. 0000	-64. 5867
K(For)2-	1. 915e-065	2. 472e-060	0. 9320	-64. 7484
S208--	4. 438e-066	8. 524e-061	0. 7535	-65. 4757
Ba(For)2(aq)	4. 321e-066	9. 821e-061	1. 0000	-65. 3644
Cl O2-	3. 767e-066	2. 540e-061	0. 9320	-65. 4546
Formal dehyde(aq)	8. 256e-068	2. 478e-063	1. 0000	-67. 0832
N3-	2. 786e-073	1. 171e-068	0. 9320	-72. 5855
Cu(NH3)2++	1. 033e-073	1. 008e-068	0. 7557	-73. 1076
HN3(aq)	6. 608e-076	2. 843e-071	1. 0000	-75. 1799
S205--	1. 176e-076	1. 695e-071	0. 7535	-76. 0524
Urea(aq)	4. 305e-078	2. 585e-073	1. 0000	-77. 3660
Cl O3-	2. 616e-078	2. 183e-073	0. 9313	-77. 6132
HCN(aq)	1. 818e-080	4. 913e-076	1. 0000	-79. 7403
CN-	8. 490e-083	2. 208e-078	0. 9306	-82. 1023
Methanol (aq)	4. 560e-083	1. 461e-078	1. 0000	-82. 3410
HS-	9. 329e-088	3. 085e-083	0. 9313	-87. 0611
H2S(aq)	5. 909e-088	2. 013e-083	1. 0000	-87. 2284
Gl ycol ate	3. 281e-089	2. 461e-084	0. 9320	-88. 5146
Ca(Gl yc) +	7. 623e-091	8. 773e-086	0. 9320	-90. 1485
UO2Cl O3+	3. 407e-091	1. 204e-085	0. 9320	-90. 4983
Mg(Gl yc) +	1. 250e-091	1. 242e-086	0. 9320	-90. 9335
Na(Gl yc) (aq)	4. 886e-092	4. 789e-087	1. 0000	-91. 3110
K(Gl yc) (aq)	1. 715e-092	1. 957e-087	1. 0000	-91. 7657
Mn(Gl yc) +	1. 018e-092	1. 323e-087	0. 9320	-92. 0229
Methane(aq)	7. 419e-093	1. 190e-088	1. 0000	-92. 1297
Gl ycol i c_aci d(aq)	7. 154e-093	5. 439e-088	1. 0000	-92. 1454
Cu(Gl yc) +	4. 835e-093	6. 698e-088	0. 9320	-92. 3462
S203--	1. 017e-093	1. 140e-088	0. 7535	-93. 1157
S--	9. 645e-094	3. 092e-089	0. 7578	-93. 1361
Ba(Gl yc) +	2. 068e-094	4. 389e-089	0. 9320	-93. 7151
Cl O4-	6. 217e-095	6. 181e-090	0. 9313	-94. 2373

	SpecE8_output_GSS_MW-01-15 June.txt			
S204--	8.136e-096	1.042e-090	0.7578	-95.2101
Acetate	2.160e-098	1.275e-093	0.9326	-97.6958
MgCH3COO+	8.947e-101	7.455e-096	0.9320	-100.0789
CaCH3COO+	8.268e-101	8.193e-096	0.9320	-100.1132
Aceti c_aci d(aq)	3.779e-101	2.268e-096	1.0000	-100.4227
NaCH3COO(aq)	2.204e-101	1.807e-096	1.0000	-100.6569
KCH3COO(aq)	5.447e-102	5.344e-097	1.0000	-101.2639
MnCH3COO+	2.715e-102	3.094e-097	0.9320	-101.5968
CuCH3COO+	6.232e-103	7.638e-098	0.9320	-102.2360
BaCH3COO+	1.094e-103	2.147e-098	0.9320	-102.9917
Mal onate	2.502e-107	2.553e-102	0.7535	-106.7245
Cu(NH3)3++	1.087e-107	1.245e-102	0.7557	-107.0856
CuCH3COO(aq)	6.021e-109	7.378e-104	1.0000	-108.2204
H-Mal onate	2.981e-109	3.071e-104	0.9320	-108.5562
Mal oni c_aci d(aq)	6.660e-114	6.929e-109	1.0000	-113.1765
Methanami ne(aq)	1.972e-117	6.123e-113	1.0000	-116.7050
Gl yci ne(aq)	7.311e-120	5.486e-115	1.0000	-119.1360
Cu(Gl y)+	9.714e-121	1.336e-115	0.9320	-120.0432
Mg(Gl y)+	6.006e-123	5.906e-118	0.9320	-122.2520
Mn(Gl y)+	3.656e-124	4.715e-119	0.9320	-123.4675
Ca(Gl y)+	6.985e-125	7.970e-120	0.9320	-124.1864
S306--	3.557e-126	6.834e-121	0.7535	-125.5718
Ba(Gl y)+	1.107e-127	2.339e-122	0.9320	-126.9864
SCN-	3.008e-132	1.747e-127	0.9313	-131.5526
Acetal dehyde(aq)	1.208e-132	5.319e-128	1.0000	-131.9180
NH4CH3COO(aq)	7.443e-133	5.735e-128	1.0000	-132.1283
Acetami de(aq)	5.648e-133	3.335e-128	1.0000	-132.2481
UO2SCN+	2.649e-144	8.690e-139	0.9320	-143.6075
Ethanol (aq)	1.158e-152	5.331e-148	1.0000	-151.9365
Ethyne(aq)	1.907e-155	4.965e-151	1.0000	-154.7195
Ethyl ene(aq)	1.317e-157	3.694e-153	1.0000	-156.8803
Lactate	8.864e-159	7.893e-154	0.9320	-158.0830
Ca(Lac)+	1.211e-160	1.563e-155	0.9320	-159.9476
Mg(Lac)+	3.270e-161	3.707e-156	0.9320	-160.5160
Na(Lac) (aq)	1.361e-161	1.525e-156	1.0000	-160.8660
K(Lac) (aq)	4.779e-162	6.123e-157	1.0000	-161.3207
Lacti c_aci d(aq)	2.007e-162	1.807e-157	1.0000	-161.6974
Mn(Lac)+	1.948e-162	2.805e-157	0.9320	-161.7410
Cu(Lac)+	5.941e-163	9.065e-158	0.9320	-162.2567
Ba(Lac)+	2.250e-164	5.091e-159	0.9320	-163.6785
Ethane(aq)	6.359e-167	1.912e-162	1.0000	-166.1966
Propanoate	3.309e-170	2.417e-165	0.9320	-169.5109
Propanoi c_aci d(a	7.737e-173	5.730e-168	1.0000	-172.1114
Ca(Prop)+	7.089e-173	8.019e-168	0.9320	-172.1800
Na(Prop) (aq)	5.016e-173	4.817e-168	1.0000	-172.2996
Mg(Prop)+	3.709e-173	3.610e-168	0.9320	-172.4613
K(Prop) (aq)	1.761e-173	1.974e-168	1.0000	-172.7543
Mn(Prop)+	4.947e-174	6.331e-169	0.9320	-173.3362
USCN+++	2.157e-174	6.386e-169	0.5355	-173.9374
Cu(Prop)+	9.447e-175	1.290e-169	0.9320	-174.0553
S406--	2.310e-175	5.178e-170	0.7535	-174.7593
Ba(Prop)+	2.139e-176	4.499e-171	0.9320	-175.7004
Succi nate	1.686e-177	1.957e-172	0.7535	-176.8959
Ca(Gl yc)2(aq)	5.346e-178	1.016e-172	1.0000	-177.2720
Mg(Gl yc)2(aq)	5.692e-179	9.923e-174	1.0000	-178.2447
H-Succi nate	2.001e-179	2.342e-174	0.9320	-178.7294
Cu(Gl yc)2(aq)	1.477e-179	3.155e-174	1.0000	-178.8306
Mn(Gl yc)2(aq)	2.920e-180	5.985e-175	1.0000	-179.5346
Na(Gl yc)2-	2.120e-180	3.668e-175	0.9320	-179.7043
Seri ne(aq)	8.063e-181	8.471e-176	1.0000	-180.0935
K(Gl yc)2-	6.086e-181	1.151e-175	0.9320	-180.2462
Ba(Gl yc)2(aq)	3.088e-182	8.872e-177	1.0000	-181.5103
Succi ni c_aci d(aq)	1.101e-182	1.299e-177	1.0000	-181.9583

SpecE8_output_GSS_MW-01-15 June.txt				
Ethanamine(aq)	5.061e-189	2.281e-184	1.0000	-188.2957
Alanine(aq)	2.681e-190	2.388e-185	1.0000	-189.5717
Cu(Ala)+	8.088e-192	1.226e-186	0.9320	-191.1228
Mg(Ala)+	1.355e-195	1.522e-190	0.9320	-194.8987
Mn(Ala)+	1.074e-195	1.536e-190	0.9320	-194.9995
Ca(Ala)+	4.806e-196	6.158e-191	0.9320	-195.3488
Ca(CH3COO)2(aq)	3.685e-197	5.827e-192	1.0000	-196.4336
Mg(CH3COO)2(aq)	1.824e-197	2.597e-192	1.0000	-196.7389
Mn(CH3COO)2(aq)	5.407e-199	9.352e-194	1.0000	-198.2671
Cu(CH3COO)2(aq)	5.370e-199	9.751e-194	1.0000	-198.2700
Na(CH3COO)2-	3.226e-199	4.550e-194	0.9320	-198.5220
Ba(Ala)+	2.757e-199	6.212e-194	0.9320	-198.5902
Acetone(aq)	2.063e-199	1.198e-194	1.0000	-198.6856
K(CH3COO)2-	4.955e-200	7.786e-195	0.9320	-199.3356
Aspartic acid(aq)	9.147e-201	1.217e-195	1.0000	-200.0387
Ba(CH3COO)2(aq)	6.832e-201	1.744e-195	1.0000	-200.1655
Propanal(aq)	6.968e-204	4.046e-199	1.0000	-203.1569
Cu(CH3COO)2-	2.265e-206	4.112e-201	0.9320	-205.6756
S3--	2.523e-220	2.427e-215	0.7535	-219.7209
1-Propyne(aq)	6.120e-223	2.451e-218	1.0000	-222.2133
1-Propanol(aq)	2.035e-224	1.223e-219	1.0000	-223.6913
1-Propene(aq)	4.545e-227	1.912e-222	1.0000	-226.3424
NH4(CH3COO)2-	1.761e-230	2.397e-225	0.9320	-229.7848
2-Hydroxybutanoate	5.375e-231	5.540e-226	0.9320	-230.3002
Asparagine(aq)	1.569e-232	2.072e-227	1.0000	-231.8045
2-Hydroxybutanone	1.074e-234	1.118e-229	1.0000	-233.9690
Cu(Gly)2(aq)	4.542e-235	9.610e-230	1.0000	-234.3428
Propane(aq)	4.325e-239	1.907e-234	1.0000	-238.3640
Diglycine(aq)	1.121e-241	1.481e-236	1.0000	-240.9503
Mg(Gly)2(aq)	8.881e-242	1.531e-236	1.0000	-241.0515
Butanoate	1.593e-242	1.387e-237	0.9320	-241.8285
Mn(Gly)2(aq)	3.950e-243	8.017e-238	1.0000	-242.4035
Diethylperazine	3.188e-244	3.636e-239	1.0000	-243.4965
Butanonecarboxylic acid(aq)	2.928e-245	2.579e-240	1.0000	-244.5334
Na(But)(aq)	2.338e-245	2.573e-240	1.0000	-244.6311
Ca(But)+	2.337e-245	2.971e-240	0.9320	-244.6620
Mg(But)+	1.164e-245	1.296e-240	0.9320	-244.9646
K(But)(aq)	8.208e-246	1.036e-240	1.0000	-245.0858
Ca(Gly)2(aq)	2.861e-246	5.383e-241	1.0000	-245.5435
Mn(But)+	2.038e-246	2.894e-241	0.9320	-245.7213
Cu(But)+	3.607e-247	5.432e-242	0.9320	-246.4734
Glytarate	7.032e-249	9.145e-244	0.7535	-248.2758
Ba(But)+	6.873e-249	1.542e-243	0.9320	-248.1935
Ba(Gly)2(aq)	4.661e-249	1.330e-243	1.0000	-248.3315
H-Glytarate	4.745e-251	6.219e-246	0.9320	-250.3543
Threonine(aq)	3.944e-253	4.696e-248	1.0000	-252.4041
Glytaric acid(aq)	3.208e-254	4.237e-249	1.0000	-253.4938
Ethyl acetate(aq)	1.356e-254	1.194e-249	1.0000	-253.8678
S506--	8.226e-255	2.108e-249	0.7535	-254.2077
1-Propanamine(aq)	3.024e-260	1.787e-255	1.0000	-259.5195
alpha-Aminobutyric acid	2.799e-262	2.886e-257	1.0000	-261.5530
Glytaminecarboxylic acid(aq)	4.211e-271	6.194e-266	1.0000	-270.3756
UO2(SCN)2(aq)	3.856e-276	1.489e-270	1.0000	-275.4139
Butanal(aq)	1.063e-277	7.663e-273	1.0000	-276.9734
S4--	7.332e-284	9.401e-279	0.7535	-283.2577
Cu(CH3COO)3-	5.804e-295	1.396e-289	0.9320	-294.2669
1-Butyne(aq)	2.738e-295	1.480e-290	1.0000	-294.5626
Mn(CH3COO)3-	5.437e-296	1.261e-290	0.9320	-295.2952
1-Butanol(aq)	1.786e-297	1.324e-292	1.0000	-296.7481
1-Butene(aq)	1.292e-299	7.246e-295	1.0000	-298.8888
2-Hydroxypentanoate	4.949e-303	5.795e-298	0.9320	-300.0000
Glytamine(aq)	8.314e-305	1.215e-299	1.0000	-300.0000
U(SCN)2++	6.744e-305	2.388e-299	0.7557	-300.0000

SpecE8_output_GSS_MW-01-15 June.txt				
2-Hydroxypentano	5.747e-307	6.787e-302	1.0000	-300.0000
n-Butane(aq)	2.559e-311	1.487e-306	1.0000	-300.0000
Al anyl gl yci ne(aq	2.159e-312	3.154e-307	1.0000	-300.0000
Pentanoate	9.110e-315	9.210e-310	0.9320	-300.0000
Pentanoi c_aci d(a	1.816e-317	1.854e-312	1.0000	-300.0000
Na(Pent) (aq)	1.453e-317	1.802e-312	1.0000	-300.0000
Ca(Lac)2(aq)	1.452e-317	3.168e-312	1.0000	-300.0000
Ca(Pent)+	8.004e-318	1.130e-312	0.9320	-300.0000
K(Pent) (aq)	5.098e-318	7.147e-313	1.0000	-300.0000
Mg(Lac)2(aq)	4.271e-318	8.645e-313	1.0000	-300.0000
Mg(Pent)+	3.894e-318	4.883e-313	0.9320	-300.0000
Mn(Pent)+	9.183e-319	1.433e-313	0.9320	-300.0000
Cu(Lac)2(aq)	3.109e-319	7.511e-314	1.0000	-300.0000
Mn(Lac)2(aq)	2.583e-319	6.018e-314	1.0000	-300.0000
Cu(Pent)+	2.134e-319	3.512e-314	0.9320	-300.0000
Na(Lac)2-	1.635e-319	3.288e-314	0.9320	-300.0000
K(Lac)2-	4.695e-320	1.020e-314	0.9320	-300.0000
Ba(Pent)+	2.604e-321	6.207e-316	0.9320	-300.0000
Ba(Lac)2(aq)	4.298e-322	1.356e-316	1.0000	-300.0000
Adi pate	9.387e-323	1.353e-317	0.7535	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Methi oni ne(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hexanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexyne(aq)	0.0000	0.0000	1.0000	-300.0000
Azel ate	0.0000	0.0000	0.7535	-300.0000
Leucyl gl yci ne(aq	0.0000	0.0000	1.0000	-300.0000
Leuci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Azel ai c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
2-Heptanone(aq)	0.0000	0.0000	1.0000	-300.0000
K(Prop)2-	0.0000	0.0000	0.9320	-300.0000
2-Butanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexene(aq)	0.0000	0.0000	1.0000	-300.0000
K(Pent)2-	0.0000	0.0000	0.9320	-300.0000
p-Tol ui c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
p-Tol uate	0.0000	0.0000	0.9320	-300.0000
Ca(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
o-Tol uate	0.0000	0.0000	0.9320	-300.0000
n-Propyl benzene(0.0000	0.0000	1.0000	-300.0000
K(But)2-	0.0000	0.0000	0.9320	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
I sol euci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Adi pi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoi c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9320	-300.0000
Hexanal (aq)	0.0000	0.0000	1.0000	-300.0000
m-Tol ui c_aci d(aq	0.0000	0.0000	1.0000	-300.0000
m-Tol uate	0.0000	0.0000	0.9320	-300.0000
Heptanoi c_aci d(a	0.0000	0.0000	1.0000	-300.0000
1-Hexanol (aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9320	-300.0000
Undecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
Heptanal (aq)	0.0000	0.0000	1.0000	-300.0000
Undecanoate	0.0000	0.0000	0.9320	-300.0000
HSb2S4-	0.0000	0.0000	0.9320	-300.0000
UO2(SCN)3-	0.0000	0.0000	0.9320	-300.0000
1-Hexanami ne(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_MW-01-15 June.txt

Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Tyrosine(aq)	0.0000	0.0000	1.0000	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
Toluene(aq)	0.0000	0.0000	1.0000	-300.0000
Suberic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Suberate	0.0000	0.0000	0.7535	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacate	0.0000	0.0000	0.7535	-300.0000
Sb2S4--	0.0000	0.0000	0.7535	-300.0000
H-Suberate	0.0000	0.0000	0.9320	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
H-Sebacate	0.0000	0.0000	0.9320	-300.0000
H-Pimelate	0.0000	0.0000	0.9320	-300.0000
S5--	0.0000	0.0000	0.7535	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacic acid(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
H-Azelate	0.0000	0.0000	0.9320	-300.0000
H-Adipate	0.0000	0.0000	0.9320	-300.0000
Pimelic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Pimelate	0.0000	0.0000	0.7535	-300.0000
Phenylalanine(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol(aq)	0.0000	0.0000	1.0000	-300.0000
Benzic acid(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoic	0.0000	0.0000	1.0000	-300.0000
1-Pentanol(aq)	0.0000	0.0000	1.0000	-300.0000
Pentanal(aq)	0.0000	0.0000	1.0000	-300.0000
Benzoate	0.0000	0.0000	0.9345	-300.0000
Octanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoate	0.0000	0.0000	0.9320	-300.0000
Octanoate	0.0000	0.0000	0.9320	-300.0000
Octanal(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoate	0.0000	0.0000	0.9320	-300.0000
Nonanal(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoic	0.0000	0.0000	1.0000	-300.0000
1-Pentanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
Ethylbenzene(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Prop)2-	0.0000	0.0000	0.9320	-300.0000
2-Hydroxynonanoate	0.0000	0.0000	0.9320	-300.0000
Na(Pent)2-	0.0000	0.0000	0.9320	-300.0000
Dodecanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoate	0.0000	0.0000	0.9320	-300.0000
Ba(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoic	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Na(But)2-	0.0000	0.0000	0.9320	-300.0000
Decanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9320	-300.0000
Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoate	0.0000	0.0000	0.9320	-300.0000
Decanal(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptanoic	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_MW-01-15 June.txt

2-Hydroxyheptano	0.0000	0.0000	0.9320	-300.0000
Valine(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecano	0.0000	0.0000	1.0000	-300.0000
1-Octanol(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(a	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecano	0.0000	0.0000	0.9320	-300.0000
n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexyl benzene(a	0.0000	0.0000	1.0000	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octyl benzene(a	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Pentyl benzene(0.0000	0.0000	1.0000	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
o-Phthalate	0.0000	0.0000	0.7535	-300.0000
1-Octanamine(aq)	0.0000	0.0000	1.0000	-300.0000
o-Toluidine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butanamine(aq)	0.0000	0.0000	1.0000	-300.0000
(NH4)2Sb2S4(aq)	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states
log Q/K

log Q/K

Fluorapatite	5.5120s/sat	Tremolite	-16.8320
Witherite	1.6380s/sat	Natrosilite	-16.9484
Quartz	0.7738s/sat	Ningyolite	-17.5644
Baryte	0.6143s/sat	MnCl2·2H2O	-17.6389
Tridymite	0.5758s/sat	MgOHCl	-17.7819
Chalcedony	0.4822s/sat	Na2SiO3	-17.8039
Cristobalite(alp	0.1738s/sat	Ba2Si3O8	-18.0405
Ice	-0.0609	CuCl2	-19.2890
Coesite	-0.0935	MnCl2·H2O	-19.3889
Cristobalite(bet	-0.3322	UOF0H: .5H2O	-19.5740
Tenorite	-0.4876	MgCl2·4H2O	-19.7238
Rhodochrosite	-0.6399	Hillebrandite	-19.9457
SiO2(am)	-0.6928	UO2S04	-19.9754
Calcite	-0.9079	UOF0H	-20.1087
Aragonite	-1.0535	Hydromagnesite	-20.3401
Dolomite-ord	-1.1900	Dicalciumsilica	-20.4838
Dolomite	-1.1901	UO2(NO3)2·6H2O	-21.5323
Malachite	-1.2442	Na2UO4(alpha)	-21.5660
Monohydrocalcite	-1.6978	Larnite	-21.9019
Magnesite	-2.0461	Akermanite	-22.0572
Diopside	-2.0742	UOF2·H2O	-22.1338
Gypsum	-2.4023	UO2Cl	-22.2361
Anhydrite	-2.7884	UOF2	-22.8298
Fluorite	-2.8289	Scacchite	-22.8930
Dolomite-di	-2.8962	UO2(NO3)2·3H2O	-23.3073
Bassanite	-3.4406	Lime	-23.5226
Sb2O3	-3.4598	Hydrophilite	-24.1812
CaSO4·0.5H2O(bet	-3.6363	Anthophyllite	-24.2980
UO3·2H2O	-3.9508	UO2(NO3)2·2H2O	-24.7401
Schoepite	-3.9528	U(HP04)2·4H2O	-25.3725
Brochantite	-4.0416	UO2(P03)2	-25.6159
UO2(OH)2(beta)	-4.1413	MgCl2·2H2O	-25.7602
Schoepite-dehy(.	-4.1971	UP2O7	-25.9364
UO3·.9H2O(alpha)	-4.2174	(UO2)3(P04)2	-25.9630
Talc	-4.2460	K2UO4	-26.1869
Sellaite	-4.2528	(UO2)2P2O7	-26.2530
Schoepite-dehy(.	-4.3051	UO2Cl2·3H2O	-27.6399

SpecE8_output_GSS_MW-01-15 June.txt

Schoepi te-dehy(1	-4. 3271	Ranki ni te	-28. 3833
Rhodoni te	-4. 4644	UO2(NO3)2: H2O	-28. 7154
CaUO4	-4. 5077	MgCl 2: H2O	-29. 4573
Cupri te	-4. 6456	UO2Cl 2: H2O	-30. 7291
Enstati te	-4. 9128	Tobermori te-14A	-30. 8774
UO2CO3	-4. 9201	UF4: 2. 5H2O	-31. 3523
Rutherfordi ne	-4. 9246	UO2(NO3)2	-32. 5124
Whi tlocki te	-5. 1600	UO2SO3	-32. 5898
Nesquehoni te	-5. 2666	Tobermori te-11A	-33. 3366
Nahcol i te	-5. 4040	Ba2Si O4	-33. 8571
Hydroxyl apati te	-5. 4252	KMgCl 3: 2H2O	-34. 0260
Schoepi te-dehy(.	-5. 5318	UO2Cl 2	-34. 9547
Ni ter	-5. 6020	Foshagi te	-35. 2174
Schoepi te-dehy(.	-6. 0973	BaSi F6	-35. 4029
Okeni te	-6. 3835	Merwi ni te	-35. 5080
Bruci te	-6. 6056	Chl oromagnesi te	-35. 7154
UO2HPO4: 4H2O	-6. 6237	UF4	-36. 0769
Wol l astoni te	-6. 8197	Afwi l l i te	-36. 7229
Mn(OH)2(am)	-6. 8297	Tobermori te-9A	-37. 3548
Bi xbyi te	-6. 8546	U(SO4)2: 8H2O	-40. 6987
Atacami te	-6. 9355	KMgCl 3	-42. 0283
Pseudowol l astoni	-7. 0980	U(SO4)2: 4H2O	-42. 1566
Pyrol usi te	-7. 1140	U(SO4)2	-42. 5218
Mi rabi l i te	-7. 2149	BaO	-42. 6063
UO3(gamma)	-7. 2176	C	-42. 8205
Di opsi de	-7. 2734	Ba2U2O7	-44. 1981
Azuri te	-7. 3849	UOCl 2	-46. 0105
Nantoki te	-7. 6155	U(CO3)2	-46. 2182
Chal canthi te	-7. 8064	UCI F3	-47. 6463
Chrysoti l e	-7. 8339	U5O12Cl	-47. 9193
Syl vi te	-7. 8510	Hatruri te	-47. 9706
UO3(beta)	-7. 8679	Xonotl i te	-50. 7442
Cu	-7. 8882	UOF4	-54. 2945
Arcani te	-7. 8980	Na3UO4	-57. 3586
UO3(al pha)	-8. 2342	UF3	-58. 0532
Hal i te	-8. 4026	(UO2)2Cl 3	-58. 4343
UO2FOH: 2H2O	-8. 4315	Na	-58. 5898
Sanborni te	-8. 4429	UOCl 3	-59. 3634
Hunti te	-8. 7749	UF5(beta)	-60. 0586
UO2FOH: H2O	-8. 9248	Na4Si O4	-60. 2508
Thenardi te	-8. 9945	UF5(al pha)	-60. 4971
UO2. 25	-9. 0532	Mn	-61. 3455
UO2. 25(beta)	-9. 1253	UCI 2F2	-61. 5295
Hausmanni te	-9. 2259	Covel l i te	-62. 3370
Ni trobari te	-9. 2503	K	-62. 7230
Tephroi te	-9. 2980	Na2O	-62. 8778
Arti ni te	-9. 3921	UOCl	-63. 8666
UO2FOH	-9. 4892	Chal coci te	-63. 9896
BaUO4	-9. 6544	S	-64. 3852
Manganosi te	-9. 7428	U2O3F6	-70. 7747
Sepi ol i te	-9. 8481	Ba3UO6	-74. 0109
Natron	-10. 1348	UCI 3F	-75. 2242
Sal eei te	-10. 4461	U(SO3)2	-79. 1458
Na2CO3: 7H2O	-10. 8149	K2O	-81. 2223
Mg1. 25SO4(OH)0. 5	-11. 0515	Al abandi te	-84. 6651
UO2SO4: H2O	-11. 4066	UCI 4	-86. 7629
(UO2)3(PO4)2: 4H2	-11. 6659	Na6Si 2O7	-87. 5980
Gyrol i te	-11. 7807	UCI 3	-89. 6568
UO2. 6667	-11. 8272	U2F9	-92. 7604
Forsteri te	-11. 9031	UF6	-93. 2455
Urani ni te	-11. 9462	U3O5F8	-94. 5936
Uranoci rci te	-11. 9937	Mg	-102. 7848
Coffi ni te	-12. 0279	BaS	-103. 2570

SpecE8_output_GSS_MW-01-15 June.txt

Periclase	-12.1502	UN1.73(alpha)	-106.4319
MnSO4	-12.1956	U2O2Cl5	-106.7122
UP05	-12.3439	UCI5	-108.9382
Thermonatrite	-12.4739	UN1.59(alpha)	-109.3512
Torbernite	-12.5522	Antigorite	-110.2060
Na2U2O7	-12.5604	P	-118.5649
Mg1.5SO4(OH)	-12.6324	Ca	-120.6693
Portlandite	-12.6474	UN	-123.3331
UO2Cl(OH)·2H2O	-12.6918	Ba	-124.8449
BaU2O7	-12.6948	Si	-127.1337
Na2CO3	-12.8500	UCI6	-129.8696
MgUO4	-12.9075	U4F17	-163.9804
Monticellite	-13.0498	U	-172.9200
MgSO4	-13.3717	US	-176.8263
BaCl2·2H2O	-14.1988	UC	-197.1193
Chalcocyanite	-14.3398	US1.9	-199.2629
Autunite-H	-14.6971	UH3(beta)	-200.9508
BaCl2·H2O	-14.9889	US2	-203.8512
NaUO3	-15.4952	Stibnite	-239.5054
UO2SO4·3.5H2O	-15.8471	UC1.94(alpha)	-239.5509
UO2F2·3H2O	-15.9310	UP	-241.3346
UO2SO4·2.5H2O	-15.9408	US3	-264.8275
UO2SO4·3H2O	-16.0141	UP2	-354.4305
MnCl2·4H2O	-16.0978	U2S3	-373.2549
UO2F2	-16.3301	o-Phthalic acid	-381.7928
Ba(OH)2·8H2O	-16.5668	U2C3	-438.7104
BaCl2	-16.6059	U3S5	-572.1616
UO2.3333(beta)	-16.6822	U3P4	-837.1014

Gases	fugacity	log fug.
N2(g)	0.1393	-0.856
H2O(g)	0.006731	-2.172
CO2(g)	0.002191	-2.659
NO2(g)	3.139e-011	-10.503
HF(g)	5.503e-015	-14.259
HCl(g)	1.723e-019	-18.764
H2(g)	5.370e-029	-28.270
NO(g)	2.062e-034	-33.686
O2(g)	7.522e-035	-34.124
CO(g)	4.063e-035	-34.391
SiF4(g)	6.486e-039	-38.188
NH3(g)	4.511e-040	-39.346
SO2(g)	1.004e-042	-41.998
Cl2(g)	8.270e-046	-45.083
UO2F2(g)	6.241e-063	-62.205
Cu(g)	3.174e-065	-64.498
Na(g)	3.178e-074	-73.498
UO2Cl2(g)	4.247e-075	-74.372
K(g)	3.174e-075	-74.498
UO3(g)	8.530e-077	-76.069
UOF4(g)	7.294e-077	-76.137
UF5(g)	1.257e-082	-81.901
UF4(g)	9.766e-085	-84.010
H2S(g)	3.129e-087	-86.505
CH4(g)	3.228e-090	-89.491
UF6(g)	1.891e-095	-94.723
F2(g)	1.097e-104	-103.960
UCI4(g)	1.456e-114	-113.837
UO2(g)	3.506e-117	-116.455
Mg(g)	3.578e-125	-124.446
UCI5(g)	5.956e-129	-128.225
UF3(g)	2.473e-131	-130.607

SpecE8_output_GSS_MW-01-15 June.txt

U2F10(g)	8.017e-137	-136.096
UCl 6(g)	1.559e-137	-136.807
UCl 3(g)	2.520e-140	-139.599
S2(g)	4.386e-145	-144.358
Ca(g)	5.557e-149	-148.255
C2H4(g)	1.546e-155	-154.811
C(g)	1.297e-170	-169.887
UF2(g)	1.853e-174	-173.732
UCl 2(g)	1.394e-184	-183.856
UO(g)	1.939e-191	-190.712
Si (g)	7.475e-205	-204.126
UF(g)	1.392e-210	-209.856
U2Cl 8(g)	1.649e-216	-215.783
UCl (g)	3.394e-226	-225.469
U2Cl 10(g)	3.371e-228	-227.472
U(g)	3.001e-266	-265.523

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
<hr/>						
Ba++	7.59e-007	7.59e-007	0.104			
Ca++	0.000559	0.000559	22.4			
Cl -	0.000110	0.000110	3.89			
Cu++	1.38e-007	1.38e-007	0.00876			
F-	1.54e-005	1.54e-005	0.292			
H+	0.000146	0.000146	0.147			
H2O	55.5	55.5	1.00e+006			
HC03-	0.00166	0.00166	101.			
HP04--	7.86e-007	7.86e-007	0.0754			
K+	0.000573	0.000573	22.4			
Mg++	0.000200	0.000200	4.87			
Mn++	8.28e-006	8.28e-006	0.455			
NH3(aq)	0.000268	0.000268	4.56			
N03-	0.00101	0.00101	62.5			
Na+	0.00136	0.00136	31.2			
O2(aq)	0.000201	0.000201	6.43			
S2--	6.07e-007	6.07e-007	0.0389			
S04--	0.000416	0.000416	39.9			
Sb(OH)3(aq)	2.32e-007	2.32e-007	0.0401			
Si O2(aq)	0.000185	0.000185	11.1			
U02++	7.36e-009	7.36e-009	0.00199			

Elemental composition	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg
<hr/>					
Antimony	2.320e-007	2.320e-007	0.02823		
Barium	7.588e-007	7.588e-007	0.1042		
Calcium	0.0005589	0.0005589	22.39		
Carbon	0.001660	0.001660	19.93		
Chlorine	0.0001099	0.0001099	3.894		
Copper	1.379e-007	1.379e-007	0.008762		
Fluorine	1.538e-005	1.538e-005	0.2921		
Hydrogen	111.0	111.0	1.119e+005		
Magnesium	0.0002003	0.0002003	4.868		
Manganese	8.278e-006	8.278e-006	0.4546		
Nitrogen	0.001276	0.001276	17.87		
Oxygen	55.52	55.52	8.880e+005		
Phosphorus	7.860e-007	7.860e-007	0.02434		
Potassium	0.0005729	0.0005729	22.39		
Silicon	0.0001848	0.0001848	5.188		
Sodium	0.001356	0.001356	31.15		
Sulfur	0.0004168	0.0004168	13.36		
Uranium	7.364e-009	7.364e-009	0.001752		

SpecE8_output_GSS_MW-01-15 June. txt

SpecE8_output_GSS_MW-01-15 May. txt

Temperature = 10.9 C Pressure = 1.013 bars
 pH = 7.300 log fO2 = -29.990
 Eh = 0.3980 volts pe = 7.0633
 Ionic strength = 0.004293
 Charge imbalance = -0.000046 eq/kg (-1.379% error)
 Activity of water = 0.999996
 Solvent mass = 1.000000 kg
 Solution mass = 1.000303 kg
 Solution density = 1.023 g/cm3
 Chlorinity = 0.000110 molal
 Dissolved solids = 303 mg/kg sol'n
 Hardness = 69.83 mg/kg sol'n as CaCO3
 carbonate = 69.83 mg/kg sol'n as CaCO3
 non-carbonate = 0.00 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 72.13 mg/kg sol'n as CaCO3
 Water type = Na-HCO3

Nernst redox couples	Eh (volts)	pe
$e^- + .25*O_2(aq) + H^+ = .5*H_2O$	0.3980	7.0633
$8*e^- + 9*H^+ + NO_3^- = 3*H_2O + NH_3(aq)$	0.5981	10.6148
$14*e^- + 16*H^+ + 2*S^{4-} = 8*H_2O + S^{2-}$	-0.2362	-4.1924

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HC03-	0.001428	87.13	0.9332	-2.8751
Na+	0.001358	31.22	0.9332	-2.8970
NO3-	0.001173	72.72	0.9319	-2.9612
K+	0.0005495	21.48	0.9319	-3.2907
Ca++	0.0005189	20.79	0.7657	-3.4009
S04--	0.0002987	28.69	0.7576	-3.6453
CO2(aq)	0.0001932	8.499	1.0000	-3.7141
SiO2(aq)	0.0001788	10.74	1.0000	-3.7477
Mg++	0.0001541	3.744	0.7732	-3.9239
N2(aq)	0.0001416	3.966	1.0000	-3.8489
Cl-	0.0001103	3.910	0.9319	-3.9880
CaSO4(aq)	1.087e-005	1.479	1.0000	-4.9638
F-	1.025e-005	0.1947	0.9326	-5.0196
Mn++	7.834e-006	0.4302	0.7657	-5.2220
CaHC03+	6.561e-006	0.6631	0.9332	-5.2131
MgS04(aq)	4.851e-006	0.5837	1.0000	-5.3142
NaHC03(aq)	3.187e-006	0.2677	1.0000	-5.4966
MgHC03+	1.908e-006	0.1628	0.9332	-5.7494
HP04--	1.322e-006	0.1268	0.7576	-5.9995
CO3--	1.155e-006	0.06930	0.7597	-6.0567
AlO2-	1.102e-006	0.06499	0.9332	-5.9877
Cu++	1.047e-006	0.06649	0.7657	-6.0962
Fe++	1.040e-006	0.05804	0.7657	-6.0991
H2P04-	9.837e-007	0.09538	0.9332	-6.0371
KS04-	9.332e-007	0.1261	0.9332	-6.0600
S2--	7.626e-007	0.04889	0.7576	-6.2383
Ba++	7.192e-007	0.09874	0.7617	-6.2614
CaCO3(aq)	5.617e-007	0.05620	1.0000	-6.2505
HAIO2(aq)	3.159e-007	0.01894	1.0000	-6.5005
HSiO3-	2.692e-007	0.02075	0.9332	-6.5999
MnS04(aq)	2.662e-007	0.04019	1.0000	-6.5748
Sb(OH)3(aq)	2.571e-007	0.04440	1.0000	-6.5899
Zn++	1.457e-007	0.009522	0.7657	-6.9526
MgCO3(aq)	8.697e-008	0.007330	1.0000	-7.0606

	SpecE8_output_GSS_MW-01-15 May. txt			
OH-	6. 694e-008	0. 001138	0. 9326	-7. 2046
H+	5. 337e-008	5. 378e-005	0. 9391	-7. 3000
Al (OH) 2+	2. 960e-008	0. 001805	0. 9332	-7. 5587
MgF+	2. 806e-008	0. 001215	0. 9332	-7. 5820
NaCl (aq)	2. 020e-008	0. 001180	1. 0000	-7. 6946
NaHSi O3(aq)	1. 952e-008	0. 001953	1. 0000	-7. 7095
CaF+	1. 849e-008	0. 001092	0. 9332	-7. 7630
FeCO3+	1. 116e-008	0. 001293	0. 9332	-7. 9824
MgCl +	1. 045e-008	0. 0006244	0. 9332	-8. 0108
CaCl +	8. 897e-009	0. 0006718	0. 9332	-8. 0808
Cd++	7. 830e-009	0. 0008799	0. 7617	-8. 2245
UO2(CO3) 2--	6. 046e-009	0. 002358	0. 7576	-8. 3391
NaCO3-	5. 792e-009	0. 0004806	0. 9332	-8. 2672
Pb++	5. 182e-009	0. 001073	0. 7597	-8. 4049
ZnSO4(aq)	3. 914e-009	0. 0006317	1. 0000	-8. 4074
Al OH++	2. 101e-009	9. 240e-005	0. 7597	-8. 7969
UO2(CO3) 3----	1. 790e-009	0. 0008052	0. 3291	-9. 2299
KCl (aq)	1. 277e-009	9. 521e-005	1. 0000	-8. 8936
NaF(aq)	1. 086e-009	4. 557e-005	1. 0000	-8. 9643
MnCl +	9. 690e-010	8. 756e-005	0. 9332	-9. 0437
HSO4-	8. 167e-010	7. 925e-005	0. 9332	-9. 1180
HF(aq)	5. 519e-010	1. 104e-005	1. 0000	-9. 2581
UO2CO3(aq)	3. 774e-010	0. 0001245	1. 0000	-9. 4232
NaAl O2(aq)	2. 031e-010	1. 664e-005	1. 0000	-9. 6923
BaCO3(aq)	1. 579e-010	3. 115e-005	1. 0000	-9. 8017
FeCl +	6. 005e-011	5. 481e-006	0. 9332	-10. 2515
Al +++	3. 406e-011	9. 188e-007	0. 5662	-10. 7148
Cu+	2. 423e-011	1. 539e-006	0. 9332	-10. 6456
BaCl +	1. 595e-011	2. 755e-006	0. 9332	-10. 8272
NaOH(aq)	1. 387e-011	5. 546e-007	1. 0000	-10. 8579
PO4---	1. 269e-011	1. 205e-006	0. 5352	-11. 1681
PbCl +	1. 129e-011	2. 738e-006	0. 9332	-10. 9774
ZnCl +	8. 544e-012	8. 613e-007	0. 9332	-11. 0984
H3PO4(aq)	5. 932e-012	5. 811e-007	1. 0000	-11. 2268
UO2OH+	5. 185e-012	1. 488e-006	0. 9332	-11. 3153
BaF+	3. 208e-012	5. 013e-007	0. 9332	-11. 5238
Sb(OH) 2F(aq)	2. 576e-012	4. 501e-007	1. 0000	-11. 5891
Sb(OH) 4-	2. 205e-012	4. 184e-007	0. 9332	-11. 6865
HNO3(aq)	2. 006e-012	1. 264e-007	1. 0000	-11. 6976
CdSO4(aq)	1. 364e-012	2. 842e-007	1. 0000	-11. 8653
CaCl 2(aq)	1. 172e-012	1. 301e-007	1. 0000	-11. 9310
HCl (aq)	1. 093e-012	3. 983e-008	1. 0000	-11. 9614
Fe+++	9. 710e-013	5. 421e-008	0. 5662	-12. 2599
Sb(OH) 2+	7. 391e-013	1. 151e-007	0. 9332	-12. 1613
MgP2O7--	2. 268e-013	4. 495e-008	0. 7576	-12. 7649
CdCl 2(aq)	1. 261e-013	2. 310e-008	1. 0000	-12. 8994
UO2++	1. 248e-013	3. 368e-008	0. 7597	-13. 0233
UO2F+	1. 079e-013	3. 119e-008	0. 9332	-12. 9968
FeF++	7. 570e-014	5. 664e-009	0. 7597	-13. 2403
KHSO4(aq)	2. 273e-014	3. 094e-009	1. 0000	-13. 6434
HP2O7---	1. 811e-014	3. 168e-009	0. 5352	-14. 0134
UO2SO4(aq)	1. 767e-014	6. 468e-009	1. 0000	-13. 7527
FeF2+	8. 119e-015	7. 617e-010	0. 9332	-14. 1205
FeSO4+	7. 202e-015	1. 094e-009	0. 9332	-14. 1726
PbCl 2(aq)	3. 787e-015	1. 053e-009	1. 0000	-14. 4217
UO2F2(aq)	3. 027e-015	9. 320e-010	1. 0000	-14. 5190
H2P2O7--	2. 862e-015	5. 034e-010	0. 7576	-14. 6639
ZnCl 2(aq)	1. 234e-015	1. 682e-010	1. 0000	-14. 9086
HF2-	1. 167e-015	4. 549e-011	0. 9332	-14. 9630
P2O7----	3. 292e-016	5. 725e-011	0. 3291	-15. 9652
UO2(SO4) 2--	3. 275e-017	1. 513e-011	0. 7576	-16. 6054
FeCl 2(aq)	2. 835e-017	3. 592e-012	1. 0000	-16. 5474
KP2O7---	1. 950e-017	4. 152e-012	0. 5352	-16. 9815

	SpecE8_output_GSS_MW-01-15 May. txt			
UO2Cl +	1. 308e-017	3. 993e-012	0. 9332	-16. 9135
FeCl ++	5. 857e-018	5. 346e-013	0. 7597	-17. 3517
UO2F3-	5. 323e-018	1. 740e-012	0. 9332	-17. 3038
(UO2)2(OH)2++	4. 944e-018	2. 838e-012	0. 7597	-17. 4253
CdCl 3-	2. 675e-018	5. 850e-013	0. 9332	-17. 6027
PbCl 3-	2. 056e-019	6. 446e-014	0. 9332	-18. 7169
UO2+	2. 041e-019	5. 510e-014	0. 9332	-18. 7202
ZnCl 3-	1. 074e-019	1. 843e-014	0. 9332	-18. 9992
(UO2)3(OH)5+	9. 726e-020	8. 703e-014	0. 9332	-19. 0421
H3P2O7-	2. 613e-020	4. 623e-015	0. 9332	-19. 6128
(UO2)3(CO3)6(6-)	2. 024e-020	2. 367e-014	0. 0820	-20. 7802
UO2F4--	3. 845e-022	1. 330e-016	0. 7576	-21. 5357
ZnCl 4--	1. 253e-022	2. 596e-017	0. 7576	-22. 0226
UO2Cl 2(aq)	5. 649e-023	1. 925e-017	1. 0000	-22. 2480
U(OH)4(aq)	3. 983e-023	1. 219e-017	1. 0000	-22. 3998
PbCl 4--	2. 354e-023	8. 212e-018	0. 7576	-22. 7488
FeCl 4--	1. 196e-024	2. 363e-019	0. 7576	-24. 0429
Mn+++	1. 119e-024	6. 144e-020	0. 5418	-24. 2174
H4P2O7(aq)	3. 221e-026	5. 730e-021	1. 0000	-25. 4920
NO2-	9. 318e-032	4. 285e-027	0. 9319	-31. 0613
Formate	2. 947e-032	1. 326e-027	0. 9326	-31. 5609
H2(aq)	7. 513e-033	1. 514e-029	1. 0000	-32. 1242
O2(aq)	1. 724e-033	5. 515e-029	1. 0000	-32. 7635
Ca(For)+	3. 493e-034	2. 972e-029	0. 9332	-33. 4867
Mg(For)+	1. 261e-034	8. 738e-030	0. 9332	-33. 9293
Na(For)(aq)	4. 129e-035	2. 807e-030	1. 0000	-34. 3842
K(For)(aq)	1. 453e-035	1. 222e-030	1. 0000	-34. 8378
Mn(For)+	1. 058e-035	1. 057e-030	0. 9332	-35. 0056
HN02(aq)	1. 037e-035	4. 873e-031	1. 0000	-34. 9843
Formic acid(aq)	8. 100e-036	3. 727e-031	1. 0000	-35. 0915
Cu(For)+	2. 916e-036	3. 165e-031	0. 9332	-35. 5652
Fe(For)+	2. 062e-036	2. 079e-031	0. 9332	-35. 7157
Si F6--	1. 142e-036	1. 622e-031	0. 7576	-36. 0630
Cl O-	9. 214e-037	4. 739e-032	0. 9332	-36. 0656
SO3--	7. 715e-037	6. 175e-032	0. 7597	-36. 2321
NH4+	6. 829e-037	1. 232e-032	0. 9313	-36. 1965
HSO3-	4. 482e-037	3. 633e-032	0. 9332	-36. 3785
Ba(For)+	3. 674e-037	6. 697e-032	0. 9332	-36. 4649
Zn(For)+	2. 516e-037	2. 777e-032	0. 9332	-36. 6292
Cd(For)+	1. 474e-038	2. 320e-033	0. 9332	-37. 8616
CO(aq)	1. 339e-038	3. 750e-034	1. 0000	-37. 8732
Pb(For)+	8. 908e-039	2. 246e-033	0. 9332	-38. 0802
H02-	2. 909e-039	9. 599e-035	0. 9332	-38. 5662
NH3(aq)	2. 563e-039	4. 363e-035	1. 0000	-38. 5913
Oxal ate	9. 441e-040	8. 308e-035	0. 7576	-39. 1455
UOH+++	1. 721e-040	4. 388e-035	0. 5418	-40. 0304
UF3+	1. 300e-040	3. 833e-035	0. 9332	-39. 9162
UF2++	8. 261e-041	2. 280e-035	0. 7597	-40. 2023
UF4(aq)	1. 211e-041	3. 803e-036	1. 0000	-40. 9168
U(CO3)5(6-)	4. 432e-042	2. 384e-036	0. 0820	-42. 4396
UF+++	1. 517e-042	3. 899e-037	0. 5418	-42. 0851
SO2(aq)	1. 143e-042	7. 319e-038	1. 0000	-41. 9420
H-Oxal ate	6. 356e-043	5. 657e-038	0. 9332	-42. 2268
MnO4--	3. 549e-043	4. 220e-038	0. 7576	-42. 5704
USO4++	3. 647e-044	1. 218e-038	0. 7597	-43. 5574
U(SO4)2(aq)	2. 763e-044	1. 188e-038	1. 0000	-43. 5587
MnO4-	2. 222e-046	2. 642e-041	0. 9326	-45. 6835
HSO5-	1. 642e-046	1. 856e-041	0. 9332	-45. 8146
U++++	1. 268e-046	3. 018e-041	0. 3398	-46. 3656
UCl +++	6. 269e-049	1. 714e-043	0. 5418	-48. 4690
Oxalic acid(aq)	5. 399e-049	4. 859e-044	1. 0000	-48. 2677
S2O6--	1. 299e-060	2. 079e-055	0. 7576	-60. 0070
Cl O2-	1. 444e-061	9. 736e-057	0. 9332	-60. 8705

	SpecE8_output_GSS_MW-01-15 May. txt			
S208--	4. 783e-063	9. 187e-058	0. 7576	-62. 4408
U+++	2. 615e-064	6. 223e-059	0. 5418	-63. 8487
Ca(For) 2(aq)	7. 891e-065	1. 026e-059	1. 0000	-64. 1029
Mg(For) 2(aq)	3. 517e-065	4. 021e-060	1. 0000	-64. 4538
Mn(For) 2(aq)	4. 276e-066	6. 197e-061	1. 0000	-65. 3690
Cu(For) 2(aq)	2. 175e-066	3. 340e-061	1. 0000	-65. 6625
Fe(For) 2(aq)	1. 145e-066	1. 671e-061	1. 0000	-65. 9410
Na(For) 2-	8. 677e-067	9. 804e-062	0. 9332	-66. 0917
K(For) 2-	2. 569e-067	3. 316e-062	0. 9332	-66. 6203
Zn(For) 2(aq)	1. 283e-067	1. 994e-062	1. 0000	-66. 8918
Ba(For) 2(aq)	6. 307e-068	1. 434e-062	1. 0000	-67. 2002
Cd(For) 2(aq)	5. 378e-068	1. 088e-062	1. 0000	-67. 2694
Pb(For) 2(aq)	4. 142e-069	1. 231e-063	1. 0000	-68. 3828
Formal dehyde(aq)	2. 349e-069	7. 050e-065	1. 0000	-68. 6292
Cl O3-	7. 899e-072	6. 590e-067	0. 9326	-71. 1327
N3-	1. 603e-072	6. 735e-068	0. 9332	-71. 8250
HN3(aq)	5. 267e-075	2. 266e-070	1. 0000	-74. 2784
Cu(NH3) 2++	4. 775e-076	4. 659e-071	0. 7597	-75. 4404
S205--	3. 691e-078	5. 319e-073	0. 7576	-77. 5534
Cd(NH3) 2++	6. 643e-081	9. 728e-076	0. 7597	-80. 2970
Urea(aq)	4. 064e-081	2. 440e-076	1. 0000	-80. 3910
HCN(aq)	1. 577e-082	4. 261e-078	1. 0000	-81. 8021
UO2Cl O3+	2. 544e-084	8. 988e-079	0. 9332	-83. 6246
CN-	7. 610e-085	1. 979e-080	0. 9319	-84. 1492
Methanol (aq)	7. 320e-086	2. 345e-081	1. 0000	-85. 1355
Cl O4-	1. 685e-086	1. 675e-081	0. 9326	-85. 8037
HS-	5. 389e-092	1. 782e-087	0. 9326	-91. 2988
Gl ycol ate	5. 195e-092	3. 897e-087	0. 9332	-91. 3145
H2S(aq)	4. 105e-092	1. 399e-087	1. 0000	-91. 3867
Ca(Gl yc) +	1. 071e-093	1. 232e-088	0. 9332	-93. 0003
Mg(Gl yc) +	1. 491e-094	1. 480e-089	0. 9332	-93. 8567
Na(Gl yc) (aq)	7. 436e-095	7. 288e-090	1. 0000	-94. 1287
Cu(Gl yc) +	4. 746e-095	6. 575e-090	0. 9332	-94. 3537
K(Gl yc) (aq)	2. 616e-095	2. 985e-090	1. 0000	-94. 5823
Gl ycol i c_aci d(aq)	1. 733e-095	1. 318e-090	1. 0000	-94. 7611
Mn(Gl yc) +	1. 432e-095	1. 861e-090	0. 9332	-94. 8740
Fe(Gl yc) +	8. 649e-096	1. 132e-090	0. 9332	-95. 0931
Zn(Gl yc) +	1. 980e-096	2. 780e-091	0. 9332	-95. 7333
Methane(aq)	5. 027e-097	8. 063e-093	1. 0000	-96. 2987
Ba(Gl yc) +	2. 995e-097	6. 358e-092	0. 9332	-96. 5536
S203--	6. 041e-098	6. 772e-093	0. 7576	-97. 3394
S--	5. 701e-098	1. 827e-093	0. 7617	-97. 3623
Pb(Gl yc) +	4. 328e-098	1. 221e-092	0. 9332	-97. 3938
Cd(Gl yc) +	2. 855e-098	5. 350e-093	0. 9332	-97. 5744
S204--	2. 011e-098	2. 576e-093	0. 7617	-97. 8148
Acetate	1. 519e-102	8. 964e-098	0. 9339	-101. 8483
CaCH3COO+	5. 420e-105	5. 371e-100	0. 9332	-104. 2960
MgCH3COO+	4. 429e-105	3. 690e-100	0. 9332	-104. 3837
Aceti c_aci d(aq)	4. 130e-105	2. 480e-100	1. 0000	-104. 3840
NaCH3COO(aq)	1. 502e-105	1. 232e-100	1. 0000	-104. 8233
KCH3COO(aq)	3. 755e-106	3. 684e-101	1. 0000	-105. 4254
CuCH3COO+	2. 824e-106	3. 461e-101	0. 9332	-105. 5792
MnCH3COO+	1. 733e-106	1. 975e-101	0. 9332	-105. 7912
FeCH3COO+	2. 739e-107	3. 145e-102	0. 9332	-106. 5925
BaCH3COO+	7. 512e-108	1. 475e-102	0. 9332	-107. 1543
ZnCH3COO+	6. 127e-108	7. 622e-103	0. 9332	-107. 2428
PbCH3COO+	1. 600e-108	4. 260e-103	0. 9332	-107. 8258
CdCH3COO+	9. 597e-109	1. 645e-103	0. 9332	-108. 0479
Al CH3COO++	6. 755e-111	5. 809e-106	0. 7597	-110. 2897
Mal onate	2. 029e-111	2. 070e-106	0. 7576	-110. 8133
Cu(NH3) 3++	1. 246e-111	1. 428e-106	0. 7597	-111. 0239
CuCH3COO(aq)	7. 907e-113	9. 691e-108	1. 0000	-112. 1020
H-Mal onate	3. 872e-113	3. 989e-108	0. 9332	-112. 4420

	SpecE8_output_GSS_MW-01-15 May. txt			
Maloni c_acid(aq)	1. 330e-117	1. 384e-112	1. 0000	-116. 8761
Methanamine(aq)	7. 337e-122	2. 278e-117	1. 0000	-121. 1345
Glucine(aq)	2. 698e-124	2. 025e-119	1. 0000	-123. 5689
Cu(Gl y) +	2. 037e-124	2. 802e-119	0. 9332	-123. 7210
Mg(Gl y) +	1. 605e-127	1. 578e-122	0. 9332	-126. 8246
Zn(Gl y) +	1. 404e-128	1. 957e-123	0. 9332	-127. 8827
Mn(Gl y) +	1. 230e-128	1. 586e-123	0. 9332	-127. 9400
Fe(Gl y) +	7. 443e-129	9. 665e-124	0. 9332	-128. 1583
Ca(Gl y) +	2. 771e-129	3. 162e-124	0. 9332	-128. 5873
S3O6--	1. 147e-129	2. 204e-124	0. 7576	-129. 0610
Pb(Gl y) +	4. 922e-130	1. 384e-124	0. 9332	-129. 3378
Cd(Gl y) +	1. 411e-130	2. 630e-125	0. 9332	-129. 8805
Ba(Gl y) +	4. 611e-132	9. 743e-127	0. 9332	-131. 3662
Acetaldehyde(aq)	1. 877e-137	8. 266e-133	1. 0000	-136. 7266
SCN-	8. 366e-138	4. 858e-133	0. 9326	-137. 1078
NH4CH3COO(aq)	1. 475e-138	1. 137e-133	1. 0000	-137. 8311
Acetamide(aq)	1. 226e-138	7. 238e-134	1. 0000	-137. 9116
UO2SCN+	1. 938e-149	6. 358e-144	0. 9332	-148. 7426
Zn(NH3)4++	4. 313e-153	5. 756e-148	0. 7597	-152. 4846
Cd(NH3)4++	1. 694e-155	3. 056e-150	0. 7597	-154. 8906
Ethanol(aq)	1. 297e-158	5. 973e-154	1. 0000	-157. 8871
Ethyne(aq)	1. 174e-159	3. 056e-155	1. 0000	-158. 9303
Ethylene(aq)	2. 137e-163	5. 992e-159	1. 0000	-162. 6703
Lactate	1. 004e-164	8. 941e-160	0. 9332	-164. 0282
Ca(Lac) +	1. 217e-166	1. 571e-161	0. 9332	-165. 9447
Mg(Lac) +	2. 908e-167	3. 296e-162	0. 9332	-166. 5664
Na(Lac) (aq)	1. 467e-167	1. 643e-162	1. 0000	-166. 8337
K(Lac) (aq)	5. 160e-168	6. 612e-163	1. 0000	-167. 2873
Cu(Lac) +	4. 209e-168	6. 422e-163	0. 9332	-167. 4058
Lactate c_acid(aq)	3. 517e-168	3. 167e-163	1. 0000	-167. 4539
Mn(Lac) +	1. 958e-168	2. 819e-163	0. 9332	-167. 7382
Fe(Lac) +	1. 478e-168	2. 141e-163	0. 9332	-167. 8604
Zn(Lac) +	1. 862e-169	2. 875e-164	0. 9332	-168. 7601
Ba(Lac) +	2. 395e-170	5. 420e-165	0. 9332	-169. 6507
Pb(Lac) +	8. 123e-171	2. 406e-165	0. 9332	-170. 1203
Cd(Lac) +	3. 732e-171	7. 518e-166	0. 9332	-170. 4580
Ethane(aq)	3. 938e-174	1. 184e-169	1. 0000	-173. 4048
Propanoate	1. 877e-177	1. 371e-172	0. 9332	-176. 7565
Propanoate c_acid(aq)	6. 864e-180	5. 083e-175	1. 0000	-179. 1634
Ca(Prop) +	3. 733e-180	4. 223e-175	0. 9332	-179. 4579
Na(Prop) (aq)	2. 704e-180	2. 597e-175	1. 0000	-179. 5680
Mg(Prop) +	1. 483e-180	1. 444e-175	0. 9332	-179. 8589
USCN+++	1. 038e-180	3. 074e-175	0. 5418	-180. 2498
K(Prop) (aq)	9. 514e-181	1. 067e-175	1. 0000	-180. 0216
Cu(Prop) +	3. 416e-181	4. 666e-176	0. 9332	-180. 4965
Mn(Prop) +	2. 510e-181	3. 212e-176	0. 9332	-180. 6303
Fe(Prop) +	7. 936e-182	1. 023e-176	0. 9332	-181. 1304
S4O6--	2. 736e-182	6. 134e-177	0. 7576	-181. 6834
Zn(Prop) +	4. 683e-183	6. 482e-178	0. 9332	-182. 3595
Pb(Prop) +	1. 732e-183	4. 853e-178	0. 9332	-182. 7914
Ba(Prop) +	1. 230e-183	2. 587e-178	0. 9332	-182. 9402
Cd(Prop) +	1. 089e-183	2. 019e-178	0. 9332	-182. 9930
Ca(Gl yc)2(aq)	1. 049e-183	1. 994e-178	1. 0000	-182. 9792
Cu(Gl yc)2(aq)	1. 860e-184	3. 972e-179	1. 0000	-183. 7306
Succinate	1. 026e-184	1. 191e-179	0. 7576	-184. 1093
Mg(Gl yc)2(aq)	8. 891e-185	1. 550e-179	1. 0000	-184. 0510
Fe(Gl yc)2(aq)	2. 723e-185	5. 607e-180	1. 0000	-184. 5649
Zn(Gl yc)2(aq)	9. 723e-186	2. 094e-180	1. 0000	-185. 0122
Mn(Gl yc)2(aq)	5. 718e-186	1. 172e-180	1. 0000	-185. 2428
Na(Gl yc)2-	4. 300e-186	7. 440e-181	0. 9332	-185. 3965
H-Succinate	1. 856e-186	2. 173e-181	0. 9332	-185. 7613
K(Gl yc)2-	1. 306e-186	2. 469e-181	0. 9332	-185. 9142
Serine(aq)	4. 799e-187	5. 042e-182	1. 0000	-186. 3189

SpecE8_output_GSS_MW-01-15 May.txt				
Pb(Gl yc)2(aq)	1. 281e-187	4. 575e-182	1. 0000	-186. 8925
Ba(Gl yc)2(aq)	6. 926e-188	1. 990e-182	1. 0000	-187. 1595
Cd(Gl yc)2(aq)	4. 536e-188	1. 190e-182	1. 0000	-187. 3433
Succi ni c_aci d(aq)	1. 529e-189	1. 805e-184	1. 0000	-188. 8156
Ethanami ne(aq)	1. 458e-196	6. 573e-192	1. 0000	-195. 8361
Al ani ne(aq)	7. 099e-198	6. 323e-193	1. 0000	-197. 1488
Cu(Al a)+	1. 213e-198	1. 839e-193	0. 9332	-197. 9461
Fe(Al a)+	2. 351e-202	3. 383e-197	0. 9332	-201. 6587
Zn(Al a)+	3. 502e-203	5. 374e-198	0. 9332	-202. 4856
Mg(Al a)+	2. 878e-203	3. 234e-198	0. 9332	-202. 5708
Mn(Al a)+	2. 777e-203	3. 970e-198	0. 9332	-202. 5865
Ca(Al a)+	1. 399e-203	1. 793e-198	0. 9332	-202. 8842
Cd(Al a)+	2. 090e-204	4. 189e-199	0. 9332	-203. 7099
Pb(Al a)+	1. 041e-204	3. 072e-199	0. 9332	-204. 0127
Ca(CH3COO)2(aq)	1. 473e-205	2. 329e-200	1. 0000	-204. 8319
Mg(CH3COO)2(aq)	5. 092e-206	7. 249e-201	1. 0000	-205. 2931
Cu(CH3COO)2(aq)	1. 385e-206	2. 514e-201	1. 0000	-205. 8587
Ba(Al a)+	8. 629e-207	1. 944e-201	0. 9332	-206. 0941
Acetone(aq)	2. 105e-207	1. 223e-202	1. 0000	-206. 6766
Mn(CH3COO)2(aq)	2. 087e-207	3. 610e-202	1. 0000	-206. 6805
Na(CH3COO)2-	1. 301e-207	1. 835e-202	0. 9332	-206. 9156
Zn(CH3COO)2(aq)	9. 395e-208	1. 723e-202	1. 0000	-207. 0271
Fe(CH3COO)2(aq)	8. 662e-208	1. 506e-202	1. 0000	-207. 0624
Asparti c_aci d(aq)	3. 995e-208	5. 316e-203	1. 0000	-207. 3984
K(CH3COO)2-	2. 145e-208	3. 370e-203	0. 9332	-207. 6986
Ba(CH3COO)2(aq)	3. 253e-209	8. 305e-204	1. 0000	-208. 4878
Cd(CH3COO)2(aq)	2. 900e-209	6. 683e-204	1. 0000	-208. 5375
Pb(CH3COO)2(aq)	1. 875e-209	6. 097e-204	1. 0000	-208. 7270
Al (CH3COO)2+	1. 080e-210	1. 566e-205	0. 9332	-209. 9966
Propanal (aq)	8. 582e-212	4. 983e-207	1. 0000	-211. 0664
Cu(CH3COO)2-	1. 681e-214	3. 053e-209	0. 9332	-213. 8044
1-Propyne(aq)	2. 437e-230	9. 760e-226	1. 0000	-229. 6132
S3--	6. 341e-231	6. 098e-226	0. 7576	-230. 3184
1-Propanol (aq)	1. 831e-233	1. 100e-228	1. 0000	-232. 7374
1-Propene(aq)	5. 388e-236	2. 266e-231	1. 0000	-235. 2686
2-Hydroxybutanoa	5. 086e-240	5. 242e-235	0. 9332	-239. 3236
NH4(CH3COO)2-	2. 127e-240	2. 895e-235	0. 9332	-239. 7022
Asparagi ne(aq)	1. 414e-241	1. 867e-236	1. 0000	-240. 8496
Cu(Gl y)2(aq)	2. 472e-243	5. 230e-238	1. 0000	-242. 6070
2-Hydroxybutanoi	1. 577e-243	1. 642e-238	1. 0000	-242. 8021
Propane(aq)	2. 227e-249	9. 816e-245	1. 0000	-248. 6524
Zn(Gl y)2(aq)	1. 861e-250	3. 973e-245	1. 0000	-249. 7302
Di gl yci ne(aq)	1. 631e-250	2. 155e-245	1. 0000	-249. 7875
Mg(Gl y)2(aq)	7. 470e-251	1. 288e-245	1. 0000	-250. 1267
Fe(Gl y)2(aq)	5. 147e-252	1. 049e-246	1. 0000	-251. 2884
Mn(Gl y)2(aq)	4. 507e-252	9. 150e-247	1. 0000	-251. 3461
Butanoate	7. 508e-253	6. 538e-248	0. 9332	-252. 1545
Di ketopi perazi ne	6. 447e-253	7. 354e-248	1. 0000	-252. 1907
Pb(Gl y)2(aq)	3. 821e-253	1. 357e-247	1. 0000	-252. 4178
Cd(Gl y)2(aq)	3. 225e-253	8. 398e-248	1. 0000	-252. 4915
Ca(Gl y)2(aq)	4. 537e-255	8. 536e-250	1. 0000	-254. 3432
Butanoi c_aci d(aq)	2. 199e-255	1. 937e-250	1. 0000	-254. 6578
Na(But) (aq)	1. 040e-255	1. 145e-250	1. 0000	-254. 9829
Ca(But)+	1. 024e-255	1. 302e-250	0. 9332	-255. 0198
Mg(But)+	3. 876e-256	4. 317e-251	0. 9332	-255. 4416
K(But) (aq)	3. 660e-256	4. 618e-251	1. 0000	-255. 4365
Cu(But)+	1. 081e-256	1. 629e-251	0. 9332	-255. 9960
Mn(But)+	8. 559e-257	1. 215e-251	0. 9332	-256. 0976
Fe(But)+	3. 283e-257	4. 692e-252	0. 9332	-256. 5137
Ba(Gl y)2(aq)	8. 446e-258	2. 410e-252	1. 0000	-257. 0733
Zn(But)+	3. 066e-258	4. 674e-253	0. 9332	-257. 5434
Gl utarate	3. 439e-259	4. 472e-254	0. 7576	-258. 5842
Ba(But)+	3. 289e-259	7. 380e-254	0. 9332	-258. 5129

SpecE8_output_GSS_MW-01-15 May.txt				
Pb(But)+	1. 944e-259	5. 720e-254	0. 9332	-258. 7413
Cd(But)+	1. 844e-259	3. 678e-254	0. 9332	-258. 7642
H-Glutarate	3. 627e-261	4. 754e-256	0. 9332	-260. 4704
Threonine(aq)	1. 969e-262	2. 345e-257	1. 0000	-261. 7057
Glutaric acid(aq)	3. 797e-264	5. 014e-259	1. 0000	-263. 4206
Ethyl acetate(aq)	1. 621e-264	1. 428e-259	1. 0000	-263. 7902
S506--	1. 046e-264	2. 681e-259	0. 7576	-264. 1009
1-Propanamine(aq)	6. 937e-271	4. 099e-266	1. 0000	-270. 1588
alpha-Aminobutyric acid(aq)	6. 120e-273	6. 309e-268	1. 0000	-272. 2133
Glutamic acid(aq)	1. 552e-281	2. 283e-276	1. 0000	-280. 8091
UO2(SCN)2(aq)	8. 345e-287	3. 222e-281	1. 0000	-286. 0786
Butanal (aq)	1. 193e-288	8. 601e-284	1. 0000	-287. 9233
S4--	1. 179e-297	1. 512e-292	0. 7576	-297. 0490
1-Butyne(aq)	9. 175e-306	4. 962e-301	1. 0000	-300. 0000
Cu(CH3COO)3-	7. 411e-307	1. 783e-301	0. 9332	-300. 0000
Zn(CH3COO)3-	1. 592e-308	3. 861e-303	0. 9332	-300. 0000
Mn(CH3COO)3-	1. 129e-308	2. 619e-303	0. 9332	-300. 0000
Pb(CH3COO)3-	3. 226e-309	1. 240e-303	0. 9332	-300. 0000
1-Butanol (aq)	1. 386e-309	1. 027e-304	1. 0000	-300. 0000
Cd(CH3COO)3-	1. 316e-310	3. 810e-305	0. 9332	-300. 0000
1-Butene(aq)	1. 302e-311	7. 306e-307	1. 0000	-300. 0000
2-Hydroxypentano	3. 873e-315	4. 535e-310	0. 9332	-300. 0000
U(SCN)2++	9. 761e-317	3. 456e-311	0. 7597	-300. 0000
Gluconic acid(aq)	6. 305e-317	9. 211e-312	1. 0000	-300. 0000
2-Hydroxypentano	7. 074e-319	8. 355e-314	1. 0000	-300. 0000
Zn(CN)4--	0. 0000	0. 0000	0. 7576	-300. 0000
Cd(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Zn(But)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Zn(Ala)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Valine(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Undecanoic acid(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Undecanoate	0. 0000	0. 0000	0. 9332	-300. 0000
Cd(Ala)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Alanine(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Hexanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexyne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
H-Suberate	0. 0000	0. 0000	0. 9332	-300. 0000
H-Sebacate	0. 0000	0. 0000	0. 9332	-300. 0000
H-Pimelate	0. 0000	0. 0000	0. 9332	-300. 0000
H-Azelate	0. 0000	0. 0000	0. 9332	-300. 0000
H-Adipate	0. 0000	0. 0000	0. 9332	-300. 0000
2-Heptanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Tryptophan(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Toluene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Suberic acid(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Suberate	0. 0000	0. 0000	0. 7576	-300. 0000
Ca(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Sebacic acid(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Sebacate	0. 0000	0. 0000	0. 7576	-300. 0000
Sb2S4--	0. 0000	0. 0000	0. 7576	-300. 0000
Ca(Pent)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Butanone(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Hexene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ca(Pent)+	0. 0000	0. 0000	0. 9332	-300. 0000
S5--	0. 0000	0. 0000	0. 7576	-300. 0000
Ca(Lac)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pimelic acid(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Adipic acid(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Phenylalanine(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Phenol (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pentanoic acid(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Pentanoate	0. 0000	0. 0000	0. 9332	-300. 0000
Pentanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000

SpecE8_output_GSS_MW-01-15 May. txt

Fe(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Adipate	0.0000	0.0000	0.7576	-300.0000
1-Hexanol(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Pent)+	0.0000	0.0000	0.9332	-300.0000
Pb(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Pent)+	0.0000	0.0000	0.9332	-300.0000
Pb(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pimelate	0.0000	0.0000	0.7576	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Pb(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Fe(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9332	-300.0000
Fe(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanal(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexanamine(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoate	0.0000	0.0000	0.9332	-300.0000
Nonanal(aq)	0.0000	0.0000	1.0000	-300.0000
Ethylbenzene(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Prop)2-	0.0000	0.0000	0.9332	-300.0000
Na(Pent)2-	0.0000	0.0000	0.9332	-300.0000
Na(Pent)(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Lac)2-	0.0000	0.0000	0.9332	-300.0000
Dodecanoic_acid(0.0000	0.0000	1.0000	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoate	0.0000	0.0000	0.9332	-300.0000
Na(But)2-	0.0000	0.0000	0.9332	-300.0000
Decanoic_acid(aq)	0.0000	0.0000	1.0000	-300.0000
Benzoin_acid(aq)	0.0000	0.0000	1.0000	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9332	-300.0000
Decanal(aq)	0.0000	0.0000	1.0000	-300.0000
Benzoate	0.0000	0.0000	0.9357	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Pent)+	0.0000	0.0000	0.9332	-300.0000
Mn(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)+	0.0000	0.0000	0.9332	-300.0000
Tyrosine(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoic	0.0000	0.0000	1.0000	-300.0000
1-Pentanol(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mn(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoic	0.0000	0.0000	0.9332	-300.0000
Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoic	0.0000	0.0000	1.0000	-300.0000
1-Pentanamine(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_MW-01-15 May. txt

Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)+	0.0000	0.0000	0.9332	-300.0000
UO2(SCN)3-	0.0000	0.0000	0.9332	-300.0000
Mg(Pent)+	0.0000	0.0000	0.9332	-300.0000
Mg(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9332	-300.0000
Cu(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoi	0.0000	0.0000	1.0000	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Methi oni ne(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoa	0.0000	0.0000	0.9332	-300.0000
Leucyl gl yci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Leuci ne(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	1.0000	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol (aq)	0.0000	0.0000	1.0000	-300.0000
K(Prop)2-	0.0000	0.0000	0.9332	-300.0000
Cd(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
K(Pent)2-	0.0000	0.0000	0.9332	-300.0000
K(Pent)(aq)	0.0000	0.0000	1.0000	-300.0000
K(Lac)2-	0.0000	0.0000	0.9332	-300.0000
Cd(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptano	0.0000	0.0000	0.9332	-300.0000
Cd(Pent)+	0.0000	0.0000	0.9332	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
K(But)2-	0.0000	0.0000	0.9332	-300.0000
p-Tol uate	0.0000	0.0000	0.9332	-300.0000
I sol euci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoi	0.0000	0.0000	1.0000	-300.0000
1-Octanol (aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
n-Propyl benzene(0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9332	-300.0000
Hexanal (aq)	0.0000	0.0000	1.0000	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoi c_aci d(a	0.0000	0.0000	1.0000	-300.0000
Ba(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9332	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanal (aq)	0.0000	0.0000	1.0000	-300.0000
n-Butane(aq)	0.0000	0.0000	1.0000	-300.0000
m-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
m-Tol uate	0.0000	0.0000	0.9332	-300.0000
HSb2S4-	0.0000	0.0000	0.9332	-300.0000
2-Hydroxydecano	0.0000	0.0000	0.9332	-300.0000
Azel ate	0.0000	0.0000	0.7576	-300.0000
Zn(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Zn(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Zn(Pent)+	0.0000	0.0000	0.9332	-300.0000
Zn(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(a	0.0000	0.0000	1.0000	-300.0000
Azel ai c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexyl benzene(a	0.0000	0.0000	1.0000	-300.0000
1-Octanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octyl benzene(a	0.0000	0.0000	1.0000	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000

	SpecE8_output_GSS_MW-01-15 May. txt			
n-Pentyl benzene(0. 0000	0. 0000	1. 0000	-300. 0000
1-Heptanami ne(aq	0. 0000	0. 0000	1. 0000	-300. 0000
o-Phthal ate	0. 0000	0. 0000	0. 7576	-300. 0000
o-Tol uate	0. 0000	0. 0000	0. 9332	-300. 0000
o-Tol ui c_aci d(aq	0. 0000	0. 0000	1. 0000	-300. 0000
1-Butanami ne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
p-Tol ui c_aci d(aq	0. 0000	0. 0000	1. 0000	-300. 0000
(NH4) 2Sb2S4(aq)	0. 0000	0. 0000	1. 0000	-300. 0000

Mineral saturation states
log Q/K

log Q/K

Nontronite-Ca	22. 5179s/sat	U02S04: H2O	-10. 9589
Nontronite-Mg	22. 4428s/sat	Mg1. 25S04(OH)0. 5	-10. 9763
Nontronite-K	22. 3771s/sat	Na2CO3: 7H2O	-11. 1091
Nontronite-Na	22. 1240s/sat	Cotunite	-11. 2863
Nontronite-H	21. 6949s/sat	Sepiolite	-11. 3114
Hematite	18. 0566s/sat	Clinochlore-7A	-11. 4276
Ferri te-Cu	15. 6318s/sat	Periclase	-11. 9744
Magnetite	15. 4316s/sat	Forsterite	-12. 0074
Cronstedtite-7A	14. 2024s/sat	MnS04	-12. 0560
Ferri te-Zn	13. 2270s/sat	UP05	-12. 1227
Delafossite	12. 5533s/sat	Na2U207	-12. 1524
Clinoptilolite-h	10. 8820s/sat	U02. 6667	-12. 1687
Clinoptilolite-C	10. 8010s/sat	BaU207	-12. 1915
Muscovite	10. 7775s/sat	CdS04	-12. 1938
Clinoptilolite-K	10. 4740s/sat	U02Cl OH: 2H2O	-12. 3153
Clinoptilolite-h	10. 1509s/sat	Gyrolite	-12. 3379
Stilbite	9. 4227s/sat	Przhevalskite	-12. 3873
Goethite	8. 5681s/sat	Portlandite	-12. 4866
Clinoptilolite-N	7. 4472s/sat	MgU04	-12. 4978
Clinoptilolite-h	7. 4455s/sat	Mg1. 5S04(OH)	-12. 5531
Illite	7. 3442s/sat	Thermonatrite	-12. 5675
Beidellite-Ca	7. 3407s/sat	Bassetite	-12. 7820
Beidellite-Mg	7. 2655s/sat	Na2CO3	-12. 8916
Beidellite-K	7. 2000s/sat	Autunite-H	-12. 9484
Epidote-ord	7. 0263s/sat	FeS04	-13. 0037
Epidote	7. 0253s/sat	Uraninite	-13. 0216
Paragonite	6. 9505s/sat	Monticellite	-13. 1205
Beidellite-Na	6. 9467s/sat	MgS04	-13. 2049
Kaolinite	6. 8452s/sat	PbF2	-13. 2565
Ferri te-Ca	6. 6568s/sat	Coffinite	-13. 2700
Beidellite-H	6. 5181s/sat	Chalcocyanite	-13. 2740
Ferri te-Mg	6. 4695s/sat	Mordenite-dehy	-14. 0418
Mesolite	6. 2770s/sat	FeF2	-14. 1784
Pyrophyllite	6. 2608s/sat	BaCl 2: 2H2O	-14. 3104
Fluorapatite	6. 1990s/sat	CdCl 2: H2O	-14. 5739
Montmor-Ca	5. 8990s/sat	ZnS04	-14. 8177
Montmor-Mg	5. 8987s/sat	BaCl 2: H2O	-15. 0422
Montmor-K	5. 8331s/sat	U02S04: 3. 5H2O	-15. 4251
Andradite	5. 8325s/sat	U02S04: 2. 5H2O	-15. 4856
Montmor-Na	5. 5751s/sat	U02S04: 3H2O	-15. 5630
Scolecite	4. 8121s/sat	CdCl 2	-15. 6928
Jarosite	4. 1942s/sat	U02F2: 3H2O	-15. 8035
Maximum_Microcline	4. 1780s/sat	NaU03	-15. 8500
K-Feldspar	4. 1761s/sat	MnCl 2: 4H2O	-16. 0469
Laumontite	3. 5056s/sat	Zn(NO3)2: 6H2O	-16. 0858
Fe(OH)3	3. 2515s/sat	U02F2	-16. 1231
Smectite-low-Fe-	3. 1200s/sat	Alum-K	-16. 2500
Diaspore	3. 0632s/sat	Pb4S07	-16. 3633
Celadonite	2. 9747s/sat	BaCl 2	-16. 5905
Margarite	2. 9303s/sat	Ba(OH)2: 8H2O	-16. 6849
Sandineigh	2. 8816s/sat	ZnF2	-17. 0448

SpecE8_output_GSS_MW-01-15 May.txt

Boehmi te	2. 6348s/sat	Natrosil i te	-17. 1798
Gi bbsi te	2. 5380s/sat	CaAl 204	-17. 1978
Mordeni te	1. 9827s/sat	MnCl 2: 2H2O	-17. 4939
Smecti te-hi gh-Fe	1. 8589s/sat	CdF2	-17. 4979
Corundum	1. 8347s/sat	MgOHCl	-17. 5534
Lawsoni te	1. 8281s/sat	Ni ngoyi te	-17. 6631
Wi theri te	1. 3148s/sat	Na2Si O3	-17. 8619
Al bi te_low	1. 2846s/sat	UO2. 3333(beta)	-17. 9614
Al bi te	1. 2845s/sat	CuCl 2	-18. 2058
Kyani te	0. 9900s/sat	Ba2Si 308	-18. 5712
Andal usi te	0. 6821s/sat	Tremol i te	-18. 5909
Anal ci me	0. 6217s/sat	CaAl 407	-18. 9408
Quartz	0. 5914s/sat	MnCl 2: H2O	-19. 1767
Dawsoni te	0. 4022s/sat	Ferri te-Di cal ci u	-19. 3015
Tri dymi te	0. 3911s/sat	UO2S04	-19. 3347
Bari te	0. 3313s/sat	Gehl eni te	-19. 5063
Chal cedony	0. 3066s/sat	MgCl 2: 4H2O	-19. 6273
Si l l i mani te	0. 2901s/sat	Hi l l ebrandi te	-19. 9018
Tenori te	0. 2825s/sat	Di cal ci um_si l l i ca	-20. 3788
Mal achi te	0. 2344s/sat	UOFOH: . 5H2O	-20. 6806
Strengi te	0. 0976s/sat	Hydromagnesi te	-20. 7662
Cri stobal i te(al p	0. 0080s/sat	UO2(N03) 2: 6H2O	-21. 0805
Natrol i te	-0. 0692	Na2UO4(al pha)	-21. 1686
I ce	-0. 0850	UOFOH	-21. 1894
Al bi te_hi gh	-0. 1306	Larni te	-21. 7553
Coesi te	-0. 2566	Akermani te	-22. 1649
Kal si l i te	-0. 3438	UO2Cl	-22. 3631
Chamosi te-7A	-0. 3466	Scacchi te	-22. 5904
Cri stobal i te(bet	-0. 4768	ZnCl 2	-22. 6273
Al uni te	-0. 5227	UO2(N03) 2: 3H2O	-22. 7303
Cerussi te	-0. 6325	Li me	-23. 0754
Rhodochrosi te	-0. 7784	UOF2: H2O	-23. 3757
Si O2(am)	-0. 8136	(UO2) 3(P04) 2	-23. 6964
Jadei te	-0. 9287	UO2(P03) 2	-23. 7870
Hercyni te	-0. 9929	Hydrophi l i te	-23. 8525
Cal ci te	-1. 0365	Lawrenci te	-23. 8812
Brochanti te	-1. 0919	UOF2	-24. 0261
Aragoni te	-1. 1816	UO2(N03) 2: 2H2O	-24. 0952
Wai raki te	-1. 3658	(UO2) 2P207	-24. 2072
Dol omi te-ord	-1. 4869	U(HP04) 2: 4H2O	-25. 3050
Dol omi te	-1. 4869	MgCl 2: 2H2O	-25. 4576
Di optase	-1. 5130	UP207	-25. 6390
Si deri te	-1. 7518	KAl (S04) 2	-25. 8045
Monohydrocal ci te	-1. 8371	K2UO4	-25. 8108
Prehni te	-1. 9717	Anthophyl l i te	-25. 8709
Anni te	-2. 1285	UO2Cl 2: 3H2O	-27. 0142
Magnesi te	-2. 1689	Pb	-27. 0422
Amesi te-14A	-2. 3712	Cl i nopti l ol i te-d	-27. 0498
Cl i nozoi si te	-2. 3838	Cl i nopti l ol i te-d	-27. 4133
Zoi si te	-2. 4319	UO2(N03) 2: H2O	-27. 9267
Gypsum	-2. 5475	Ranki ni te	-28. 3522
Anhydri te	-2. 8643	Pargasi te	-28. 8060
Berl i ni te	-3. 0094	MgCl 2: H2O	-29. 0362
Daphni te-14A	-3. 0210	UO2Cl 2: H2O	-29. 9665
Anorthi te	-3. 0426	Cl i nopti l ol i te-d	-30. 4168
Nephel i ne	-3. 0869	Ettri ngi te	-31. 1412
Dol omi te-di s	-3. 1387	UO2(N03) 2	-31. 6083
Ferrosil i te	-3. 1471	Tobermori te-14A	-32. 0936
Smi thsoni te	-3. 2451	UF4: 2. 5H2O	-32. 8176
Fl uori te	-3. 2654	UO2S03	-33. 1911
Al amosi te	-3. 3328	Ba2Si O4	-33. 6865
Zn2Si O4	-3. 3337	KMgCl 3: 2H2O	-33. 8030
Bassani te	-3. 5136	UO2Cl 2	-34. 0727

SpecE8_output_GSS_MW-01-15 May. txt

Sb203	-3. 6565	Tobermori te-11A	-34. 3274
CaS04: 0. 5H20(bet	-3. 7001	Zn30(S04)2	-34. 8752
Schoepi te	-3. 7066	Chl oromagnesi te	-35. 1216
U03: 2H20	-3. 7066	Foshagi te	-35. 3740
U02(OH)2(beta)	-3. 8716	Merwi ni te	-35. 4006
Schoepi te-dehy(.	-3. 9327	Al 2(S04)3: 6H20	-35. 7729
U03: . 9H20(al pha)	-3. 9390	BaSi F6	-36. 2070
Atacami te	-4. 0006	Afwil li te	-36. 6216
Phl ogopi te	-4. 0197	Cd	-36. 9583
Schoepi te-dehy(.	-4. 0244	UF4	-37. 3562
Schoepi te-dehy(1	-4. 0407	Fe	-37. 4012
Angl esi te	-4. 0568	Tobermori te-9A	-38. 1626
Saponi te-Ca	-4. 1107	Mol ysi te	-39. 0486
Saponi te-Mg	-4. 1841	Fe2(S04)3	-40. 7282
Mi nnesotai te	-4. 1842	U(S04)2: 8H20	-41. 4045
Saponi te-K	-4. 2514	KMgCl 3	-41. 5674
Hydroxyl apati te	-4. 3039	BaO	-41. 8308
CaU04	-4. 3206	U(S04)2: 4H20	-42. 7254
Whi tlocki te	-4. 3262	U(S04)2	-42. 9794
Zi nci te	-4. 3390	Ba2U207	-44. 7571
Saponi te-Na	-4. 5045	C	-44. 9919
Cupri te	-4. 5486	UOCl 2	-46. 4224
Rhodoni te	-4. 6163	U(C03)2	-46. 5293
U02C03	-4. 6383	Hatruri te	-47. 3155
Rutherfordi ne	-4. 6523	UCI F3	-48. 5028
Sel lai te	-4. 6573	Zn	-48. 5629
Zn(OH)2(epsil on)	-4. 7375	U5012Cl	-49. 6996
Saponi te-H	-4. 9334	Xonotli te	-51. 2462
Zn(OH)2(beta)	-5. 0216	UOF4	-53. 5437
Azuri te	-5. 0483	Al 2(S04)3	-54. 4862
Enstati te	-5. 0719	Na3U04	-57. 0495
Wusti te	-5. 1352	(U02)2Cl 3	-57. 6348
Tal c	-5. 1993	Na	-58. 1163
Schoepi te-dehy(.	-5. 2118	UOCl 3	-58. 7584
Paral auri oni te	-5. 2124	UF3	-59. 4035
Nesquehoni te	-5. 2775	Na4Si 04	-59. 6109
CdSi 03	-5. 2914	UF5(beta)	-60. 1853
Bi xbyi te	-5. 3554	UF5(al pha)	-60. 5887
Nahcol i te	-5. 5044	Mn	-61. 5690
Pyrol usi te	-5. 5103	Na20	-61. 7075
Lanarki te	-5. 5491	Pb2Cl 5NH4	-61. 9315
Ri pi dol i te-14A	-5. 6618	UCI 2F2	-61. 9638
Ni ter	-5. 7204	K	-62. 2170
Schoepi te-dehy(.	-5. 7618	Ca4Al 2Fe2010	-64. 0088
U02HP04: 4H20	-5. 8660	Ca3Al 206	-64. 3608
Fe0	-5. 9584	UOCl	-64. 5648
Greenal i te	-6. 0107	Covel li te	-66. 3551
Anal ci me-dehy	-6. 0967	S	-67. 5649
Fe(OH)2	-6. 2402	Chal coci te	-68. 7373
Daphni te-7A	-6. 5255	U203F6	-69. 8148
Bruci te	-6. 5994	Ba3U06	-72. 7219
Okeni te	-6. 7647	UCI 3F	-75. 2408
Mn(OH)2(am)	-6. 7894	CdS	-75. 6954
U03(gamma)	-6. 8429	Gal ena	-76. 8007
Wol l astoni te	-6. 9288	Sphal eri te	-79. 1840
Li tharge	-7. 0191	K20	-79. 7627
Chal canthi te	-7. 0641	Wurtzi te	-81. 5989
Fayal i te	-7. 1373	U(S03)2	-81. 8483
Pseudowol l astoni	-7. 1939	CdCl 2(NH3)2	-84. 0325
Massi cot	-7. 2161	ZnCl 2(NH3)2	-84. 8687
NaFe02	-7. 2720	Troi li te	-86. 3262
Mel anteri te	-7. 2738	UCI 4	-86. 3566
U03(beta)	-7. 4768	Pyrrhoti te	-86. 4301

SpecE8_output_GSS_MW-01-15 May. txt

Nantoki te	-7. 4987	Na6Si 207	-86. 7443
Cordi eri te_hydr	-7. 5218	PbS04(NH3) 2	-86. 9491
Mi rabi li te	-7. 6389	Al abandi te	-89. 0150
Di opsi de	-7. 6543	UCI 3	-89. 8928
Hausmanni te	-7. 7845	UF6	-91. 9793
U03(al pha)	-7. 8307	U305F8	-93. 2646
PbC03. Pb0	-7. 9306	U2F9	-94. 2487
Cl i nochl ore-14A	-7. 9351	Mg	-102. 0075
Syl vi te	-7. 9552	U202Cl 5	-106. 4670
Spi nel	-8. 0569	UCI 5	-107. 2916
PbFCI	-8. 1152	BaS	-107. 3176
Cd(OH) 2	-8. 1311	UN1. 73(al pha)	-108. 5115
FeF3	-8. 1575	UN1. 59(al pha)	-111. 3506
Arcani te	-8. 1808	P	-119. 1270
U02F0H: 2H2O	-8. 2810	Ca	-119. 4719
Hedenbergi te	-8. 3890	Anti gori te	-119. 9565
Chrysoti le	-8. 4042	Al	-122. 5509
Hal i te	-8. 4240	Ba	-123. 6820
Cu	-8. 4943	UN	-124. 8551
ZnS04: 7H2O	-8. 6183	UCI 6	-127. 0148
Cd(OH)Cl	-8. 7144	Si	-127. 4112
U02F0H: H2O	-8. 7370	Pb(N3) 2(orth)	-142. 5498
Sanborni te	-8. 7998	Pyri te	-142. 9572
Al F3	-8. 8116	Pb(N3) 2(mono)	-143. 0062
Sal eei te	-8. 8384	Chal copyri te	-146. 4809
ZnS04: 6H2O	-8. 9230	ZnCl 2(NH3) 4	-162. 1114
Thenardi te	-9. 1177	CdCl 2(NH3) 4	-163. 0509
Ri pi dol i te-7A	-9. 1583	PbS04(NH3) 4	-167. 6436
Hunti te	-9. 2192	U4F17	-168. 0511
Ca-Al_Pyroxene	-9. 2569	U	-173. 3167
U02F0H	-9. 2619	US	-181. 8316
Ni trobari te	-9. 3433	Zn(Cl 04) 2: 6H2O	-184. 1124
Tephroi te	-9. 4263	UC	-200. 1224
BaU04	-9. 4641	UH3(beta)	-203. 0792
Manganosi te	-9. 6083	US1. 9	-207. 9627
Arti ni te	-9. 6158	US2	-212. 9154
Monteponi te	-9. 6278	Mayeni te	-238. 6686
U02. 25	-9. 8039	ZnCl 2(NH3) 6	-241. 0863
(U02) 3(P04) 2: 4H2	-9. 8150	CdCl 2(NH3) 6	-242. 1634
U02. 25(beta)	-9. 8724	UP	-243. 4900
Torberni te	-10. 0297	UC1. 94(al pha)	-244. 5394
Pb2Si 04	-10. 0904	Sti bni te	-252. 9960
Cordi eri te_anhyd	-10. 1525	US3	-277. 1565
CdS04: 2. 667H2O	-10. 2190	Pb(SCN) 2	-282. 4818
Phosgeni te	-10. 2618	Borni te	-283. 6417
ZnS04: H2O	-10. 4262	UP2	-357. 3001
CdS04: H2O	-10. 4734	U2S3	-387. 4864
Uranoci rci te	-10. 4898	o-Phthal i c_aci d	-396. 7500
Natron	-10. 5542	U2C3	-446. 8323
Grossul ar	-10. 8156	U3S5	-595. 5970
Pb3S06	-10. 9216	U3P4	-844. 2789

Gases	fugaci ty	l og fug.
N2(g)	0. 1690	-0. 772
H2O(g)	0. 01061	-1. 974
CO2(g)	0. 003611	-2. 442
NO2(g)	8. 919e-012	-11. 050
HF(g)	1. 084e-014	-13. 965
HCl (g)	5. 792e-019	-18. 237
H2(g)	8. 575e-030	-29. 067
O2(g)	1. 023e-030	-29. 990
NO(g)	5. 533e-032	-31. 257

SpecE8_output_GSS_MW-01-15 May. txt

C0(g)	1. 032e-035	-34. 986
Si F4(g)	9. 321e-039	-38. 031
NH3(g)	1. 988e-041	-40. 702
S02(g)	4. 360e-043	-42. 361
Cl 2(g)	3. 874e-043	-42. 412
Cd(g)	3. 437e-052	-51. 464
Pb(g)	6. 741e-058	-57. 171
U02F2(g)	2. 179e-061	-60. 662
Cu(g)	2. 537e-064	-63. 596
Zn(g)	4. 906e-067	-66. 309
U02Cl 2(g)	5. 216e-073	-72. 283
Na(g)	2. 919e-073	-72. 535
K(g)	2. 592e-074	-73. 586
U03(g)	1. 551e-074	-73. 809
U0F4(g)	2. 123e-075	-74. 673
UF5(g)	5. 523e-082	-81. 258
UF4(g)	1. 258e-084	-83. 900
H2S(g)	2. 681e-091	-90. 572
UF6(g)	5. 799e-094	-93. 237
CH4(g)	2. 597e-094	-93. 586
F2(g)	7. 310e-101	-100. 136
UCl 4(g)	2. 888e-113	-112. 539
U02(g)	1. 454e-115	-114. 838
Mg(g)	1. 016e-123	-122. 993
UCl 5(g)	1. 311e-126	-125. 883
UF3(g)	1. 065e-130	-129. 972
UCl 6(g)	2. 512e-134	-133. 600
U2F10(g)	1. 984e-136	-135. 703
UCl 3(g)	4. 118e-139	-138. 385
Ca(g)	5. 737e-147	-146. 241
S2(g)	7. 113e-151	-150. 148
C2H4(g)	3. 077e-161	-160. 512
C(g)	1. 387e-169	-168. 858
UF2(g)	2. 114e-173	-172. 675
Al (g)	7. 644e-177	-176. 117
UCl 2(g)	4. 857e-183	-182. 314
U0(g)	1. 238e-189	-188. 907
Si (g)	4. 080e-203	-202. 389
UF(g)	2. 622e-209	-208. 581
U2Cl 8(g)	2. 033e-214	-213. 692
U2Cl 10(g)	2. 060e-224	-223. 686
UCl (g)	1. 993e-224	-223. 700
U(g)	2. 784e-264	-263. 555

Original	basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
<hr/>							
Al +++		1. 45e-006	1. 45e-006	0. 0391			
Ba++		7. 19e-007	7. 19e-007	0. 0988			
Ca++		0. 000537	0. 000537	21. 5			
Cd++		7. 83e-009	7. 83e-009	0. 000880			
Cl -		0. 000110	0. 000110	3. 91			
Cu++		1. 05e-006	1. 05e-006	0. 0665			
F-		1. 03e-005	1. 03e-005	0. 196			
Fe++		1. 05e-006	1. 05e-006	0. 0587			
H+		0. 000187	0. 000187	0. 188			
H2O		55. 5	55. 5	1. 00e+006			
HC03-		0. 00164	0. 00164	99. 7			
HP04--		2. 31e-006	2. 31e-006	0. 221			
K+		0. 000550	0. 000550	21. 5			
Mg++		0. 000161	0. 000161	3. 91			
Mn++		8. 10e-006	8. 10e-006	0. 445			
NH3(aq)		0. 000283	0. 000283	4. 82			

SpecE8_output_GSS_MW-01-15 May.txt

N03-	0.00117	0.00117	72.7
Na+	0.00136	0.00136	31.3
O2(aq)	0.000212	0.000212	6.80
Pb++	5.19e-009	5.19e-009	0.00108
S2--	7.63e-007	7.63e-007	0.0489
S04--	0.000316	0.000316	30.3
Sb(OH)3(aq)	2.57e-007	2.57e-007	0.0444
Si O2(aq)	0.000179	0.000179	10.8
U02++	8.22e-009	8.22e-009	0.00222
Zn++	1.50e-007	1.50e-007	0.00978

Elemental composition	In fluid		Sorbed	
	total moles	moles	mg/kg	moles mg/kg
Al umi num	1.450e-006	1.450e-006	0.03911	
Anti mony	2.571e-007	2.571e-007	0.03129	
Bari um	7.194e-007	7.194e-007	0.09876	
Cadmi um	7.831e-009	7.831e-009	0.0008801	
Cal ci um	0.0005369	0.0005369	21.51	
Carbon	0.001635	0.001635	19.63	
Chl ori ne	0.0001104	0.0001104	3.911	
Copper	1.047e-006	1.047e-006	0.06649	
Fl uori ne	1.030e-005	1.030e-005	0.1956	
Hydrogen	111.0	111.0	1.119e+005	
Iron	1.051e-006	1.051e-006	0.05867	
Lead	5.193e-009	5.193e-009	0.001076	
Magnesi um	0.0001610	0.0001610	3.911	
Manganese	8.101e-006	8.101e-006	0.4449	
Ni trogen	0.001456	0.001456	20.39	
Oxygen	55.52	55.52	8.880e+005	
Phosphorus	2.305e-006	2.305e-006	0.07138	
Potassi um	0.0005504	0.0005504	21.51	
Si l i con	0.0001791	0.0001791	5.028	
Sodi um	0.001361	0.001361	31.29	
Sul fur	0.0003172	0.0003172	10.17	
Urani um	8.219e-009	8.219e-009	0.001956	
Zi nc	1.496e-007	1.496e-007	0.009778	

SpecE8_output_GSS_MW-95-8 July.txt

Temperature = 10.3 C Pressure = 1.013 bars
 pH = 7.700 log fO2 = -26.089
 Eh = 0.4320 volts pe = 7.6810
 Ionic strength = 0.007290
 Charge imbalance = 0.000189 eq/kg (3.77% error)
 Activity of water = 0.999996
 Solvent mass = 1.000000 kg
 Solution mass = 1.000427 kg
 Solution density = 1.023 g/cm3
 Chlorinity = 0.000110 molal
 Dissolved solids = 427 mg/kg sol'n
 Hardness = 198.19 mg/kg sol'n as CaCO3
 carbonate = 128.90 mg/kg sol'n as CaCO3
 non-carbonate = 69.30 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 128.90 mg/kg sol'n as CaCO3
 Water type = Ca-HCO3

Nernst redox couples	Eh (volts)	pe
e- + .25*O2(aq) + H+ = .5*H2O	0.4320	7.6810

No minerals in system.

Aqueous species	molality	mg/kg sol'n	act. coef.	log act.
HC03-	0.002517	153.5	0.9160	-2.6372
Ca++	0.001584	63.44	0.7146	-2.9463
NO3-	0.001543	95.63	0.9139	-2.8508
Na+	0.001101	25.31	0.9160	-2.9962
SO4--	0.0003711	35.64	0.7026	-3.5837
Mg++	0.0003047	7.401	0.7257	-3.6555
K+	0.0001997	7.804	0.9139	-3.7387
SiO2(aq)	0.0001897	11.39	1.0000	-3.7220
CO2(aq)	0.0001345	5.918	1.0000	-3.8712
Cl-	0.0001103	3.908	0.9139	-3.9966
CaSO4(aq)	3.560e-005	4.844	1.0000	-4.4486
CaHCO3+	3.300e-005	3.335	0.9160	-4.5196
MgSO4(aq)	1.023e-005	1.231	1.0000	-4.9900
CaCO3(aq)	6.785e-006	0.6788	1.0000	-5.1685
MgHCO3+	6.250e-006	0.5330	0.9160	-5.2422
CO3--	5.327e-006	0.3195	0.7057	-5.4249
F-	5.095e-006	0.09675	0.9150	-5.3315
NaHCO3(aq)	4.435e-006	0.3724	1.0000	-5.3531
Ba++	9.604e-007	0.1318	0.7087	-6.1670
HSiO3-	7.167e-007	0.05523	0.9160	-6.1828
MgCO3(aq)	6.879e-007	0.05798	1.0000	-6.1625
KS04-	3.904e-007	0.05275	0.9160	-6.4466
OH-	1.637e-007	0.002782	0.9150	-6.8246
Cu++	1.077e-007	0.006842	0.7146	-7.1136
NaHSiO3(aq)	4.111e-008	0.004113	1.0000	-7.3860
CaF+	2.613e-008	0.001543	0.9160	-7.6209
MgF+	2.590e-008	0.001121	0.9160	-7.6248
CaCl+	2.533e-008	0.001913	0.9160	-7.6344
H+	2.157e-008	2.173e-005	0.9251	-7.7000
NaCO3-	2.041e-008	0.001693	0.9160	-7.7283
MgCl+	1.944e-008	0.001161	0.9160	-7.7494
NaCl(aq)	1.572e-008	0.0009184	1.0000	-7.8035
UO2(CO3)3----	1.278e-009	0.0005749	0.2433	-9.5074
BaCO3(aq)	8.312e-010	0.0001640	1.0000	-9.0803
UO2(CO3)2--	7.655e-010	0.0002985	0.7026	-9.2693
KCl(aq)	4.417e-010	3.291e-005	1.0000	-9.3549

	SpecE8_output_GSS_MW-95-8 Jul y. txt			
NaF(aq)	4. 196e-010	1. 761e-005	1. 0000	-9. 3772
HSO4-	3. 767e-010	3. 655e-005	0. 9160	-9. 4621
HF(aq)	1. 062e-010	2. 124e-006	1. 0000	-9. 9739
NaOH(aq)	2. 659e-011	1. 063e-006	1. 0000	-10. 5752
BaCl +	1. 966e-011	3. 396e-006	0. 9160	-10. 7444
UO2CO3(aq)	1. 042e-011	3. 438e-006	1. 0000	-10. 9820
CaCl 2(aq)	3. 238e-012	3. 592e-007	1. 0000	-11. 4897
BaF+	1. 971e-012	3. 080e-007	0. 9160	-11. 7434
HNO3(aq)	1. 018e-012	6. 414e-008	1. 0000	-11. 9921
Cu+	5. 610e-013	3. 563e-008	0. 9160	-12. 2892
HCl (aq)	4. 273e-013	1. 557e-008	1. 0000	-12. 3693
UO2OH+	8. 267e-014	2. 372e-008	0. 9160	-13. 1208
KHSO4(aq)	3. 646e-015	4. 963e-010	1. 0000	-14. 4382
UO2++	8. 660e-016	2. 337e-010	0. 7057	-15. 2138
UO2F+	3. 459e-016	9. 994e-011	0. 9160	-15. 4991
UO2SO4(aq)	1. 296e-016	4. 742e-011	1. 0000	-15. 8875
HF2-	1. 108e-016	4. 321e-012	0. 9160	-15. 9934
UO2F2(aq)	4. 649e-018	1. 431e-012	1. 0000	-17. 3327
UO2(SO4)2--	2. 955e-019	1. 365e-013	0. 7026	-18. 6826
UO2Cl +	8. 382e-020	2. 560e-014	0. 9160	-19. 1147
UO2F3-	4. 058e-021	1. 327e-015	0. 9160	-20. 4297
(UO2) 2(OH) 2++	1. 353e-021	7. 763e-016	0. 7057	-21. 0201
UO2+	3. 202e-022	8. 643e-017	0. 9160	-21. 5327
(UO2) 3(CO3) 6(6-)	7. 126e-023	8. 335e-017	0. 0415	-23. 5288
(UO2) 3(OH) 5+	2. 451e-024	2. 193e-018	0. 9160	-23. 6488
UO2Cl 2(aq)	3. 461e-025	1. 180e-019	1. 0000	-24. 4607
UO2F4--	1. 519e-025	5. 252e-020	0. 7026	-24. 9719
U(OH) 4(aq)	1. 533e-026	4. 689e-021	1. 0000	-25. 8145
O2(aq)	1. 390e-029	4. 446e-025	1. 0000	-28. 8570
Formate	4. 727e-034	2. 127e-029	0. 9150	-33. 3640
Cl O-	7. 832e-035	4. 028e-030	0. 9160	-34. 1442
H2(aq)	6. 751e-035	1. 360e-031	1. 0000	-34. 1706
Ca(For) +	1. 602e-035	1. 363e-030	0. 9160	-34. 8334
Mg(For) +	3. 793e-036	2. 629e-031	0. 9160	-35. 4591
H02-	6. 008e-037	1. 982e-032	0. 9160	-36. 2593
Na(For) (aq)	5. 182e-037	3. 522e-032	1. 0000	-36. 2855
K(For) (aq)	8. 135e-038	6. 840e-033	1. 0000	-37. 0897
Formi c_aci d(aq)	5. 084e-038	2. 339e-033	1. 0000	-37. 2938
SO3--	8. 638e-039	6. 913e-034	0. 7057	-38. 2150
Ba(For) +	7. 305e-039	1. 331e-033	0. 9160	-38. 1745
Cu(For) +	4. 536e-039	4. 923e-034	0. 9160	-38. 3814
HSO3-	1. 883e-039	1. 526e-034	0. 9160	-38. 7632
Si F6--	4. 712e-040	6. 692e-035	0. 7026	-39. 4800
CO(aq)	8. 307e-041	2. 326e-036	1. 0000	-40. 0806
Oxal ate	2. 757e-041	2. 426e-036	0. 7026	-40. 7128
U(CO3) 5(6-)	1. 336e-043	7. 183e-038	0. 0415	-44. 2560
H-Oxal ate	6. 963e-045	6. 196e-040	0. 9160	-44. 1953
HSO5-	6. 191e-045	6. 997e-040	0. 9160	-44. 2464
UOH+++	5. 017e-045	1. 279e-039	0. 4605	-44. 6363
SO2(aq)	1. 857e-045	1. 189e-040	1. 0000	-44. 7312
UF2++	2. 179e-046	6. 013e-041	0. 7057	-45. 8130
UF3+	1. 580e-046	4. 660e-041	0. 9160	-45. 8394
UF+++	8. 970e-048	2. 305e-042	0. 4605	-47. 3839
UF4(aq)	7. 106e-048	2. 230e-042	1. 0000	-47. 1484
USO4++	4. 622e-049	1. 544e-043	0. 7057	-48. 4865
U(SO4) 2(aq)	3. 684e-049	1. 584e-043	1. 0000	-48. 4337
Oxal i c_aci d(aq)	2. 310e-051	2. 079e-046	1. 0000	-50. 6364
U++++	1. 724e-051	4. 102e-046	0. 2561	-51. 3551
UCl +++	7. 533e-054	2. 059e-048	0. 4605	-53. 4598
Cl O2-	1. 062e-057	7. 157e-053	0. 9160	-57. 0122
S2O8--	8. 396e-062	1. 612e-056	0. 7026	-61. 2292
S2O6--	2. 493e-063	3. 990e-058	0. 7026	-62. 7566
Cl O3-	5. 345e-066	4. 459e-061	0. 9150	-65. 3106

	SpecE8_output_GSS_MW-95-8 Jul y. txt			
Ca(For)2(aq)	5. 627e-068	7. 319e-063	1. 0000	-67. 2497
Mg(For)2(aq)	1. 659e-068	1. 896e-063	1. 0000	-67. 7801
U+++	6. 909e-070	1. 644e-064	0. 4605	-69. 4974
Na(For)2-	1. 766e-070	1. 996e-065	0. 9160	-69. 7910
Cu(For)2(aq)	5. 294e-071	8. 128e-066	1. 0000	-70. 2762
K(For)2-	2. 322e-071	2. 997e-066	0. 9160	-70. 6722
Ba(For)2(aq)	1. 936e-071	4. 400e-066	1. 0000	-70. 7130
Formal dehyde(aq)	1. 324e-073	3. 974e-069	1. 0000	-72. 8780
Cl O4-	1. 039e-078	1. 032e-073	0. 9150	-78. 0222
UO2Cl O3+	1. 114e-080	3. 935e-075	0. 9160	-79. 9914
S2O5--	6. 781e-083	9. 770e-078	0. 7026	-82. 3220
Methanol (aq)	3. 999e-092	1. 281e-087	1. 0000	-91. 3981
Gl ycol ate	4. 989e-098	3. 742e-093	0. 9160	-97. 3401
Ca(Gl yc)+	2. 948e-099	3. 393e-094	0. 9160	-98. 5685
Mg(Gl yc)+	2. 670e-100	2. 651e-095	0. 9160	-99. 6117
HS-	1. 970e-100	6. 513e-096	0. 9150	-99. 7441
H2S(aq)	5. 984e-101	2. 039e-096	1. 0000	-100. 2230
Na(Gl yc) (aq)	5. 597e-101	5. 484e-096	1. 0000	-100. 2521
K(Gl yc) (aq)	8. 787e-102	1. 003e-096	1. 0000	-101. 0562
Gl ycol i c_aci d(aq)	6. 521e-102	4. 957e-097	1. 0000	-101. 1857
Cu(Gl yc)+	4. 446e-102	6. 159e-097	0. 9160	-101. 3902
Ba(Gl yc)+	3. 583e-103	7. 606e-098	0. 9160	-102. 4839
S2O4--	3. 477e-105	4. 454e-100	0. 7087	-104. 6083
Methane(aq)	2. 720e-105	4. 362e-101	1. 0000	-104. 5654
S--	5. 315e-106	1. 704e-101	0. 7087	-105. 4240
S2O3--	1. 084e-106	1. 215e-101	0. 7026	-106. 1181
Acetate	1. 439e-110	8. 492e-106	0. 9170	-109. 8796
CaCH3COO+	1. 467e-112	1. 454e-107	0. 9160	-111. 8716
MgCH3COO+	7. 869e-113	6. 556e-108	0. 9160	-112. 1422
Aceti c_aci d(aq)	1. 532e-113	9. 195e-109	1. 0000	-112. 8148
NaCH3COO(aq)	1. 115e-113	9. 143e-109	1. 0000	-112. 9527
KCH3COO(aq)	1. 243e-114	1. 219e-109	1. 0000	-113. 9055
CuCH3COO+	2. 603e-115	3. 189e-110	0. 9160	-114. 6227
BaCH3COO+	8. 825e-116	1. 732e-110	0. 9160	-115. 0924
Mal onate	3. 471e-119	3. 540e-114	0. 7026	-118. 6128
H-Mal onate	2. 489e-121	2. 564e-116	0. 9160	-120. 6421
CuCH3COO(aq)	1. 682e-122	2. 061e-117	1. 0000	-121. 7743
Mal oni c_aci d(aq)	3. 348e-126	3. 482e-121	1. 0000	-125. 4752
S3O6--	3. 468e-139	6. 662e-134	0. 7026	-138. 6132
Acetal dehyde(aq)	6. 281e-148	2. 766e-143	1. 0000	-147. 2020
Ethyne(aq)	3. 534e-170	9. 197e-166	1. 0000	-169. 4518
Ethanol (aq)	4. 127e-171	1. 900e-166	1. 0000	-170. 3844
Ethyl ene(aq)	6. 606e-176	1. 852e-171	1. 0000	-175. 1801
Lactate	5. 606e-177	4. 991e-172	0. 9160	-176. 2894
Ca(Lac)+	1. 948e-178	2. 514e-173	0. 9160	-177. 7486
Mg(Lac)+	3. 018e-179	3. 420e-174	0. 9160	-178. 5584
Na(Lac) (aq)	6. 422e-180	7. 193e-175	1. 0000	-179. 1924
K(Lac) (aq)	1. 008e-180	1. 292e-175	1. 0000	-179. 9964
Lacti c_aci d(aq)	7. 685e-181	6. 920e-176	1. 0000	-180. 1143
Cu(Lac)+	2. 290e-181	3. 494e-176	0. 9160	-180. 6782
Ba(Lac)+	1. 662e-182	3. 761e-177	0. 9160	-181. 8175
Ethane(aq)	1. 216e-188	3. 654e-184	1. 0000	-187. 9152
Propanoate	1. 025e-191	7. 486e-187	0. 9160	-191. 0274
Ca(Prop)+	5. 821e-194	6. 584e-189	0. 9160	-193. 2731
Mg(Prop)+	1. 517e-194	1. 477e-189	0. 9160	-193. 8571
Propanoi c_aci d(a	1. 465e-194	1. 085e-189	1. 0000	-193. 8340
Na(Prop) (aq)	1. 158e-194	1. 112e-189	1. 0000	-193. 9364
Ca(Gl yc)2(aq)	2. 697e-195	5. 126e-190	1. 0000	-194. 5691
K(Prop) (aq)	1. 818e-195	2. 038e-190	1. 0000	-194. 7405
Cu(Prop)+	1. 815e-196	2. 478e-191	0. 9160	-195. 7792
Mg(Gl yc)2(aq)	1. 494e-196	2. 605e-191	1. 0000	-195. 8256
Cu(Gl yc)2(aq)	1. 637e-197	3. 496e-192	1. 0000	-196. 7859
Ba(Prop)+	8. 296e-198	1. 745e-192	0. 9160	-197. 1192

SpecE8_output_GSS_MW-95-8 July.txt				
Na(Gl yc)2-	3. 148e-198	5. 446e-193	0. 9160	-197. 5401
S406--	1. 366e-198	3. 063e-193	0. 7026	-198. 0178
Succinate	1. 017e-198	1. 180e-193	0. 7026	-198. 1460
K(Gl yc)2-	4. 247e-199	8. 032e-194	0. 9160	-198. 4100
Ba(Gl yc)2(aq)	7. 677e-200	2. 205e-194	1. 0000	-199. 1148
H-Succinate	6. 938e-201	8. 120e-196	0. 9160	-200. 1968
Succinic acid(aq)	2. 242e-204	2. 647e-199	1. 0000	-203. 6493
Ca(CH3COO)2(aq)	3. 678e-221	5. 815e-216	1. 0000	-220. 4344
Mg(CH3COO)2(aq)	8. 398e-222	1. 195e-216	1. 0000	-221. 0758
Cu(CH3COO)2(aq)	1. 184e-223	2. 149e-218	1. 0000	-222. 9267
Na(CH3COO)2-	9. 273e-224	1. 308e-218	0. 9160	-223. 0709
Acetone(aq)	4. 123e-224	2. 394e-219	1. 0000	-223. 3848
K(CH3COO)2-	6. 785e-225	1. 066e-219	0. 9160	-224. 2066
Ba(CH3COO)2(aq)	3. 491e-225	8. 914e-220	1. 0000	-224. 4570
Propanal (aq)	1. 657e-228	9. 617e-224	1. 0000	-227. 7808
Cu(CH3COO)2-	3. 444e-232	6. 252e-227	0. 9160	-231. 5011
1-Propyne(aq)	4. 296e-247	1. 721e-242	1. 0000	-246. 3669
1-Propanol (aq)	3. 356e-252	2. 016e-247	1. 0000	-251. 4741
S3--	1. 482e-252	1. 425e-247	0. 7026	-251. 9823
1-Propene(aq)	9. 668e-255	4. 067e-250	1. 0000	-254. 0146
2-Hydroxybutanoic acid	1. 631e-258	1. 681e-253	0. 9160	-257. 8256
2-Hydroxybutanoic acid	1. 980e-262	2. 060e-257	1. 0000	-261. 7034
Propane(aq)	3. 950e-270	1. 741e-265	1. 0000	-269. 4034
Butanoate	2. 356e-273	2. 051e-268	0. 9160	-272. 6660
Ca(But)+	9. 173e-276	1. 166e-270	0. 9160	-275. 0756
Butanoic acid(aq)	2. 694e-276	2. 373e-271	1. 0000	-275. 5695
Na(But) (aq)	2. 561e-276	2. 818e-271	1. 0000	-275. 5916
Mg(But)+	2. 279e-276	2. 538e-271	0. 9160	-275. 6804
K(But) (aq)	4. 021e-277	5. 072e-272	1. 0000	-276. 3957
Cu(But)+	3. 302e-278	4. 972e-273	0. 9160	-277. 5193
Gluconate	1. 963e-279	2. 553e-274	0. 7026	-278. 8603
Ba(But)+	1. 275e-279	2. 860e-274	0. 9160	-278. 9326
H-Gluconate	7. 797e-282	1. 022e-276	0. 9160	-281. 1462
Gluconic acid(aq)	3. 194e-285	4. 218e-280	1. 0000	-284. 4957
Ethyl acetate(aq)	1. 916e-285	1. 687e-280	1. 0000	-284. 7176
S506--	7. 947e-288	2. 036e-282	0. 7026	-287. 2530
Butanal (aq)	1. 314e-311	9. 473e-307	1. 0000	-300. 0000
Pimelate	0. 0000	0. 0000	0. 7026	-300. 0000
Dodecanoic acid	0. 0000	0. 0000	1. 0000	-300. 0000
2-Hydroxyhexanoic acid	0. 0000	0. 0000	1. 0000	-300. 0000
Dodecanoate	0. 0000	0. 0000	0. 9160	-300. 0000
Pentanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Decanoic acid(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Decanoate	0. 0000	0. 0000	0. 9160	-300. 0000
Decanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ba(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Hydroxyhexanoic acid	0. 0000	0. 0000	0. 9160	-300. 0000
1-Pentanol (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Octanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Nonanoic acid(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Nonanal (aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ba(Pent)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Hydroxyheptanoic acid	0. 0000	0. 0000	1. 0000	-300. 0000
Cu(Prop)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Ba(Pent)+	0. 0000	0. 0000	0. 9160	-300. 0000
Cu(Pent)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Na(Prop)2-	0. 0000	0. 0000	0. 9160	-300. 0000
Cu(Pent)+	0. 0000	0. 0000	0. 9160	-300. 0000
Ba(Lac)2(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
2-Hydroxyheptanoic acid	0. 0000	0. 0000	0. 9160	-300. 0000
1-Octyne(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
1-Heptene(aq)	0. 0000	0. 0000	1. 0000	-300. 0000
Na(Pent)2-	0. 0000	0. 0000	0. 9160	-300. 0000

SpecE8_output_GSS_MW-95-8 Jul y. txt

Na(Pent) (aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Lac)2-	0.0000	0.0000	0.9160	-300.0000
2-Hydroxydecanoic	0.0000	0.0000	1.0000	-300.0000
Na(But)2-	0.0000	0.0000	0.9160	-300.0000
Cu(CH3COO)3-	0.0000	0.0000	0.9160	-300.0000
2-Hydroxydecanoic	0.0000	0.0000	0.9160	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoate	0.0000	0.0000	0.9160	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)+	0.0000	0.0000	0.9160	-300.0000
Mg(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9160	-300.0000
Octanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octanol (aq)	0.0000	0.0000	1.0000	-300.0000
Pentanoate	0.0000	0.0000	0.9160	-300.0000
Pentanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol (aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol (aq)	0.0000	0.0000	1.0000	-300.0000
Picric acid(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butanol (aq)	0.0000	0.0000	1.0000	-300.0000
Azelate	0.0000	0.0000	0.7026	-300.0000
Azelaic acid(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hexanone(aq)	0.0000	0.0000	1.0000	-300.0000
K(Prop)2-	0.0000	0.0000	0.9160	-300.0000
Adipic acid(aq)	0.0000	0.0000	1.0000	-300.0000
K(Pent)2-	0.0000	0.0000	0.9160	-300.0000
K(Pent) (aq)	0.0000	0.0000	1.0000	-300.0000
K(Lac)2-	0.0000	0.0000	0.9160	-300.0000
Adipate	0.0000	0.0000	0.7026	-300.0000
2-Heptanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexyne(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
K(But)2-	0.0000	0.0000	0.9160	-300.0000
Ca(Pent)+	0.0000	0.0000	0.9160	-300.0000
2-Butanone(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9160	-300.0000
Ca(Lac)2(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanal (aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9160	-300.0000
p-Toluidic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanal (aq)	0.0000	0.0000	1.0000	-300.0000
p-Toluate	0.0000	0.0000	0.9160	-300.0000
o-Toluidic acid(aq)	0.0000	0.0000	1.0000	-300.0000
o-Toluate	0.0000	0.0000	0.9160	-300.0000
o-Phthalate	0.0000	0.0000	0.7026	-300.0000
1-Hexene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butyne(aq)	0.0000	0.0000	1.0000	-300.0000
n-Pentyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_MW-95-8 Jul y. txt

n-Butane(aq)	0.0000	0.0000	1.0000	-300.0000
m-Tol u i c_ a c i d(aq)	0.0000	0.0000	1.0000	-300.0000
Undecano i c_ a c i d(0.0000	0.0000	1.0000	-300.0000
Undecanoate	0.0000	0.0000	0.9160	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxypentano	0.0000	0.0000	1.0000	-300.0000
H-Suberate	0.0000	0.0000	0.9160	-300.0000
H-Sebacate	0.0000	0.0000	0.9160	-300.0000
H-Pi mel ate	0.0000	0.0000	0.9160	-300.0000
2-Hydroxypentano	0.0000	0.0000	0.9160	-300.0000
1-Hexanol (aq)	0.0000	0.0000	1.0000	-300.0000
H-Azel ate	0.0000	0.0000	0.9160	-300.0000
H-Adi pate	0.0000	0.0000	0.9160	-300.0000
2-Hydroxyoctanoi	0.0000	0.0000	1.0000	-300.0000
Tol uene(aq)	0.0000	0.0000	1.0000	-300.0000
Suberi c_ a c i d(aq)	0.0000	0.0000	1.0000	-300.0000
Suberate	0.0000	0.0000	0.7026	-300.0000
m-Tol uate	0.0000	0.0000	0.9160	-300.0000
Benzoi c_ a c i d(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoa	0.0000	0.0000	0.9160	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
Sebaci c_ a c i d(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacate	0.0000	0.0000	0.7026	-300.0000
Benzoate	0.0000	0.0000	0.9199	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
S5--	0.0000	0.0000	0.7026	-300.0000
Benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoi	0.0000	0.0000	1.0000	-300.0000
n-Propyl benzene(0.0000	0.0000	1.0000	-300.0000
S4--	0.0000	0.0000	0.7026	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoa	0.0000	0.0000	0.9160	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Butene(aq)	0.0000	0.0000	1.0000	-300.0000

Mi neral saturati on states

log Q/K

log Q/K

Wi theri te	2.0534s/sat	UO2S04: 2. 5H2O	-17.6267
Quartz	0.6298s/sat	UO2S04: 3H2O	-17.7038
Bar i te	0.4980s/sat	Di cal ci um_ si l i ca	-17.9168
Dol omi te-ord	0.4931s/sat	NaUO3	-17.9812
Dol omi te	0.4930s/sat	Anthophyl l i te	-18.3401
Tri dymi te	0.4295s/sat	UO2F2: 3H2O	-18.6222
Chal cedony	0.3445s/sat	Akermani te	-18.6315
Cal ci te	0.0480s/sat	UO2F2	-18.9482
Cri stoba l i te(al p	0.0451s/sat	CuCl 2	-19.2559
Tenori te	0.0429s/sat	Larni te	-19.2965
I ce	-0.0831	MgCl 2: 4H2O	-19.3904
Aragoni te	-0.0971	UO2S04	-21.4904
Coesi te	-0.2197	Li me	-21.8870
Mal achi te	-0.3882	Na2UO4(al pha)	-22.0171
Cri stoba l i te(bet	-0.4413	UO2(N03)2: 6H2O	-23.0428
Monohydrocal ci te	-0.7516	Hydrophi l i te	-23.4414
Si O2(am)	-0.7799	UO2. 3333(beta)	-23.4549
Dol omi te-di s	-1.1629	Ranki ni te	-24.6372
Magnesi te	-1.2770	UO2(N03)2: 3H2O	-24.7023
Di optase	-1.7122	UOFOH: . 5H2O	-24.8044
Tal c	-1.9364	UO2Cl	-25.1909
Gypsum	-2.0304	MgCl 2: 2H2O	-25.2366

SpecE8_output_GSS_MW-95-8 Jul y. txt

Anhydri te	-2. 3526	UOF0H	-25. 3152
Brochanti te	-2. 7568	Tobermori te-14A	-25. 7389
Bassani te	-3. 0021	UO2(N03)2: 2H2O	-26. 0724
CaS04: 0. 5H20(bet	-3. 1893	K2U04	-27. 3583
Fluori te	-3. 4288	Tobermori te-11A	-27. 9904
Enstati te	-4. 0046	UOF2: H2O	-28. 1991
Nesquehoni te	-4. 3929	MgCl 2: H2O	-28. 8244
CaU04	-4. 5015	UOF2	-28. 8530
Sellai te	-5. 0154	UO2Cl 2: 3H2O	-29. 2379
U03: 2H2O	-5. 1147	UO2(N03)2: H2O	-29. 9148
Schoepi te	-5. 1148	Foshagi te	-30. 3973
Chrysotil e	-5. 2194	Merwi ni te	-30. 6608
U02(OH)2(beta)	-5. 2817	Tobermori te-9A	-31. 8398
Di opsi de	-5. 3227	Ba2Si 04	-31. 9514
Schoepi te-dehy(.	-5. 3424	UO2Cl 2: H2O	-32. 2008
U03: . 9H20(al pha)	-5. 3496	Afwil li te	-32. 9119
Nahcol i te	-5. 3592	UO2(N03)2	-33. 6054
Schoepi te-dehy(.	-5. 4351	KMgCl 3: 2H2O	-34. 0336
Schoepi te-dehy(1	-5. 4519	Chl oromagnesi te	-34. 9231
Hunti te	-5. 4630	UO2Cl 2	-36. 3164
Okeni te	-5. 4635	U02S03	-37. 3609
Bruci te	-5. 5692	UF4: 2. 5H2O	-39. 0437
Woll astoni te	-5. 6731	BaSi F6	-39. 5187
Atacami te	-5. 7339	BaO	-41. 0286
Pseudowoll astoni	-5. 9393	KMgCl 3	-41. 8165
Ni ter	-6. 0451	UF4	-43. 5967
Azuri te	-6. 0656	Hatruri te	-43. 6729
U02C03	-6. 1975	Xonotli te	-43. 7243
Rutherfordi ne	-6. 2109	U(S04)2: 8H2O	-46. 2827
Schoepi te-dehy(.	-6. 6256	U(S04)2: 4H2O	-47. 6144
Sepi ol i te	-6. 9297	U(S04)2	-47. 8773
Cupri te	-7. 0265	Ba2U207	-47. 8784
Schoepi te-dehy(.	-7. 1768	C	-49. 1926
Arti ni te	-7. 6853	U(C03)2	-50. 3008
Mi rabil i te	-7. 7511	UOCl 2	-50. 6764
Sanborni te	-7. 8617	UCI F3	-54. 4591
Chal canthi te	-8. 0182	UOF4	-57. 8305
U03(gamma)	-8. 2611	Na4Si 04	-58. 4929
Syl vi te	-8. 4051	Na3U04	-58. 6604
Hal i te	-8. 5299	Na	-58. 9220
U03(beta)	-8. 8963	Na2O	-61. 2262
Arcani te	-9. 0054	U5012Cl	-62. 1767
Ni trobari te	-9. 0125	UOCl 3	-62. 4290
Nantoki te	-9. 1361	(U02)2Cl 3	-62. 7097
U03(al pha)	-9. 2512	K	-63. 3755
Thenardi te	-9. 2538	UF3	-65. 9835
Gyrol i te	-9. 7760	UF5(beta)	-66. 1695
Forsteri te	-9. 9137	UF5(al pha)	-66. 5756
BaU04	-10. 0062	UCI 2F2	-67. 6359
Natron	-10. 0984	Anti gori te	-68. 9036
U02F0H: 2H2O	-10. 3916	UOCl	-69. 4584
Mg1. 25S04(OH)0. 5	-10. 4107	Ba3U06	-71. 6243
Tremol i te	-10. 6614	S	-74. 3666
Na2C03: 7H2O	-10. 6626	Covel li te	-75. 3824
Cu	-10. 7378	U203F6	-76. 9268
Monti cel li te	-10. 8371	Chal coci te	-79. 9986
U02F0H: H2O	-10. 8505	K2O	-80. 0044
Peri cl ase	-10. 9572	UCI 3F	-80. 6284
Portl andi te	-11. 2761	Na6Si 207	-85. 0470
U02F0H	-11. 3786	U(S03)2	-90. 7931
Mg1. 5S04(OH)	-11. 7302	UCI 4	-91. 4600
Thermonatri te	-12. 1371	UCI 3	-95. 6084
Na2C03	-12. 4651	UF6	-97. 7257

SpecE8_output_GSS_MW-95-8 July.txt

UO2. 25	-12. 7179	Mg	-103. 1465
UO2. 25(beta)	-12. 7866	U3O5F8	-103. 2144
MgUO4	-12. 8880	U2F9	-106. 4670
MgSO4	-12. 9041	UCI 5	-111. 8288
UO2SO4: H2O	-13. 0997	U2O2CI 5	-114. 3955
UO2CI OH: 2H2O	-14. 1266	BaS	-115. 2994
BaU2O7	-14. 1467	Ca	-120. 4510
BaCl 2: 2H2O	-14. 2263	Ba	-125. 0194
Chal cocyani te	-14. 2533	UCI 6	-130. 9834
Na2U2O7	-14. 3911	Si	-131. 5851
BaCl 2: H2O	-14. 9627	U	-181. 0049
Ba(OH)2: 8H2O	-15. 8132	U4F17	-192. 7485
Hydromagnesi te	-16. 1727	US	-196. 2116
UO2. 6667	-16. 3275	UC	-211. 9779
Urani ni te	-16. 4380	UH3(beta)	-213. 7985
BaCl 2	-16. 5162	US1. 9	-228. 4001
Natrosi l i te	-16. 5418	US2	-234. 0294
Coffi ni te	-16. 6493	UC1. 94(al pha)	-260. 3479
Ba2Si 308	-16. 7335	US3	-305. 0657
MgOHCl	-16. 9336	U2S3	-422. 9679
Na2Si O3	-17. 2617	o-Phthal i c_aci d	-428. 4317
Hi l l ebrandi te	-17. 4350	U2C3	-474. 7483
UO2SO4: 3. 5H2O	-17. 5636	U3S5	-652. 1820

Gases	fugaci ty	l og fug.
H2O(g)	0. 01025	-1. 989
CO2(g)	0. 002470	-2. 607
NO2(g)	4. 631e-013	-12. 334
HF(g)	2. 007e-015	-14. 697
HCl (g)	2. 134e-019	-18. 671
O2(g)	8. 152e-027	-26. 089
H2(g)	7. 670e-032	-31. 115
CO(g)	6. 332e-038	-37. 198
Si F4(g)	1. 386e-041	-40. 858
Cl 2(g)	5. 082e-042	-41. 294
SO2(g)	6. 926e-046	-45. 159
UO2F2(g)	2. 572e-064	-63. 590
Cu(g)	1. 109e-066	-65. 955
Na(g)	4. 189e-074	-73. 378
UO2Cl 2(g)	2. 403e-075	-74. 619
K(g)	1. 675e-075	-74. 776
UO3(g)	4. 240e-076	-75. 373
UOF4(g)	9. 665e-080	-79. 015
UF5(g)	4. 998e-088	-87. 301
UF4(g)	5. 652e-091	-90. 248
UF6(g)	9. 998e-100	-99. 000
H2S(g)	3. 847e-100	-99. 415
F2(g)	1. 819e-100	-99. 740
CH4(g)	1. 387e-102	-101. 858
UCI 4(g)	1. 943e-118	-117. 711
UO2(g)	3. 458e-119	-118. 461
Mg(g)	6. 551e-125	-124. 184
UCI 5(g)	3. 363e-131	-130. 473
UF3(g)	1. 972e-137	-136. 705
UCI 6(g)	2. 538e-138	-137. 596
UCI 3(g)	6. 131e-145	-144. 212
Ca(g)	5. 213e-148	-147. 283
U2F10(g)	1. 902e-148	-147. 721
S2(g)	1. 603e-164	-163. 795
C2H4(g)	9. 370e-174	-173. 028
C(g)	4. 957e-174	-173. 305
UF2(g)	1. 640e-180	-179. 785

SpecE8_output_GSS_MW-95-8 Jul y. txt

UCI 2(g)	1. 487e-189	-188. 828
UO(g)	2. 213e-195	-194. 655
Si (g)	1. 914e-207	-206. 718
UF(g)	8. 846e-217	-216. 053
U2Cl 8(g)	1. 007e-224	-223. 997
UCI (g)	1. 277e-231	-230. 894
U2Cl 10(g)	1. 590e-233	-232. 799
U(g)	3. 754e-272	-271. 426

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
Ba++	9. 61e-007	9. 61e-007	0. 132			
Ca++	0. 00166	0. 00166	66. 5			
Cl -	0. 000110	0. 000110	3. 91			
Cu++	1. 08e-007	1. 08e-007	0. 00684			
F-	5. 15e-006	5. 15e-006	0. 0977			
H+	0. 000121	0. 000121	0. 122			
H2O	55. 5	55. 5	1. 00e+006			
HC03-	0. 00271	0. 00271	165.			
K+	0. 000200	0. 000200	7. 82			
Mg++	0. 000322	0. 000322	7. 82			
N03-	0. 00154	0. 00154	95. 6			
Na+	0. 00111	0. 00111	25. 4			
O2(aq)	-5. 21e-009	-5. 21e-009	-0. 000167			
S04--	0. 000417	0. 000417	40. 1			
Si O2(aq)	0. 000190	0. 000190	11. 4			
U02++	2. 05e-009	2. 05e-009	0. 000554			

Elemental composition	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg
Bari um	9. 613e-007	9. 613e-007	0. 1320		
Cal ci um	0. 001659	0. 001659	66. 47		
Carbon	0. 002708	0. 002708	32. 52		
Chl ori ne	0. 0001103	0. 0001103	3. 910		
Copper	1. 077e-007	1. 077e-007	0. 006842		
Fl uori ne	5. 147e-006	5. 147e-006	0. 09775		
Hydrogen	111. 0	111. 0	1. 118e+005		
Magnesi um	0. 0003219	0. 0003219	7. 820		
Ni trogen	0. 001543	0. 001543	21. 60		
Oxygen	55. 52	55. 52	8. 880e+005		
Potassi um	0. 0002001	0. 0002001	7. 820		
Si li con	0. 0001904	0. 0001904	5. 346		
Sodi um	0. 001106	0. 001106	25. 41		
Sul fur	0. 0004174	0. 0004174	13. 38		
Urani um	2. 054e-009	2. 054e-009	0. 0004887		

SpecE8_output_GSS_MW-95-8 June.txt

Temperature = 10.2 C Pressure = 1.013 bars
 pH = 7.600 log fO2 = -32.704
 Eh = 0.3450 volts pe = 6.1367
 Ionic strength = 0.007230
 Charge imbalance = 0.000233 eq/kg (4.652% error)
 Activity of water = 0.999996
 Solvent mass = 1.000000 kg
 Solution mass = 1.000427 kg
 Solution density = 1.023 g/cm3
 Chlorinity = 0.000110 molal
 Dissolved solids = 427 mg/kg sol'n
 Hardness = 198.18 mg/kg sol'n as CaCO3
 carbonate = 122.62 mg/kg sol'n as CaCO3
 non-carbonate = 75.55 mg/kg sol'n as CaCO3
 Rock mass = 0.000000 kg
 Carbonate alkalinity = 122.62 mg/kg sol'n as CaCO3
 Water type = Ca-HCO3

Nernst redox couples

Eh (volts)

pe

e- + .25*O2(aq) + H+ = .5*H2O	0.3450	6.1367
8*e- + 9*H+ + NO3- = 3*H2O + NH3(aq)	0.5698	10.1349

No minerals in system.

Aqueous species

molality

mg/kg sol'n

act. coef.

log act.

HCO3-	0.002399	146.3	0.9163	-2.6580
NO3-	0.001711	106.0	0.9142	-2.8057
Ca++	0.001591	63.73	0.7155	-2.9438
Na+	0.001059	24.34	0.9163	-3.0130
SO4--	0.0003256	31.26	0.7036	-3.6400
Mg++	0.0003063	7.440	0.7265	-3.6527
K+	0.0002247	8.781	0.9142	-3.6874
SiO2(aq)	0.0001801	10.81	1.0000	-3.7446
CO2(aq)	0.0001618	7.120	1.0000	-3.7909
Cl-	0.0001103	3.907	0.9142	-3.9965
CaHCO3+	3.164e-005	3.197	0.9163	-4.5377
CaSO4(aq)	3.144e-005	4.278	1.0000	-4.5026
N2(aq)	2.846e-005	0.7971	1.0000	-4.5457
MgSO4(aq)	9.020e-006	1.085	1.0000	-5.0448
MgHCO3+	5.996e-006	0.5114	0.9163	-5.2601
CaCO3(aq)	5.138e-006	0.5140	1.0000	-5.2892
F-	5.094e-006	0.09673	0.9153	-5.3314
NaHCO3(aq)	4.077e-006	0.3423	1.0000	-5.3897
CO3--	4.015e-006	0.2408	0.7067	-5.5471
Ba++	1.060e-006	0.1455	0.7097	-6.1236
HSiO3-	5.379e-007	0.04145	0.9163	-6.3073
MgCO3(aq)	5.220e-007	0.04399	1.0000	-6.2824
KS04-	3.859e-007	0.05214	0.9163	-6.4514
Cu++	1.692e-007	0.01075	0.7155	-6.9169
OH-	1.286e-007	0.002186	0.9153	-6.9292
NaHSiO3(aq)	2.978e-008	0.002979	1.0000	-7.5261
H+	2.715e-008	2.735e-005	0.9253	-7.6000
CaF+	2.627e-008	0.001551	0.9163	-7.6185
MgF+	2.607e-008	0.001128	0.9163	-7.6218
Cd++	2.574e-008	0.002893	0.7097	-7.7383
CaCl+	2.548e-008	0.001924	0.9163	-7.6317
MgCl+	1.958e-008	0.001170	0.9163	-7.7461
NaCl(aq)	1.512e-008	0.0008832	1.0000	-7.8205
NaCO3-	1.486e-008	0.001233	0.9163	-7.8660
UO2(CO3)3----	1.139e-009	0.0005126	0.2446	-9.5549

	SpecE8_output_GSS_MW-95-8	June.txt	
UO2(CO3)2--	8.981e-010	0.0003502	0.7036 -9.1993
BaCO3(aq)	6.917e-010	0.0001364	1.0000 -9.1601
KCl(aq)	4.961e-010	3.697e-005	1.0000 -9.3044
HSO4-	4.152e-010	4.029e-005	0.9163 -9.4197
NaF(aq)	4.033e-010	1.693e-005	1.0000 -9.3944
HF(aq)	1.334e-010	2.668e-006	1.0000 -9.8747
Cu+	3.075e-011	1.953e-006	0.9163 -10.5500
BaCl +	2.170e-011	3.748e-006	0.9163 -10.7015
NaOH(aq)	2.013e-011	8.048e-007	1.0000 -10.6962
UO2CO3(aq)	1.625e-011	5.361e-006	1.0000 -10.7891
CdSO4(aq)	4.228e-012	8.811e-007	1.0000 -11.3738
CaCl 2(aq)	3.265e-012	3.622e-007	1.0000 -11.4861
BaF+	2.176e-012	3.400e-007	0.9163 -11.7003
HNO3(aq)	1.418e-012	8.934e-008	1.0000 -11.8482
HCl(aq)	5.383e-013	1.962e-008	1.0000 -12.2689
CdCl 2(aq)	3.700e-013	6.779e-008	1.0000 -12.4318
UO2OH+	1.346e-013	3.861e-008	0.9163 -12.9090
KHSO4(aq)	4.519e-015	6.151e-010	1.0000 -14.3449
UO2++	1.787e-015	4.823e-010	0.7067 -14.8987
UO2F+	7.147e-016	2.065e-010	0.9163 -15.1838
UO2SO4(aq)	2.345e-016	8.580e-011	1.0000 -15.6299
HF2-	1.391e-016	5.422e-012	0.9163 -15.8948
UO2F2(aq)	9.611e-018	2.959e-012	1.0000 -17.0172
CdCl 3-	7.772e-018	1.700e-012	0.9163 -17.1474
UO2(SO4)2--	4.682e-019	2.163e-013	0.7036 -18.4822
UO2Cl +	1.730e-019	5.282e-014	0.9163 -18.7999
UO2+	2.311e-020	6.237e-015	0.9163 -19.6742
UO2F3-	8.388e-021	2.742e-015	0.9163 -20.1143
(UO2)2(OH)2++	3.613e-021	2.073e-015	0.7067 -20.5929
(UO2)3(CO3)6(6-)	1.167e-022	1.365e-016	0.0420 -23.3094
U(OH)4(aq)	3.906e-023	1.195e-017	1.0000 -22.4082
(UO2)3(OH)5+	6.709e-024	6.003e-018	0.9163 -23.2113
UO2Cl 2(aq)	7.137e-025	2.432e-019	1.0000 -24.1465
UO2F4--	3.139e-025	1.086e-019	0.7036 -24.6559
Formate	8.728e-031	3.927e-026	0.9153 -30.0975
H2(aq)	1.304e-031	2.628e-028	1.0000 -30.8846
Ca(For) +	2.978e-032	2.533e-027	0.9163 -31.5641
Mg(For) +	7.067e-033	4.897e-028	0.9163 -32.1887
Na(For) (aq)	9.210e-034	6.261e-029	1.0000 -33.0357
NO2-	7.481e-034	3.440e-029	0.9142 -33.1650
K(For) (aq)	1.691e-034	1.421e-029	1.0000 -33.7720
Formic acid(aq)	1.183e-034	5.441e-030	1.0000 -33.9271
Ba(For) +	1.490e-035	2.716e-030	0.9163 -34.8647
SO3--	1.466e-035	1.173e-030	0.7067 -34.9847
Cu(For) +	1.320e-035	1.433e-030	0.9163 -34.9172
NH4+	1.269e-035	2.287e-031	0.9132 -34.9361
HSO3-	4.024e-036	3.261e-031	0.9163 -35.4333
O2(aq)	3.378e-036	1.080e-031	1.0000 -35.4713
Cd(For) +	1.349e-036	2.123e-031	0.9163 -35.9080
CO(aq)	1.927e-037	5.396e-033	1.0000 -36.7151
NH3(aq)	8.861e-038	1.509e-033	1.0000 -37.0525
Oxalate	4.842e-038	4.260e-033	0.7036 -37.4676
HNO2(aq)	4.147e-038	1.949e-033	1.0000 -37.3822
ClO-	3.814e-038	1.962e-033	0.9163 -37.4566
SiF6--	1.141e-039	1.620e-034	0.7036 -39.0956
HO2-	2.296e-040	7.574e-036	0.9163 -39.6771
U(CO3)5(6-)	2.111e-040	1.135e-034	0.0420 -41.0519
UOH+++	2.555e-041	6.514e-036	0.4619 -40.9281
H-Oxalate	1.540e-041	1.370e-036	0.9163 -40.8505
SO2(aq)	4.985e-042	3.192e-037	1.0000 -41.3024
UF2++	1.414e-042	3.902e-037	0.7067 -42.0003
UF3+	1.026e-042	3.026e-037	0.9163 -42.0267
UF+++	5.810e-044	1.493e-038	0.4619 -43.5713

	SpecE8_output_GSS_MW-95-8 June.txt			
UF4(aq)	4. 625e-044	1. 452e-038	1. 0000	-43. 3349
USO4++	2. 629e-045	8. 778e-040	0. 7067	-44. 7311
U(SO4)2(aq)	1. 836e-045	7. 893e-040	1. 0000	-44. 7362
U++++	1. 112e-047	2. 647e-042	0. 2574	-47. 5431
Oxal i c_aci d(aq)	6. 434e-048	5. 790e-043	1. 0000	-47. 1915
HSO5-	3. 293e-048	3. 722e-043	0. 9163	-47. 5204
UCl +++	4. 892e-050	1. 337e-044	0. 4619	-49. 6460
S2O6--	5. 803e-060	9. 288e-055	0. 7036	-59. 3890
Ca(For)2(aq)	1. 935e-061	2. 517e-056	1. 0000	-60. 7132
Mg(For)2(aq)	5. 730e-062	6. 549e-057	1. 0000	-61. 2418
Na(For)2-	5. 813e-064	6. 567e-059	0. 9163	-63. 2736
Cu(For)2(aq)	2. 856e-064	4. 384e-059	1. 0000	-63. 5443
Cl O2-	2. 527e-064	1. 704e-059	0. 9163	-63. 6354
U+++	1. 531e-064	3. 642e-059	0. 4619	-64. 1505
Cd(For)2(aq)	1. 429e-064	2. 891e-059	1. 0000	-63. 8451
K(For)2-	8. 924e-065	1. 152e-059	0. 9163	-64. 0874
Ba(For)2(aq)	7. 297e-065	1. 658e-059	1. 0000	-64. 1369
S2O8--	4. 888e-065	9. 387e-060	0. 7036	-64. 4636
Formal dehyde(aq)	5. 951e-067	1. 786e-062	1. 0000	-66. 2254
N3-	9. 159e-073	3. 847e-068	0. 9163	-72. 0761
Cu(NH3)2++	9. 655e-074	9. 420e-069	0. 7067	-73. 1660
HN3(aq)	1. 500e-075	6. 453e-071	1. 0000	-74. 8238
Cl O3-	6. 307e-076	5. 261e-071	0. 9153	-75. 2386
S2O5--	3. 094e-076	4. 458e-071	0. 7036	-75. 6621
Cd(NH3)2++	2. 678e-077	3. 921e-072	0. 7067	-76. 7230
Urea(aq)	4. 169e-078	2. 503e-073	1. 0000	-77. 3799
HCN(aq)	7. 657e-080	2. 068e-075	1. 0000	-79. 1160
CN-	7. 176e-082	1. 866e-077	0. 9142	-81. 1831
Methanol (aq)	3. 532e-082	1. 131e-077	1. 0000	-81. 4519
HS-	3. 163e-087	1. 046e-082	0. 9153	-86. 5383
H2S(aq)	1. 216e-087	4. 142e-083	1. 0000	-86. 9151
Gl ycol ate	4. 197e-088	3. 148e-083	0. 9163	-87. 4150
Ca(Gl yc)+	2. 498e-089	2. 875e-084	0. 9163	-88. 6403
UO2Cl O3+	2. 717e-090	9. 599e-085	0. 9163	-89. 6039
Mg(Gl yc)+	2. 263e-090	2. 247e-085	0. 9163	-89. 6833
Na(Gl yc)(aq)	4. 534e-091	4. 443e-086	1. 0000	-90. 3435
K(Gl yc)(aq)	8. 323e-092	9. 496e-087	1. 0000	-91. 0797
Gl ycol i c_aci d(aq)	6. 913e-092	5. 255e-087	1. 0000	-91. 1604
Cl O4-	6. 060e-092	6. 024e-087	0. 9153	-91. 2560
Cu(Gl yc)+	5. 904e-092	8. 178e-087	0. 9163	-91. 2668
Methane(aq)	4. 747e-092	7. 612e-088	1. 0000	-91. 3236
S--	6. 716e-093	2. 153e-088	0. 7097	-92. 3218
Ba(Gl yc)+	3. 334e-093	7. 077e-088	0. 9163	-92. 5150
S2O3--	1. 929e-093	2. 162e-088	0. 7036	-92. 8673
Cd(Gl yc)+	7. 144e-094	1. 339e-088	0. 9163	-93. 1840
S2O4--	3. 101e-095	3. 971e-090	0. 7097	-94. 6575
Acetate	2. 389e-097	1. 410e-092	0. 9173	-96. 6592
CaCH3COO+	2. 452e-099	2. 429e-094	0. 9163	-98. 6485
MgCH3COO+	1. 318e-099	1. 098e-094	0. 9163	-98. 9180
Aceti c_aci d(aq)	3. 204e-100	1. 923e-095	1. 0000	-99. 4943
NaCH3COO(aq)	1. 783e-100	1. 462e-095	1. 0000	-99. 7489
KCH3COO(aq)	2. 323e-101	2. 279e-096	1. 0000	-100. 6339
CuCH3COO+	6. 817e-102	8. 354e-097	0. 9163	-101. 2043
BaCH3COO+	1. 619e-102	3. 177e-097	0. 9163	-101. 8288
CdCH3COO+	4. 679e-103	8. 019e-098	0. 9163	-102. 3678
Mal onate	5. 470e-106	5. 580e-101	0. 7036	-105. 4147
CuCH3COO(aq)	1. 534e-107	1. 880e-102	1. 0000	-106. 8142
Cu(NH3)3++	8. 886e-108	1. 018e-102	0. 7067	-107. 2021
H-Mal onate	4. 943e-108	5. 092e-103	0. 9163	-107. 3440
Mal oni c_aci d(aq)	8. 376e-113	8. 712e-108	1. 0000	-112. 0770
Methanami ne(aq)	1. 256e-116	3. 901e-112	1. 0000	-115. 9009
Gl yci ne(aq)	3. 974e-119	2. 982e-114	1. 0000	-118. 4007
Cu(Gl y)+	9. 063e-120	1. 247e-114	0. 9163	-119. 0807

	SpecE8_output_GSS_MW-95-8 June.txt			
Mg(Gl y) +	8. 674e-122	8. 529e-117	0. 9163	-121. 0997
Ca(Gl y) +	2. 252e-123	2. 569e-118	0. 9163	-122. 6854
Cd(Gl y) +	1. 252e-124	2. 333e-119	0. 9163	-123. 9404
S306--	8. 433e-126	1. 620e-120	0. 7036	-125. 2267
Ba(Gl y) +	1. 785e-126	3. 771e-121	0. 9163	-125. 7864
Acetal dehyde(aq)	2. 541e-131	1. 119e-126	1. 0000	-130. 5950
SCN-	1. 397e-131	8. 112e-127	0. 9153	-130. 8932
NH4CH3COO(aq)	4. 155e-132	3. 202e-127	1. 0000	-131. 3814
Acetami de(aq)	3. 424e-132	2. 021e-127	1. 0000	-131. 4655
UO2SCN+	4. 306e-145	1. 412e-139	0. 9163	-144. 4039
Cd(NH3) 4++	8. 306e-149	1. 499e-143	0. 7067	-148. 2314
Ethanol (aq)	3. 267e-151	1. 505e-146	1. 0000	-150. 4858
Ethyne(aq)	1. 395e-153	3. 632e-149	1. 0000	-152. 8553
Ethyl ene(aq)	5. 197e-156	1. 457e-151	1. 0000	-155. 2843
Lactate	4. 225e-157	3. 762e-152	0. 9163	-156. 4121
Ca(Lac) +	1. 478e-158	1. 908e-153	0. 9163	-157. 8682
Mg(Lac) +	2. 290e-159	2. 595e-154	0. 9163	-158. 6781
Na(Lac) (aq)	4. 662e-160	5. 222e-155	1. 0000	-159. 3315
K(Lac) (aq)	8. 558e-161	1. 096e-155	1. 0000	-160. 0676
Lacti c_aci d(aq)	7. 297e-161	6. 571e-156	1. 0000	-160. 1368
Cu(Lac) +	2. 724e-161	4. 156e-156	0. 9163	-160. 6027
Ba(Lac) +	1. 385e-162	3. 133e-157	0. 9163	-161. 8966
Cd(Lac) +	4. 864e-163	9. 796e-158	0. 9163	-162. 3510
Ethane(aq)	1. 892e-165	5. 688e-161	1. 0000	-164. 7230
Propanoate	1. 521e-168	1. 111e-163	0. 9163	-167. 8557
Ca(Prop) +	8. 697e-171	9. 836e-166	0. 9163	-170. 0986
Propanoi c_aci d(a	2. 740e-171	2. 029e-166	1. 0000	-170. 5622
Mg(Prop) +	2. 272e-171	2. 211e-166	0. 9163	-170. 6816
Na(Prop) (aq)	1. 655e-171	1. 590e-166	1. 0000	-170. 7811
K(Prop) (aq)	3. 039e-172	3. 407e-167	1. 0000	-171. 5173
Cu(Prop) +	4. 251e-173	5. 805e-168	0. 9163	-172. 4095
Cd(Prop) +	2. 736e-174	5. 073e-169	0. 9163	-173. 6008
Ba(Prop) +	1. 360e-174	2. 859e-169	0. 9163	-173. 9046
S406--	3. 508e-175	7. 863e-170	0. 7036	-174. 6076
Ca(Gl yc) 2(aq)	1. 928e-175	3. 664e-170	1. 0000	-174. 7150
Succi nate	1. 435e-175	1. 665e-170	0. 7036	-174. 9959
USCN+++	1. 365e-175	4. 039e-170	0. 4619	-175. 2004
Mg(Gl yc) 2(aq)	1. 070e-176	1. 865e-171	1. 0000	-175. 9708
Cu(Gl yc) 2(aq)	1. 837e-177	3. 922e-172	1. 0000	-176. 7359
H-Succi nate	1. 234e-177	1. 445e-172	0. 9163	-176. 9465
Na(Gl yc) 2-	2. 152e-178	3. 722e-173	0. 9163	-177. 7052
K(Gl yc) 2-	3. 391e-179	6. 413e-174	0. 9163	-178. 5076
Seri ne(aq)	1. 890e-179	1. 985e-174	1. 0000	-178. 7236
Cd(Gl yc) 2(aq)	8. 971e-180	2. 354e-174	1. 0000	-179. 0472
Ba(Gl yc) 2(aq)	6. 015e-180	1. 728e-174	1. 0000	-179. 2208
Succi ni c_aci d(aq)	5. 028e-181	5. 935e-176	1. 0000	-180. 2986
Ethanami ne(aq)	1. 291e-187	5. 818e-183	1. 0000	-186. 8891
Al ani ne(aq)	5. 444e-189	4. 848e-184	1. 0000	-188. 2641
Cu(Al a) +	2. 810e-190	4. 260e-185	0. 9163	-189. 5892
Mg(Al a) +	8. 022e-194	9. 012e-189	0. 9163	-193. 1337
Ca(Al a) +	5. 906e-194	7. 566e-189	0. 9163	-193. 2667
Ca(CH3COO) 2(aq)	1. 024e-194	1. 618e-189	1. 0000	-193. 9899
Cd(Al a) +	9. 651e-195	1. 934e-189	0. 9163	-194. 0534
Mg(CH3COO) 2(aq)	2. 346e-195	3. 339e-190	1. 0000	-194. 6296
Cu(CH3COO) 2(aq)	5. 171e-197	9. 388e-192	1. 0000	-196. 2864
Na(CH3COO) 2-	2. 469e-197	3. 481e-192	0. 9163	-196. 6455
Ba(Al a) +	1. 731e-197	3. 901e-192	0. 9163	-196. 7996
Acetone(aq)	1. 497e-197	8. 689e-193	1. 0000	-196. 8249
Cd(CH3COO) 2(aq)	2. 179e-198	5. 021e-193	1. 0000	-197. 6617
K(CH3COO) 2-	2. 109e-198	3. 314e-193	0. 9163	-197. 7138
Ba(CH3COO) 2(aq)	1. 064e-198	2. 717e-193	1. 0000	-197. 9729
Asparti c_aci d(aq)	2. 511e-199	3. 341e-194	1. 0000	-198. 6001
Propanal (aq)	5. 993e-202	3. 479e-197	1. 0000	-201. 2223

SpecE8_output_GSS_MW-95-8 June.txt				
Cu(CH3COO)2-	5. 235e-204	9. 505e-199	0. 9163	-203. 3190
S3--	2. 047e-219	1. 968e-214	0. 7036	-218. 8417
1-Propyne(aq)	1. 523e-220	6. 098e-216	1. 0000	-219. 8174
1-Propanol (aq)	2. 376e-222	1. 427e-217	1. 0000	-221. 6242
1-Propene(aq)	6. 811e-225	2. 865e-220	1. 0000	-224. 1668
2-Hydroxybutanoa	1. 099e-228	1. 132e-223	0. 9163	-227. 9971
NH4(CH3COO)2-	9. 550e-229	1. 299e-223	0. 9163	-228. 0580
Asparagi ne(aq)	3. 189e-231	4. 211e-226	1. 0000	-230. 4964
2-Hydroxybutanoi	1. 679e-232	1. 748e-227	1. 0000	-231. 7749
Cu(Gl y)2(aq)	3. 139e-233	6. 642e-228	1. 0000	-232. 5031
Propane(aq)	5. 495e-237	2. 422e-232	1. 0000	-236. 2601
Mg(Gl y)2(aq)	1. 132e-239	1. 951e-234	1. 0000	-238. 9462
Di gl yci ne(aq)	3. 516e-240	4. 643e-235	1. 0000	-239. 4540
Butanoate	3. 125e-240	2. 721e-235	0. 9163	-239. 5431
Cd(Gl y)2(aq)	8. 012e-242	2. 086e-236	1. 0000	-241. 0963
Di ketopi perazi ne	1. 347e-242	1. 536e-237	1. 0000	-241. 8707
Ca(But)+	1. 224e-242	1. 557e-237	0. 9163	-241. 9500
Butanoi c_aci d(aq)	4. 501e-243	3. 964e-238	1. 0000	-242. 3467
Na(But) (aq)	3. 272e-243	3. 601e-238	1. 0000	-242. 4852
Mg(But)+	3. 049e-243	3. 395e-238	0. 9163	-242. 5538
Ca(Gl y)2(aq)	1. 011e-243	1. 901e-238	1. 0000	-242. 9954
K(But) (aq)	6. 007e-244	7. 577e-239	1. 0000	-243. 2214
Cu(But)+	6. 911e-245	1. 041e-239	0. 9163	-244. 1984
Gl utarate	2. 477e-246	3. 221e-241	0. 7036	-245. 7588
Cd(But)+	2. 376e-246	4. 739e-241	0. 9163	-245. 6620
Ba(But)+	1. 867e-246	4. 188e-241	0. 9163	-245. 7668
Ba(Gl y)2(aq)	8. 891e-247	2. 537e-241	1. 0000	-246. 0510
H-Gl utarate	1. 240e-248	1. 625e-243	0. 9163	-247. 9447
Threoni ne(aq)	3. 979e-251	4. 737e-246	1. 0000	-250. 4003
Gl utari c_aci d(aq)	6. 397e-252	8. 448e-247	1. 0000	-251. 1940
Ethyl acetate(aq)	3. 175e-252	2. 796e-247	1. 0000	-251. 4983
S506--	2. 115e-254	5. 418e-249	0. 7036	-253. 8274
1-Propanami ne(aq)	3. 166e-258	1. 871e-253	1. 0000	-257. 4995
a-Ami nobutyri c_a	2. 412e-260	2. 486e-255	1. 0000	-259. 6177
Gl utami c_aci d(aq)	5. 003e-269	7. 359e-264	1. 0000	-268. 3007
Butanal (aq)	4. 242e-275	3. 057e-270	1. 0000	-274. 3724
UO2(SCN)2(aq)	2. 967e-276	1. 145e-270	1. 0000	-275. 5277
S4--	6. 298e-283	8. 075e-278	0. 7036	-282. 3535
Cu(CH3COO)3-	4. 498e-292	1. 082e-286	0. 9163	-291. 3850
1-Butyne(aq)	2. 941e-292	1. 590e-287	1. 0000	-291. 5315
Cd(CH3COO)3-	1. 594e-294	4. 614e-289	0. 9163	-293. 8354
1-Butanol (aq)	9. 202e-295	6. 818e-290	1. 0000	-294. 0361
1-Butene(aq)	8. 438e-297	4. 732e-292	1. 0000	-296. 0738
2-Hydroxypentano	4. 298e-300	5. 032e-295	0. 9163	-299. 4047
Gl utami ne(aq)	7. 293e-303	1. 065e-297	1. 0000	-300. 0000
2-Hydroxypentano	3. 865e-304	4. 564e-299	1. 0000	-300. 0000
U(SCN)2++	1. 914e-305	6. 776e-300	0. 7067	-300. 0000
n-Butane(aq)	1. 392e-308	8. 090e-304	1. 0000	-300. 0000
Al anyl gl yci ne(aq)	2. 416e-310	3. 530e-305	1. 0000	-300. 0000
Pentanoate	7. 636e-312	7. 718e-307	0. 9163	-300. 0000
Ca(Lac)2(aq)	7. 350e-314	1. 603e-308	1. 0000	-300. 0000
Ca(Pent)+	1. 775e-314	2. 505e-309	0. 9163	-300. 0000
Pentanoi c_aci d(a	1. 196e-314	1. 221e-309	1. 0000	-300. 0000
Mg(Lac)2(aq)	1. 174e-314	2. 377e-309	1. 0000	-300. 0000
Na(Pent) (aq)	8. 492e-315	1. 054e-309	1. 0000	-300. 0000
Mg(Pent)+	4. 318e-315	5. 414e-310	0. 9163	-300. 0000
K(Pent) (aq)	1. 559e-315	2. 185e-310	1. 0000	-300. 0000
Cu(Lac)2(aq)	5. 471e-316	1. 322e-310	1. 0000	-300. 0000
Na(Lac)2-	2. 276e-316	4. 576e-311	0. 9163	-300. 0000
Cu(Pent)+	1. 710e-316	2. 814e-311	0. 9163	-300. 0000
K(Lac)2-	3. 587e-317	7. 789e-312	0. 9163	-300. 0000
Cd(Pent)+	7. 484e-318	1. 597e-312	0. 9163	-300. 0000
Cd(Lac)2(aq)	4. 441e-318	1. 290e-312	1. 0000	-300. 0000

	SpecE8_output_GSS_MW-95-8 June.txt			
Ba(Pent)+	2.983e-318	7.110e-313	0.9163	-300.0000
Ba(Lac)2(aq)	1.218e-318	3.840e-313	1.0000	-300.0000
Adi pate	1.539e-319	2.217e-314	0.7036	-300.0000
H-Adi pate	7.559e-322	1.097e-316	0.9163	-300.0000
Leucyl gl yci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Leuci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
K(Prop)2-	0.0000	0.0000	0.9163	-300.0000
Ba(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
1-Octanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
K(Pent)2-	0.0000	0.0000	0.9163	-300.0000
Cd(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Azel ate	0.0000	0.0000	0.7036	-300.0000
2-Hexanone(aq)	0.0000	0.0000	1.0000	-300.0000
K(But)2-	0.0000	0.0000	0.9163	-300.0000
I sol euci ne(aq)	0.0000	0.0000	1.0000	-300.0000
Azel ai c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
Hexanoate	0.0000	0.0000	0.9163	-300.0000
Hexanal (aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoi c_aci d(a	0.0000	0.0000	1.0000	-300.0000
2-Heptanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexyne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
Heptanoate	0.0000	0.0000	0.9163	-300.0000
Heptanal (aq)	0.0000	0.0000	1.0000	-300.0000
2-Butanone(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Adi pi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexene(aq)	0.0000	0.0000	1.0000	-300.0000
p-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
p-Tol uate	0.0000	0.0000	0.9163	-300.0000
o-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
o-Phthal ate	0.0000	0.0000	0.7036	-300.0000
n-Propyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Pentyl benzene(0.0000	0.0000	1.0000	-300.0000
H-Suberate	0.0000	0.0000	0.9163	-300.0000
H-Sebacate	0.0000	0.0000	0.9163	-300.0000
n-Hexyl benzene(a	0.0000	0.0000	1.0000	-300.0000
H-Pi mel ate	0.0000	0.0000	0.9163	-300.0000
n-Hexane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Heptyl benzene(0.0000	0.0000	1.0000	-300.0000
n-Heptane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Butyl benzene(a	0.0000	0.0000	1.0000	-300.0000
m-Tol ui c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000
m-Tol uate	0.0000	0.0000	0.9163	-300.0000
H-Azel ate	0.0000	0.0000	0.9163	-300.0000
Val i ne(aq)	0.0000	0.0000	1.0000	-300.0000
Undecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
Undecanoate	0.0000	0.0000	0.9163	-300.0000
1-Hexanol (aq)	0.0000	0.0000	1.0000	-300.0000
1-Butanami ne(aq)	0.0000	0.0000	1.0000	-300.0000
Ca(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
UO2(SCN)3-	0.0000	0.0000	0.9163	-300.0000
Ca(Al a)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ethyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Pentanone(aq)	0.0000	0.0000	1.0000	-300.0000
Tryptophan(aq)	0.0000	0.0000	1.0000	-300.0000
Tol uene(aq)	0.0000	0.0000	1.0000	-300.0000
Dodecanoi c_aci d(0.0000	0.0000	1.0000	-300.0000
Suberi c_aci d(aq)	0.0000	0.0000	1.0000	-300.0000

SpecE8_output_GSS_MW-95-8 June.txt

Suberate	0.0000	0.0000	0.7036	-300.0000
Dodecanoate	0.0000	0.0000	0.9163	-300.0000
2-Octanone(aq)	0.0000	0.0000	1.0000	-300.0000
1-Pentyne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Hexanamine(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Sebacate	0.0000	0.0000	0.7036	-300.0000
Decanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Tyrosine(aq)	0.0000	0.0000	1.0000	-300.0000
Decanoate	0.0000	0.0000	0.9163	-300.0000
Decanal(aq)	0.0000	0.0000	1.0000	-300.0000
S5--	0.0000	0.0000	0.7036	-300.0000
1-Pentene(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Benzoinic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Benzoate	0.0000	0.0000	0.9202	-300.0000
2-Hydroxyoctanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Phenylalanine(aq)	0.0000	0.0000	1.0000	-300.0000
Phenol(aq)	0.0000	0.0000	1.0000	-300.0000
Pentanal(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyoctanoic acid(aq)	0.0000	0.0000	0.9163	-300.0000
1-Pentanol(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptyne(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Octanoate	0.0000	0.0000	0.9163	-300.0000
Octanal(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Nonanoate	0.0000	0.0000	0.9163	-300.0000
Nonanal(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxynonanoic acid(aq)	0.0000	0.0000	0.9163	-300.0000
1-Pentanamine(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Prop)2-	0.0000	0.0000	0.9163	-300.0000
Ba(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Na(Pent)2-	0.0000	0.0000	0.9163	-300.0000
Cu(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyhexanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Cu(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
Ba(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Na(But)2-	0.0000	0.0000	0.9163	-300.0000
2-Hydroxyhexanoic acid(aq)	0.0000	0.0000	0.9163	-300.0000
1-Octyne(aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptene(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxyheptanoic acid(aq)	0.0000	0.0000	0.9163	-300.0000
1-Octene(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Prop)2(aq)	0.0000	0.0000	1.0000	-300.0000
Cd(Pent)2(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octane(aq)	0.0000	0.0000	1.0000	-300.0000
n-Octyl benzene(aq)	0.0000	0.0000	1.0000	-300.0000
n-Pentane(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoic acid(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(But)2(aq)	0.0000	0.0000	1.0000	-300.0000
Mg(Ala)2(aq)	0.0000	0.0000	1.0000	-300.0000
o-Toluate	0.0000	0.0000	0.9163	-300.0000
Methionine(aq)	0.0000	0.0000	1.0000	-300.0000
2-Hydroxydecanoic acid(aq)	0.0000	0.0000	0.9163	-300.0000

SpecE8_output_GSS_MW-95-8 June.txt

1-Octanol (aq)	0.0000	0.0000	1.0000	-300.0000
1-Heptanol (aq)	0.0000	0.0000	1.0000	-300.0000

Mineral saturation states

log Q/K

log Q/K

Witherite	1.9775s/sat	UO2S04: 3H2O	-17.4476
Quartz	0.6102s/sat	Na2SiO3	-17.5240
Baryte	0.4876s/sat	CdF2	-17.6531
Tridymite	0.4099s/sat	Hillebrandite	-17.8680
Chalcedony	0.3248s/sat	UO2F2: 3H2O	-18.3079
Dolomite-ord	0.2524s/sat	Dicalciumsilica	-18.3509
Dolomite	0.2523s/sat	UO2F2	-18.6354
Tenorite	0.0346s/sat	UO2.3333(beta)	-18.8393
Cristobalite(alp)	0.0252s/sat	CuCl2	-19.0624
Calcite	-0.0722	Akermanite	-19.2905
Ice	-0.0827	MgCl2: 4H2O	-19.3906
Aragonite	-0.2173	Larnite	-19.7313
Coesite	-0.2397	Anthophyllite	-19.9357
Malachite	-0.3214	UO2S04	-21.2376
Cristobalite(bet)	-0.4615	UOFOH: .5H2O	-21.2975
SiO2(am)	-0.8006	UOFOH	-21.8087
Monohydrocalcite	-0.8715	Lime	-22.0995
Magnesite	-1.3983	Na2UO4(alpha)	-22.1492
Dolomite-dis	-1.4046	UO2(N03)2: 6H2O	-22.6359
Dioptase	-1.7397	UO2Cl	-23.3338
Gypsum	-2.0840	Hydrophillite	-23.4446
Anhydrite	-2.4074	UO2(N03)2: 3H2O	-24.2975
Talc	-2.6284	UO2F2: H2O	-24.5892
Brochantite	-2.6390	MgCl2: 2H2O	-25.2404
Bassanite	-3.0570	UO2F2	-25.2440
CaSO4: 0.5H2O(bet)	-3.2443	Rankinite	-25.2976
Fluorite	-3.4249	UO2(N03)2: 2H2O	-25.6689
Cuprite	-3.7462	Tobermorite-14A	-26.8782
CdSiO3	-4.2212	K2UO4	-27.3542
Enstatite	-4.2305	MgCl2: H2O	-28.8303
Nesquehoniite	-4.5159	UO2Cl2: 3H2O	-28.9261
CaUO4	-4.5941	Tobermorite-11A	-29.1336
UO3: 2H2O	-5.0035	UO2(N03)2: H2O	-29.5137
Schoepite	-5.0037	Foshagite	-31.2820
Sellaite	-5.0130	Merwinite	-31.5283
UO2(OH)2(beta)	-5.1710	UO2Cl2: H2O	-31.8914
Schoepite-dehy(.UO3: .9H2O(alpha))	-5.2316	Ba2SiO4	-32.3050
Schoepite-dehy(.UO3: .9H2O(alpha))	-5.2390	Tobermorite-9A	-32.9862
Schoepite-dehy(.UO3: .9H2O(alpha))	-5.3246	UO2(N03)2	-33.2063
Schoepite-dehy(1UO3: .9H2O(alpha))	-5.3414	Afwillite	-33.5735
Nahcolite	-5.3954	UO2S03	-33.8147
Atacamite	-5.5572	KMgCl3: 2H2O	-33.9848
Okenite	-5.7072	Cd	-34.6667
Dioptide	-5.7722	Chloromagnesite	-34.9320
Brucite	-5.7751	UF4: 2.5H2O	-35.2289
Chrysotile	-5.8724	UO2Cl2	-36.0092
Wollastonite	-5.8988	BaSiF6	-39.0882
Azurite	-5.9263	UF4	-39.7851
Niter	-5.9458	BaO	-41.2060
Huntite	-5.9480	KMgCl3	-41.7718
UO2C03	-6.0046	C	-42.5290
Rutherfordine	-6.0179	U(SO4)2: 8H2O	-42.5860
Pseudowollastonite	-6.1651	U(SO4)2: 4H2O	-43.9201
Schoepite-dehy(.UO3: .9H2O(alpha))	-6.5158	U(SO4)2	-44.1850
Schoepite-dehy(.UO3: .9H2O(alpha))	-7.0672	Hatrurite	-44.3213
Cd(OH)2	-7.0824	Ba2U207	-44.6938
Nantokite	-7.3936	Xonotlite	-45.0809

SpecE8_output_GSS_MW-95-8 June.txt

Cu	-7.4503	U(CO3)2	-46.7436
Mirabilite	-7.8355	UOCl2	-47.0748
Sepiolite	-7.8642	UCIF3	-50.6519
Chalcanthite	-7.8773	U5O12Cl	-53.2856
Cd(OH)Cl	-7.9491	Na3UO4	-57.2754
Artinite	-8.0106	UOF4	-57.3260
Sanbornite	-8.0650	Na	-57.4147
UO3(gamma)	-8.1522	Na4SiO4	-59.0080
Sylvite	-8.3521	UOCl3	-60.3774
Halite	-8.5462	(UO2)2Cl3	-60.5461
Monteponite	-8.5856	UF3	-60.6354
UO3(beta)	-8.7877	Na2O	-61.4872
Nitrobarite	-8.8753	K	-61.8009
Arcanite	-8.9567	Covelite	-62.0718
UO3(al pha)	-9.1428	Chalcocite	-63.3984
Thenardite	-9.3436	UCI2F2	-63.8329
CdSO4: 2.667H2O	-9.7347	UF5(beta)	-63.9134
CdSO4: H2O	-9.9938	UOCl	-64.3196
BaUO4	-10.0580	UF5(al pha)	-64.3202
UO2.25	-10.1356	S	-64.3475
UO2FOH: 2H2O	-10.1782	CdS	-70.1181
UO2.25(beta)	-10.2044	Ba3UO6	-72.0229
Gyrolite	-10.2442	U2O3F6	-76.1096
Natron	-10.2492	UCI3F	-76.8296
Forsterite	-10.3463	Antigorite	-79.3871
Mg1.25SO4(OH)0.5	-10.5207	K2O	-80.1347
UO2FOH: H2O	-10.6378	CdCl2(NH3)2	-80.4583
Na2CO3: 7H2O	-10.8156	U(SO3)2	-80.5183
Periclase	-11.1661	Na6Si2O7	-85.8292
UO2FOH	-11.1666	UCI4	-87.6655
Monticellite	-11.2692	UCI3	-90.2705
Portlandite	-11.4836	UF6	-97.0293
CdSO4	-11.7224	Mg	-100.0942
Mg1.5SO4(OH)	-11.8916	U2F9	-100.3980
Tremolite	-12.2517	U3O5F8	-102.0872
Thermonatrite	-12.2937	BaS	-102.1571
Na2CO3	-12.6226	UN1.73(al pha)	-106.7039
UO2SO4: H2O	-12.8435	U2O2Cl5	-108.7434
MgSO4	-12.9642	UN1.59(al pha)	-109.5019
MgUO4	-12.9855	UCI5	-109.5882
Uraniinite	-13.0321	Ca	-117.4049
Coffinite	-13.2634	Ba	-121.9319
UO2.6667	-13.9099	UN	-122.8450
UO2ClOH: 2H2O	-13.9140	Si	-125.0598
BaU2O7	-14.0885	UCI6	-130.2961
CdCl2: H2O	-14.1067	CdCl2(NH3)4	-156.3925
Chalcocyanite	-14.1182	U	-171.0676
BaCl2: 2H2O	-14.1810	US	-176.2301
Na2U2O7	-14.4080	U4F17	-179.0560
BaCl2: H2O	-14.9183	UC	-195.3693
CdCl2	-15.2298	UH3(beta)	-198.9233
Ba(OH)2: 8H2O	-15.9748	US1.9	-199.3869
NaUO3	-16.3441	US2	-204.0135
BaCl2	-16.4731	CdCl2(NH3)6	-232.4220
Natrosilite	-16.8240	UC1.94(al pha)	-237.4766
Hydromagnesite	-16.8649	US3	-265.0291
MgOHCl	-17.0397	U2S3	-372.9677
Ba2Si3O8	-17.1204	o-Phthalic_acid	-378.4368
UO2SO4: 3.5H2O	-17.3069	U2C3	-434.8687
UO2SO4: 2.5H2O	-17.3706	U3S5	-572.1633

Gases

fugacity

log fug.

SpecE8_output_GSS_MW-95-8 June.txt

N2(g)	0.03354	-1.474
H2O(g)	0.01017	-1.993
CO2(g)	0.002960	-2.529
NO2(g)	2.886e-011	-10.540
HF(g)	2.500e-015	-14.602
HCl(g)	2.652e-019	-18.576
H2(g)	1.480e-028	-27.830
O2(g)	1.976e-033	-32.704
NO(g)	1.008e-033	-32.997
CO(g)	1.465e-034	-33.834
NH3(g)	6.640e-040	-39.178
SiF4(g)	3.298e-041	-40.482
SO2(g)	1.850e-042	-41.733
Cl2(g)	3.935e-045	-44.405
Cd(g)	6.028e-050	-49.220
Cu(g)	2.023e-063	-62.694
UO2F2(g)	5.009e-064	-63.300
Na(g)	1.321e-072	-71.879
K(g)	6.189e-074	-73.208
UO2Cl2(g)	4.644e-075	-74.333
UO3(g)	5.051e-076	-75.297
UOF4(g)	3.000e-079	-78.523
UF5(g)	8.738e-086	-85.059
H2S(g)	7.787e-087	-86.109
UF4(g)	3.463e-087	-86.461
CH4(g)	2.414e-089	-88.617
UF6(g)	4.926e-099	-98.307
F2(g)	1.326e-103	-102.878
UCl4(g)	1.168e-114	-113.933
UO2(g)	7.899e-116	-115.102
Mg(g)	7.193e-122	-121.143
UCl5(g)	5.691e-129	-128.245
UF3(g)	4.057e-132	-131.392
UCl6(g)	1.218e-137	-136.914
UCl3(g)	1.259e-139	-138.900
U2F10(g)	6.028e-144	-143.220
S2(g)	1.711e-144	-143.767
Ca(g)	5.610e-145	-144.251
C2H4(g)	7.345e-154	-153.134
C(g)	2.009e-167	-166.697
UF2(g)	1.138e-173	-172.944
UCl2(g)	1.026e-182	-181.989
UO(g)	9.378e-189	-188.028
Si(g)	5.918e-201	-200.228
UF(g)	2.087e-208	-207.680
U2Cl8(g)	3.711e-217	-216.430
UCl(g)	2.976e-223	-222.526
U2Cl10(g)	4.722e-229	-228.326
U(g)	2.955e-262	-261.529

Original basis	total moles	In fluid moles	mg/kg	Sorbed moles	mg/kg	Kd L/kg
<hr/>						
Ba++	1.06e-006	1.06e-006	0.146			
Ca++	0.00166	0.00166	66.5			
Cd++	2.57e-008	2.57e-008	0.00289			
Cl-	0.000110	0.000110	3.91			
Cu++	1.69e-007	1.69e-007	0.0108			
F-	5.15e-006	5.15e-006	0.0977			
H+	0.000151	0.000151	0.152			
H2O	55.5	55.5	1.00e+006			
HCO3-	0.00261	0.00261	159.			
K+	0.000225	0.000225	8.80			

SpecE8_output_GSS_MW-95-8 June.txt

Mg++	0.000322	0.000322	7.82
NH3(aq)	5.69e-005	5.69e-005	0.969
NO3-	0.00171	0.00171	106.
Na+	0.00106	0.00106	24.4
O2(aq)	4.25e-005	4.25e-005	1.36
SO4--	0.000366	0.000366	35.2
SiO2(aq)	0.000181	0.000181	10.8
UO2++	2.05e-009	2.05e-009	0.000554

Elemental composition	In fluid		Sorbed	
	total moles	moles	mg/kg	moles mg/kg
Barium	1.061e-006	1.061e-006	0.1456	
Cadmium	2.575e-008	2.575e-008	0.002893	
Calcium	0.001659	0.001659	66.46	
Carbon	0.002612	0.002612	31.36	
Chlorine	0.0001103	0.0001103	3.909	
Copper	1.693e-007	1.693e-007	0.01075	
Fluorine	5.147e-006	5.147e-006	0.09774	
Hydrogen	111.0	111.0	1.118e+005	
Magnesium	0.0003218	0.0003218	7.819	
Nitrogen	0.001768	0.001768	24.75	
Oxygen	55.52	55.52	8.880e+005	
Potassium	0.0002251	0.0002251	8.796	
Silicon	0.0001806	0.0001806	5.071	
Sodium	0.001063	0.001063	24.43	
Sulfur	0.0003664	0.0003664	11.74	
Uranium	2.054e-009	2.054e-009	0.0004887	